

REPORT DOCUMENTATION PAGE			Form Approved OMB No. 0704-0188	
Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing the collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden, to Washington Headquarters Services, Directorate for Information Operations and Reports, 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302, and to the Office of Management and Budget, Paperwork Reduction Project (0704-0188), Washington, DC 20503.				
1. AGENCY USE ONLY (Leave blank)	2. REPORT DATE 21 May 1996	3. REPORT TYPE AND DATES COVERED Technical		
4. TITLE AND SUBTITLE SYNTHESIS OF NOVEL TETRAMETALLIC AND TRIMETALLIC COMPOUNDS FROM REACTIONS OF MMe_3 ($\text{M} = \text{Al}, \text{Ga}$) WITH $1,2\text{-(H}_2\text{E)}_2\text{C}_6\text{H}_4$ ($\text{E} = \text{N}, \text{P}$)		5. FUNDING NUMBERS •N00014-95-1-0194 R&T Project 3135008---16 •Dr. Harold E. Guard		
6. AUTHOR(S) Richard L. Wells, Hamid Rahbarnoohi, Paul B. Glaser Louise M. Liable-Sands and Arnold L. Rheingold				
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) Department of Chemistry Duke University Durham, NC 27708-0346		8. PERFORMING ORGANIZATION REPORT NUMBER Technical Report No. DU/DC/TR-56		
9. SPONSORING/MONITORING AGENCY NAME(S) AND ADDRESS(ES) Office of Naval Research 300 North Quincy Street Arlington, VA 22217-5000		10. SPONSORING/MONITORING AGENCY REPORT NUMBER		
11. SUPPLEMENTARY NOTES Accepted for Publication <i>Organometallics</i>				
12a. DISTRIBUTION/AVAILABILITY STATEMENT Approved for Public Release Distribution Unlimited		12b. DISTRIBUTION CODE		
13. ABSTRACT (Maximum 200 words) The independent 2:1 reactions of MMe_3 ($\text{M} = \text{Al}, \text{Ga}$) with $1,2\text{-(H}_2\text{P)}_2\text{C}_6\text{H}_4$ resulted in the formation of the novel 8-membered tetrametallic ring compounds with the general formula $\{(\text{Me}_2\text{M})_4[(\mu\text{-PH})_2(\text{C}_6\text{H}_4)]_2\}$ ($\text{M} = \text{Al}$ (1), Ga (2)) in a nearly quantitative yield. Combining AlMe_3 and $1,2\text{-(H}_2\text{P)}_2\text{C}_6\text{H}_4$ in a 1:1 ratio also afforded 1. The reaction of AlMe_3 with $1,2\text{-(H}_2\text{N)}_2\text{C}_6\text{H}_4$ (2:1) yielded the novel asymmetric compound $[(\text{Me}_2\text{Al})_2\text{AlMe}(\text{C}_6\text{H}_4(\text{NH})_2)_2] \cdot \text{AlMe}_3$ (3). Compound 3 is composed of four aluminum centers which are chemically inequivalent. Reaction of GaMe_3 with $1,2\text{-(H}_2\text{N)}_2\text{C}_6\text{H}_4$ in a 3:2 ratio gave $\{(\text{Me}_2\text{Ga})_3[(\mu\text{-NH})_2(\text{C}_6\text{H}_4)(\mu\text{-NH})(\text{C}_6\text{H}_4\text{NH}_2)]\}$ (4). The synthesis and characterization of 1 through 4, including their solid-state structures, are presented.				
14. SUBJECT TERMS aluminum-nitrogen, gallium-nitrogen, aluminum-phosphorus, gallium-phosphorus, trimetallic and tetrametallic compounds synthesis, structures		15. NUMBER OF PAGES 29		
		16. PRICE CODE		
17. SECURITY CLASSIFICATION OF REPORT Unclassified	18. SECURITY CLASSIFICATION OF THIS PAGE Unclassified	19. SECURITY CLASSIFICATION OF ABSTRACT Unclassified	20. LIMITATION OF ABSTRACT Unlimited	

19960603 017

OFFICE OF NAVAL RESEARCH

Grant N00014-95-1-0194
R&T Project 3135008---16

Dr. Harold E. Guard

Technical Report No. DU/DC/TR-56

**SYNTHESIS OF NOVEL TETRAMETALLIC AND TRIMETALLIC COMPOUNDS
FROM REACTIONS OF MMe_3 ($M = Al, Ga$) WITH $1,2-(H_2E)_2C_6H_4$ ($E = N, P$)**

RICHARD L. WELLS,¹ HAMID RAHBARNOOHI,² PAUL B. GLASER,³
LOUISE M. LIABLE-SANDS² and ARNOLD L. RHEINGOLD²

¹Department of Chemistry, Duke University, Durham, NC 27708-0354

²Department of Chemistry, University of Delaware, Newark, DE 19716

Accepted for Publication in *Organometallics*

Duke University
Department of Chemistry,
P. M. Gross Chemical Laboratory
Box 90346
Durham, NC 27708-0346

21 May 1996

Reproduction in whole or in part is permitted for any purpose of the United States Government.

This document has been approved for public release and sale; its distribution is unlimited.

2

SYNTHESIS OF NOVEL TETRAMETALLIC AND TRIMETALLIC COMPOUNDS
FROM REACTIONS OF MMe_3 ($M = Al, Ga$) WITH $1,2-(H_2E)_2C_6H_4$ ($E = N, P$)
RICHARD L. WELLS*, HAMID RAHBARNOOHI, and PAUL B. GLASER

Department of Chemistry, Paul M. Gross Chemical Laboratory, Duke University, Durham,
North Carolina 27708, U.S.A.

LOUISE M. LIABLE-SANDS and ARNOLD L. RHEINGOLD

Department of Chemistry, University of Delaware,
Newark, Delaware 19716, U.S.A.

Abstract- The independent 2:1 reactions of MMe_3 ($M = Al, Ga$) with $1,2-(H_2P)_2C_6H_4$ resulted in the formation of the novel 8-membered tetrametallic ring compounds with the general formula $\{(Me_2M)_4[(\mu-PH)_2(C_6H_4)]_2\}$ ($M = Al$ (1), Ga (2)) in a nearly quantitative yield. Combining $AlMe_3$ and $1,2-(H_2P)_2C_6H_4$ in a 1:1 ratio also afforded 1. The reaction of $AlMe_3$ with $1,2-(H_2N)_2C_6H_4$ (2:1) yielded the novel asymmetric compound $[(Me_2Al)_2AlMe(C_6H_4(NH)_2)_2] \cdot AlMe_3$ (3). Compound 3 is composed of four aluminum centers which are chemically inequivalent. Reaction of $GaMe_3$ with $1,2-(H_2N)_2C_6H_4$ in a 3:2 ratio gave $\{(Me_2Ga)_3[(\mu-NH)_2(C_6H_4)(\mu-NH)(C_6H_4NH_2)]\}$ (4). The synthesis and characterization of 1 through 4, including their solid-state structures, are presented.

*Author for correspondence

Introduction

The syntheses and characterization of organometallic compounds which contain both Group 13 and Group 15 elements have been a subject of interest for many years.¹⁻⁵ These compounds have the potential to serve as intermediates to ceramic, semi-conductor or electronic materials.⁶⁻¹⁶ Considering this, we were interested in exploring the feasibility of forming large-aggregate or polymeric 13-15 compounds by allowing bidentate ligands of Group 15 to react with Group 13 alkyl derivatives. To this end, we investigated the reactivity of 1,2-(H₂N)₂C₆H₄ and 1,2-(H₂P)₂C₆H₄ towards AlMe₃ and GaMe₃. Chelating ligands such as diethylenetriamine, diazabutadiene, and ethylenediamine have been used successfully to stabilize electrophilic Group 13 elements,^{12,17-23} and recent work on the chemistry of bidentate ligands of Group 15 has produced a number of interesting compounds. For example, Al(PhNCH₂CH₂NH₂)Me₂,²⁴ [Al(μ-PhNCH₂CH₂NH)Me]₃,²⁴ [^tBu₂PCH₂CH₂C₅H₄]InCl₂,²⁵ [^tBu₂PCH₂CH₂C₅H₄]GaMe₂,²⁵ [{HN(^tBu)CH(^tBu)CH₂N(^tBu)}AlH₂],²⁶ [(ⁱPr)₂ATI]AlH₂,²⁷ and [(ⁱPr)₂ATI]AlMe₂²⁷ (ATI = N-ⁱPr-2-(ⁱPr-amino)troponimine) have been prepared and characterized.

As early as 1957, Ulmschneider and Goubeau²⁸ reported the reaction of BMe₃ with 1,2-(H₂N)₂C₆H₄, and Norman and co-workers have also used this ligand in the reaction with RPCl₂ (R = Me, Ph) to produce [C₆H₄N₂(PR)₂]₂.²⁹ Also Isseleib and co-workers have investigated the reactivity of 1,2-(H₂P)₂C₆H₄ with main group and transition metal compounds.³⁰ To our knowledge, no structurally authenticated compounds of Group 13 elements with the 1,2-(H₂E)₂C₆H₄ (E = N, P) ligands have been reported. In this paper, we present the synthesis and solid-state structural determination of {(Me₂M)₄[(μ-PH)₂(C₆H₄)]₂} (M = Al (1), Ga (2)), [(Me₂Al)₂AlMe(C₆H₄(NH)₂)₂]•AlMe₃ (3), and {(Me₂Ga)₃[(μ-NH)₂(C₆H₄)(μ-NH)(C₆H₄NH₂)]} (4).

Experimental Section

General Considerations. All manipulations of air- and moisture-sensitive materials were performed in a Vacuum Atmospheres HE-493 Dri-Lab containing an argon atmosphere and by general Schlenk techniques. Toluene and pentane were distilled over Na/K alloy.

