



(2+2) Resonance-Enhanced
Multiphoton Ionization (REMPI) and
Photoacoustic (PA) Spectroscopic
Detection of Nitric Oxide (NO)
and Nitrogen Dioxide (NO₂)
Near 454 nm

by R. C. Sausa
and R. L. Pastel

ARL-TR-1418

July 1997

DTIC QUALITY INSPECTED 2

Approved for public release; distribution is unlimited.

19970815 039

The findings in this report are not to be construed as an official Department of the Army position unless so designated by other authorized documents.

Citation of manufacturer's or trade names does not constitute an official endorsement or approval of the use thereof.

Destroy this report when it is no longer needed. Do not return it to the originator.

Army Research Laboratory

Aberdeen Proving Ground, MD 21005-5066

ARL-TR-1418

July 1997

(2+2) Resonance-Enhanced Multiphoton Ionization (REMPI) and Photoacoustic (PA) Spectroscopic Detection of Nitric Oxide (NO) and Nitrogen Dioxide (NO₂) Near 454 nm

R. C. Sausa, R. L. Pastel

Weapons and Materials Research Directorate, ARL

Abstract

Trace concentrations of NO and NO₂ are detected with a dye laser operating near 454 nm. NO is detected by a (2 + 2) resonance-enhanced multiphoton ionization process by means of NO A ²Σ⁺–X ²Π(0, 0) transitions with miniature electrodes, and NO₂ is detected by a one-photon absorption photoacoustic process by means of NO₂ $\tilde{A}'$ ²B₁(0, 8, 0) – $\tilde{X}$ ²A₁(0, 0, 0) transitions with a miniature microphone. Rotationally resolved excitation spectra show that the spectral resolution is sufficiently high to identify these species at 1 atm. The technique's analytical merits are evaluated as functions of concentration, pressure, and laser intensities. Low laser intensities favor NO₂ photoacoustic detection whereas high laser intensities favor NO ionization. Limits of detection (signal-to-noise ratio 3) of 160 parts in 10⁹ for NO and 400 parts in 10⁹ for NO₂ are determined at 1 atm for a 10-s integration time. Signal response and noise analyses show that three decades of NO/NO₂ mixtures can be measured with a computational relative error in concentration that is three times the relative error in measuring the NO and NO₂ signals.

ACKNOWLEDGMENTS

This work was supported by the U.S. Army Research Laboratory Director's Research Initiative (ARL-DRI) and the Strategic Environmental Research and Development Program (SERDP) on cleanup, ARMY-713-94 and ARMY-729-94. Purchase of equipment through the Productivity Capital Investment Program (R. C. Sausa) and the ARL-SCEEE Postdoctoral Research Fellowship Program (R. L. Pastel) is gratefully acknowledged.

INTENTIONALLY LEFT BLANK.

TABLE OF CONTENTS

	<u>Page</u>
ACKNOWLEDGMENTS	iii
LIST OF FIGURES	vii
LIST OF TABLES	vii
1. INTRODUCTION	1
2. EXPERIMENTAL	3
3. RESULTS AND DISCUSSION	5
4. CONCLUSION	15
5. REFERENCES	17
DISTRIBUTION LIST	19
REPORT DOCUMENTATION PAGE	27

INTENTIONALLY LEFT BLANK.

LIST OF FIGURES

<u>Figure</u>	<u>Page</u>
1. Schematic of the experimental apparatus	4
2. Partial potential energy diagram of NO and NO ₂	6
3. Laser excitation spectra of NO and NO ₂ obtained by (2+2) REMPI and PA spectroscopy, respectively, near 454 nm	8
4. Signal dependence on laser intensity for 147-ppm NO ₂ PA (▲) without lens, 86-ppm NO REMPI (●), 78-ppm NO ₂ REMPI (◆) using a 12-cm lens, and 333-ppm NO REMPI using a 50-cm lens (■)	10
5. Plots of NO REMPI and NO ₂ PA signals as a function of N ₂ pressure at (a) constant density and (b) constant mixture	11
6. Sensitivity plots of (a) NO PA (●) and NO ₂ PA (■), and (b) NO REMPI (●) and NO ₂ REMPI (■), at 453.856 nm	12

LIST OF TABLES

<u>Table</u>	<u>Page</u>
1. NO and NO ₂ Responses by REMPI and PA Detection at Laser Wavelengths of 453.856 and 454.348 nm	13
2. NO and NO ₂ LOD (ppm) by REMPI and PA Detection at Laser Wavelengths of 453.856 and 454.348 nm	14

INTENTIONALLY LEFT BLANK.

1. INTRODUCTION

There is a growing interest in laser-based analytical techniques for remote or *in situ* trace detection of nitric oxide (NO) and nitrogen dioxide (NO₂) [1, 3]. Much of this interest stems from concerns related to public health and the environment. These compounds play key roles in catalytic ozone destruction and in acid rain and photochemical smog formation. NO and NO₂ are hazardous pollutants emitted predominantly from motor vehicle exhaust and stationary sources, such as electric utilities and industrial boilers. The U.S. Federal Environmental Protection Agency (EPA) has established a 25-ppm National Ambient Air Quality Standard (NAAQS) threshold limit for NO with concentrated exposures not to exceed 100 ppm for 15 min [4]. NO₂ is estimated to be 30 times more toxic than NO [4]. The detection of NO and NO₂ is also important to laser photofragmentation/fragment detection techniques being developed for the chemical analysis of propellants and explosives because they are generated in the photolysis of many energetic materials [5-7].

Laser photoacoustic (PA) and resonance-enhanced multiphoton ionization (REMPI) spectroscopy are sensitive techniques for NO and NO₂ trace detection. PA detection is based on sensing the pressure wave generated by the heat released from an excited analyte caused by collisional deactivation. PA detection is particularly effective at high pressures for weak fluorescers or species that predissociate with laser absorption. High sensitivity and selectivity can be achieved using a narrow band, high-power pulsed laser. REMPI detection is based on sensing the ion signal following excitation of the analyte. The ionization process is enhanced if there are electronic states resonant with the energy of one or more of the photons. Ion detection can be accomplished using a mass spectrometer or a pair of miniature electrodes. With electrode-based detection, the bulky mass spectrometer is no longer needed. However, species selectivity is then based solely on excitation wavelength. In both PA and REMPI detection, using a laser frequently increases the response without increasing the noise.

The analytical application of PA for NO₂ detection using visible laser radiation has been reported. Fried measured the 488-nm PA detection limit of NO₂ in NO, N₂, H₂O and O₂

matrices with identical sensitivities of 5 ppb ($S/N=1$) in all matrix gases except O_2 [8]. Oxygen decreased the NO_2 signal by 20% because of electronic energy transfer between the two molecules. From low-resolution excitation NO_2 spectra recorded in the range 480–625 nm, Claspy et al. [9] estimated a 4-ppb sensitivity ($S/N=1$) at 600 nm for a laser power of 1 W [9]; Terhune and Anderson measured NO_2 sensitivities better than 0.1 ppb using a 514.5-nm Ar^+ ion laser [10]; and Angus et al. estimated a 10-ppb NO_2 detection limit from low-resolution cw dye laser excitation scans over the range of 580–610 nm [11].

Recently, we have studied NO and NO_2 detection by ultraviolet (UV) laser spectroscopy [12, 3]. NO was detected by (1+1) REMPI via its $A^2\Sigma^+-X^2\Pi$ (0,0) transitions near 226 nm, while NO_2 was detected by laser photofragmentation with subsequent fragment NO ionization also using 226-nm radiation. The NO and NO_2 limits of detection (LOD) at 100 Torr were 1 and 22 ppb, respectively, for a $S/N=3$, a laser energy of 10 μJ , and 10-s integration time [12]. The two species could not be differentiated, however, because only the total NO ion signal was measured.

The detection and discrimination of NO and NO_2 by a single, laser-based apparatus are important analytical challenges. Part of the challenge stems from NO and NO_2 absorbing in different spectral regions. NO_2 absorbs in the visible; whereas, NO absorbs in the UV. NO_2 predissociates at wavelengths less than 400 nm making ionization and PA detection difficult. Although both NO and NO_2 absorb in the infrared, few lasers can be tuned in the region where both species absorb. Also, H_2O is a major spectral interferant in the infrared.

In the present study we employ a single laser operating near 454 nm to detect trace NO and NO_2 concentrations in N_2 at atmospheric pressure by (2+2) REMPI and PA spectroscopy. Rotationally resolved excitation spectra are recorded and characterized. To the best of our knowledge, this is the first time that a high-resolution NO_2 visible PA spectrum has been reported. The effects of laser energy, pressure, and species concentration on the REMPI and PA signals are determined and discussed. A single laser coupled to a pair of miniature electrodes and a

microphone minimizes the instrumental complexity and makes the technique attractive for NO and NO₂ detection.

2. EXPERIMENTAL

The experimental schematic is depicted in Figure 1. An excimer-pumped dye laser (Lumonics, HYPER EX-400, and HYPER DYE-300) operating at 10 Hz with a line width of $\sim 0.07 \text{ cm}^{-1}$ was used to excite NO or NO₂ around 454 nm. The energy per pulse was typically 15 mJ with better than 10% shot-to-shot variation, and the pulse duration was $\sim 20 \text{ ns}$ (FWHM). A calorimeter (Scientech) and Joulemeter (Molelectron Detector, J4-05) monitored the laser energy. The laser beam diameter was 3 mm. Two 5-cm-diameter quartz lenses with focal lengths of 12 and 50 cm focused the laser beam for REMPI detection. The laser beam was not focused for PA detection.

