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Summary

This project was carried out at the National Renewable Energy Laboratory (NREL), Sandia National Laboratory (SNL), and by a group at the University of Idaho (UI) working under subcontract to NREL. Project staff met with engineering and environmental compliance staff at Longhorn and Iowa Army Ammunition Plants to become familiar with current pink water treatment technology and requirements. The initial focus was on photocatalytic oxidation and reduction chemistry that could use sunlight as the photon source. This work demonstrated the technical feasibility of the photocatalytic processes; however, the process removed the hazardous components of pink water at such low efficiency that it would be expensive to operate lamps in a treatment system. Technical success in removing trinitrotoluene (TNT) from water by photocatalytic reduction chemistry led us to test that as a pretreatment step for a biological treatment process. The photocatalytic reduction of TNT did render the simulated pink water more amenable to mineralization by *Phanerochaete chrysosporium*. This result led us to explore other chemistry for removal of TNT and related explosives from water.

The new process that was investigated in this project very effectively removed TNT from pink water. The operating cost is estimated to be about half that of conventional treatment using granular activated carbon (GAC). The capital costs are estimated to be about the same for the new process and GAC treatment. The new process is sufficiently novel to warrant initiating steps that would lead to a patent. Reporting in an open literature document constitutes disclosure. For this reason the details of the new process are being withheld until the novelty and potential for a patent application are determined. Patents resulting from this work will be submitted to the Strategic Environmental Research and Development Program (SERDP) personnel when issued. In this report a general description of the results of work on the new process are presented. The details will be made available to parties willing to sign a non-disclosure agreement with NREL.

1.0 Introduction

1.1 Objective of Work

The objective of this project was to develop new photocatalytic or other innovative process chemistry for the treatment of pink water and related contaminated water.

1.2 General Background

Pink water is the term applied to waste water from Load, Assemble, and Pack (LAP) operations at Army Ammunition Plants (AAPs) where 2,4,6-trinitrotoluene (TNT) in combination with other explosives is loaded into various munitions. 1,3,5-Trinitro-1,3,5-triazacyclohexane (RDX) and 1,3,5,7-tetranitro-1,3,5,7-tetrazacyclooctane (HMX) are used with TNT in the composite explosives.

The abbreviation HE will be used to designate high explosives individually or collectively.¹ The

chemical structures of the compounds of interest in this work are shown in Figure 1.1. Water from washing and condensed steam used to melt composite explosives contain low levels of TNT (about 100 ppm), HMX (about 10 ppm) and RDX (about 25 ppm). When exposed to sunlight the water becomes colored owing to the formation of a number of products of photochemical reactions of TNT. The photochemistry of TNT in water has been elucidated by a number of workers.^{2,3,4,5}

Pink water also arises from cutting and milling operations and during demilitarization of munitions being

taken out of service, in which the explosive is melted using steam. Contaminated ground water with components like those in pink water is found at AAPs where soil has been contaminated owing to past disposal practices or where waste disposal ponds have leaked into aquifers. Environmental contamination by these compounds is also found at depots and bases within the Department of Defense complex and at some Defense Program sites operated by the Department of Energy. Pink water is considered hazardous because of the toxicity of the original explosive compounds or the presence of other compounds that were by-products of the production process. Dinitrotoluene and trinitrobenzene are examples of such compounds.

The location of AAPs in the United States are shown on the map in Figure 1.2.¹ Most of the AAPs are in various stages of shut down and environmental clean up activities. Pink water is produced at Iowa, Longhorn, and Holston.

The current state of the art for treating pink water at AAPs is adsorption on granulated activated carbon (GAC). The flow diagram for the treatment system in use at Iowa AAP is shown in Figure 1.3. The system as operated meets current discharge standards for TNT, HMX, and RDX of <1, <5, and <5 ppm, respectively. Iowa AAP currently produces about 18,000 gal/day of pink water and would produce about 100,000 gal/day at full mobilization. The spent carbon is sent for regeneration and can be recycled about four times before it becomes unusable due to particle size reduction and fragility.⁶ Experience with GAC regeneration and disposal at AAPs was surveyed by A. D. Little, Inc. for the Army.⁷ There is concern that the system at Iowa AAP will not be able to meet lower discharge standards currently under discussion. The new limits would be 600 ppb TNT and 1 ppm for HMX and RDX.⁶

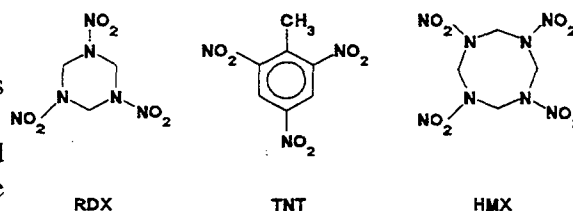


Figure 1.1. Structure of HE Compounds Most Commonly Found in Pink Water.

- 1 Alabama
- 2 Badger
- 3 Cornhusker
- 4 Crane
- 5 Gateway
- 6 Hawthorne
- 7 Hayes
- 8 Holston
- 9 Indiana
- 10 Iowa
- 11 Joliet
- 12 Kansas
- 13 LakeCity
- 14 LoneStar

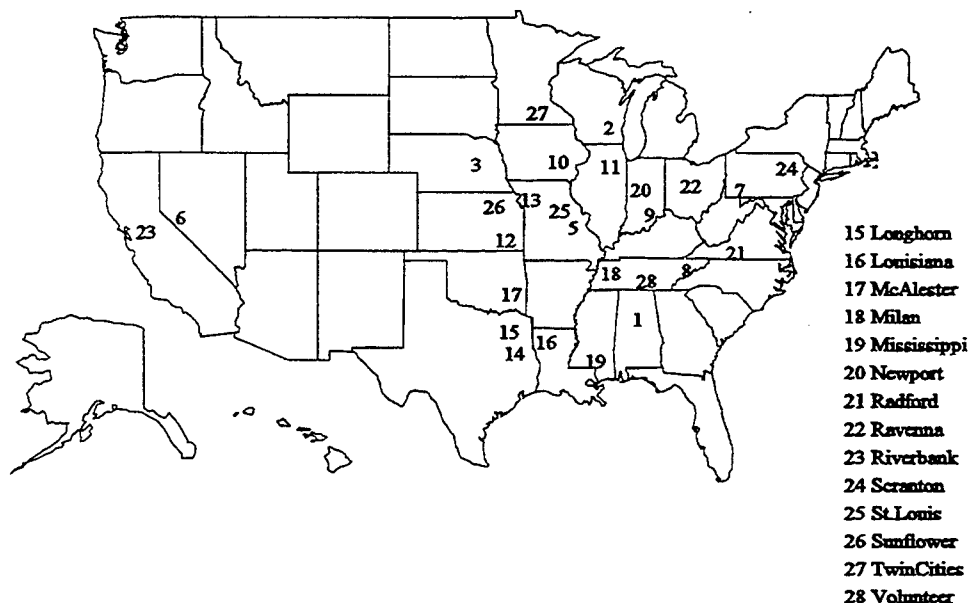


Figure 1.2. Location of Army Ammunition Plants

The ability of GAC treatment to meet proposed or potential discharge standards having lower levels of contaminants has been investigated for the Army at a bench and pilot scale. The work concluded that GAC would have a problem meeting new discharge standards if they were to be set as low as 300 ppb for TNT and 30 ppb for HMX and RDX as was being discussed for some locations.⁸

The requirement for regeneration and ultimate disposal of contaminated GAC, the associated costs, and potential that the carbon adsorption process might not treat pink water to meet future standards has resulted in a continuing search for new treatment processes.