1,2-(H₂P)₂C₆H₄, AlMe₃, and GaMe₃ were purchased from Strem and were used as received. 1,2-(H₂N)₂C₆H₄ was purchased from Aldrich and was used as received. ¹H and ¹³C{¹H}NMR spectra were recorded on a QE-300 spectrometer operating at 300 and 75.4 MHz, respectively. ¹H and ¹³C{¹H} spectra were referenced to TMS by using the residual protons or carbons of deuterated benzene at δ 7.15 or 128 ppm, respectively; the upfield pentet of C₇D₈ at δ 2.09 ppm for ¹H NMR spectra and δ 20.4 ppm for ¹³C{¹H} spectra, the upfield broad peak of THF-d₈ at δ 1.73 ppm for ¹H NMR spectra and the upfield pentet at δ 25.3 ppm for ¹³C{¹H} spectra. All NMR samples were prepared in 5-mm tubes which were septum-sealed under argon. Melting points (uncorrected) were obtained with a Thomas-Hoover Uni-melt apparatus, and capillaries were flame-sealed under argon. Elemental Analyses were performed by E+R Microanalytical Laboratory, Inc., Corona, NY. Mass spectral data were collected on a JEOL JMS-SX 102A spectrometer operating in the electron ionization mode at 20 eV. IR Spectra were acquired for KBr pellets on a BOMEM Michelson MB-100 FT-IR spectrometer. X-ray crystallographic data were obtained at 25 °C on a Siemens P4 diffractometer utilizing graphite-monochromated Mo-Kα (λ = 0.71073 Å) radiation.

Preparation of {[Me₂Al]₄[(μ-PH)₂(C₆H₄)₂]} (1). Inside the Dri-Lab, a 250 mL Schlenk flask equipped with a magnetic stirbar, was charged with 1,2-(H₂P)₂C₆H₄ (0.20 g, 1.41 mmol) and ca. 5 mL of toluene. AlMe₃ (0.20 g, 2.82 mmol) dissolved in ca. 5 mL of toluene was added to the stirring solution of 1,2-(H₂P)₂C₆H₄. The reaction flask was then brought out of the Dri-Lab and left undisturbed at room temperature for 24 h. A slow evolution of gas from the solution was observed. X-ray quality colorless crystals of **1** were deposited at the bottom of the flask. Yield: 0.32 g, 90% based on Al. mp 165 °C (decomp.) Anal. Calcd. (found) for C₂₀H₃₆Al₄P₄: C, 47.26 (47.11); H, 7.14 (6.98). ¹H NMR (THF-d₈): δ 7.54(4H, m, C₆H₄); 7.21(4H, m, C₆H₄); 4.33, 3.64 [4H, m, P-H(¹J_{P-H} = 207 Hz)]; -0.96 (24H, s, Al(CH₃)₃). ¹³C{¹H} NMR (THF-d₈): δ 136.2, 135.4, 129.3 (s, C₆H₄); -9.1(br, Al-C). Mass spectral data (EI mode): peaks at *m/e*, 508, 493, 463, 436, 421, 390, 354, 239, 198, 183 corresponding to fragments of C₂₀H₃₆Al₄P₄ = M⁺, [M - Me]⁺, [M - 3Me]⁺, [M - AlMe₃]⁺, [M - AlMe₄]⁺, [M - Al₂Me₄]⁺,

$[\text{C}_6\text{H}_4(\text{PH})_2\text{Al}_2\text{Me}_4]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{Al}_2\text{Me}_3]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{AlMe}_2]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{AlMe}]^+$, respectively. IR (cm^{-1}): 2314 (m), P-H.

Note: Reaction of AlMe_3 with 1,2- $(\text{H}_2\text{P})_2\text{C}_6\text{H}_4$ in a 1:1 molar ratio resulted in the formation of **1** but in a lower yield.

Preparation of $\{[\text{Me}_2\text{Ga}]_4[(\mu\text{-PH})_2(\text{C}_6\text{H}_4)]_2\}$ (**2**).

[Compound **2** was synthesized using a procedure similar to that used for **1**]

Reactants: 1,2- $(\text{H}_2\text{P})_2\text{C}_6\text{H}_4$ (0.30 g, 2.11 mmol), GaMe_3 (0.48 g, 4.2 mmol).

Yield: 0.67 g, 94% based on Ga. mp (decomp.) 136 °C. Anal. Calcd. (found) for $\text{C}_{20}\text{H}_{36}\text{Ga}_4\text{P}_4$: C, 35.36 (35.29); H, 5.34 (5.32). ^1H NMR (C_6D_6): δ 7.54 (4H, m, C_6H_4), 6.90 (4H, m, C_6H_4); 4.18, 3.25 [4H, m, P-H ($^1J_{\text{P-H}} = 205$ Hz)]; -0.73 (24H, pentet, Ga- $(\text{CH}_3)_3$, $^3J_{\text{PGaCH}} = 3.3$ Hz). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 136.2, 128.8, 127.8 (s, C_6H_4); -5.8 (br, Ga-C). Mass spectral data (EI mode): peaks at m/e , 678, 663, 565, 550, 339, 324, 240, 225 corresponding to fragments of $\text{C}_{20}\text{H}_{36}\text{Ga}_4\text{P}_4 = \text{M}^+$, $[\text{M} - \text{Me}]^+$, $[\text{M} - \text{GaMe}_3]^+$, $[\text{M} - \text{GaMe}_4]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{Ga}_2\text{Me}_4]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{Ga}_2\text{Me}_3]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{GaMe}_2]^+$, $[\text{C}_6\text{H}_4(\text{PH})_2\text{GaMe}]^+$, respectively. IR (cm^{-1}): 2314 (m), P-H.

Preparation of $[(\text{Me}_2\text{Al})_2\text{AlMe}(\text{C}_6\text{H}_4(\text{NH})_2)_2] \cdot \text{AlMe}_3$ (3**).** Inside the Dri-Lab, a 250 mL Schlenk flask equipped with a magnetic stirbar was charged with 1,2- $(\text{H}_2\text{N})_2\text{C}_6\text{H}_4$ (0.50 g, 4.62 mmol) and ca. 10 mL of toluene. To the resulting suspension was added AlMe_3 (0.67 g, 9.24 mmol) in ca. 10 mL of toluene. Stirring the reaction mixture a few minutes resulted in a clear-pinkish solution with evolution of gas. The reaction flask was brought out of the Dri-Lab, slightly warmed with a heat gun, and covered completely with aluminum foil to allow gradual cooling to room temperature. X-ray quality crystals of **3** were deposited at the bottom of the flask overnight. Yield: 0.86 g, 85% based on Al. mp (decomp.) 120 °C. Anal. Calcd. (found) for $\text{C}_{20}\text{H}_{36}\text{Al}_4\text{N}_4$: C, 54.54 (54.26); H, 8.24 (8.02); N, 12.72 (12.50). ^1H NMR (C_6D_6): δ 6.11-7.54 (8H, m, C_6H_4); 2.87, 2.73, 2.10, 1.93 (4H, s, N-H); -0.28 (9H, s, Al- Me_3), -0.33, -0.40 (6H s, AlMe_2), -0.53, -0.96 (6H, s, AlMe_2), -1.96 (3H, s, AlMe). $^{13}\text{C}\{^1\text{H}\}$ NMR (C_6D_6): δ 136.5, 134.3, 126.4, 126.1, 124.3, 121.1 (s, C_6H_4); -6.6, -9.2 (br, Al-C). Mass spectral data (EI

mode): peaks at m/e , 368, 352, 337, 296, 281, 205, 165, 148 corresponding to fragments $[M - AlMe_3]^+$, $[M - AlMe_4]^+$, $[M - AlMe_5]^+$, $[M - Al_2Me_6]^+$, $[M - Al_2Me_7]^+$, $[C_6H_4(NH)_2Al_2Me_3]^+$, $[C_6H_4(NH)_2AlMe_2]^+$, $[C_6H_4(NH)_2AlMe]^+$, respectively. IR (cm^{-1}): 3245(s), N-H.

Preparation of $\{(Me_2Ga)_3[(\mu-NH)_2(C_6H_4)(\mu-NH)(C_6H_4NH_2)]\}$ (4).

[Compound 4 was synthesized using a procedure similar to that used for 3]

Reactants: 1,2-(H_2N) $_2C_6H_4$ (0.37 g, 3.41 mmol), $GaMe_3$ (0.58 g, 5.04 mmol).

Yield: 0.77 g, 90% based on Ga. mp. 125 °C. X-ray quality crystals were grown by a method similar to that given for 3. Anal. Calcd. (found) for $C_{18}H_{31}Ga_3N_4$: C, 42.17(41.97); H, 6.10 (6.16); N, 10.93 (10.73). 1H NMR (C_6D_6): δ 6.40-6.74 (8H, m, C_6H_4); 2.68 (5H, br, N-H, NH_2); -0.38 (18H, m, Ga-Me $_3$). $^{13}C\{^1H\}$ NMR (C_6D_6): δ 135.3, 120.2, 116.7 (s, C_6H_4); -7.2(br, Ga-C). Mass spectral data (EI mode): peaks at m/e , 496, 382, 291, 206, 190 corresponding to fragments $[M - Me]^+$, $[M - GaMe_4]^+$, $[C_6H_4(NH)_2Ga_2Me_3]^+$, $[C_6H_4(NH)_2GaMe_2]^+$, $[C_6H_4(NH)_2GaMe]^+$, respectively. IR (cm^{-1}): 3285, 3427, and 3358 cm^{-1} , N-H stretching mode; 1624 and 1592 cm^{-1} , N-H bending mode.

X-ray structural solution and refinement

Crystal, data collection, and refinement parameters are given in Table 1. Suitable crystals of 1-4 were mounted in thin-walled capillaries and temporarily sealed with silicone grease under an argon atmosphere and then flame-sealed.

The photographic data, unit cell parameters, occurrences of equivalent reflections and systematic absences in the diffraction data are uniquely consistent with the $P2_1/n$ space group for 1 and 2, $P2_1/c$ for 3, and C_2 , Cm or $C2/m$ for 4. The E-statistic suggested the centrosymmetric option for 4 and the space group choices were confirmed by chemically reasonable results of refinement. The unit-cell parameters were obtained by least-square refinement of the angular setting of 24 reflections ($20^\circ \leq 2\theta \leq 24^\circ$).