The sample cell was a six-arm (4-cm diameter) stainless steel cross. Opposing arms with quartz Brewster windows provided optical access to the sample. Two other arms provided access for a pair of laboratory-constructed stainless steel electrodes and a directional electret condenser microphone (Radio Shack 270-090), both mounted on separate vacuum feedthrough flanges. A mechanical pump drew the sample gas through the cell, and a needle valve upstream of the cell regulated the flow rate. The sample gas pressure was varied to 1 atm and monitored with a capacitance manometer (Edwards 600A-1000T-R16-H21X). Various concentration of NO and NO₂ were prepared by serial dilution of 0.1% NO in N₂ or 500-ppm NO₂ in N₂ with N₂ buffer gas. All the gases were obtained from Matheson with a purity of >99.999%. The NO measurements were recorded using N₂ as a diluent instead of air to eliminate reduction of NO by the $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ three-body reaction. At room temperature, 50% of NO (100 ppm) oxidizes to NO₂ in 40 min, the time required for a typical experimental run. The NO₂ measurements were also recorded using N₂ buffer gas for experimental consistency. Samples were flowed through the cell to prevent accumulation of photolysis products.

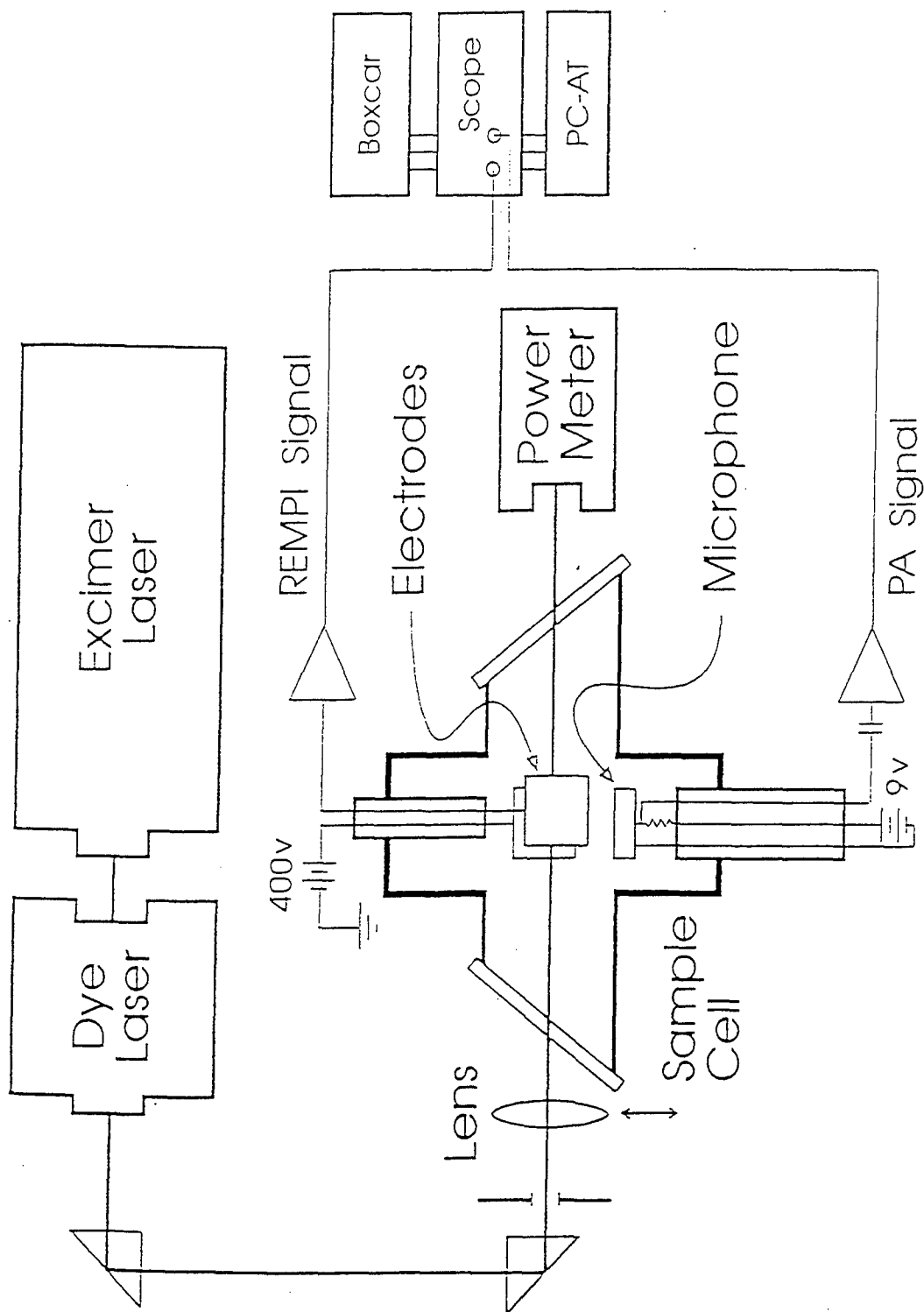


Figure 1. Schematic of the experimental apparatus.

For REMPI detection, the laser beam was focused at the center of two square electrodes (1.25×1.25 cm) separated by 3 mm. The collection voltage was 400 V for all measurements. The ion signal was amplified by a current amplifier (Keithly 427, gain 10^5 – 10^6 V/A, time constant 0.01 ms) and displayed on a 125-MHz digital oscilloscope (LeCroy 9400). For PA detection, the electrodes were removed from the laser-interaction region and a small microphone was positioned 2–3 mm from the laser beam. The focusing lens was also removed from the beam path. The microphone was encased in a small cylinder (9-mm diameter by 6 mm long) and had a 3-mm^2 active area. Its frequency response was flat from 20 to 1500 Hz. A 9-V battery and 1-k Ω resistor biased the microphone, and a 3.3- μ F capacitor coupled the signal to the current amplifier. Signals from the amplifier were then directed to the digital oscilloscope. A PC-AT computer interfaced to a gated integrator (Stanford Research Systems SR250) with 3-shot averaging recorded and stored the spectra. The sensitivity signals were integrated for 10 s at 10 Hz for 100-shot averaging and read from the oscilloscope. The background noise was measured with the laser operating at the excitation frequencies and N₂ flowing through the sample cell. Twenty independent measurements of the noise with 100-shot averaging were made. The limiting noise in this study is due to fluctuations in background ionization and laser energy for the photoionization detection and window vibrations for the PA detection.

3. RESULTS AND DISCUSSION

A potential-energy diagram that shows the physical processes in NO₂ PA and NO (2+2) REMPI detection is presented in Figure 2. In PA detection, 454-nm radiation excites NO₂ into the $\tilde{A}'$ (2B_1) or $\tilde{A}$ (2B_2) states. Radiationless deactivation of excited NO₂ by inter- and intramolecular interactions causes heat release inducing a pressure wave that generates the PA signal. Loss mechanisms for PA detection include ionization, dissociation, energy transfer, and fluorescence. PA detection is favored over ionization or dissociation at low laser intensities. At low pressures, fluorescence competes effectively with PA detection. The 454-nm laser excitation of NO₂ results in both discrete fluorescence and continuum emission [13–15]. The continuum emission is prompt (<100 ns) and decays by intramolecular radiationless transitions [14]. Longer time-scale quenching mechanisms include intermolecular and buffer gas radiationless collisions.