Advanced oxidation processes have received the most attention among the chemical treatment methods. The common feature of these processes is the involvement of hydroxyl radicals in oxidation. Ultraviolet (UV) light in combination with ozone, UV-O₃,^{9,10,11} and hydrogen peroxide, UV-H₂O₂,^{12,13,14,15} have been studied at laboratory and pilot plant scale. The Fenton reaction (Fe⁺²/H₂O₂) has also been investigated.^{16,17} Heterogeneous photocatalytic oxidation has received some attention and will be discussed in Section 3 in connection with the description of one of the processes being investigated in this project. Other physical or chemical processes that have been studied include use of advanced adsorption materials,^{18,19} surfactant precipitation,²⁰ and photosensitized reactions.²¹

Biological treatment of pink water and water having related contaminants has been investigated by number of groups.^{22,23,24,25} Anaerobic biodegradation has been studied at both laboratory and pilot scale. Of concern in the biotreatment processes is the general observation that only a small fraction of carbon-14 labeling incorporated

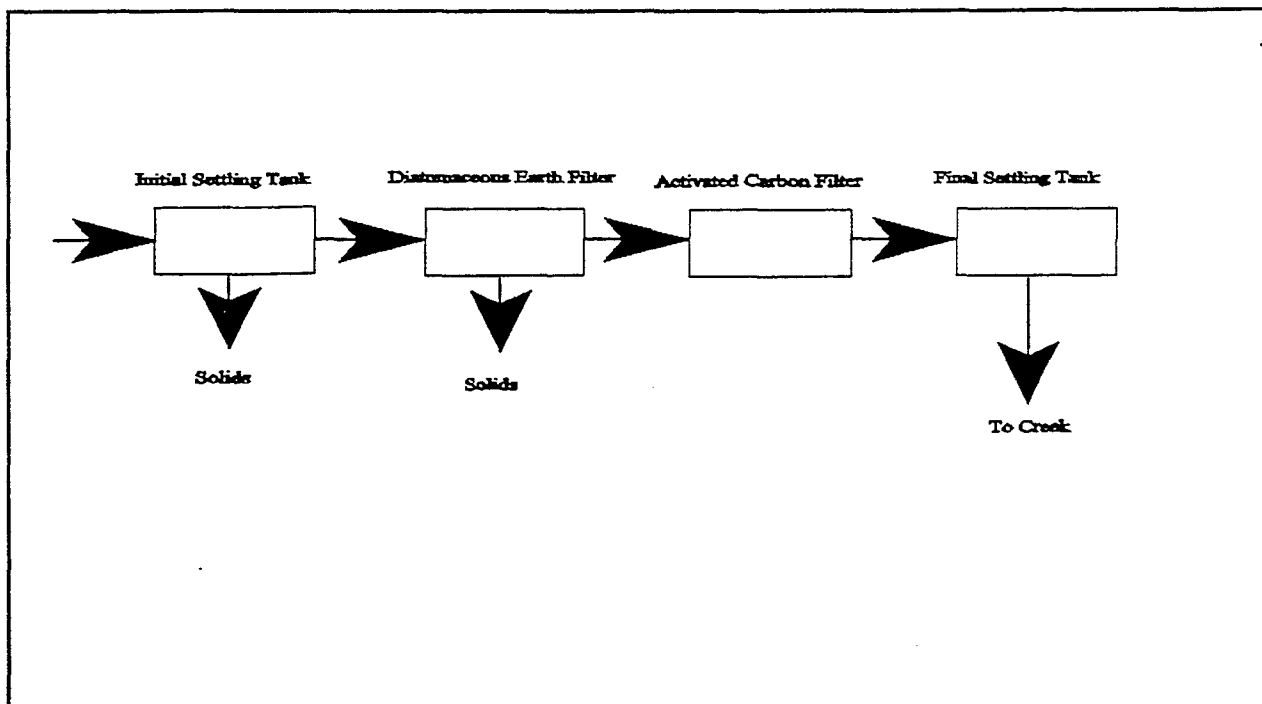


Figure 1.3. Overall Process Flow Diagram for Pink Water Treatment Unit at Iowa AAP

into TNT is released as carbon dioxide. This indicates that TNT is converted to other compounds that remain in the water or soil being treated. Failure to achieve mineralization and the potential that some of the by-products of biological treatment may be hazardous has been a barrier to adoption of the technology.²⁵

Reduction chemistry for purposes of treating pink water has been investigated mainly in the context of biological treatment under anaerobic conditions. However, there is an extensive body of work on reduction of nitro groups in organic compounds for synthetic purposes.²⁶ Denitrification (removal of nitro groups) of unsymmetrical isomers of TNT by reaction with sodium borohydride was investigated as a means of purifying or recycling waste material.²⁷ A study of photochemical reactions of sodium borohydride with dinitrotoluenes and related compounds was recently reported.²⁸

2.0 Photocatalytic Oxidation and Reduction Processes and the Combination of Photocatalytic Reduction with Biological Treatment

2.1 Background

Heterogeneous photocatalytic chemistry as applied to oxidation is considered to be an advanced oxidation process because as usually applied under aerobic conditions, hydroxyl radicals are believed to be formed. The general reaction scheme for action of a photocatalyst in water is as follows:



The valence band hole functions as an oxidizing agent and the conduction band electron as a reducing agent. The semiconductor absorbs light at all wavelengths having energy greater than the band gap, which is 3.2 eV ($\lambda < 385 \text{ nm}$) for anatase. This is shown pictorially in Figure 2.1. The reduction potentials of valence and conduction bands depend on the particular semiconductor. Most work done on the application of this chemistry to environmental problems to date is based on titanium dioxide in the anatase form as the photocatalyst. The oxidizing power of the electron hole in TiO_2 is sufficient to oxidize water to hydroxyl radical, OH^\bullet , which then initiates the oxidation of the organic compound.^{29,30} The conduction band electron must be transferred to an oxidizing agent to maintain electroneutrality. The oxidizing agent is usually molecular oxygen or hydrogen peroxide. The net chemical reaction is the oxidation of the target compound by oxygen or hydrogen peroxide. A lesser amount of work has been done in which a sacrificial reducing agent, usually an organic compound, is used as a hole (or OH^\bullet) acceptor and the conduction band electron is used to reduce metal ions for the purpose of removing them in the elemental form or as a reduced hydroxide precipitate.³¹ Some studies have demonstrated electron transfer from the excited semiconductor particles to dye molecules. The process has been suggested as a method to decolor water contaminated with dyes.³²

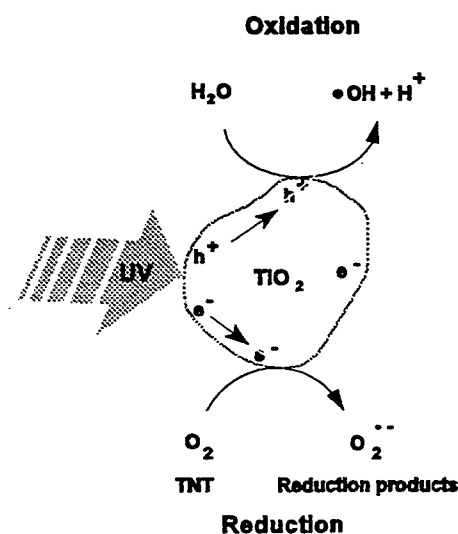


Figure 2.1. Oxidation and Reduction on TiO_2 .

Some work on photocatalytic oxidation of TNT had previously been carried out under aerobic conditions.^{33,34,35,36,37} Other related compounds such as nitrophenols and nitrobenzenes have also been studied.^{38,39} This and other work on photocatalytic processes suggested that the process has potential for treating pink water and that there is opportunity for innovation.