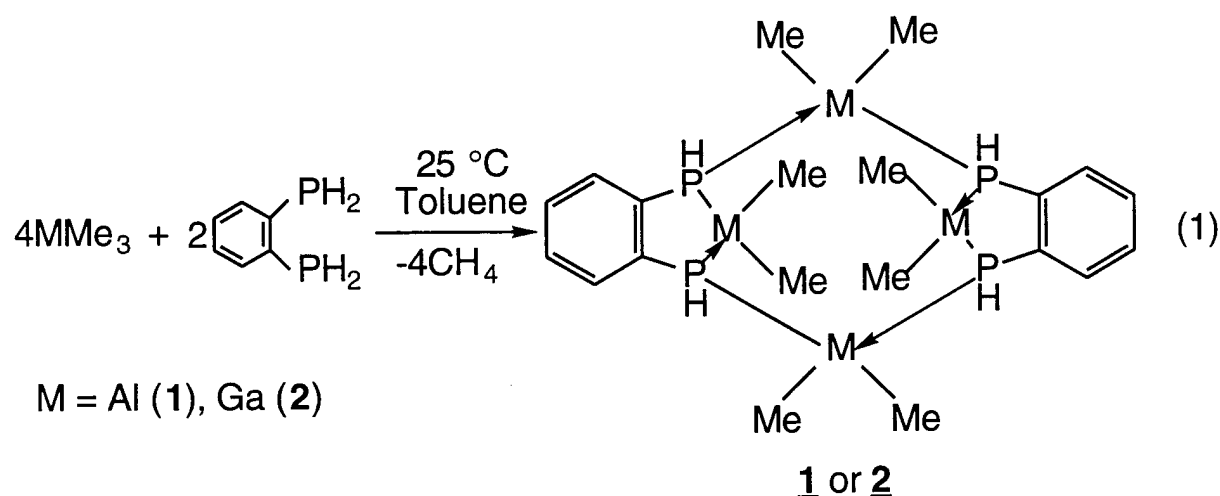
Structures of 1, 3 and 4 were solved using direct methods, completed by difference Fourier syntheses and refined by full-matrix least square procedures. The atomic coordinates of 1

were used as a trial solution for the isomorphous structure **2**. Semi-empirical absorption corrections were applied for **2**, but were not required for **1**, **3**, and **4** because there was less than 10% variation observed in the ψ -scan data. All non-hydrogen atoms were refined with anisotropic displacement coefficients. Structures **1** and **2** have an inversion center. Structure **4** is located on a mirror plane. Hydrogen atoms, bonded to phosphorus atoms in **1** and **2**, were located from the difference map, constrained with an average refined P-H distance, and treated with idealized isotropic parameters. The hydrogen atoms on N (1) and N (2) in **4** are disordered over a mirror plane and were ignored. All other hydrogen atoms were treated as idealized contributions. The largest remaining peak in the difference map of **4** ($1.35 \text{ e}\text{\AA}^{-3}$) occurs at a chemically unreasonable position and was considered as noise.

All software and sources of the scattering factors are contained in the SHELXTL (5.3) program libraries.³¹

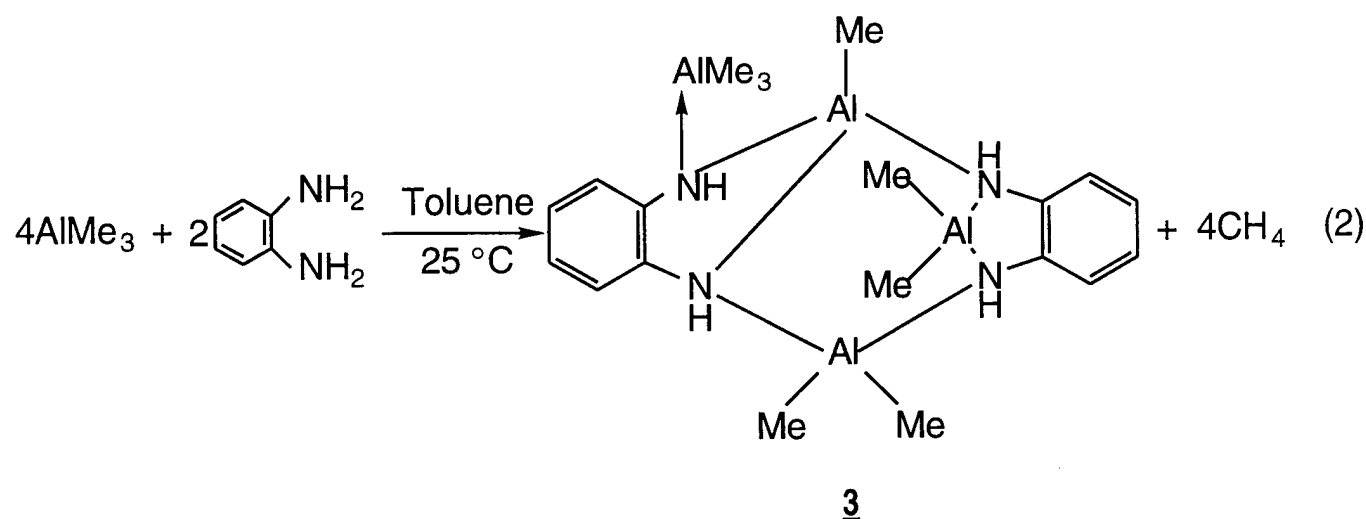
Results and Discussion

Syntheses. Independent reactions of MMe_3 ($\text{M} = \text{Al}, \text{Ga}$) with $1,2\text{-(H}_2\text{P)}_2\text{C}_6\text{H}_4$ in a 2:1 ratio resulted in the formation of $\{(\text{Me}_2\text{M})_4[(\mu\text{-P(H)})_2(\text{C}_6\text{H}_4)]_2\}$ ($\text{M} = \text{Al}$ (**1**), Ga (**2**)) in nearly quantitative yields (eq. 1). Slow evolution of gas was apparent after mixing the reagents in toluene at room temperature. The possible reaction pathway could include the formation of the adduct followed by the elimination of methane, to give **1** or **2**; however, under the applied conditions, no intermediate adducts were isolated from these reactions. In contrast to this, in a related study, the reaction of AlMe_3 with $\text{PH}_2\text{CH}_2\text{CH}_2\text{PH}_2$ in a 2:1 ratio and under similar mild conditions afforded the bis-adduct, $\text{Me}_3\text{Al}\cdot\text{P(H)}_2\text{CH}_2\text{CH}_2(\text{H}_2)\text{P}\cdot\text{AlMe}_3$ which will be discussed elsewhere.³²

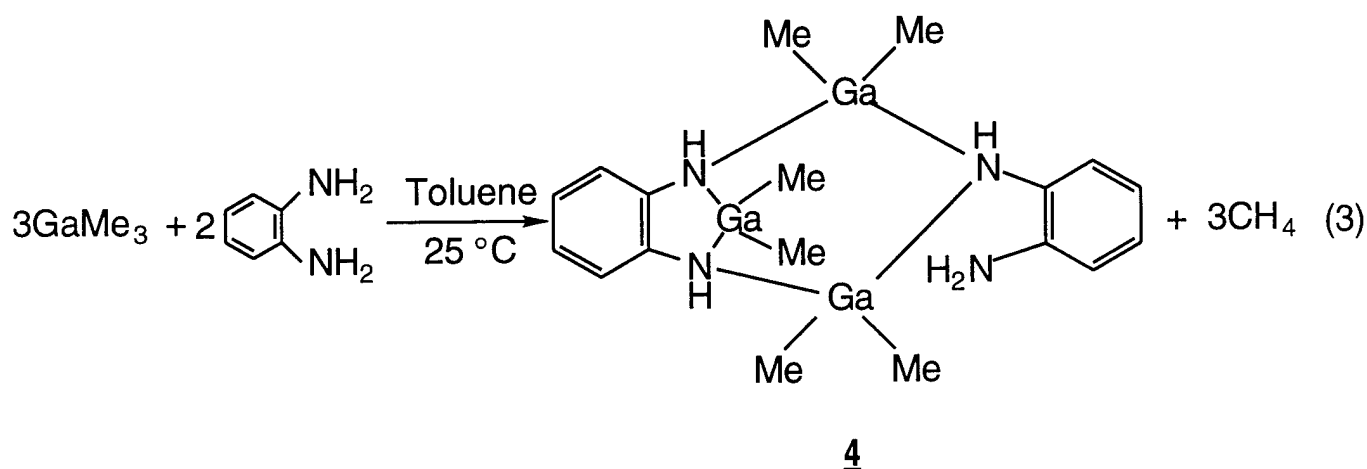


It is interesting to note that compound 1 was isolated even when a 1:1 ratio of AlMe_3 and $1,2\text{-(H}_2\text{P)}_2\text{C}_6\text{H}_4$ was used. Compounds 1 and 2 crystallized from toluene but show very little solubility in toluene, once crystallized. Compound 1 was soluble in the donor solvent THF, and the ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR spectra were obtained in THF-d_8 solutions to account for the ratio of AlMe_3 protons vs. the aromatic protons. Compound 2 was insoluble in THF, thus the ^1H and $^{13}\text{C}\{^1\text{H}\}$ NMR spectra were obtained from a very dilute C_6D_6 solution.

Compounds 3 and 4 were prepared in a manner similar to 1 and 2, but showed better solubility in toluene. Unlike 1 and 2, compound 3 is an asymmetric molecule (*vide infra*) (eq 2).