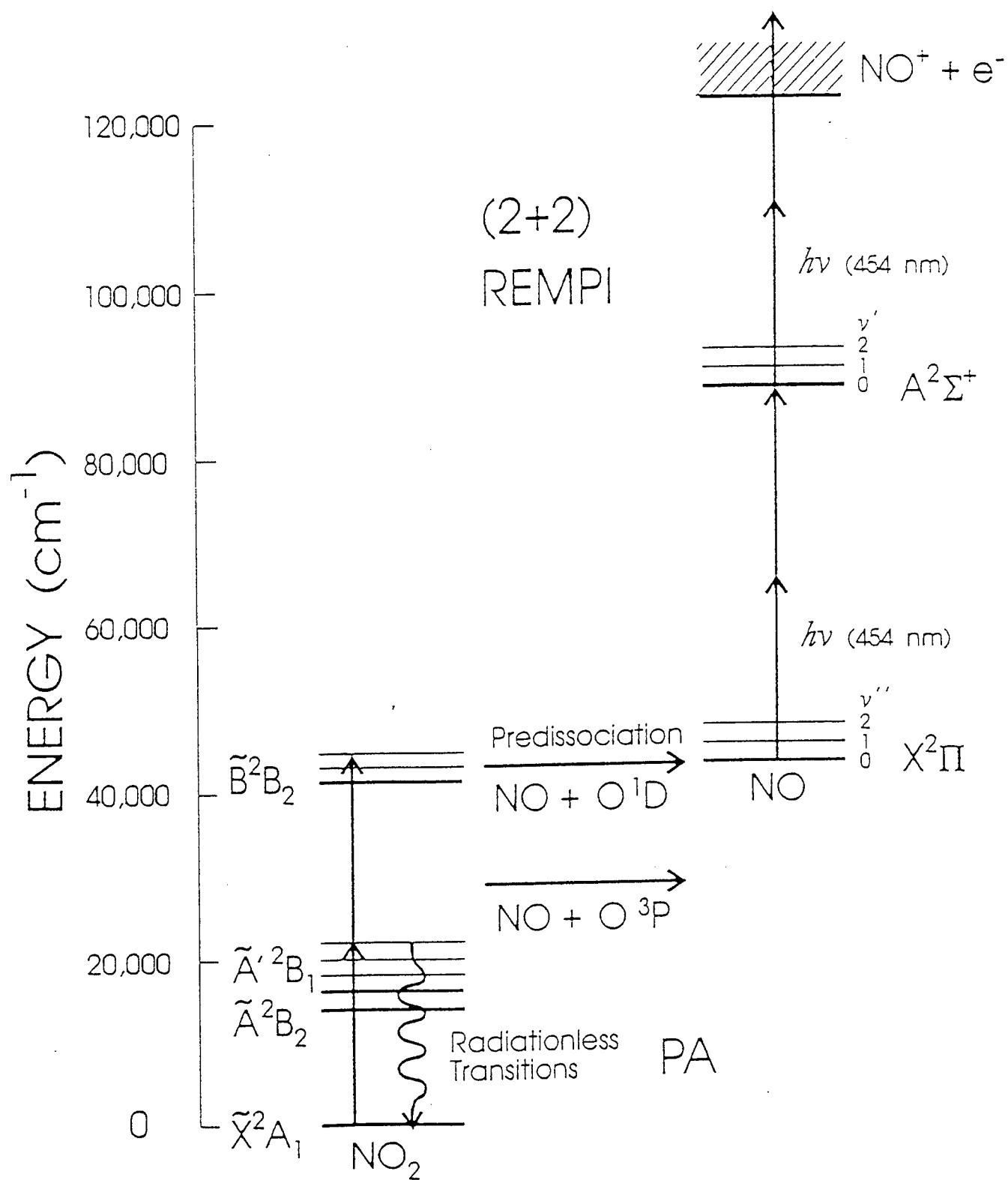


Figure 2. Partial potential energy diagram of NO and NO₂.

In the visible and at low laser intensities, NO has a small absorption coefficient, hence, no PA response. At high laser intensities, NO is ionized by 4-photon absorption, (2+2) REMPI. The ionization process is enhanced at 454 nm by 2-photon resonance with the $A^2\Sigma^+$ state. At high laser intensities NO_2 also contributes to the “ambient” NO REMPI signal because it photodissociates. As shown in Figure 2, two 454-nm photons excite NO_2 into the $\tilde{B}(^2B_2)$ state. Electronically excited NO_2 then quickly [16] (<40 psec) predissociates [17] into NO $X^2\Pi + O(^3P)$ and NO $X^2\Pi + O(^1D)$. For the NO $X^2\Pi + O(^1D)$ dissociation pathway, NO is predominantly formed in the $v''=0$ and 1 vibrational levels because of intramolecular relaxation in the intermediate $\tilde{A}'$ and $\tilde{A}$ states that are resonant with the energy of one 454-nm photon [18]. As a result, the nascent NO ($v''=0$) from NO_2 photoionizes just like ambient NO ($v''=0$).

NO (2+2) REMPI and NO_2 PA laser excitation spectra were recorded between 440–470 nm at various concentrations and pressures. Presented in Figure 3 are typical NO_2 PA and NO REMPI spectra recorded in the 453.5–454.5-nm region. The laser was focused for NO REMPI detection, but not for NO_2 PA detection. The concentrations of NO and NO_2 in N_2 were 95 and 65 ppm, respectively, and the cell pressure was approximately 1 atm. The prominent features of spectrum A are attributed to the $P_{22}+O_{12}$, $P_{12}+O_{22}$, and O_{12} branches of the $A^2\Sigma^+-X^2\Pi(0,0)$ band. These features are the result of 2-photon selection rules governing the (2+2) NO REMPI by means of the resonant $A^2\Sigma^+$ state. Thus, the visible REMPI spectrum exhibits strong, main O and S branches ($\Delta J=\pm 2$) in addition to the P, Q, and R branches ($\Delta J=0,\pm 1$) observed in 226-nm (1+1) NO REMPI. An enhancement of the rotational lines of the O_{12} branch (e.g., $J = 10.5$) that is due to double resonance processes is not observed as in low-pressure environments since these processes are quenched at 1 atm. The O and S branches enhance the NO REMPI selectivity because they contribute additional “fingerprints” for NO identification. In addition, the O_{12} branch is red-shifted from all the other congested branches and exhibits a larger rotational energy spacing. These features are particularly attractive for NO detection at atmospheric conditions because collisions degrade the spectral resolution.

The visible spectrum of NO_2 is complex due to Coriolis and Renner-Teller rovibronic interactions. Imasaka et al. characterized the vibrational levels of NO_2 2B_1 and 2B_2 states by

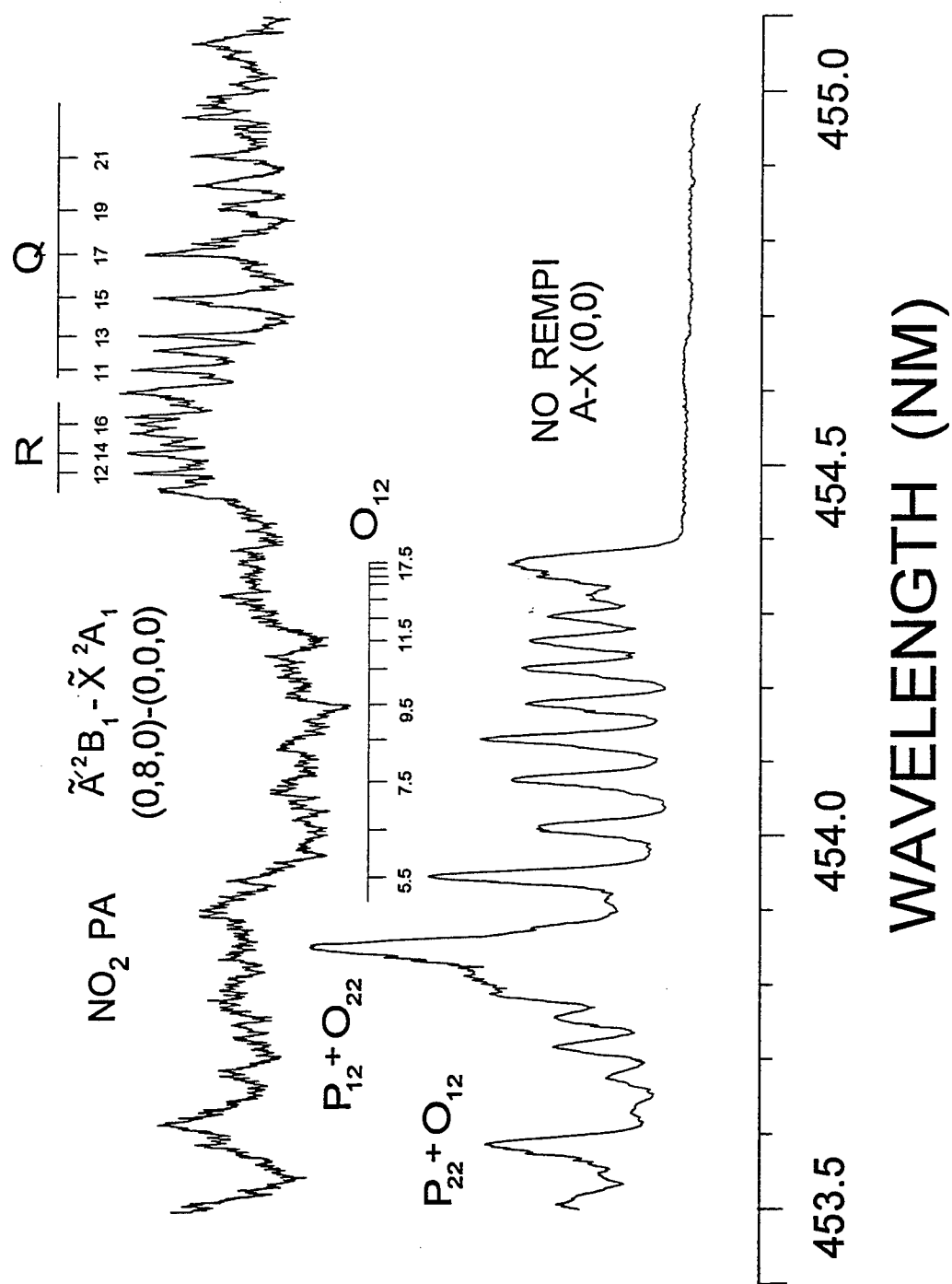


Figure 3. Laser excitation spectra of NO and NO₂ obtained by (2+2) REMPI and PA spectroscopy, respectively, near 454 nm.

time-resolved spectroscopy [19]. Based on their work, and the works of Hardwick [20] and Douglas and Huber [21], we assign the sharp features of the NO₂ to P and Q rotational lines of the $\tilde{A}'^2B_1-X^2A_1$ [(0,8,0)-(0,0,0)] band. Other vibrational bands with higher intensities were observed, but the (0,8,0) band was studied because of its proximity to the NO O₁₂ branch. As shown in Figure 3, the rotational resolution at 1 atm is sufficiently high to discriminate NO and NO₂ by laser wavelength excitation.

Presented in Figure 4 is the laser energy dependence of NO₂ PA, NO REMPI, and NO from NO₂ REMPI signals with 12- and 50-cm lenses. Fitting the data to powers of the laser energy yields powers of 1 for NO₂ PA, 3.9 for NO REMPI when a 50-cm lens is used, and 1.9 and 3.0 for NO and NO₂ REMPI, respectively, when a 12-cm lens is used. The NO₂ PA signal response is approximately linear with laser energy and suggests an unsaturated 1-photon absorption process, as expected for PA detection. The NO (2+2) REMPI quartic dependence indicates a 4-photon process at low laser intensity; whereas, the near quadratic dependence indicates saturation from the intermediate $A^2\Sigma^+$ state into the ionization continuum subsequent to 2-photon absorption. A geometric effect caused by strong conical focusing of the laser beam is ruled out, as it would yield a 3/2-power dependence [22]. The REMPI response using the 12-cm lens was 100 times greater than that using the 50-cm lens. Thus, the 12-cm lens was used for the pressure and LOD studies. Also, the laser beam was not focused for the NO₂ PA studies because focusing enhanced multiphoton dissociation and decreased the signal.