TNT and other explosives are based on chemical structures which contain both oxidizing and reducing sites. This presents the possibility that they might be amenable to a photocatalytic destruction process under conditions where no added oxidant or reductant is required. A second possibility is that the oxidizing nature of the substances may make them amenable to photocatalytic reduction chemistry. Reaction with a sacrificial reductant may be a means to convert them to harmless substances or forms that can be readily be treated by another process step. These innovative chemical approaches have been investigated in this work. A further consideration in this

project is that a combination of a photocatalytic treatment operation with a biological treatment process might be more effective than either alone.

To these ends, the research reported here was done. Characterization of the photocatalytic chemistry under aerobic and anaerobic conditions and with various oxidizing and reducing agents and semiconductors was carried out at NREL and SNL. Work on the combination of photocatalytic reduction chemistry and biotreatment was done under subcontract by a group at the University of Idaho that has extensive experience with biotreatment of soil and water contaminated with TNT and herbicides based on nitroaromatic compounds.

2.2 Methods

General methods and reagents. Synthetic pink water was prepared in the laboratory from deionized water and the appropriate high explosive. It was prepared by dissolving the explosive material(s) in 18 M Ω water with stirring over a period of days. The stock solutions were colorless and stored in amber glass bottles which were kept in dark cabinets until used. The concentrations of the high explosives approached their saturation limits in water (1ppm = 1mg/L): TNT (90 ppm), RDX (40 ppm) and HMX (4 ppm). Actual pinkwater was obtained from the Louisiana Army Ammunition Plant (LAAP), Thiokol Corporation, Shreveport, LA and from the Iowa AAP (IAAP), Mason and Hanger-Silas Mason Co., Inc., Middletown, OH. The sample from the LAAP contained 63 ppm of TNT, 52 ppm RDX and a number of other components that may have resulted from photo-reactions of TNT. The color of this solution is deep pink or orange. The IAAP sample contained approximately 45 ppm C of dissolved organic carbon and approximately 210 ppm C of dissolved inorganic carbon. Three organic species were identified and quantified as HMX (~10 ppm), RDX (~30 ppm) and TNT (~80 ppm), which accounts for about 32 ppm C of the dissolved organic carbon. Additionally, 2-amino-4,6-dinitrotoluene was detected in the sample but was not quantified. Degussa P25 and anatase from Tioxide were used in the photocatalysts.

Batch reactor. Most experiments performed at Sandia National Laboratories were performed in a 1.5 L batch photocatalytic reactor assembly, shown schematically in Figure 2-2. In each experiment, 400 mL of the high explosive (HE) mixtures were placed in the glass pot reactor and illuminated with an Oriel 1000 Watt Hg-Xe ozone-free arc lamp. The pot was maintained at 20°–25°C through the use of a cooling jacket. All external optics employed were constructed of quartz (however, the lamp bulb is composed of Pyrex to forestall ozone generation) to maximize transmittance. A Spectronics Spectroline band pass filter, 300–420 nm, was used to restrict the wavelengths passed to the near UV region. Actinometry measurements showed that about 5.2 watts of incident light that is available for TiO₂ photocatalysis (300–400 nm). The pot reactor was sealed gas-tight with a quartz lid and a Viton O-ring. Anaerobic reactions were carried out by vigorously sparging the reactor with nitrogen gas for 30 min before the reaction and then slowly during the reaction through a fritted glass tube. A similar procedure was employed for aerobic reactions using zero-air. All reactions were stirred using a magnetic stirrer. In some experiments the initial pH was adjusted by addition of nitric acid or sodium hydroxide solutions.

A typical run was performed for 4–5 h and had samples removed for total organic carbon (TOC) (5mL), (HPLC) (2mL) and UV-Vis (2mL) analyses at predetermined times. TOC samples were covered with Parafilm and refrigerated until analyzed. HPLC samples were sealed and refrigerated until analyzed. In order to maintain the UV-Vis spectra within the linear range of the spectrophotometer (2.0 absorbance units), all UV samples were either analyzed with a 1-mm-pathlength cell or diluted by a factor of four with deionized water. All of the samples were withdrawn through a sealable sampling port. Each of the samples was filtered prior to analysis through a 0.2 μ m Teflon syringe filter. Catalyst was recovered for analysis by filtration using nylon filters and was allowed to air dry prior to solid state analysis.

Experiments requiring concentrated light were done using the Oriel 1000 watt lamp fitted with a 4" convex, quartz focussing lens and appropriate neutral density filters.

Recirculating reactor. A schematic diagram of the recirculating, batch, annular photoreactor reactor assembly, showing the major components, is presented in Figure 2-3. The assembly consists of three main components: a jacketed water-cooled reservoir, a magnetic-drive centrifugal pump, and an annular photoreactor. The three parts are connected by 3/8" corrugated flexible Teflon tubing. The reservoir is a 1 L, water-jacketed, round-bottomed flask, modified to contain three ground-glass fittings at the top and one 1/2" (OD) glass tube protruding from the bottom. The centrifugal pump has a magnetic drive with a plastic impeller. Two different pump models were used in the various reactor assemblies. Both had maximum flowrates of approximately 10 L per minute. This flow rate provides a residence time of approximately 1.5 s per pass in the annular reactor.

The annular reactor design is similar in size and design to those used by Turchi⁴⁰ and consists of six parts: two identical endpieces, machined out of Teflon by technicians at NREL, two glass pipes which constituted the inner and outer walls of the reactor, and two Teflon flow dispersion disks. The outer glass pipe is constructed from 54 mm standard-wall Pyrex tubing. The inner glass pipe is constructed of 35 mm standard-wall Pyrex tubing. The dispersion disks are used to initiate turbulent flow and decrease any dead space of the reaction mixture and to guarantee a well-distributed flow through the reactor. A satisfactory design of the distributor holes was determined by trial and error. Several different designs were tested and the flow distribution was visually examined using a methylene blue solution in the reactor to improve contrast.

The reactor assembly is illuminated by placing a 1" diameter, 18" long fluorescent lamp (NEC 15W F15T8-BLB "Blacklight Blue") in the center of the annulus. Actinometry measurements on this lamp yielded 1.7 watts of light available for photocatalysis (300-400 nm). The temperature was controlled at $25 \pm 1^\circ\text{C}$ by the circulation of water through the reservoir water jacket from a controlled temperature water bath (VWR scientific).

Samples were taken at intervals throughout the experiment. Each sample (5.0 mL in volume) was drawn from the reactor reservoir and filtered through a 0.2- μm -pore-size nylon syringe filter. The first 2 mL of filtrate were discarded and the remainder was split into two amber HPLC vials and sealed with a crimped aluminum cap and septum. Following the experiment, the catalyst was recovered by filtration using a 0.2- μm -pore-size Nalgene cellulose acetate filter for eventual analysis to determine the amount of carbon and nitrogen adsorbed on the used catalyst. A fraction of the supernatant was also kept for use in other analyses, if required.

2.2.1 Analytical methods

Spectrophotometric methods. Ultraviolet-Visible spectroscopy (UV-Vis) was performed using an HP8452A photodiode array spectrometer. Scans were made between 190 and 820 nm using a 1-mm-path-length cell digitally zeroed against 18 M Ω deionized water.