Compound 4 was prepared according to eq. 3. Initially the stoichiometry of 4:2 was used



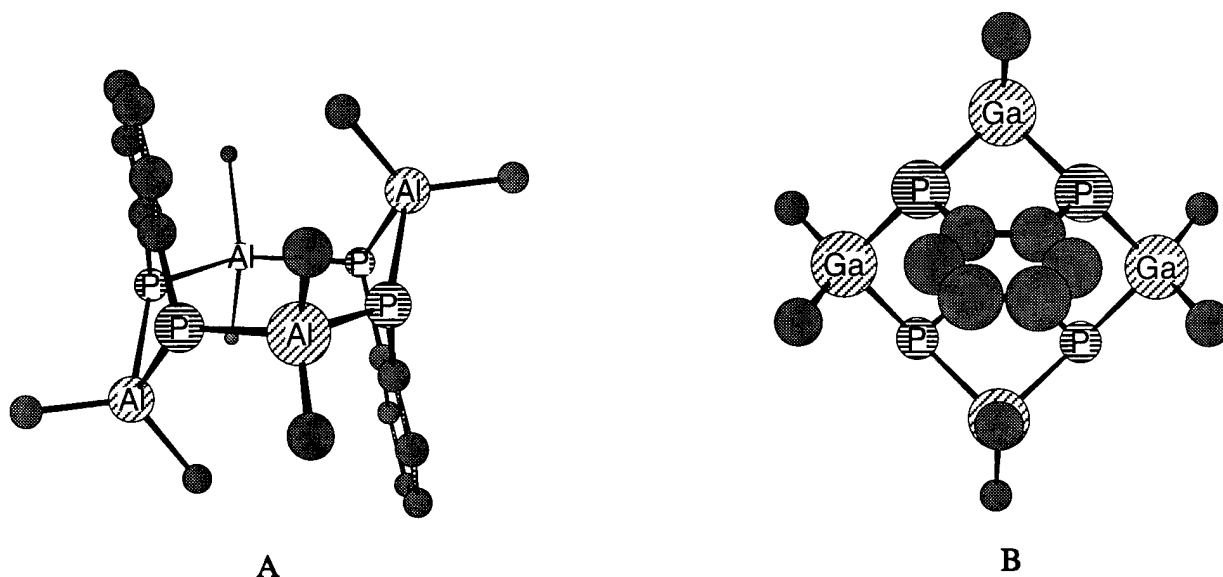
and a white, toluene insoluble powder was isolated. The partial elemental analysis (C, H, N) of this powder is in accordance with the chemical formula $\{[\text{Me}_2\text{Ga}]_4[(\mu\text{-NH})_2(\text{C}_6\text{H}_4)]_2\}$, analogous to **1** and **2**. However, reaction in a 3:2 molar ratio (eq. 3) resulted in quantitative formation of **4**, which is of different aggregation than that for **1-3** (*vide infra*). Compounds **1-4** are colorless crystals which decompose slowly when exposed to air and moisture.

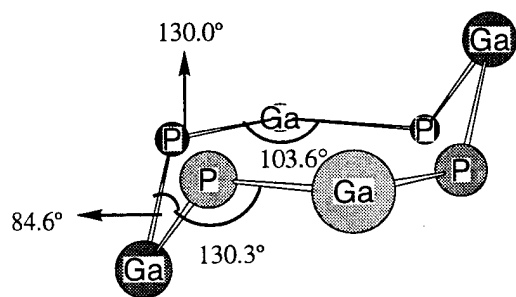
Structure and Bonding. Thermal ellipsoid diagrams of **1-4** are shown in Figures 1-4. Crystal data and structure refinement for **1-4** are given in Table 1. Selected interatomic bond distances and bond angles for **1-4** are presented in Tables 2-5, respectively. Compounds **1** and **2** belong to the monoclinic system $\text{P}2_1/\text{n}$ with two molecules per unit cell and an inversion center for each molecule. The geometry of the $(\text{MP})_4$ core is that of an eight-membered ring in a chair conformation with approximately equal M-P bond lengths. The Al-P bond distances in **1** range from 2.4310(14)-2.447(2) Å, avg. 2.44 Å, which can be compared to six-membered ring compounds such as $[(\text{Me}_3\text{Si})_2\text{AlP}(\text{c-C}_6\text{H}_{11})(\text{H})]_3$ (avg. 2.44 Å),³³ $[\text{Me}_2\text{AlPMe}_2]_3$ (2.43 Å),³⁴ and $[\text{Me}_2\text{AlCH}_2\text{PMe}_2]_2$ (2.45 Å).³⁵ The chair conformation of the eight-membered ring in **1** is shown in I. The internal angles of Al-P-Al are almost identical (130.07(5)° and 130.79(5)°). These values are very close to the Al-P-Al angles reported for $[\text{Me}_2\text{AlPMe}_2]_3$ (131.7(8)°), which was studied in the gas phase.³⁴ The P(2A)-Al(2)-P(1) and P(2)-Al(2A)-P(1A) angles in **1** are smaller than the corresponding angles in $[\text{Me}_2\text{AlPMe}_2]_3$ (83.95(5)° vs. 96.4(7)°) and in $[(\text{Me}_3\text{Si})_2\text{AlP}(\text{c-C}_6\text{H}_{11})(\text{H})]_3$ (83.95(5)° vs. 99.6(1)°). This could be due to the fact that the five-

membered ring fragment (which forms part of the eight-membered ring) $\overbrace{\text{Al-P-C-C-P}}$ imposes a ring strain, since the P...P separation is held rigid by the aromatic ring. The Al-C bond distances in **1** range from 1.944(3) Å to 1.959(4) Å, and are similar to reported values in the literature.⁴

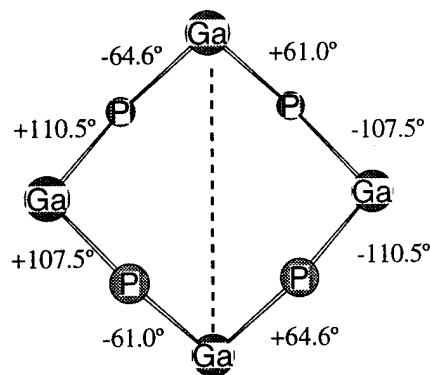
Compound **2** is the gallium analog of **1**, and is isomorphous with **1**. Due to the close similarity in the atomic radii of aluminum and gallium atoms, the angles and bond lengths in **1** and **2** were identical and within experimental errors. The Ga-P bond lengths in **2** range from 2.419(3) Å to 2.433(3) Å with an average of 2.424 (3) Å. This is in order with the trimeric (GaP)₃ compounds [Me₂GaPPh₂]₃ (2.433(1) Å),³⁶ [^tBuGaP(μ-PH₂)]₃ (2.439 (3) Å),³⁷ [Me₂Ga(μ-P-ⁱPr₂)₃ (avg. 2.44 Å),³⁸ [Me₂GaP(Me)Ph]₃ (avg. 2.41 Å),³⁹ and [(Me₃CCH₂)ClGaPPh₂]₃ (avg. 2.42 Å).⁴⁰ The Ga-C bond distances in **2** range from 1.952(10) Å to 1.983(10) Å and are similar to reported values in the literature.³⁶⁻⁴⁰ The torsion and dihedral angles of compounds **1** or **2** are presented in I.

I. (A) Chair conformation of **1**. (B) View of **2** through the phenyl ring. (C) Internal angles in **2**. (D) Torsion angles of **2** (The torsion angle a-b-c-d is defined as positive if, when viewed along the b-c bond, atom a must be rotated clockwise to eclipse atom d).