Figure 5(a) shows the N₂ buffer gas pressure dependence on NO₂ PA and NO REMPI signals at a fixed NO or NO₂ density. The NO₂ PA signal increases until around 400 Torr and then levels off. The PA signal is proportional to $kM/(kM+A)$, where A is the spontaneous transition rate, k is the quenching rate constant, and M is the density of the buffer gas [23]. Thus, a leveling off occurs when the radiationless quenching rate, kM, is greater than radiation rate, A. For NO REMPI, the signal increases to a maximum near 100 Torr and then decreases nonexponentially. The increase is due to charge amplification by N₂; whereas, the decrease is due to NO $A^2\Sigma^+$ quenching by N₂ and three-body recombination of NO⁺, N₂, and electrons.

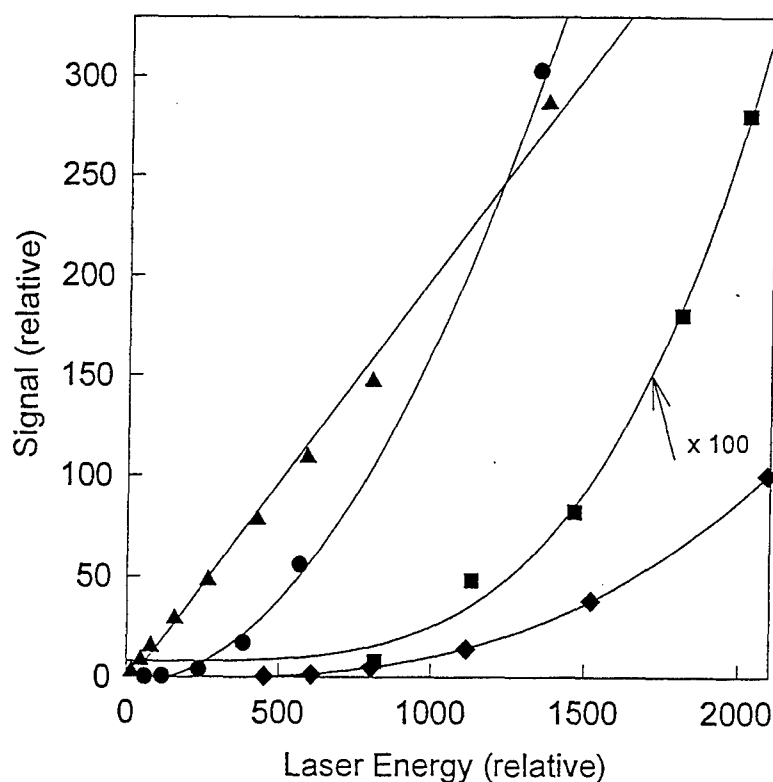


Figure 4. Signal dependence on laser intensity for 147-ppm NO₂ PA (▲) without lens, 86-ppm NO REMPI (●), 78-ppm NO₂ REMPI (◆) using a 12-cm lens, and 333-ppm NO REMPI using a 50-cm lens (■).

Presented in Figure 5(b) is the PA and REMPI signal dependence on pressure for 86-ppm NO and 147-ppm NO₂. As the density of NO or NO₂ increases with the pressure, these plots represent the combined pressure and density effects. The NO₂ PA signal has a 0.7 power dependence with pressure. This tendency to saturate is due to the pressure saturation observed in Figure 5(a). The NO REMPI response is approximately linear with pressure and suggests that the pressure effects at fixed density, illustrated in Figure 5(a), are small compared to the NO REMPI response from increased density.

Figures 6(a) and 6(b) show sensitivity plots of the NO₂ and NO PA and NO and NO₂ REMPI, respectively, at 453.856 nm. Measurements were made under flowing conditions at ~1 atm by serial dilution of NO or NO₂ with N₂. The ionization signals were measured at the peak ion signal; whereas, the PA signals were measured as the difference between the first maximum and

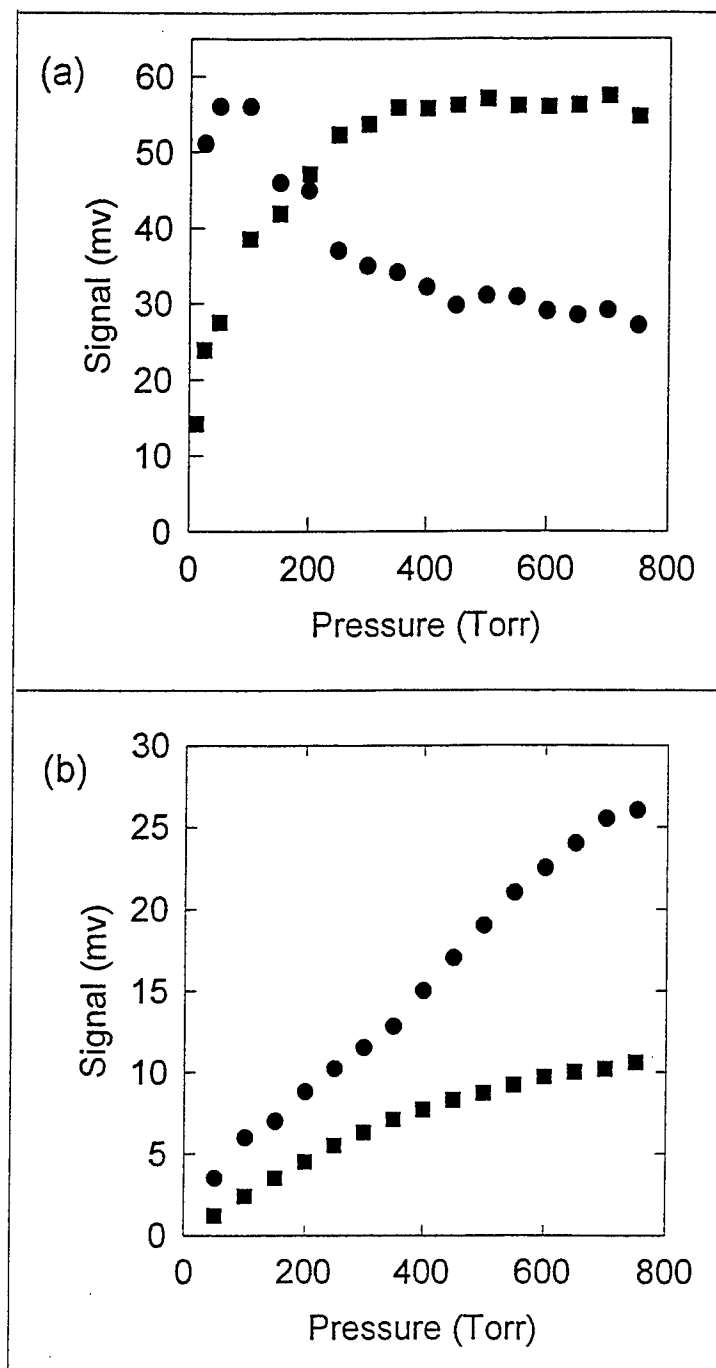


Figure 5. Plots of NO REMPI and NO₂ PA signals as a function of N₂ pressure at (a) constant density and (b) constant mixture. For (a), the densities are [NO₂] = 3.7 × 10¹⁴ cm⁻³ (■) and [NO] = 1.5 × 10¹⁵ cm⁻³ (●). For (b), the mixtures are 86-ppm NO in N₂ (●) 147-ppm NO₂ in N₂ (■).