High performance liquid chromatography (HPLC). At NREL, analyses for parent HE compounds and detection of intermediates were performed using a Hewlett Packard 1050 HPLC system with an online vacuum degasser and gradient pump, autosampler, and single wavelength 240 nm detector, which was interfaced to a PC-based data collection system sampling at 5 Hz. The system was calibrated through a range of two orders of magnitude with six calibration standards analyzed in triplicate. A quantitation limit for TNT in water was established to be below 20 ppb. The chromatography was performed on a 250- x 4-mm column packed with 5 μm C18 Spherisorb ODS2 stationary phase. The mobile phase was a water/methanol ($\text{H}_2\text{O}/\text{MeOH}$) gradient from 5% methanol (0 to 3 min) to 100% MeOH from 3 to 10 min and held at the final level for the final 10 min of elapsed time. Between each chromatographic run, the column was equilibrated with MeOH for 5 min. The flow rate was 0.75 mL/min, the column was maintained at 27.5°C , and the sample volume was 25.0 μL . All intermediates were reported in units of UV-Vis response unless they were specifically identified and could be quantitated.

At SNL, a Waters HPLC system with an online helium degasser, Waters 600E gradient pump, WISP 715 autosampler, and a Waters UV-Vis 486 detector was interfaced with a PC data collection and control system running Water's Millennium 2010 software, sampling at 80 Hz and data averaging to manageable data files of 5 Hz. The system was maintained at 1 mL/min and ambient temperature. It was calibrated with multiple volume injections of two standards for each HE compound. The chromatography was performed on a 25-cm by 3.9-mm column packed with 5 μ m reverse-phase Spherisorb stationary phase. The mobile phase was a water/methanol gradient from 90/10% to 30/70% during 40 min. The typical sample volume was 40 μ L and the detector was set at 231 nm. The detection limit was 150 ppb with no change in sensitivity range and 10 ppb with a change in sensitivity range.

Total organic carbon (TOC) analysis. At NREL, a Dohrmann DC-180 TOC analyzer measured total dissolved organic carbon using the UV/persulfate method. The instrument was calibrated as per manufacturer specifications at 100 ppm C using sodium bicarbonate, NaHCO_3 , for inorganic carbon (IC) and potassium hydrogen phthalate, KHP, for organic carbon (OC). The calibration was verified and corrected daily with a two point curve (plus a blank) analyzed in triplicate. The total dissolved organic carbon (TDOC) was determined by analyzing the total IC and the total dissolved carbon in a 150 μ L sample and taking the difference.

At SNL, TOC was performed using a Shimadzu TOC-5000 with an ASI-5000 autosampler. The instrument measures total carbon by passing the sample through a platinum oxidation catalyst and detection of resultant CO_2 with an infrared analyzer. The inorganic carbon was measured by the outgassing of an acidified solution. The instrument was calibrated as per manufacturer's instructions with three points and 1 blank of 18 M Ω Millipore Ultrapure water. The ranges were two organic carbon ranges (10-50 ppm and 1-10 ppm) using the compounds as above. The instrument was also calibrated for two inorganic carbon ranges (0-1 ppm and 1-10 ppm). All of these calibrations can be stored within the instrument. The instrument was calibrated a minimum of once a month and standards were performed alongside experimental samples to ensure that the calibration was still valid.

Carbon, hydrogen, and nitrogen (CHN) analysis. CHN analysis of solid samples was performed on a Perkin-Elmer 2400 CHN Elemental Analyzer. Oxidative combustion is performed at 925°C with a mixture of catalysts including silver tungstate on magnesium oxide, silver vanadate as well as chromium and nickel oxides. The reduction analysis is performed at 640°C and employs a mix of Cuprox and a copper catalyst. The samples are placed in a crimped foil holder and can be varied from 2.5 mg to 7.5 mg in size. Standards are performed prior to each analysis and use high-carbon-iron alloy for carbon standardization, sucrose for hydrogen and carbon calibration and acetaniline for nitrogen calibration. All samples were performed as triplicates and amounts were reported as percentages.

Ion Chromatography. Data were collected using a Dionex 2000i/SP ion chromatography system fitted with a Dionex autosampler and attached through a Dionex Advanced Computer Interface to an IBM 386 PC running Dionex AI-450 Acquisition Software (rev. 3.32). Anions were eluted in an aqueous 1.8 mM Na_2CO_3 /1.7 mM NaHCO_3 solution through a Dionex AS4A 4-mm column. Cations were eluted in an aqueous 20 mM methane sulfonic acid solution through a Dionex CS12 4-mm column. Both solutions were He degassed and suppressed inline for complementary ions. A 0.5 mL sample was injected into the system and known standard solutions were used to calibrate the instrument.

Carbon Dioxide Analysis. Off-gassing of carbon dioxide was measured via a $\text{Ba}(\text{OH})_2$ solution connected to the outlet purge line of the pot reactor. This solution contained a two times excess of reagent based on the total carbon in the reaction system. Background checks showed no appreciable CO_2 in the zero-air over the time period of the pot reactions. Pot reactions were taken "off-lamp" and allowed to stir and purge for 30 min after the end of the reaction to fully sweep out the solution and head space. The barium carbonate precipitate was recovered on ashless filtered paper and then washed with degassed deionized water. The sample was allowed to air dry for

a minimum of two days and ashed in a 500°C muffle furnace. The crucibles used were first zeroed at 500°C. The samples were ignited for four hours total. They were removed and weighed after 2 h of heating and 4 h (cooling in a desiccator) with no significant difference in the two weights. This analysis was not performed for HMX reactions owing to the low level of carbon in the samples.

2.2.2 Photocatalytic Process Under Oxidative Conditions

2.2.2.1 Results

It is well known that TNT is light sensitive. Photochemistry may be initiated by irradiation with light that has an overlap with the absorption spectrum shown in Figure 2.4. Photochemical degradation of TNT in water is illustrated in Figure 2.5 for a solution of 90 ppm TNT that was illuminated in a batch reactor in the absence of any photocatalyst. Two types of data are shown in the figure. The left ordinate with the bottom abscissa corresponds to UV-Vis difference spectra obtained by subtracting the initial spectrum ($t=0$) from the spectra taken at later times. The dashed spectrum shows what the difference spectrum would look like if all of the TNT initially present were destroyed, and the horizontal line at zero absorbance units is the difference spectrum for $t=0$. The spectra between these two extremes show the progress of the reaction. Difference spectra were used because they highlight the formation of light-absorbing reaction intermediates. Any part of the spectrum that is greater than zero results from these intermediates. As shown in Figure 2.5, the conversion of TNT is slow in the absence of catalyst and significant intermediates was formed. This result was confirmed by the TOC data which are plotted against the right ordinate and the top abscissa. There is some scatter in the TOC data but it is evident that the extent of mineralization (conversion of carbon content to CO_2) is very low during the five hours of light exposure.

Photocatalytic destruction of TNT, HMX, and RDX individually and as components in actual pink water were carried out in both the batch and recirculating reactor systems. Aerobic reactions were performed by sparging with compressed air. A representative example of the photocatalytic oxidation of TNT in the batch reactor is shown in Figure 2.6. Complete destruction of TNT was achieved in five hours of irradiation in the batch reactor system. The TOC was reduced by only about 50%-60% under these conditions. A similar experiment was done in the recirculating reactor, Figure 2.7, with comparable results. Because pH can have a significant effect on TNT chemistry, experiments were done to test the effect of initial pH on photocatalytic oxidation. These results are shown in Figure 2.8. The final pH was usually in the range of three to four independent of the initial pH. When the pH was held constant at nine by the periodic addition of base there was a smaller decrease in the TOC and a larger number of detectable by-products were formed. Figure 2.9 provides information on the fate of the nitrogen under aerobic and anaerobic photocatalytic conditions. Nitrate, nitrite, and ammonium ion were found.