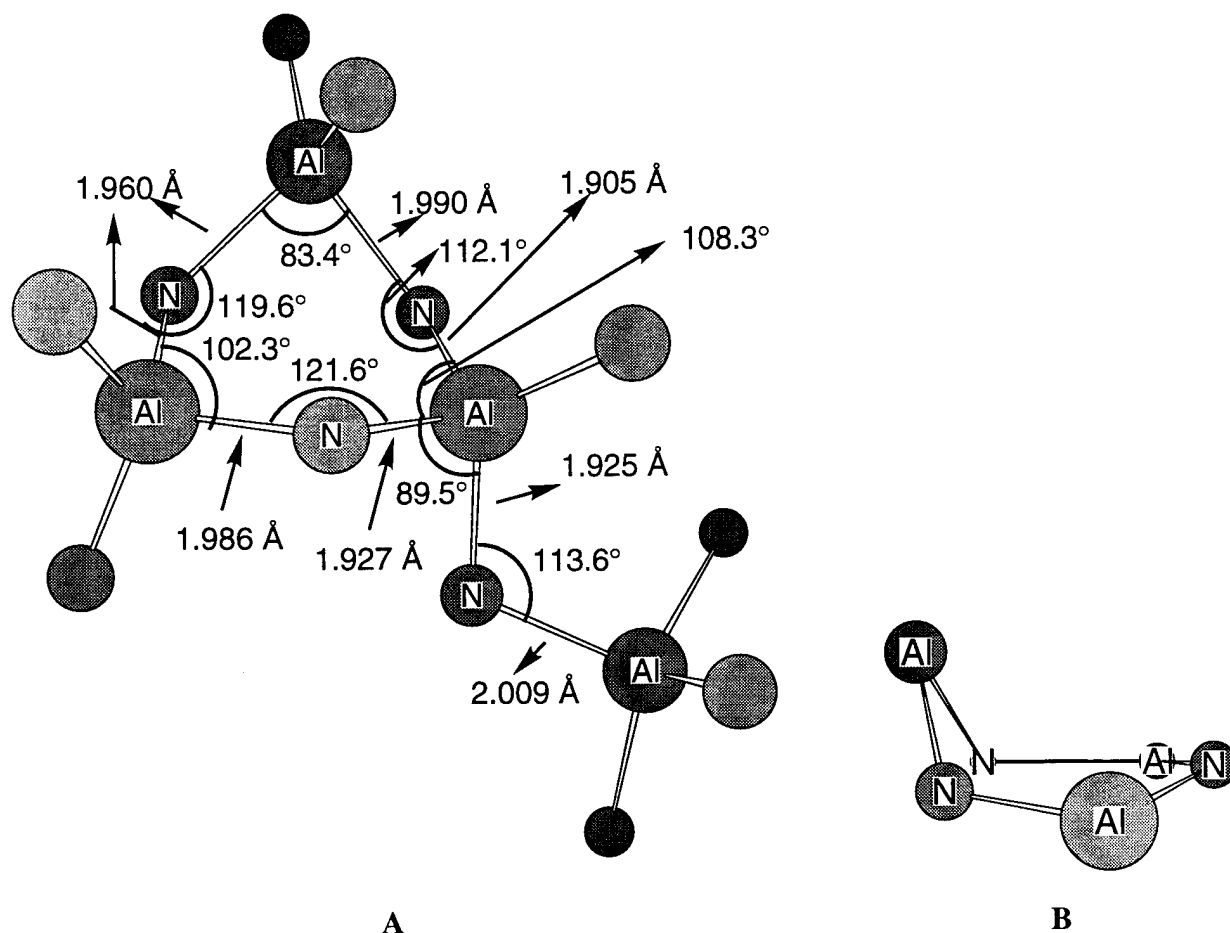
C



D

Compound **3** crystallizes in the monoclinic system $P2_1/c$ with four molecules in the unit cell. The asymmetric compound **3** has four Al centers which are chemically inequivalent and all the Al-N bonds are crystallographically independent. To our knowledge, this is the first example of a 13-15 compound in which there are three different M linkages of the type RM, R_2M , and R_3M existing within the same molecule. The Al centers in **3** are all four-coordinate with distorted tetrahedral geometry. Both Al(3) and Al(4) have two methyl groups while Al(1) possesses only one methyl group. Al(2) has three methyl groups and is bonded in a dative fashion to N(1) with a distance of 2.009(1) Å. This distance is similar to the dative bonds in $[AlEt_2]_2[C_{10}H_{22}N_4][AlEt_3]_2$ (2.073 (4) and 2.025 (4) Å)⁴¹ and $[AlMe_2]_2[C_{10}H_{20}N_4][AlMe_3]_2$ (2.042 (4) Å).⁴² The shortest Al-N bond in **3** is 1.905(2) Å for Al(1)-N(4). This bond length could be considered a normal covalent bond when compared to the corresponding Al-N bond distances in $[AlEt_2]_2[C_{10}H_{22}N_4][AlEt_3]_2$ (1.906 (4) Å),⁴¹ $[AlMe_2]_2[C_{10}H_{20}N_4][AlMe_3]_2$ (1.886 (4) Å),⁴² and $AlCl_4(\mu-NMe_2)_2(\mu-NMe_2)_4$ (1.76 Å and 1.92 Å).⁴³ The Al(1)-N(1) and Al(1)-N(2) bond lengths are 1.925 (2) and 1.927 (2) Å, respectively. These distances are also within normal bonding with very little or no dative character. The remaining Al-N bond lengths in **3**, Al(3)-N(3) (1.960 (2) Å) and Al(4)-N(3) (1.960 (2) Å), are slightly longer than the Al-N bond lengths discussed above. On the other hand, Al(3)-N(4) (1.990 (2) Å) and Al(4)-N(2) (1.986 (2) Å) are even longer and may have some dative character. The overall geometry around the nitrogen atoms is pyramidal. The core of **3** is composed of a six-membered (AlN)₃ ring which is attached to a five-membered AlN₂C₂ ring through Al(1)-N(2). The six-membered (AlN)₃ ring has a distorted boat conformation (**IIB**). The internal angles and Al-N bond distances are also given in **IIA**. The

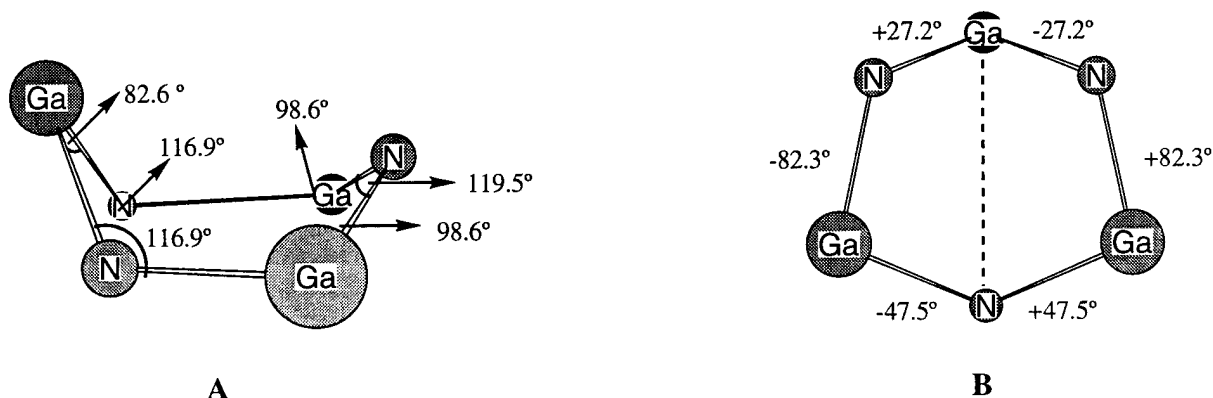
Al-C bond distances in **1** range from 1.931(3) Å to 1.983(3) Å and are similar to reported values in the literature.⁴



II. (A) Internal angles and Al-N distances given for **3**. C₆H₄ rings and the hydrogen atoms have been omitted for clarity. (B) Trimeric distorted Boat conformation of the core in **3**.

Compound **4** crystallizes in the monoclinic system C2/m. There are four molecules in the unit cell, where each molecule is located on a mirror plane. The Ga-N bond distances range from 2.027 (10) Å to 2.048 (9) Å with an average value of 2.04 Å. The Ga-N bonds are slightly shorter than the Ga-N bond distances in [o-(Me₂NCH₂)C₆H₄]₂GaCl₂ (2.071 (2) Å) and [o-(Me₂NCH₂)C₆H₄]₂GaH₂ (2.087 (7) Å).⁴⁴ To our knowledge, there are only two trimeric gallium-nitrogen systems which have been structurally characterized in the solid state. One such is the trimer [(Me₂GaNH₂)₃], which possesses a twisted chair conformation for the (GaN)₃ core and an

average Ga-N bond length of 1.98 Å.⁴⁵ In another trimer, $[(\text{CH}_2)_2\text{NGaH}_2]_3$, the $(\text{GaN})_3$ core has a chair conformation with an average Ga-N bond length of 1.97 Å.⁴⁶ These values are slightly shorter than the Ga-N bond length in **4**. The six-membered core $(\text{GaN})_3$ has a boat geometry which is shown in **IIIA**. Torsion angles in **4** are presented in **IIIB**.



III. (A) Trimeric Boat conformation of the core in 4. (B) Torsion angles in 4
(The torsion angle a-b-c-d is defined as positive if, when viewed along the b-c bond, atom a must be rotated clockwise to eclipse atom d).

Acknowledgment. We are grateful for the financial support of this work provided by the Office of Naval Research. P. B. G. wishes to thank the National Science Foundation REU summer program for support. We also thank Dr. George R. Dubay at Duke University for performing the electron ionization mass spectroscopy of our compounds.

Supplementary Materials Available: Complete listings of crystal and X-ray data collection parameters, bond distances and angles, anisotropic thermal parameters for the non hydrogen atoms, and atomic coordinates and isotropic thermal parameters for the hydrogen atoms (57 pages). Ordering information is given on any current masthead page.

REFERENCES

- (1) Coates, G. E.; Green, M. L. H.; Wade, K. *Organometallic Compounds, Volume 1: The Main Group Elements*; Butler & Tanner Ltd.: London, 1967.
- (2) Mole, T.; Jeffrey, E. A. *Organoaluminum Compounds*; Elsevier: Amsterdam, 1972.
- (3) *Comprehensive Organometallic Chemistry*; Wilkinson, G.; Stone, F. G. A.; Abel, E. W., Eds.; Pergamon: Oxford, 1982; Vol. 1.
- (4) Robinson, G. H. *Coordination Chemistry of Aluminum*; VCH: New York, 1993.
- (5) Downs, A. J. *Chemistry of Aluminium, Gallium, Indium and Thallium*; Chapman & Hall: London, 1993.
- (6) Cowley, A. H.; Jones, R. A. *Angew. Chem., Int. Ed. Engl.* **1989**, *28*, 1208.
- (7) Interrante, L. V.; Sauls, F. C. *Coord. Chem. Rev.* **1993**, *128*, 193.
- (8) Wells, R. L. *Coord. Chem. Rev.* **1992**, *112*, 273-291.
- (9) Bent, B. E.; Nuzzo, R. G.; Dubois, L. H. *J. Am. Chem. Soc.* **1989**, *111*, 1634.
- (10) Cowley, A. H.; Benac, B. L.; Ekerdt, J. G.; Jones, R. A.; Kidd, K. B.; Lee, J. Y.; Miller, J. E. *J. Am. Chem. Soc.* **1988**, *110*, 6248.
- (11) Buhro, W. E. *Polyhedron* **1994**, *13*, 1131-1148.
- (12) Jiang, Z.; Interrante, L. V.; Kwon, D.; Tham, F. S.; Kulling, R. *Inorg. Chem.* **1991**, *30*, 995.
- (13) Paine, R. T.; Narula, C. N. *Chem. Rev.* **1990**, *90*, 73.
- (14) Paine, R. T.; Noth, H. *Chem. Rev.* **1995**, *95*, 343-379.
- (15) Aubuchon, S. R.; McPhail, A. T.; Wells, R. L.; Giambra, J. A.; Bowser, J. R. *Chem. Mater.* **1994**, *6*, 82-86.
- (16) Wells, R. L.; Aubuchon, S. R.; Kher, S. S.; Lube, M. S.; White, P. S. *Chem. Mater.* **1995**, *7*, 793.
- (17) Goebel, D. W., Jr.; Hencher, J. L.; Oliver, J. P. *Organometallics* **1983**, *2*, 746.
- (18) Wissing, E.; Jastrzebski, J. T. B. H.; Boersma, J.; Koten, G. V. *J. Organomet. Chem.* **1993**, *459*, 11.
- (19) Muller, J.; Englert, U. *Chem. Ber* **1995**, *128*, 493.
- (20) Robinson, G. H.; Sangokoya, S. A. *J. Am. Chem. Soc.* **1987**, *109*, 6862.