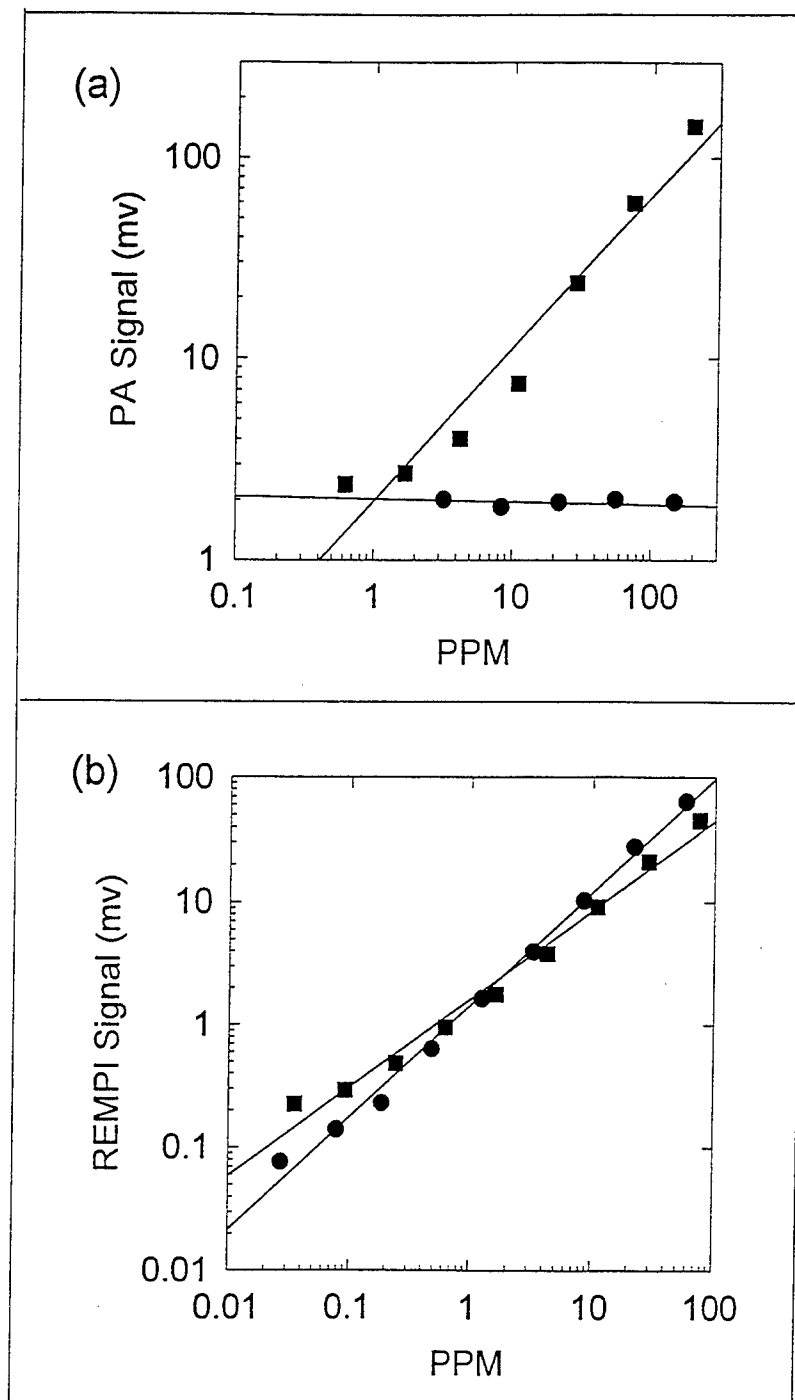


Figure 6. Sensitivity plots of (a) NO PA (●) and NO₂ PA (■), and (b) NO REMPI (●) and NO₂ REMPI (■), at 453.856 nm. The data are plotted on log scales.

minimum of the pressure wave. The slope of the sensitivity plot is the response (millivolts/parts per million [mV/ppm]). Responses for NO and NO₂ PA and REMPI are listed in Table 1 for both 453.856-nm and 454.348-nm laser excitation. The NO₂ PA signals are linear to 400 ppb, whereas the NO and NO₂ REMPI signals are linear to 20 ppb. The NO PA responses were below the sensitivity of the apparatus and are listed as zero.

Table 1. NO and NO₂ Responses by REMPI and PA Detection at Laser Wavelengths of 453.856 and 454.348 nm.

	453.856 nm		454.348 nm	
	NO	NO ₂	NO	NO ₂
PA	0.0	0.74	0.0	0.71
REMPI	1.12	0.61	0.33	0.18

In this study, the LOD is defined as the concentration that produces a signal equaling three times the standard deviation of the noise. The standard deviation of the noise was determined from 20 independent 100-shot averages with only N₂ flowing through the sample cell. Table 2 shows the LOD for NO and NO₂ in N₂ by REMPI and PA detection at the two-laser-excitation wavelengths. The NO₂ PA LOD would be approximately 5% less in air [8] compared with N₂; whereas, the NO and NO₂ REMPI LODs would not change [12]. The NO REMPI LOD at 453.856 nm is approximately a factor of 3 less than that measured at 454.348 nm. This difference is due to the difference in signals, apparent in the NO REMPI spectrum, because the noise is same for both laser wavelengths. The NO₂ PA LODs are similar for both wavelengths. We project higher NO₂ PA sensitivities by using the P and Q transitions near 454.7 nm or other vibrational bands that are more intense than the (0, 8, 0) band.

An upper bound of the relative response error due to computing the relative concentration error of the mixtures is determined by numerical analysis. The combined PA and REMPI signals can be expressed by two simultaneous linear equations, whose matrix representation is:

$$\begin{bmatrix} R_{PA}(NO) & R_{PA}(NO_2) \\ R_{REMPI}(NO) & R_{REMPI}(NO_2) \end{bmatrix} \begin{bmatrix} [NO] \\ [NO_2] \end{bmatrix} = \begin{bmatrix} S_{PA} \\ S_{REMPI} \end{bmatrix} \quad (1)$$

Table 2. NO and NO₂ LOD (ppm) by REMPI and PA Detection at Laser Wavelengths of 453.856 and 454.348 nm

Laser Wavelength	PA		(2+2) REMPI	
	NO	NO ₂	NO	NO ₂
453.856 nm	—	0.41	0.18	0.34
454.348 nm	—	0.42	0.63	1.10

Equation 1 can be rewritten in vector notation as, $\mathbf{R}\mathbf{x} = \mathbf{s}$, where the elements of the response matrix, $\mathbf{R}$, are the NO and NO₂ responses to PA and REMPI detection; the elements of the signal vector, $\mathbf{s}$, are the total PA and REMPI signals from a NO-NO₂ mixture; and the elements of unknown vector, $\mathbf{x}$, are the NO and NO₂ concentrations. The condition number of the response matrix [24], $K = \|\mathbf{R}\| \|\mathbf{R}^{-1}\|$, where $\|\mathbf{R}\|$ is the norm of $\mathbf{R}$, scales the relative signal error, $\|\Delta\mathbf{s}\|/\|\mathbf{s}\|$, to the relative concentration error, $\|\Delta\mathbf{x}\|/\|\mathbf{x}\|$. The condition numbers for the 453.856- and 454.348-nm PA and REMPI detection systems are 2.8 and 2.7, respectively, using the values presented in Table 1 and row norm. Thus, the relative error in calculating the NO and NO₂ concentrations will be no more than 2.8 times the relative signal error.

The range of NO-NO₂ mixtures that can be accurately measured by REMPI is determined by the noise in the signals while measuring the NO and NO₂ responses. The signal noise is $N_s = n(\Delta E/E)m^{-1/2}$, where n is the power of the signal dependence on laser energy (see Figure 4), $\Delta E/E$ is the variation in laser energy, and m is the number of laser pulses. The signal noises for NO and NO₂ are 0.02 and 0.03, respectively, for 10% (shot-to-shot) laser energy variation and 100-shot averaging. Requiring the ratio of signals from NO and NO₂ to be greater than three times the relative signal noise yields,

$$3N_s \left(\frac{R_{\text{REMPI}}(\text{NO}_2)}{R_{\text{REMPI}}(\text{NO})} \right) \leq \frac{[\text{NO}]}{[\text{NO}_2]} \leq \left(\frac{1}{3N_s} \right) \left(\frac{R_{\text{REMPI}}(\text{NO}_2)}{R_{\text{REMPI}}(\text{NO})} \right) \quad (2)$$

Thus, nearly three decades of NO-NO₂ mixtures centered around [NO]/[NO₂]=1/2 can be accurately measured using our present apparatus.

4. CONCLUSION

The analytical utility of a combined PA and REMPI technique has been demonstrated for the trace detection of NO and NO₂ in N₂ at atmospheric pressure using a narrow-band tunable dye-laser operating near 454 nm. Rotationally resolved spectra of the NO A²Σ⁺-X²Π (0,0) and NO₂ $\tilde{A}'^2\text{B}_1(0,8,0)$ -X²A₁(0,0,0) bands suggest that the technique is highly selective based on excitation wavelength. The sensitivity of the technique is also high with LODs (S/N=3) of 100 and 400 ppb for (2+2) NO REMPI and NO₂ PA, respectively, for 10-s integration time. Higher sensitivities are projected with an increase in laser energy, different excitation schemes, and an improved apparatus. The results reveal that the technique enables a simple instrument design to be used for the sensitive measurement of these compounds and NO/NO₂ mixtures.

INTENTIONALLY LEFT BLANK.

5. REFERENCES

1. Simeonsson, J. B., and R. C. Sausa. "A Critical Review of Laser Photofragmentation/Fragment Detection for Gas-Phase Chemical Analysis." Review of Applied Spectroscopy, vol. 31, p. 1, 1996.
2. Jaegle, L., C. R. Webster, R. D. May, D. W. Fahey, E. L. Woodbridge, E. R. Keim, R. S. Gao, M. H. Proffitt, R. M. Stimple, R. J. Salawith, S. C. Wofsy, and L. Pfister. "In Situ Measurements of the NO₂/NO Ratio for Testing Atmospheric Photochemical Models." Geophysical Research Letters, vol. 21, pp. 2555–2558, 1994.
3. Kolsch, H. J., P. Rairoux, J. P. Wolf, and L. Woste. "Simultaneous NO and NO₂ DIAL Measurement Using BBO Crystals." Applied Optics, vol. 26, pp. 2052–2056, 1989.
4. Last, J. A., W. M. Sun, and H. Witschi. "Ozone, NO, and NO₂: Oxidation and Air Pollutants and More." Environmental Health Perspectives, vol. 102, supplement 10, p. 179, 1994.
5. Lemire, G. W., J. B. Simeonsson, and R. C. Sausa. "Monitoring of Vapor-Phase Nitro-Compounds Using 226-nm Radiation: Fragmentation with Subsequent NO Resonance-Enhanced Multiphoton Ionization Detection." Analytical Chemistry, vol. 65, pp. 529–533, 1993.
6. Simeonsson, J. B., G. W. Lemire, and R. C. Sausa. "Trace Detection of Nitrocompounds by ArF Laser Photofragmentation/Ionization Spectrometry." Applied Spectroscopy, vol. 47, pp. 1907–1912, 1993.
7. Marshall, A., A. Clark, K. W. D. Ledingham, J. Sander, R. P. Singhal, C. Kosmidis, and R. M. Deas. "Detection and Identification of Explosives Compounds Using Laser Ionization Time-of-flight Techniques." Rapid Communications in Mass Spectrometry, vol. 8, pp. 521–526, 1994.
8. Fried, A. "A Study of Measurement Interference in the Optoacoustic Detection of NO₂ by Argon-Ion Laser Excitation." Applied Spectroscopy, vol. 36, pp. 562–565, 1982.
9. Claspy, P. C., C. Ha, and Y. H. Pao. "Optoacoustic Detection of NO₂ Using a Pulsed Dye Laser." Applied Optics, vol. 16, pp. 2972–2973, 1977.
10. Terhune, R. W., and J. E. Anderson. "Spectrophone Measurements of the Absorption of Visible by Aerosols in the Atmosphere." Optic Letter, vol. 1, pp. 70–72, 1973.
11. Angus, A. M., E. E. Marinero, and M. J. Colles. "Opto-acoustic Spectroscopy with a Visible CW Dye Laser." Optic Communication, vol. 14, pp. 223–225, 1975.