Analytical work was done on the aqueous, solid, and gas phases in order to identify the fate of the carbon and nitrogen in these reactions. This is summarized in Table 2-1. In this set of experiments about 40% of the carbon from the explosive TNT was converted to CO_2 (i.e., mineralized) from solutions by oxidative photocatalysis. Oxidation of TNT resulted in a number of organic intermediates remaining in the solution after the 5 h reaction time (about 38% of carbon by TOC analysis and seven intermediates as determined from HPLC analysis). The catalyst recovered from the aerobic reaction with TNT contains detectable amounts of carbon and nitrogen (about 15% and 2% of the total material introduced, respectively). However, no color is evident by visual inspection on the surface of the used catalyst. Intermediate products were not in general identified. An exception is trinitrobenzene which was identified by comparison of HPLC retention times and UV-Vis spectra with those of authentic samples.

Similar data were obtained for HMX and RDX which are also shown in Table 2-1. The destruction of RDX, about 75% conversion, yielded predominantly CO_2 and material adsorbed on the catalyst surface. In the case of HMX, the extent of destruction was about 80% in a 5 h experiment.

Table 2-1. Mole Balance Data for Carbon and Nitrogen Under Aerobic Conditions.

	TNT O ₂	RDX O ₂	HMX O ₂
Type of Carbon			
Solution Carbon			
HE Carbon	0%	30.6%	0%
Carbonate	<1	<1	0
By-product(s)	37.8	0	17.8
Carbon on TiO ₂	14.8	27.2	1.5
Carbon Dioxide	40.1	44.1	ND*
Recovered Carbon	92.7	101.9	19.3
Type of Nitrogen			
Solution Nitrogen			
HE Nitrogen	<1%	30.6%	0%
Ammonium	15.8	14.8	19.8
Nitrite	<1	<1	0
Nitrate	12.4	14.1	16.3
By-product Nitrogen**	37.8	0	17.8
Nitrogen on TiO ₂	2	0	0
Recovered Nitrogen	68.2	59.7	53.9

*ND = not done

**The numbers for by-product nitrogen are assumed from the percentage of by-product carbon. All other values are experimentally determined.

A number of experiments with mixtures of TNT and RDX were performed in the recirculating reactor assembly using a hard quartz annulus and a lamp which emitted short wavelength UV ($\lambda \sim 185$ to 254nm) illumination. In this configuration the reaction mixture, with TiO₂ photocatalyst, was illuminated by the full lamp output spectrum. Passing air through the space between the lamp and the inner reactor wall generates ozone. The mixture of oxygen and ozone was then bubbled through the reactor reservoir. The results of these experiments are presented in Figure 2.10 for TNT, HMX, and RDX. TOC analyses of these experiments are shown in Figure 2.11. In all three compounds the removal of TOC was significantly slower than the disappearance of the explosive.

The use of short-wavelength UV illumination with no catalyst and an air atmosphere results in loss of the reactants at a rate comparable to that found with the use of photocatalyst. Data on removal of the explosives and the corresponding TOC analysis is shown in Figures 2.12 and 2.13, respectively. The catalyst was found to have little effect and may have slightly reduced the rate of reaction by blocking the light available for direct photolysis. It is likely that the use of short-wavelength illumination results in direct photolysis not possible with near-UV light.³

2.2.2.2 Discussion

The photocatalytic oxidation of TNT, HMX, and RDX in both synthetic samples and in the form of pink water from AAPs has been studied under a wide range of conditions. The results generally agree with those reported by Schmelling and Gray,³³ and Schmelling et al.,^{33,34} Kang and Kutal,³⁶ and Diller, et al.³⁷ However, Schmelling and Gray report significantly more rapid mineralization of TNT than we have been able to achieve. They report >90% mineralization of TNT within 120 min whereas the best we have been able to achieve is about 50% conversion to CO₂ within 300 min.

We carried out experiments under conditions that attempted to reproduce some of the earlier work. Analytical methods were chosen which allowed us to follow the disappearance of the parent HE compound, track the appearance of intermediate oxidation products and in some cases identify them, to follow total organic carbon content of the aqueous phase, to analyze the titanium dioxide photocatalyst for adsorbed carbon and nitrogen, and to determine the amount of evolved carbon dioxide. This allowed mole balances to be determined, Table 2-1.

In general, under conditions in which the aqueous phase was saturated with oxygen by sparging with air during the experiment, the photocatalytic oxidation of TNT proceeds with formation of a number of intermediates. These were not identified but were followed by HPLC analysis and reported as total counts by the UV detector of the HPLC unit. Depending on pH, the number of detectable major intermediates ranged from four to seven; the fewest were observed at the lowest pH. After four hours of irradiation the conversion of TNT was in the range 50%-60% and the TOC in solution remained at about 40%-50% of the initial level. There was about a 40% yield of carbon dioxide. This is consistent with the existence of intermediate products that are not mineralized.

Our results are in reasonable agreement with the work of Wang and Kutal who reported about 60% mineralization of TNT during a 4 h reaction.³⁶ Their experimental system is similar to ours; they used a purged batch reactor and light of wavelength above 290 nm. They used oxygen instead of air, 0.1 wt% of catalyst, and used the Degussa P25 form of TiO₂ rather than that from Tioxide, Ltd. We tested the effects of these three variables and have observed no significant differences in either the rate of reaction, nor in the final TOC concentration. It is difficult to make more detailed comparisons because of the lack of details of the lamp and reactor geometries.

Comparison with the results of Schmelling and Gray³³ and Schnelling et al.,^{34,35} is less direct. They report about 90% mineralization within about 120 min under the following conditions. They used a closed batch reactor system, saturated with oxygen, 0.025 wt % P25 catalyst, and irradiation with a 450 W medium pressure mercury arc lamp with the output filtered to remove all but light between 340 and 390 nm.

In an attempt to duplicate the work of Schmelling and Gray, we experimented with other variables: a saturated oxygen atmosphere, the use of tap water for the HE solutions, lowered concentrations of TNT in solution (40-50 ppm), higher than normal incident light concentrations (use of a focusing lens and a small reactor), the use of P25 as the catalyst, and a band pass light filter. The results of these experiments were similar to results of our previous work. One observable difference was the appearance of a red color in the reaction solution that used a tap water/HE mixture. This red solution is reported at the beginning (after exposure to the lamp) of all of the reactions reported by Schmelling and Gray; however, we do not observe color in the solution until 2-3 h into the reaction. Although TOC dropped during the reaction to one third to half of the initial concentration, Carbon in our solutions did not disappear to the degree achieved by Schmelling and Gray. We attempted to increase the TOC removal by varying other reaction conditions. In all of our reactions the pH started out slightly acidic (ca. pH 6.5) and became more strongly acidic (ca. pH 3.5) at the end of the reaction. Experiments tested how the initial or reaction pH affected the final TOC concentrations. The reactions were started following adjustment to a pH value of 9, 7, 5, or 3 by addition of sodium hydroxide or nitric acid solution, Figure 2-8. In another test, the solution was held in the range of 8-10 by periodic addition of base throughout the. Color change in these

reactions evident only towards the 3-4 h mark. We can generate this color by artificially raising the pH at the beginning of the reaction (> 8 pH). Therefore there could be an effect of the initial pH. We have also performed experiments where the pH is varied (to address this problem) and this has no major effect on the final TOC results, but it does significantly influence the type and number of intermediates formed. In all experiments, there was no significant impact on the TOC compared with our earlier results.

2.2.3 Photocatalytic Chemistry Under Reductive Conditions

2.2.3.1 Results

Reducing conditions have also been tested in order to evaluate the full range of photocatalytic capability of irradiated TiO_2 . When a photocatalytic system is run under anaerobic conditions electron transfer from the conduction band of the irradiated semiconductor to a target compound can be enhanced since there is no oxygen to scavenge electrons.^{31,32} Under these conditions electrons and holes must both be consumed by the HE compound(s) in order for a sustained reaction to occur. Reduction of the HE can be forced by added reducing agents that would consume holes so that only electrons would be available to react with the HE compound(s). Because light absorption by the TiO_2 photocatalyst generates both a hole and an electron, it was postulated that if we could prevent electron scavenging by oxygen (by keeping a strictly anaerobic system), or by adding a sacrificial reductants (to scavenge the photo-generated holes), we could direct the chemistry toward photo-reduction of the explosive compounds. Organic and inorganic reducing agents were screened for activity.