- (21) Zaworotko, M. J.; Atwood, J. L. *Inorg. Chem.* **1980**, *19*, 268.
- (22) Jiang, Z.; Interrante, L. V.; Kwon, D.; Tham, F. S.; Kullnig, R. *Inorg. Chem.* **1992**, *31*, 4815.
- (23) Jones, C.; Koutsantonis, G. A.; Raston, C. L. *Polyhedron* **1993**, *12*, 1829.
- (24) Wang, S.; Trepanier, S. J. *J. Chem. Soc., Dalton Trans.* **1995**, 2425.
- (25) Cowley, A. H.; King, C. S.; Decken, A. *Organometallics* **1995**, *14*, 20-23.
- (26) Raston, C. L.; Lawrence, S. M.; Gardiner, M. G. *Inorg. Chem.* **1995**, *34*, 4652.
- (27) Dias, H. V. R.; Jin, W.; Ratcliff, E. R. *Inorg. Chem.* **1995**, *34*, 6100.
- (28) Ulmschneider, D.; Goubeau, J. *Chem. Ber.* **1957**, *90*, 2733.
- (29) Barendt, J. M.; Bent, E. G.; Young, S. M.; Haltiwanger, R. C.; Norman, A. D. *Inorg. Chem.* **1991**, *30*, 325.
- (30) Issleib, K.; Leissring, E.; Schmidt, H. *Phosphorus and Sulfur* **1983**, *18*, 15.
- (31) *SHELXTL PC*; Siemens Analytical X-Ray Instruments, Inc.: Madison, WI, 1990.
- (32) Rahbarnoochi, H.; Wells, R. L.; Heeg, M. J. Unpublished results.
- (33) Janik, J. F.; Duesler, E. N.; McNamara, W. F.; Westerhausen, M.; Paine, R. T. *Organometallics* **1989**, *8*, 506.
- (34) Haaland, A.; Hougen, J.; Volden, H. V.; Hanika, G.; Karasch, H. H. *J. Organomet. Chem.* **1987**, *322*, C24.
- (35) Karsch, H. H.; Appelt, A.; Kohler, F. H.; Müller, G. *Organometallics* **1985**, *4*, 231.
- (36) Robinson, G. H.; Burns, J. A.; Pennington, W. T. *Main Group Chem.* **1995**, *1*, 153.
- (37) Cowley, A. H.; Harris, P. R.; Jones, R. A.; Nunn, C. M. *Organometallics* **1991**, *10*, 652.
- (38) Cowley, A. H.; Jones, R. A.; Mardones, M. A.; Nunn, C. M. *Organometallics* **1991**, *10*, 1635.
- (39) Beachley, O. T., Jr.; Royster, T. L., Jr.; Arhar, J. R.; Rheingold, A. L. *Organometallics* **1993**, *12*, 1976.
- (40) Beachley, O. T., Jr.; Maloney, J. D.; Rogers, R. D. *J. Organomet. Chem.* **1993**, *449*, 69.

- (41) Moise, F.; Pennington, W. T.; Robinson, G. H.; Sangokoya, S. A. *Acta. Cryst.* **1990**, *C46*, 1110.
- (42) Robinson, G. H.; Rae, A. D.; Campana, C. F.; Byram, S. K. *Organometallics* **1987**, *6*, 1227.
- (43) Thewalt, U.; Kawada, I. *Chem. Ber.* **1970**, *103*, 2754.
- (44) Isom, H. S.; Cowley, A. H.; Decken, A.; Sissingh, F.; Corbelin, S.; Lagow, R. J. *Organometallics* **1995**, *14*, 2400.
- (45) Almond, M. J.; Drew, M. G. B.; Jenkins, C. E.; Rice, D. A. *J. Chem. Soc., Dalton Trans.* **1992**, 5.
- (46) Harrison, W.; Storr, A.; Trotter, J. *J. Am. Chem. Soc., Dalton Trans.* **1972**, 1554.

Captions to Figures

Figure 1. Molecular diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of $\{(\text{Me}_2\text{Al})_4[(\mu\text{-PH})_2(\text{C}_6\text{H}_4)]_2\}$ (1). Hydrogen atoms on the CH_3 and C_6H_4 groups are omitted for clarity.

Figure 2. Molecular diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of $\{(\text{Me}_2\text{M})_4[(\mu\text{-PH})_2(\text{C}_6\text{H}_4)]_2\}$ (2).

Figure 3. Molecular diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of $[(\text{Me}_2\text{Al})_2\text{AlMe}(\text{C}_6\text{H}_4(\text{NH})_2)_2] \cdot \text{AlMe}_3$ (3). Hydrogen atoms are omitted for clarity.

Figure 4. Molecular diagram (30% probability ellipsoids) showing the solid-state structure and atom numbering scheme of $\{(\text{Me}_2\text{Ga})_3[(\mu\text{-NH})_2(\text{C}_6\text{H}_4)(\mu\text{-NH})(\text{C}_6\text{H}_4\text{NH}_2)]\}$ (4). Hydrogen atoms are omitted for clarity.