12. Simeonsson, J. B., G. W. Lemire, and R. C. Sausa. "Laser-Induced Photofragmentation/Photoionization Spectrometry: A Method for Detection Ambient Oxides of Nitrogen." Analytical Chemistry, vol. 66, pp. 2272–2278, 1994.
13. Donnelly, V. M., and F. Kaufman. "Fluorescence Lifetime Studies of NO₂. II. Dependence of the Perturbed ²B₂ State Lifetimes on Excitation Energy." Journal Chemical Physics, vol. 69, pp. 1456–1460, 1978.
14. Imasake, T., T. Ogawa, and N. Ishibashi. "Inter- and Intra-molecular Radiationless Transitions of NO₂ at Around 454.6 nm." Chemical Physics, vol. 45, pp. 273–278, 1980.
15. Hass, Y., P. L. Houston, J. H. Clark, C. B. Moore, H. Rosen, and P. Robrish. "Long-Lived K_a=0, ²B₁ States of NO₂: A Direct Measurements Using a Tunable Dye Laser." Journal Chemical Physics, vol. 63, pp. 4195–4197, 1975.
16. Tapper, R. S., T. L. Whetten, G. S. Ezra, and E. R. Grant. "The Role of Near-Resonant Intermediate States in the Two-Photon Excitation of NO₂: Origin Bands in Bent-to-Linear Transitions." Journal Physical Chemistry, vol. 88, pp. 1273–1275, 1984.
17. Morrison, R. J. S., B. H. Rockney, and E. R. Grant. "Multiphoton Ionization of NO₂: Spectroscopy and Dynamics." Journal Chemical Physics, vol. 75, pp. 2643–2651, 1981.
18. Bigio, L., R. S. Tapper, and E. R. Grant. "The Role of Near-Resonant Intermediate States in the Two-Photon Excitation of NO₂: The Distinct Dynamics of Two-Photon Photofragmentation." Journal Physical Chemistry, vol. 88, pp. 1271–1273, 1984.
19. Imasaka, T., T. Ogawa, and N. Ishibashi. "Time-Resolved Spectroscopy on the Excited Electronic States of NO₂ in the Neighborhood of 454.6 nm." Journal Chemical Physics, vol. 70, pp. 881–885, 1979.
20. Hardwick, J. L. "Fluorescence from the ²B₁ State of NO₂ Excited at 4545 Å." Journal of Molecular Spectroscopy, vol. 66, pp. 248–258, 1977.
21. Douglas, A. E., and K. P. Huber. "The Absorption Spectrum of NO₂ in 3700–4600 Å Region." Canadian Journal Physics, vol. 43, p. 74, 1965.
22. Zakheim, D. S., and P. M. Johnson. "Rate Equation of Molecular Multiphoton Ionization Dynamics." Chemical Physics, vol. 46, pp. 263–272, 1980.
23. Williamson, C. K., R. L. Pastel, and R. C. Sausa. "Detection of Ambient NO by Laser-Induced Photoacoustic Spectrometry Using A²Σ–X²Π (0,0) Transitions Near 226 nm." Applied Spectroscopy, vol. 50, pp. 205–210, 1996.
24. Conte, S.D., and C. de Boor. Elementary Numerical Analysis. New York: McGraw-Hill, 1972.

<u>NO. OF COPIES</u>	<u>ORGANIZATION</u>
2	DEFENSE TECHNICAL INFORMATION CENTER DTIC DDA 8725 JOHN J KINGMAN RD STE 0944 FT BELVOIR VA 22060-6218
1	HQDA DAMO FDQ DENNIS SCHMIDT 400 ARMY PENTAGON WASHINGTON DC 20310-0460
1	CECOM SP & TRRSTRL COMMCTN DIV AMSEL RD ST MC M H SOICHER FT MONMOUTH NJ 07703-5203
1	PRIN DPTY FOR TCHNLGY HQ US ARMY MATCOM AMCDCG-T M FISETTE 5001 EISENHOWER AVE ALEXANDRIA VA 22333-0001
1	PRIN DPTY FOR ACQUSTN HQS US ARMY MATCOM AMCDCG-A D ADAMS 5001 EISENHOWER AVE ALEXANDRIA VA 22333-0001
1	DPTY CG FOR RDE HQS US ARMY MATCOM AMCRD BG BEAUCHAMP 5001 EISENHOWER AVE ALEXANDRIA VA 22333-0001
1	ASST DPTY CG FOR RDE HQS US ARMY MATCOM AMCRD COL S MANESS 5001 EISENHOWER AVE ALEXANDRIA VA 22333-0001

<u>NO. OF COPIES</u>	<u>ORGANIZATION</u>
1	DPTY ASSIST SCY FOR R&T SARD TT F MILTON THE PENTAGON RM 3E479 WASHINGTON DC 20310-0103
1	DPTY ASSIST SCY FOR R&T SARD TT D CHAIT THE PENTAGON WASHINGTON DC 20310-0103
1	DPTY ASSIST SCY FOR R&T SARD TT K KOMINOS THE PENTAGON WASHINGTON DC 20310-0103
1	DPTY ASSIST SCY FOR R&T SARD-TT B REISMAN THE PENTAGON WASHINGTON DC 20310-0103
1	DPTY ASSIST SCY FOR R&T SARD-TT T KILLION THE PENTAGON WASHINGTON DC 20310-0103
1	OSD OUSD(A&T)/ODDDR&E(R) J LUPO THE PENTAGON WASHINGTON DC 20301-7100
1	ARL ELECTROMAG GROUP CAMPUS MAIL CODE F0250 A TUCKER UNIVERSITY OF TX AUSTIN TX 78712
1	DUSD SPACE 1E765 J G MCNEFF 3900 DEFENSE PENTAGON WASHINGTON DC 20301-3900
1	USAASA MOAS AI W PARRON 9325 GUNSTON RD STE N319 FT BELVOIR VA 22060-5582

NO. OF
COPIES ORGANIZATION

- 1 CECOM
PM GPS COL S YOUNG
FT MONMOUTH NJ 07703
- 1 GPS JOINT PROG OFC DIR
COL J CLAY
2435 VELA WAY STE 1613
LOS ANGELES AFB CA 90245-5500
- 1 ELECTRONIC SYS DIV DIR
CECOM RDEC
J NIEMELA
FT MONMOUTH NJ 07703
- 3 DARPA
L STOTTS
J PENNELLA
B KASPAR
3701 N FAIRFAX DR
ARLINGTON VA 22203-1714
- 1 SPCL ASST TO WING CMNDR
50SW/CCX
CAPT P H BERNSTEIN
300 O'MALLEY AVE STE 20
FALCON AFB CO 80912-3020
- 1 USAF SMC/CED
DMA/JPO M ISON
2435 VELA WAY STE 1613
LOS ANGELES AFB CA 90245-5500
- 1 US MILITARY ACADEMY
MATH SCI CTR OF EXCELLENCE
DEPT OF MATHEMATICAL SCI
MDN A MAJ DON ENGEN
THAYER HALL
WEST POINT NY 10996-1786
- 1 DIRECTOR
US ARMY RESEARCH LAB
AMSRL CS AL TP
2800 POWDER MILL RD
ADELPHI MD 20783-1145
- 1 DIRECTOR
US ARMY RESEARCH LAB
AMSRL CS AL TA
2800 POWDER MILL RD
ADELPHI MD 20783-1145

NO. OF
COPIES ORGANIZATION

- 3 DIRECTOR
US ARMY RESEARCH LAB
AMSRL CI LL
2800 POWDER MILL RD
ADELPHI MD 20783-1145
- ABERDEEN PROVING GROUND
- 2 DIR USARL
AMSRL CI LP (305)

NO. OF
COPIES ORGANIZATION

1 HQDA
SARD TT
MR J APPEL
WASH DC 20310-0103

1 HQDA OASA RDA
DR C H CHURCH
PENTAGON ROOM 3E486
WASH DC 20310-0103

4 COMMANDER
US ARMY RESEARCH OFC
R GHIRARDELLI
D MANN
R SINGLETON
R SHAW
P O BOX 12211
RESEARCH TRIANGLE PARK NC
27709-2211