Organic reductants that prior work³¹ found to be effective in the photocatalytic reduction of metal ions were tested on TNT degradation in a batch reactor. Organic reductant (0.72 mM concentration, 2:1 molar ratio of added organic to TNT) was added to a nitrogen-purged TNT solution containing suspended TiO_2 . The results are summarized in Figure 2.14, which also includes the case of the self-redox reaction (no added organic). Citric acid and disodium ethylenediaminetetraacetic acid (Na_2EDTA) produce the greatest increase in the rate; RDX and hydrogen have no effect; and salicylic acid inhibits the reductive photocatalysis.

Mass balances were determined for the reactions of TNT, HMX and RDX under anaerobic conditions with no added reducing agent. The results are summarized in Table 2-2. At the end of the anaerobic reaction of TNT the catalyst was coated with a brown substance (60% deposited carbon and 45% nitrogen, where percentages refer to the fraction of the element initially present as TNT) or a 4/1 by mass ratio of carbon to nitrogen. This result, when compared to a 2.3/1 ratio of carbon to nitrogen in the parent compound, suggests that this is not simply adsorbed TNT, but rather some decomposition product. Analytical techniques have yet to positively determine the nature of this coating, but early results point towards a polyaniline analogue produced through the reduction of nitro-groups to amines or diazo linkages.⁴¹ We have also found that the surface of this coating, as observed with X-ray photoelectron spectroscopy (XPS), differs from the bulk, indicating that there is a change in the deposited species during the course of the reaction. The results for HMX and RDX also reveal that there is little mineralization and significant deposition of material on the catalyst under anaerobic conditions when there is no added reducing agent. Soluble nitrogen species were identified by ion chromatographic analysis, (Figure 2.9b).

The reduction with EDTA was investigated in more detail in a recirculating reactor. Reaction at an initial pH = 6.85 ("natural" pH of TNT solution) resulted in about 65% destruction of TNT and formation of four major intermediates or by-products. The latter are identified by HPLC retention time in Figure 2.15. Buffering with sodium bicarbonate to a pH = 7.0 decreases the rate of disappearance of TNT and resulted in different intermediates (Figure 2.16). Reaction at initial pH = 5.15 resulted in more rapid TNT destruction (Figure 2.17). At still lower pH, 3.05, removal of TNT was achieved within 40 min with the lowest concentration of intermediates (Figure 2.18). HPLC data showed three major reduction products in these experiments.

Table 2-2. Mole Balance Data Under Anaerobic Conditions With No Added Reductant

	TNT N ₂	RDX N ₂	HMX N ₂
Type of Carbon			
Solution Carbon			
HE Carbon	<1%	44.5%	17.7 %
Carbonate	<1	<1	<1
By-product(s)	32.9	45.3	28.0
Carbon on TiO ₂	62.0	2.4	1.9
Carbon Dioxide	4.0	10.2	ND*
Recovered Carbon	98.9	102.8	47.6
Type of Nitrogen			
Solution Nitrogen			
HE Nitrogen	<1	44.5	17.7
Ammonium	15.1	3.8	1.4
Nitrite	2.2	<1	0
Nitrate	<1	<1	<1
By-product nitrogen**	32.9	45.3	28
Nitrogen on TiO ₂	42.7	0	<1
Recovered N	94.2	93.9	47.1

*ND = not done

* The numbers for by-product nitrogen are assumed from the percentage of by-product carbon. All other values are experimentally determined.

experiments. These were identified by matching HPLC retention times, UV-Vis spectra, and HPLC-MS analyses. The three products were 2,4-dihydroxylamino-6-nitrotoluene, 2-hydroxylamino-4-amino-6-nitrotoluene, and 2,4-diamino-6-nitrotoluene.⁴²

Mass balance studies with carbon-14 labeled TNT, [U-ring, ¹⁴C]-TNT, were done on EDTA experiments to determine how well transformation products could be recovered and detected. Results showed that most of the products were soluble but a significant amount remained on the reactor and had to be washed off with NaOH or water (Figure 2.19). The solubility of reaction products when EDTA is used is in contrast to the degree to which insoluble material is deposited on the TiO₂ when no reducing agent is present in the reaction mixture, which was described above. HPLC analysis of the soluble radioactivity showed the three products mentioned above accounted for the majority of eluted radioactivity (Figure 2.20) but only about 55% of the injected radioactivity was recovered in eluent. The hydroxylamine products, if formed, apparently have a propensity to bind to the

HPLC column, or other products were formed which are not amenable to this type of analysis. Silylation and (GC/MS) did not reveal any other products and in fact was not successful in determining the hydroxylamines. The amount of soluble radioactivity decreased during the time of incubation in air (Figure 2.21) indicating polymerization or other reaction that produces insoluble products. ¹⁴C-labeled TNT was used to confirm that at acidic pH conditions most of the carbon was present as soluble compounds. At basic pH, azoxy coupling to give dimers occurred to a greater extent. The reaction of 4-amino-2,6-dinitrotoluene under photocatalytic conditions was tested. This compound was reduced to 2,4-diamino-6-nitrotoluene, which was stable under the reaction conditions indicating that it is a practical end point of reduction (Figure 2.22).⁴¹

Inorganic reductants and additional organic reductants were investigated in experiments at NREL in the recirculating annular reactor. The results of these tests are shown in Figure 2.23a. The data for EDTA under these conditions is also included as a reference point for comparison with the results described above. The rate of TNT removal with formate was nearly the same as for EDTA. Glucose removed TNT at a rate similar to that of the inorganic reductants. Sulfite and hypophosphite were active only when light and catalyst were present and dramatically increased the reductive degradation rate of TNT (Figure 2.23b). The experiment with sodium borohydride (NaBH₄), a strong reducing agent, is notable. TNT was rapidly removed upon addition of NaBH₄, before the lamp was turned on (Figure 2.23a). Subsequent illumination decreased TNT concentration only very slight further, and did not discernibly decrease TOC.

Other experiments examined the light intensity dependence of the rate of photocatalytic conversion of TNT. The pot reactor described above was used, and neutral density filters were inserted between the reactor and the light source to attenuate the incident light. In the absence of any organic reductant, the photocatalytic conversion rate of TNT shows linear behavior with respect to incident light intensity under both aerobic and anaerobic conditions (Figure 2.24a). When the hole scavenger EDTA is added under anaerobic conditions, not only does the rate increase dramatically, but it exhibits half-order dependence on light intensity (Figure 2.24b). We believe this may indicate that when the rate of the chemical reactions are slow, photo-generated holes and electrons are consumed by chemical reactions in proportion to their concentration, and that physical processes within the catalyst that govern the rate of recombination are not factors. When the chemistry is accelerated by adding hole scavengers, the competitive effect of exciton recombination becomes evident.

Experiments using an initial anaerobic photocatalytic step followed by oxidative conditions in TNT wastewater show this to be an effective way to clean the catalyst that has been coated in the reductive reaction, while producing fewer intermediate species than direct oxidation of TNT, Figure 2.25. This procedure does have the drawback of increasing the reaction time, but it may prove useful in decreasing the overall toxicity of the treated water.