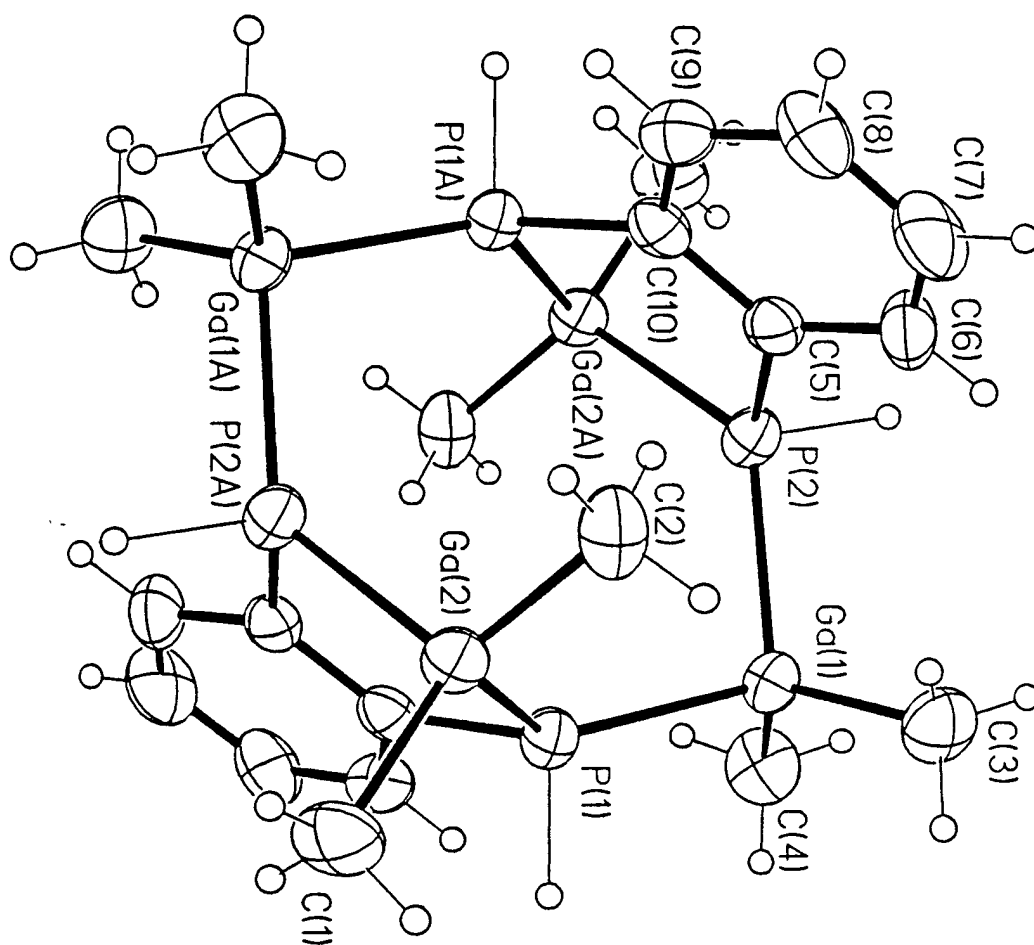


Figure 2

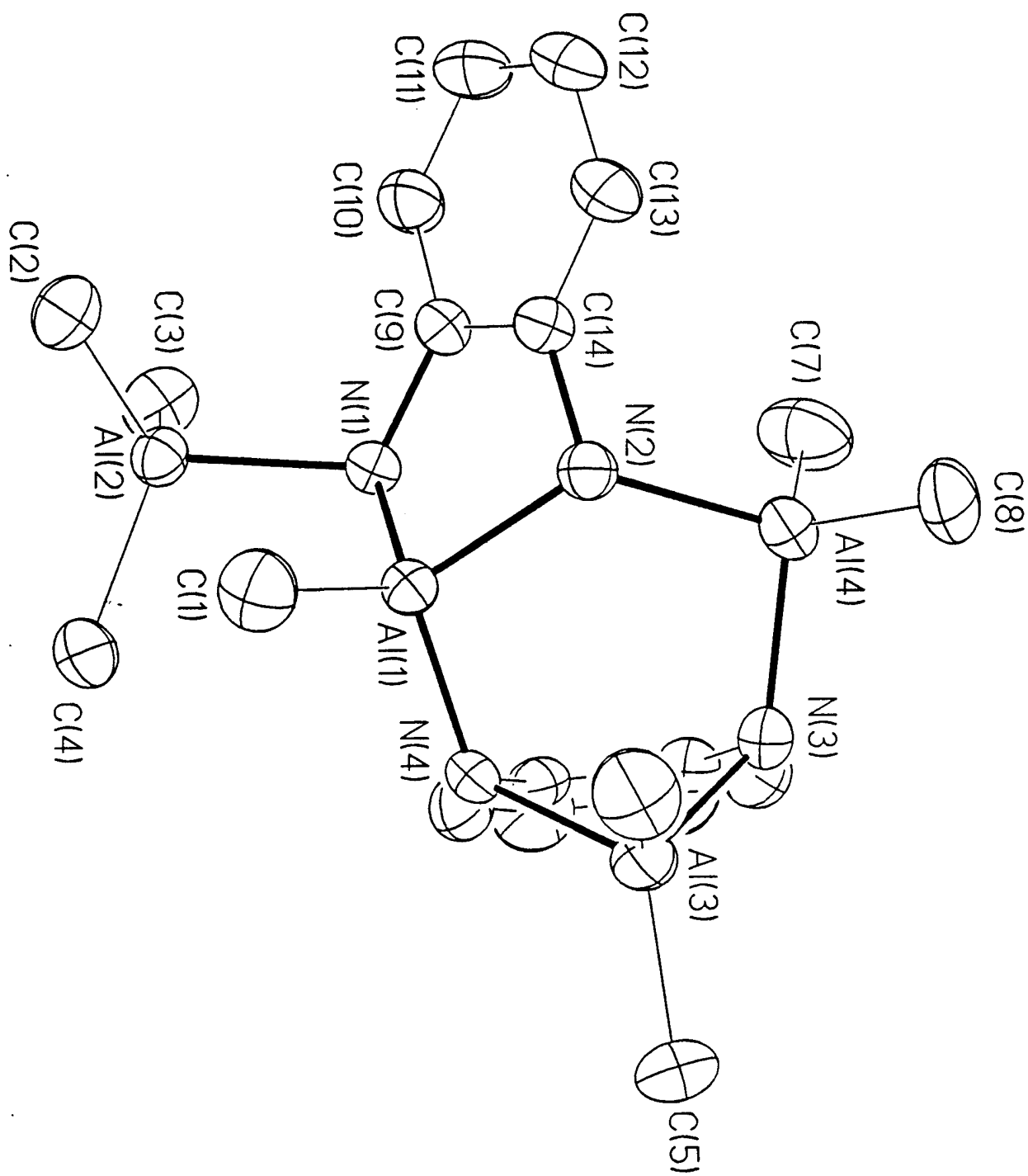


Figure 3

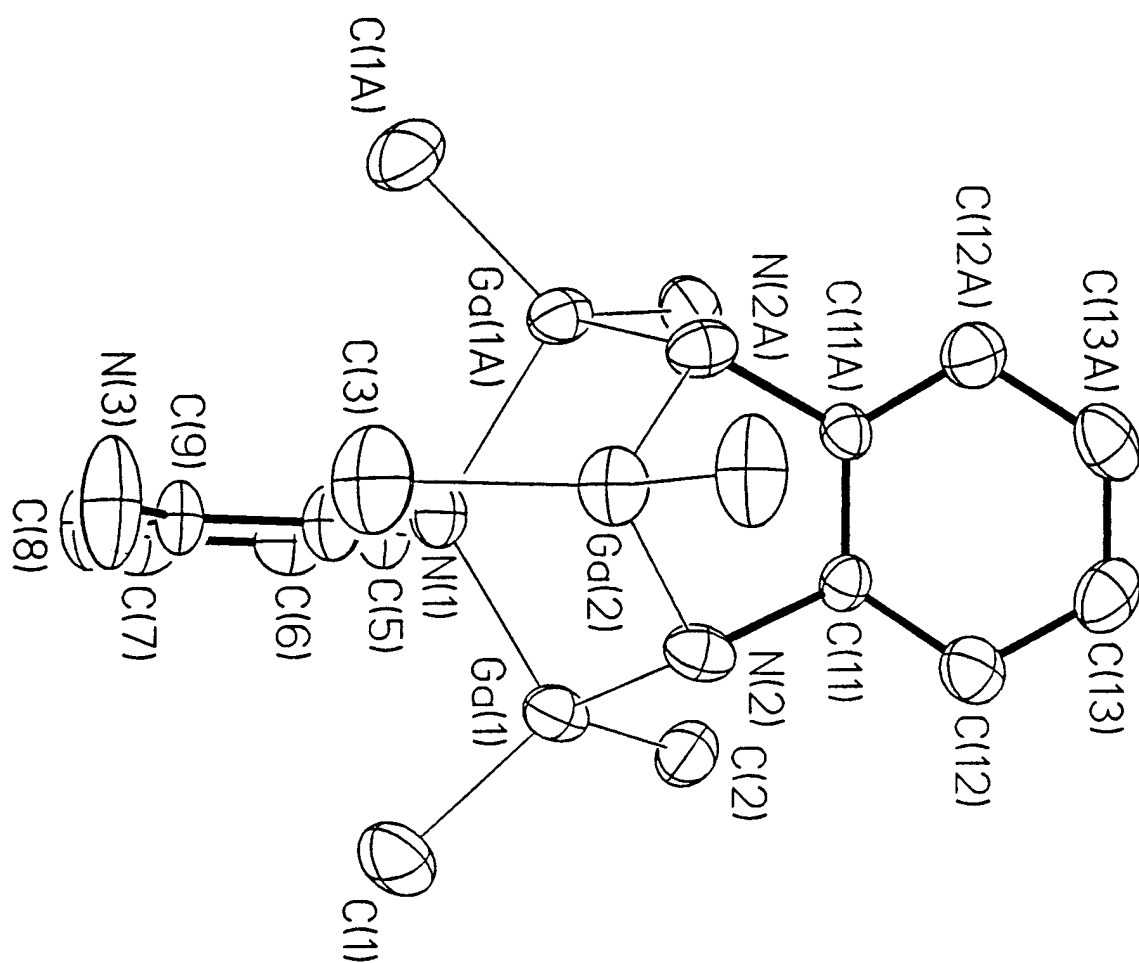


Figure 4

Table 1. Crystal Data and Structure Refinement for $\{(\text{Me}_2\text{M})_4[(\mu\text{-PH})_2(\text{C}_6\text{H}_4)]_2\}$ ($\text{M} = \text{Al}$ (1), Ga (2)), $[(\text{Me}_2\text{Al})_2\text{AlMe}(\text{C}_6\text{H}_4(\text{NH}_2)_2)_2] \cdot \text{AlMe}_3$ (3), $\{(\text{Me}_2\text{Ga})_3[(\mu\text{-NH})_2(\text{C}_6\text{H}_4)(\mu\text{-NH})(\text{C}_6\text{H}_4\text{NH}_2)]\}$ (4).

	1	2	3
Empirical formula	$\text{C}_{20}\text{H}_{36}\text{Al}_4\text{P}_4$	$\text{C}_{20}\text{H}_{36}\text{Ga}_4\text{P}_4$	$\text{C}_{20}\text{H}_{36}\text{Al}_4\text{N}_4$
formula weight	508.29	679.25	440.45
temp, K	298 (2)	298 (2)	241 (2)
radiation	Mo-K α (0.71073)	Mo-K α (0.71073)	Mo-K α (0.71073)
(wavelength, Å)			
crystal system	monoclinic	monoclinic	monoclinic
space group	P2 ₁ /n	P2 ₁ /n	P2 ₁ /c
a, Å	10.479(2)	10.466(2)	15.611 (1)
b, Å	13.975(3)	13.93(1)	9.173 (1)
c, Å	10.498(2)	10.538(2)	18.959 (2)
β , deg	108.43(3)	108.54(1)	104.079 (9)
V, Å ³	1458.5(5)	1456.8(4)	2633.5 (4)
D _{calcd} , g cm ⁻³	1.157	1.548	1.111
Z	2	2	4
absorption coeff.	0.385 mm ⁻¹	3.884 mm ⁻¹	0.189 mm ⁻¹
F(000)	536	680	944
crystal dimens., mm	0.18 x 0.20 x 0.22	0.30 x 0.30 x 0.10	0.40 x 0.35 x 0.28
crystal habit	colorless plate	colorless plate	colorless block
θ range for			
data collection	2.39 to 22.52°	2.51 to 22.50°	2.21 to 30.00°
no. of rflns collected	2489	2458	7880
Independent rflns	1900 ($R_{\text{int}} = 0.0254$)	1886 ($R_{\text{int}} = 0.0785$)	6337 ($R_{\text{int}} = 0.0469$)

Table 1 (continued)

	1	2	3
data / restraints / parameter	1900 / 2 / 135	1886 / 2 / 136	6337 / 0 / 254
goodness-of-fit on F^2	1.144	0.717	1.001
final R indices ^a [$I > 2\sigma(I)$]	R1 = 0.0389	R1 = 0.0561	R1 = 0.0492
	wR2 = 0.0865	wR1 = 0.1290	wR2 = 0.1051
R indices (all data)	R1 = 0.0600	R1 = 0.0937	R1 = 0.1022
	wR2 = 0.0961	wR2 = 0.1453	wR2 = 0.1142
extinction coeff.	0.0058 (14)	0.0014 (11)	0.0059 (4)
D(r), and hole e/Å ⁻³	0.27 and -0.193	0.910 and -1.128	0.316 and -0.300

^a Quantity minimized = $aR = S(|F_o| - |F_c|)/\sum |F_o|$; $R_w = [\sum w(|F_o| - |F_c|)^2 / \sum w |F_o|^2]^{1/2}$.

Table 1 (continued)

4	
Empirical formula	$\text{C}_{18}\text{H}_{31}\text{Ga}_3\text{N}_4$
formula weight	512.63
temp, K	298 (2)
radiation	Mo-K α (0.71073)
(wavelength, Å)	
crystal system	monoclinic
space group	C2/m
a, Å	14.184(6)
b, Å	10.454(3)
c, Å	18.926(5)
β , deg	111.37(3)
V, Å ³	2614(2)
D _{calcd} , g cm ⁻³	1.303
Z	4
absorption coeff.	3.077 mm ⁻¹
F(000)	1040
crystal dims., mm	0.40 x 0.20 x 0.20
crystal habit	colorless block
θ range for	
data collection	2.48 to 22.50°
no. of rflns collected	2316
Independent rflns	1811 ($R_{\text{int}} = 0.0617$)

Table 1 (continued)

4	
data / restraints / parameter	1803 / 0 / 130
goodness-of-fit on F^2	1.485
final R indices ^a [$I > 2\sigma(I)$]	R1 = 0.0818
	wR2 = 0.2260
R indices (all data)	R1 = 0.1109
	wR2 = 0.2675
extinction coeff.	-
D(r), and hole $e/\text{\AA}^{-3}$	1.351 and -1.081

^a Quantity minimized = $aR = S(|F_o| - |F_c|)/\sum |F_o|$; $R_w = [S_w (|F_o| - |F_c|)^2 / S_w |F_o|^2]^{1/2}$.