1 DIRECTOR
ARMY RESEARCH OFFICE
AMXRO RT IP LIB SRVCS
P O BOX 12211
RESEARCH TRIANGLE PARK NC
27709-2211

2 COMMANDER
US ARMY ARDEC
AMSTA AR AEE B
D S DOWNS
PICATINNY ARSENAL NJ
07806-5000

2 COMMANDER
US ARMY ARDEC
AMSTA AR AEE J A LANNON
PICATINNY ARSENAL NJ
07806-5000

1 COMMANDER
US ARMY ARDEC
AMSTA AR AEE BR
L HARRIS
PICATINNY ARSENAL NJ
07806-5000

NO. OF
COPIES ORGANIZATION

2 COMMANDER
US ARMY MISSILE COMMAND
AMSMI RD PR E A R MAYKUT
AMSMI RD PR P R BETTS
REDSTONE ARSENAL AL

1 OFFICE OF NAVAL RESEARCH
DEPARTMENT OF THE NAVY
R S MILLER CODE 432
800 N QUINCY STREET
ARLINGTON VA 22217

1 COMMANDER
NAVAL AIR SYSTEMS COMMAND
J RAMNARACE AIR-54111C
WASHINGTON DC 20360

2 COMMANDER
NSWC
R BERNECKER R-13
G B WILMOT R-16
SILVER SPRING MD 20903-5000

5 COMMANDER
NAVAL RSRCH LAB
M C LIN
J MCDONALD
E ORAN
J SHNUR
R J DOYLE CODE 6110
WASHINGTON DC 20375

2 COMMANDER
NAVAL WEAPONS CENTER
T BOGGS CODE 388
T PARR CODE 3895
CHINA LAKE CA 93555-6001

1 SUPERINTENDENT
NAVAL POSTGRDTE SCHL
DEPT OF AERONAUTICS
D W NETZER
MONTEREY CA 93940

3 AL LSCF
R CORLEY
R GEISLER
J LEVINE
EDWARDS AFB CA 93523-5000

<u>NO. OF</u> <u>COPIES</u>	<u>ORGANIZATION</u>	<u>NO. OF</u> <u>COPIES</u>	<u>ORGANIZATION</u>
1	AFOSR J M TISHKOFF BOLLING AIR FORCE BASE WASHINGTON DC 20332	2	PRINCETON COMBUSTION RSRCH LABORATORIES INC N A MESSINA M SUMMERFIELD PRINCETON CORPORATE PLAZA BLDG IV SUITE 119 11 DEERPARK DRIVE MONMOUTH JUNCTION NJ 08852
1	OSD SDIO IST L CAVENY PENTAGON WASHINGTON DC 20301-7100		
1	COMMANDANT USAFAS ATSF TSM CN FORT SILL OK 73503-5600	3	DIRECTOR SANDIA NATIONAL LABS DIVISION 8354 S JOHNSTON P MATTERN D STEPHENSON LIVERMORE CA 94550
1	UNIV OF DAYTON RSRCH INST D CAMPBELL AL PAP EDWARDS AFB CA 93523	1	BRIGHAM YOUNG UNIVERSITY DEPT OF CHMCL ENGNRNG M W BECKSTEAD PROVO UT 84058
1	NASA LANGLEY RESEARCH CENTER LANGLEY STATION G B NORTHAM MS 168 HAMPTON VA 23365	1	CALIFORNIA INST OF TECH JET PROPULSION LAB L STRAND MS 125 224 4800 OAK GROVE DRIVE PASADENA CA 91109
4	NTNL BUREAU OF STNDRDS J HASTIE M JACOX T KASHIWAGI H SEMERJIAN US DEPT OF COMMERCE WASHINGTON DC 20234	1	CALIFORNIA INSTITUTE OF TECHNOLOGY F E C CULICK MC 301-46 204 KARMAN LAB PASADENA CA 91125
2	DIRECTOR LLNL C WESTBROOK W TAO MS L 282 P O BOX 808 LIVERMORE CA 94550	1	UNIV OF CALIFORNIA LOS ALAMOS SCNTFC LAB P O BOX 1663 MAIL STOP B216 LOS ALAMOS NM 87545
1	DIRECTOR LOS ALAMOS NATIONAL LAB B NICHOLS T7 MS-B284 P O BOX 1663 LOS ALAMOS NM 87545	1	UNIV OF CA BERKELEY CHEMISTRY DEPARMENT C BRADLEY MOORE 211 LEWIS HALL BERKELEY CA 94720
		1	UNIV OF CA SAN DIEGO F A WILLIAMS AMES B010 LA JOLLA CA 92093

NO. OF
COPIES ORGANIZATION

2 UNIV OF CA SANTA BARBARA
QUANTUM INSTITUTE
K SCHOFIELD
M STEINBERG
SANTA BARBARA CA 93106

1 UNIV OF CO AT BOULDER
ENGINEERING CENTER
J DAILY
CAMPUS BOX 427
BOULDER CO 80309-0427

3 UNIV OF SOUTHERN CA
DEPT OF CHEMISTRY
R BEAUDET
S BENSON
C WITTIG
LOS ANGELES CA 90007

1 CORNELL UNIVERSITY
DEPT OF CHEMISTRY
T A COOL
BAKER LABORATORY
ITHACA NY 14853

1 UNIV OF DELAWARE
T BRILL
CHEMISTRY DEPARTMENT
NEWARK DE 19711

1 UNIVERSITY OF FLORIDA
DEPT OF CHEMISTRY
J WINEFORDNER
GAINESVILLE FL 32611

3 GA INST OF TECHNOLOGY
SCHL OF AERSPCE ENGNRNG
E PRICE
W C STRAHLE
B T ZINN
ATLANTA GA 30332

1 UNIVERSITY OF ILLINOIS
DEPT OF MECH ENG
H KRIER
144MEB 1206 W GREEN ST
URBANA IL 61801

NO. OF
COPIES ORGANIZATION

1 THE JOHNS HOPKINS UNIV
CPIA
T W CHRISTIAN
10630 LTLE PTXNT PKWY STE 202
COLUMBIA MD 21044-3200

1 UNIVERSITY OF MICHIGAN
GAS DYNAMICS LAB
AEROSPACE ENGNRNG BLDG
G M FAETH
ANN ARBOR MI 48109-2140

1 UNIVERSITY OF MINNESOTA
DEPT OF MCHNCL ENGNRNG
E FLETCHER
MINNEAPOLIS MN 55455

4 PA STATE UNIVERSITY
DEPT OF MCHNCL ENGNRNG
K KUO
M MICCI
S THYNELL
V YANG
UNIVERSITY PARK PA 16802

2 PRINCETON UNIVERSITY
FORRESTAL CAMPUS LIB
K BREZINSKY
I GLASSMAN
P O BOX 710
PRINCETON NJ 08540

1 PURDUE UNIVERSITY
SCHOOL OF AERO & ASTRO
J R OSBORN
GRISSOM HALL
WEST LAFAYETTE IN 47906

1 PURDUE UNIVERSITY
DEPT OF CHEMISTRY
E GRANT
WEST LAFAYETTE IN 47906

2 PURDUE UNIVERSITY
SCHL OF MCHNCL ENGNRNG
N M LAURENDEAU
S N B MURTHY
TSPC CHAFFEE HALL
WEST LAFAYETTE IN 47906

<u>NO. OF COPIES</u>	<u>ORGANIZATION</u>
1	RENSSELAER PLYTCHNC INST DEPT OF CHMCL ENGNRNG A FONTIJN TROY NY 12181
1	STANFORD UNIVERSITY DEPT OF MCHNCL ENGNRNG R HANSON STANFORD CA 94305
1	UNIVERSITY OF TEXAS DEPT OF CHEMISTRY W GARDINER AUSTIN TX 78712
1	VIRGINIA PLYTCHNC INST AND STATE UNIVERSITY A SCHETZ BLACKSBURG VA 24061
1	APPLIED COMBUSTION TECHNOLOGY INC A M VARNEY P O BOX 607885 ORLANDO FL 32860
2	APPLIED MCHNCS REVIEWS THE AMERICAN SOCIETY OF MECHANICAL ENGINEERS R E WHITE A B WENZEL 345 E 47TH STREET NEW YORK NY 10017
1	BATTELLE TWSTIAC 505 KING AVENUE COLUMBUS OH 43201-2693
1	COHEN PRFSSNL SERVICES N S COHEN 141 CHANNING STREET REDLANDS CA 92373
1	EXXON RSRCH & ENGRNG CO A DEAN ROUTE 22E ANNANDALE NJ 08801

<u>NO. OF COPIES</u>	<u>ORGANIZATION</u>
1	GENERAL APPLIED SCIENCE LABORATORIES INC 77 RAYNOR AVENUE RONKONKAMA NY 11779-6649
1	GENERAL ELECTRIC ORDNANCE SYSTEMS J MANDZY 100 PLASTICS AVENUE PITTSFIELD MA 01203
1	GENERAL MOTORS RSCH LABS PHYSCL CHMSTRY DEPT T SLOANE WARREN MI 48090-9055
2	HERCULES INC ALLEGHENY BALLISTICS LAB W B WALKUP E A YOUNT P O BOX 210 ROCKET CENTER WV 26726
1	HERCULES INC R V CARTWRIGHT 100 HOWARD BLVD KENVIL NJ 07847
1	ALLIANT TECHSYSTEMS INC MARINE SYSTEMS GROUP D E BRODEN MS MN50-2000 600 2ND STREET NE HOPKINS MN 55343
1	ALLIANT TECHSYSTEMS INC R E TOMPKINS MN 11 2720 600 SECOND ST NORTH HOPKINS MN 55343
1	IBM CORPORATION A C TAM RESEARCH DIVISION 5600 COTTLE ROAD SAN JOSE CA 95193
1	IIT RESEARCH INSTITUTE R F REMALY 10 WEST 35TH STREET CHICAGO IL 60616