2.2.3.2 Discussion

The initial rationale for testing the photocatalytic reaction of TNT under anaerobic conditions was based on the knowledge that the molecular structures of HE compounds are designed to include both oxidizing and reducing functionalities. HE compounds are designed to self destruct and the challenge is to bring this about under mild conditions. TNT is reactive under conditions in which no other oxidizer or reducing agent is present as shown in Figure 2.14 (curve for no added electron scavenger). Unfortunately, very little of the carbon content is converted to CO₂ under these conditions. The majority, about 62%, of the TNT is converted to forms which are deposited on the catalyst surface and most of the balance is in the form of soluble by-products (Table 2-2). Subsequent photocatalytic reaction of the mixture under aerobic conditions can accomplish some cleaning of the catalyst.

Table 2-3. Electrical power requirements for photocatalytic oxidation and reduction processes estimated from the results obtained in this project.

Reactor and Conditions	Time and power needed to reduce TNT by one decade		kWh/1000 gal of pink water
	min	Wh/L	
Batch (1000 W; 0.4 L)			
Reduction-EDTA	10	113	428
Reduction	230	417	1578
Oxidation	145	6014	22,868
Recirculating (15 W; 1 L)			
Reduction-EDTA	50	12.5	47
Reduction	600	150	568
Oxidation	453	9583	36.276

Equation = (Electrical Power)*(Time to destroy one decade of TNT in hrs) / (Volume of reaction solution, in liters), this is then converted to kWh/1000 gal. A cost can be calculated by assuming a cost for electricity.

Table 2-4. Apparent Photoefficiency (mole TNT eliminated/mole photons supplied) for photocatalytic oxidation and reduction processes

Lamp and Conditions	Time needed to reduce TNT by one decade (min)	UV photons available (Ei)	Mol of TNT/ Mol of photons
Oriol (9.5098E-04 Ei/min; 1.43E-04 mol TNT)			
Reduction-EDTA	10	0.00950	0.01503
Reduction	230	0.21872	0.00065
Oxidation	145	0.13789	0.00103
Black (2.7441E-04 Ei/min; 3.56E-04 mol TNT)			
Reduction-EDTA	50	0.01372	0.02594
Reduction	600	0.16464	0.00216
NREL (2.7441E-04 Ei/min; 3.56E-04 mol TNT)			
Oxidation	450	0.12348	0.00288

Equation = (Flux in Ei/min from actinometry)*(Time to degrade unit amount, 81 ppm, in minutes)/(Moles of TNT in unit amount, 81ppm*volume of reactor)

requirements are based on the use of fluorescent black lamps that have the same electrical to near-UV light output as the lamps used in this study (15:1.1 W) or on the Hg-Xe arc lamp (1000:4 W). Each decade of destruction, 100 to 10 ppm, 10 to 1 ppm, etc., would require the same amount of time to achieve under conditions which fit the assumptions. There are a number of variables that are included to determine the power requirements for photocatalysis. One method is to determine the electrical power required to create one decade of reduction (from 100%-10%) in the target contaminant in one liter. This methodology provides the results summarized in Table 2-3.

A second method is to determine the efficiency of the reaction by measuring the UV photons available expressed in Einsteins (Ei), from actinometry measurements, per molecule of compound degraded, and comparing this with the concentration decrease of the target species per unit time (100-10% is 90 ppm-9 ppm, or 81 ppm of TNT destroyed). This generates a unit-less efficiency, Table 2-4.

G. Figures

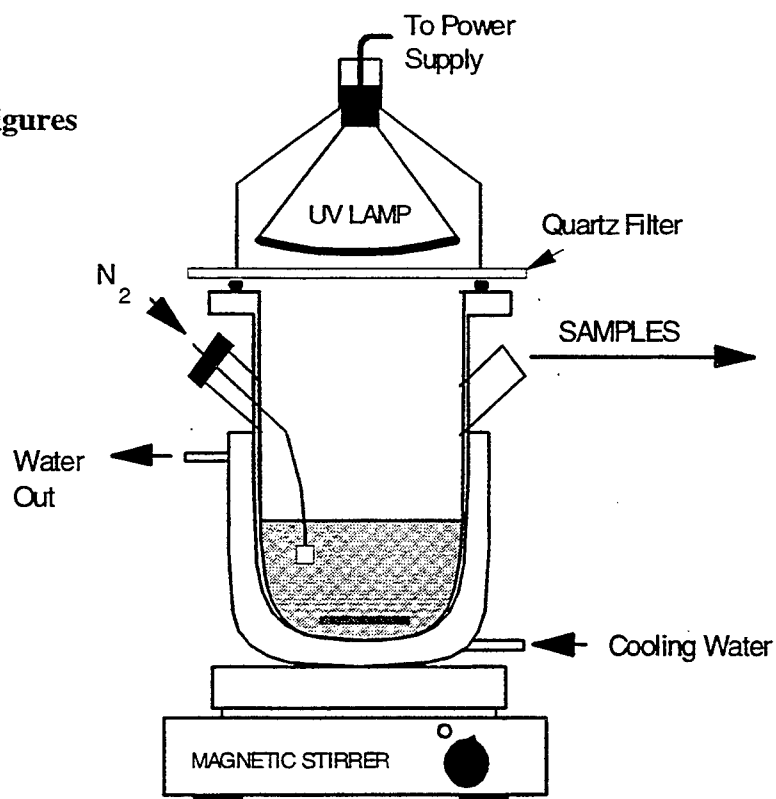


Figure 2.2. Batch photocatalytic reactor system.

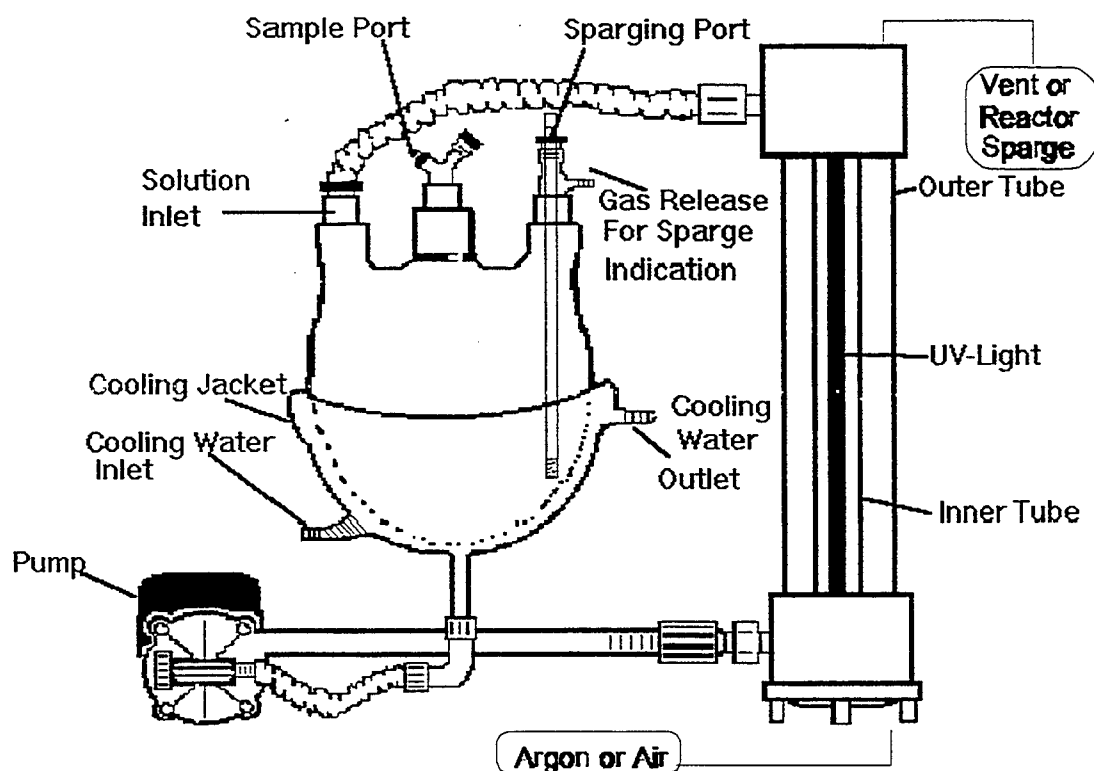


Figure 2.3. Recirculating photochemical reactor system.