Table 2. Selected bond distances (Å) and bond angles (°) for {[Me₂Al]₄[(μ-PH)₂(C₆H₄)₂] (1),
with Estimated Standard Deviations in Parentheses

Bond Lengths			
Al(1)-C(3)	1.956(4)	Al(1)-C(4)	1.959(4)
Al(1)-P(1)	2.4310(14)	Al(1)-P(2)	2.4340(14)
Al(2)-C(2)	1.947(4)	Al(2)-C(1)	1.944(3)
Al(2)-P(2A)	2.4390(14)	Al(2)-P(1)	2.447(2)
P(1)-C(10A)	1.833(3)	P(2)-C(5)	1.836(4)
P(2)-Al(2A)	2.4390(14)	C(5)-C(6)	1.396(5)
C(5)-C(10)	1.398(5)	C(10)-P(1A)	1.833(3)
Bond Angles			
C(3)-Al(1)-C(4)	119.9(2)	C(3)-Al(1)-P(1)	104.79(12)
C(4)-Al(1)-P(1)	110.69(13)	C(3)-Al(1)-P(2)	110.49(12)
C(4)-Al(1)-P(2)	106.43(13)	P(1)-Al(1)-P(2)	103.44(5)
C(2)-Al(2)-C(1)	123.2(2)	C(2)-Al(2)-P(2A)	114.50(13)
C(1)-Al(2)-P(2A)	108.51(14)	C(2)-Al(2)-P(1)	112.12(13)
C(1)-Al(2)-P(1)	107.55(13)	P(2A)-Al(2)-P(1)	83.95(5)
C(10A)-P(1)-Al(1)	113.59(11)	C(10A)-P(1)-Al(2)	102.60(12)
Al(1)-P(1)-Al(2)	130.07(5)	C(5)-P(2)-Al(1)	111.27(12)
C(5)-P(2)-Al(2A)	102.48(12)	Al(1)-P(2)-Al(2A)	130.79(5)
C(6)-C(5)-P(2)	120.3(3)	C(10)-C(5)-P(2)	120.7(3)
C(5)-C(10)-P(1A)	120.5(3)	C(9)-C(10)-P(1A)	120.4(3)

Symmetry transformation used to generate equivalent atoms: -x, -y+1, -z+2

Table 3. Selected bond distances (Å) and bond angles (°) for $\{[\text{Me}_2\text{Ga}]_4[(\mu\text{-PH})_2(\text{C}_6\text{H}_4)]_2\}$ (2), with Estimated Standard Deviations in Parentheses

Bond Lengths			
Ga(1)-C(4)	1.983(10)	Ga(1)-P(1)	2.419(3)
Ga(1)-P(2)	2.422(3)	Ga(2)-C(1)	1.952(10)
Ga(2)-C(2)	1.955(11)	Ga(2)-P(2A)	2.423(3)
Ga(2)-P(1)	2.433(3)	P(1)-C(10A)	1.832(9)
P(2)-C(5)	1.830(10)	P(2)-Ga(2A)	2.423(3)
C(5)-C(10)	1.405(13)	C(9)-C(10)	1.385(14)
C(10)-P(1A)	1.832(9)	Ga(1)-C(3)	1.972(10)
Bond Angles			
C(3)-Ga(1)-C(4)	120.6(5)	C(3)-Ga(1)-P(1)	105.0(3)
C(4)-Ga(1)-P(1)	110.3(3)	C(3)-Ga(1)-P(2)	109.9(4)
C(4)-Ga(1)-P(2)	106.3(3)	P(1)-Ga(1)-P(2)	103.59(9)
C(1)-Ga(2)-C(2)	123.8(5)	C(1)-Ga(2)-P(2A)	107.8(4)
C(2)-Ga(2)-P(2A)	114.1(3)	C(1)-Ga(2)-P(1)	107.3(3)
C(2)-Ga(2)-P(1)	112.2(3)	P(2A)-Ga(2)-P(1)	84.55(9)
C(10A)-P(1)-Ga(1)	113.8(3)	C(10A)-P(1)-Ga(2)	102.1(3)
Ga(1)-P(1)-Ga(2)	130.09(12)	C(5)-P(2)-Ga(1)	111.6(3)
C(5)-P(2)-Ga(2A)	102.6(3)	Ga(1)-P(2)-Ga(2A)	130.32(13)
C(6)-C(5)-C(10)	119.1(9)	C(6)-C(5)-P(2)	120.4(8)
C(9)-C(10)-P(1A)	120.4(8)	C(5)-C(10)-P(1A)	120.7(7)

Symmetry transformation used to generate equivalent atoms: $-x, -y+1, -z+2$

Table 4. Selected bond distances (Å) and bond angles (°) for
 $[(\text{Me}_2\text{Al})_2\text{AlMe}(\text{C}_6\text{H}_4(\text{NH})_2)_2] \cdot \text{AlMe}_3$ (**3**), with Estimated Standard Deviations in
 Parentheses

Bond Lengths			
Al(1)-N(4)	1.905(2)	Al(1)-N(1)	1.925(2)
Al(1)-N(2)	1.927(2)	Al(1)-C(1)	1.931(3)
Al(2)-C(2)	1.982(3)	Al(2)-C(3)	1.983(3)
Al(2)-C(4)	1.983(3)	Al(2)-N(1)	2.009(2)
Al(3)-C(5)	1.950(3)	Al(3)-C(6)	1.951(3)
Al(3)-N(3)	1.960(2)	Al(3)-N(4)	1.990(2)
Al(4)-C(7)	1.951(3)	Al(4)-C(8)	1.955(3)
Al(4)-N(3)	1.960(2)	Al(4)-N(2)	1.986(2)
N(1)-C(9)	1.468(3)	C(15)-C(20)	1.388(4)
Bond Angles			
N(4)-Al(1)-N(1)	110.20(9)	N(4)-Al(1)-N(2)	108.34(9)
N(1)-Al(1)-N(2)	89.50(8)	N(4)-Al(1)-C(1)	113.54(11)
N(1)-Al(1)-C(1)	119.01(10)	N(2)-Al(1)-C(1)	113.59(11)
C(2)-Al(2)-C(3)	116.43(11)	C(2)-Al(2)-C(4)	112.07(12)
C(3)-Al(2)-C(4)	111.67(12)	C(2)-Al(2)-N(1)	104.05(10)
C(3)-Al(2)-N(1)	107.38(11)	C(4)-Al(2)-N(1)	104.05(10)
C(5)-Al(3)-C(6)	118.96(14)	C(5)-Al(3)-N(3)	110.70(12)
C(6)-Al(3)-N(3)	115.57(11)	C(5)-Al(3)-N(4)	111.43(12)
C(6)-Al(3)-N(4)	111.35(11)	N(3)-Al(3)-N(4)	83.43(9)
C(7)-Al(4)-C(8)	120.63(14)	C(7)-Al(4)-N(3)	108.94(11)
C(8)-Al(4)-N(3)	107.33(11)	C(7)-Al(4)-N(2)	108.82(12)
C(8)-Al(4)-N(2)	107.28(12)	N(3)-Al(4)-N(2)	102.30(9)
C(9)-N(1)-Al(1)	103.3(2)	C(9)-N(1)-Al(2)	116.02(14)

Table 4 (continued)

Bond Angles

Al(1)-N(1)-Al(2)	113.61(10)	C(14)-N(2)-Al(1)	102.24(14)
C(14)-N(2)-Al(4)	109.1(2)	Al(1)-N(2)-Al(4)	121.57(10)
C(20)-N(3)-Al(4)	112.1(2)	C(20)-N(3)-Al(3)	100.8(2)
Al(4)-N(3)-Al(3)	119.59(11)	C(19)-N(4)-Al(1)	116.56(14)
C(19)-N(4)-Al(3)	100.33(14)	Al(1)-N(4)-Al(3)	112.12(10)
C(14)-C(9)-N(1)	116.9(2)	C(19)-C(20)-C(15)	119.8(3)

Table 5. Selected bond distances (Å) and bond angles (°) for $\{(\text{Me}_2\text{Ga})_3[(\mu\text{-NH})_2(\text{C}_6\text{H}_4)(\mu\text{-NH})(\text{C}_6\text{H}_4\text{NH}_2)]\}$ (4), with Estimated Standard Deviations in Parentheses

Bond Lengths			
Ga(1)-C(2)	1.908(12)	Ga(1)-C(1)	1.985(12)
Ga(1)-N(1)	2.043(6)	Ga(1)-N(2)	2.048(9)
Ga(2)-C(4)	1.99(2)	Ga(2)-N(2A)	2.027(10)
Ga(2)-N(2)	2.027(10)	Ga(2)-C(3)	2.01(2)
N(1)-C(10)	1.45(2)	N(1)-Ga(1A)	2.043(6)
N(2)-C(11)	1.470(14)	N(3)-C(9)	1.46(3)
C(5)-C(10)	1.46(3)	C(11)-C(12)	1.37(2)
Bond Angles			
C(2)-Ga(1)-C(1)	125.4(6)	C(2)-Ga(1)-N(1)	111.3(5)
C(1)-Ga(1)-N(1)	104.7(5)	C(2)-Ga(1)-N(2)	108.2(5)
C(1)-Ga(1)-N(2)	105.4(5)	N(1)-Ga(1)-N(2)	98.5(4)
C(4)-Ga(2)-N(2A)	109.6(5)	C(4)-Ga(2)-N(2)	109.6(5)
N(2A)-Ga(2)-N(2)	82.6(6)	C(4)-Ga(2)-C(3)	118.0(9)
N(2A)-Ga(2)-C(3)	115.9(6)	N(2)-Ga(2)-C(3)	115.9(6)
C(10)-N(1)-Ga(1A)	111.9(5)	C(10)-N(1)-Ga(1)	111.9(5)
Ga(1A)-N(1)-Ga(1)	119.5(6)	C(11)-N(2)-Ga(2)	101.3(7)
C(11)-N(2)-Ga(1)	111.0(7)	Ga(2)-N(2)-Ga(1)	116.9(5)
N(1)-C(10)-C(5)	121(2)	C(12)-C(11)-N(2)	122.5(11)

Symmetry transformation used to generate equivalent atoms: $x, -y, z$

TECHNICAL REPORT DISTRIBUTION LIST - GENERAL

Office of Naval Research (1)*
Chemistry Division, ONR 331
800 North Quincy Street
Arlington, Virginia 22217-5660

Dr. Richard W. Drisko (1)
Naval Facilities & Engineering
Service Center
Code L52
Port Hueneme, CA 93043

Defense Technical Information
Center (2)
Building 5, Cameron Station
Alexandria, VA 22314

Dr. Eugene C. Fischer (1)
Code 2840
Naval Surface Warfare Center
Carderock Division Detachment
Annapolis, MD 21402-1198

Dr. James S. Murday (1)
Chemistry Division, Code 6100
Naval Research Laboratory
Washington, D.C. 20375-5320

Dr. Bernard E. Douda (1)
Crane Division
Naval Surface Warfare Center
Crane, Indiana 47522-5000

Dr. John Fischer, Director (1)
Chemistry Division, C0235
Naval Air Weapons Center
Weapons Division
China Lake, CA 93555-6001

Dr. Peter Seligman (1)
Naval Command, Control and
Ocean Surveillance Center
RDT&E Division
San Diego, CA 92152-5000

* Number of copies to forward