<u>NO. OF COPIES</u>	<u>ORGANIZATION</u>
1	LOCKHEED MSLS & SPACE CO GEORGE LO 3251 HANOVER STREET DEPT 52-35 B204 2 PALO ALTO CA 94304
1	OLIN ORDNANCE V MCDONALD LIBRARY P O BOX 222 ST MARKS FL 32355-0222
1	PAUL GOUGH ASSOCIATES INC P S GOUGH 1048 SOUTH STREET PORTSMOUTH NH 03801-5423
1	HUGHES AIRCRAFT COMPANY T E WARD 8433 FALLBROOK AVENUE CANOGA PARK CA 91303
1	ROCKWELL INTRNTNL CORP ROCKETDYNE DIVISION J E FLANAGAN HB02 6633 CANOGA AVENUE CANOGA PARK CA 91304
1	SCIENCE APPLICATIONS INC R B EDELMAN 23146 CUMORAH CREST WOODLAND HILLS CA 91364
3	SRI INTERNATIONAL G SMITH D CROSLY D GOLDEN 333 RAVENSWOOD AVENUE MENLO PARK CA 94025
1	STEVENS INST OF TECH DAVIDSON LABORATORY R MCALEVY III HOBOKEN NJ 07030
1	SVERDRUP TCHNLGY INC LERC GROUP R J LOCKE MS SVR 2 2001 AEROSPACE PKWY BROOK PARK OH 44142

<u>NO. OF COPIES</u>	<u>ORGANIZATION</u>
1	SVERDRUP TCHNLGY INC J DEUR 2001 AEROSPACE PARKWAY BROOK PARK OH 44142
3	THIOKOL CORPORATION ELKTON DIVISION R BIDDLE R WILLER TECH LIB P O BOX 241 ELKTON MD 21921
3	THIOKOL CORPORATION WASATCH DIVISION S J BENNETT P O BOX 524 BRIGHAM CITY UT 84302
1	UNITED TCHNLGS RSRCH CTR A C ECKBRETH EAST HARTFORD CT 06108
1	UNITED TECHNOLOGIES CORP CHEMICAL SYSTEMS DIVISION R R MILLER P O BOX 49028 SAN JOSE CA 95161-9028
1	UNIVERSAL PRPLSN CO H J MCSPADDEN 25401 NORTH CENTRAL AVE PHOENIX AZ 85027-7837
1	VERITAY TECHNOLOGY INC E B FISHER 4845 MILLERSPORT HWY P O BOX 305 EAST AMHERST NY 14051-0305
1	FREEDMAN ASSOCIATES E FREEDMAN 2411 DIANA ROAD BALTIMORE MD 21209-1525
1	ALLIANT TECHSYSTEMS INC J BODE 600 SECOND ST NE HOPKINS MN 55343

<u>NO. OF</u> <u>COPIES</u>	<u>ORGANIZATION</u>	<u>NO. OF</u> <u>COPIES</u>	<u>ORGANIZATION</u>
1	ALLIANT TECHSYSTEMS INC C CANDLAND 600 SECOND ST NE HOPKINS MN 55343		<u>ABERDEEN PROVING GROUND</u>
1	ALLIANT TECHSYSTEMS INC L OSGOOD 600 SECOND ST NE HOPKINS MN 55343	41	DIR, USARL ATTN: AMSRL-WM-P, A.W. HORST AMSRL-WM-PC, B.E. FORCH G.F. ADAMS W.R. ANDERSON R.A. BEYER S.W. BUNTE C.F. CHABALOWSKI K.P. MCNEILL- BOONSTOPPEL A. COHEN R. CUMPTON R. DANIEL D. DEVYNCK R.A. FIFER J.M. HEIMERL B.E. HOMAN A. JUHASZ A.J. KOTLAR R. KRANZE E. LANCASTER W.F. MCBRATNEY K.L. MCNESBY M. MCQUAID N.E. MEAGHER M.S. MILLER A.W. MIZIOLEK J.B. MORRIS J.E. NEWBERRY S.V. PAI R.A. PESCE-RODRIGUEZ J. RASIMAS P. REEVES B.M. RICE P. SAEGAR R.C. SAUSA M.A. SCHROEDER R. SCHWEITZER L.D. SEGER J.A. VANDERHOFF D. VENIZELOS A. WHREN H.L. WILLIAMS
1	ALLIANT TECHSYSTEMS INC R BURETTA 600 SECOND ST NE HOPKINS MN 55343		
1	ALLIANT TECHSYSTEMS INC R BECKER 600 SECOND ST NE HOPKINS MN 55343		
1	ALLIANT TECHSYSTEMS INC M SWENSON 600 SECOND ST NE HOPKINS MN 55343		
1	BENET LABORATORIES SAM SOPOK AMSTA AR CCB B WATERVLIET NY 12189		

REPORT DOCUMENTATION PAGE			Form Approved OMB No. 0704-0188	
Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing the collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden, to Washington Headquarters Services, Directorate for Information Operations and Reports, 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302, and to the Office of Management and Budget, Paperwork Reduction Project (0704-0188), Washington, DC 20503.				
1. AGENCY USE ONLY (Leave blank)		2. REPORT DATE July 1997		3. REPORT TYPE AND DATES COVERED Final, Jan 96 - Jan 97
4. TITLE AND SUBTITLE (2+2) Resonance-Enhanced Multiphoton Ionization (REMPI) and Photoacoustic (PA) Spectroscopic Detection of Nitric Oxide (NO) and Nitrogen Dioxide (NO ₂) Near 454 nm			5. FUNDING NUMBERS PR: 1L161102AH43	
6. AUTHOR(S) R. C. Sausa and R. L. Pastel				
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) U.S. Army Research Laboratory ATTN: AMSRL-WM-PC Aberdeen Proving Ground, MD 21005-5066			8. PERFORMING ORGANIZATION REPORT NUMBER ARL-TR-1418	
9. SPONSORING/MONITORING AGENCY NAMES(S) AND ADDRESS(ES)			10. SPONSORING/MONITORING AGENCY REPORT NUMBER	
11. SUPPLEMENTARY NOTES				
12a. DISTRIBUTION/AVAILABILITY STATEMENT Approved for public release; distribution is unlimited.			12b. DISTRIBUTION CODE	
13. ABSTRACT (Maximum 200 words) Trace concentrations of NO and NO ₂ are detected with a dye laser operating near 454 nm. NO is detected by a (2 + 2) resonance-enhanced multiphoton ionization process by means of NO A ² Σ ⁺ - X ² Π(0,0) transitions with miniature electrodes, and NO ₂ is detected by a one-photon absorption photoacoustic process by means of NO ₂ $\tilde{A}'$ ² B ₁ (0, 8, 0) - $\tilde{X}$ ² A ₁ (0, 0, 0) transitions with a miniature microphone. Rotationally resolved excitation spectra show that the spectral resolution is sufficiently high to identify these species at 1 atm. The technique's analytical merits are evaluated as functions of concentration, pressure, and laser intensities. Low laser intensities favor NO ₂ photoacoustic detection whereas high laser intensities favor NO ionization. Limits of detection (signal-to-noise ratio 3) of 160 parts in 10 ⁹ for NO and 400 parts in 10 ⁹ for NO ₂ are determined at 1 atm for a 10-s integration time. Signal response and noise analyses show that three decades of NO/NO ₂ mixtures can be measured with a computational relative error in concentration that is three times the relative error in measuring the NO and NO ₂ signals.				
14. SUBJECT TERMS photoacoustic spectroscopy, resonance-enhanced multiphoton ionization, nitric oxide, nitrogen dioxide, visible radiation, laser			15. NUMBER OF PAGES 31	
			16. PRICE CODE	
17. SECURITY CLASSIFICATION OF REPORT UNCLASSIFIED	18. SECURITY CLASSIFICATION OF THIS PAGE UNCLASSIFIED	19. SECURITY CLASSIFICATION OF ABSTRACT UNCLASSIFIED	20. LIMITATION OF ABSTRACT UL	

INTENTIONALLY LEFT BLANK.

USER EVALUATION SHEET/CHANGE OF ADDRESS

This Laboratory undertakes a continuing effort to improve the quality of the reports it publishes. Your comments/answers to the items/questions below will aid us in our efforts.

1. ARL Report Number/Author ARL-TR-1418 (Sausa) Date of Report July 1997

2. Date Report Received _____

3. Does this report satisfy a need? (Comment on purpose, related project, or other area of interest for which the report will be used.) _____

4. Specifically, how is the report being used? (Information source, design data, procedure, source of ideas, etc.) _____

5. Has the information in this report led to any quantitative savings as far as man-hours or dollars saved, operating costs avoided, or efficiencies achieved, etc? If so, please elaborate. _____

6. General Comments. What do you think should be changed to improve future reports? (Indicate changes to organization, technical content, format, etc.) _____

CURRENT
ADDRESS

Organization

Name

E-mail Name

Street or P.O. Box No.

City, State, Zip Code

7. If indicating a Change of Address or Address Correction, please provide the Current or Correct address above and the Old or Incorrect address below.

OLD
ADDRESS

Organization

Name

Street or P.O. Box No.

City, State, Zip Code

(Remove this sheet, fold as indicated, tape closed, and mail.)
(DO NOT STAPLE)

DEPARTMENT OF THE ARMY

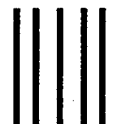
OFFICIAL BUSINESS

BUSINESS REPLY MAIL

FIRST CLASS PERMIT NO 0001,APG,MD

POSTAGE WILL BE PAID BY ADDRESSEE

DIRECTOR
US ARMY RESEARCH LABORATORY
ATTN AMSRL WM PC
ABERDEEN PROVING GROUND MD 21005-5066



NO POSTAGE
NECESSARY
IF MAILED
IN THE
UNITED STATES