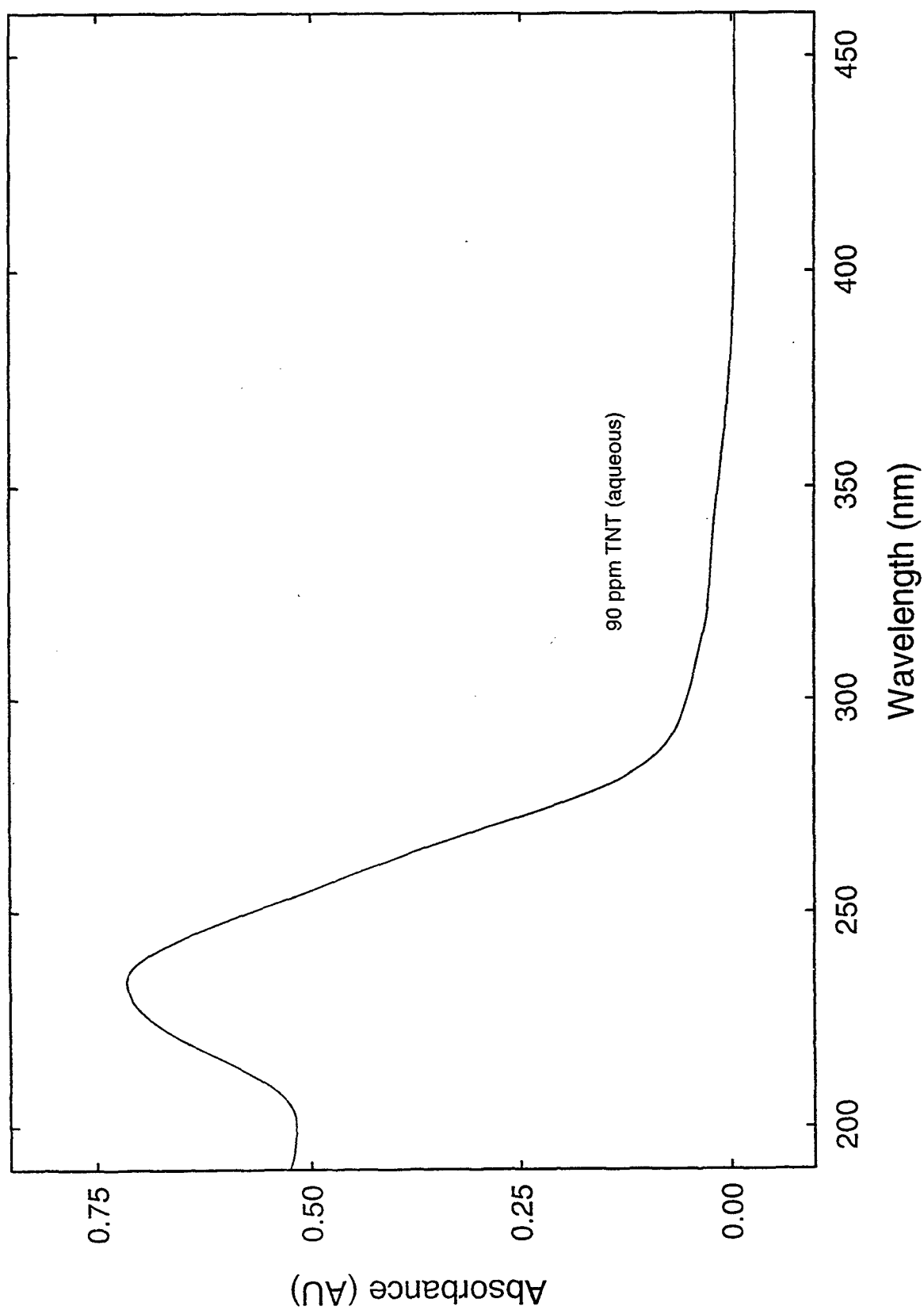


Figure 2.4. UV-Visible absorption spectrum of TNT.

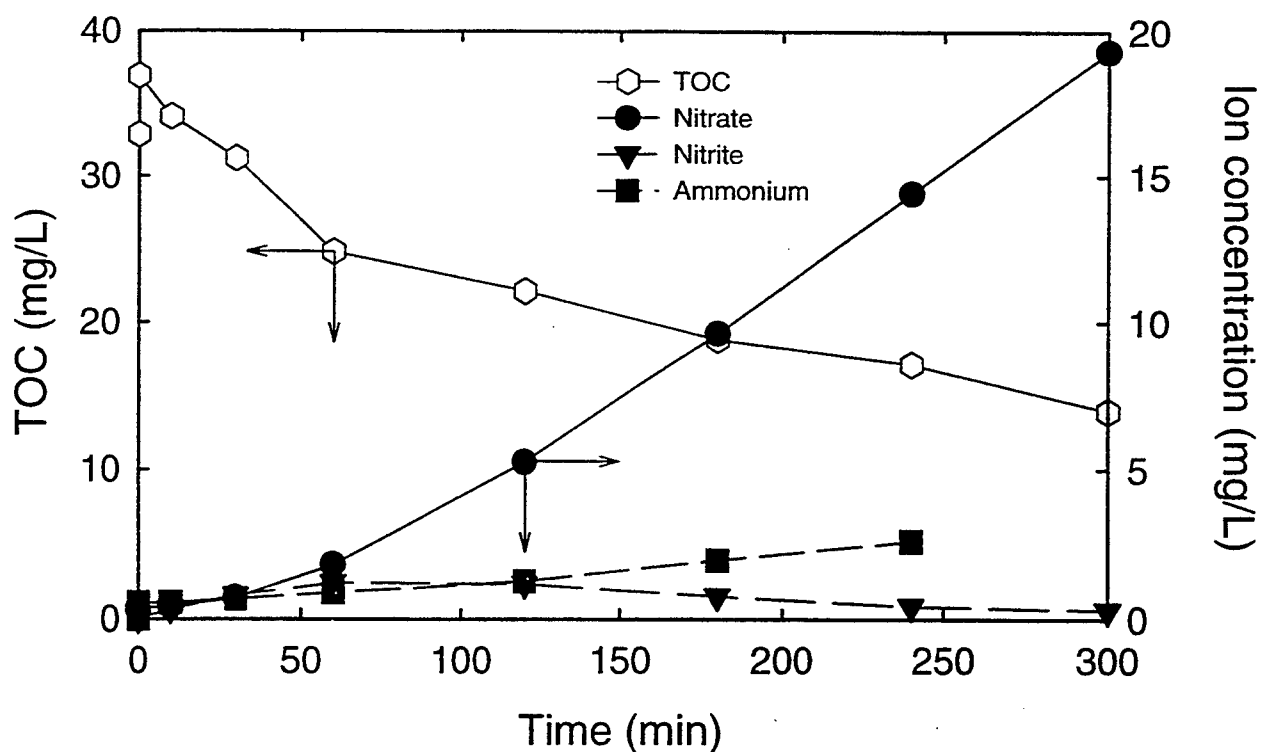


Figure 2.9a. TOC and inorganic nitrogen species formed during oxidative destruction of TNT in the batch reactor.

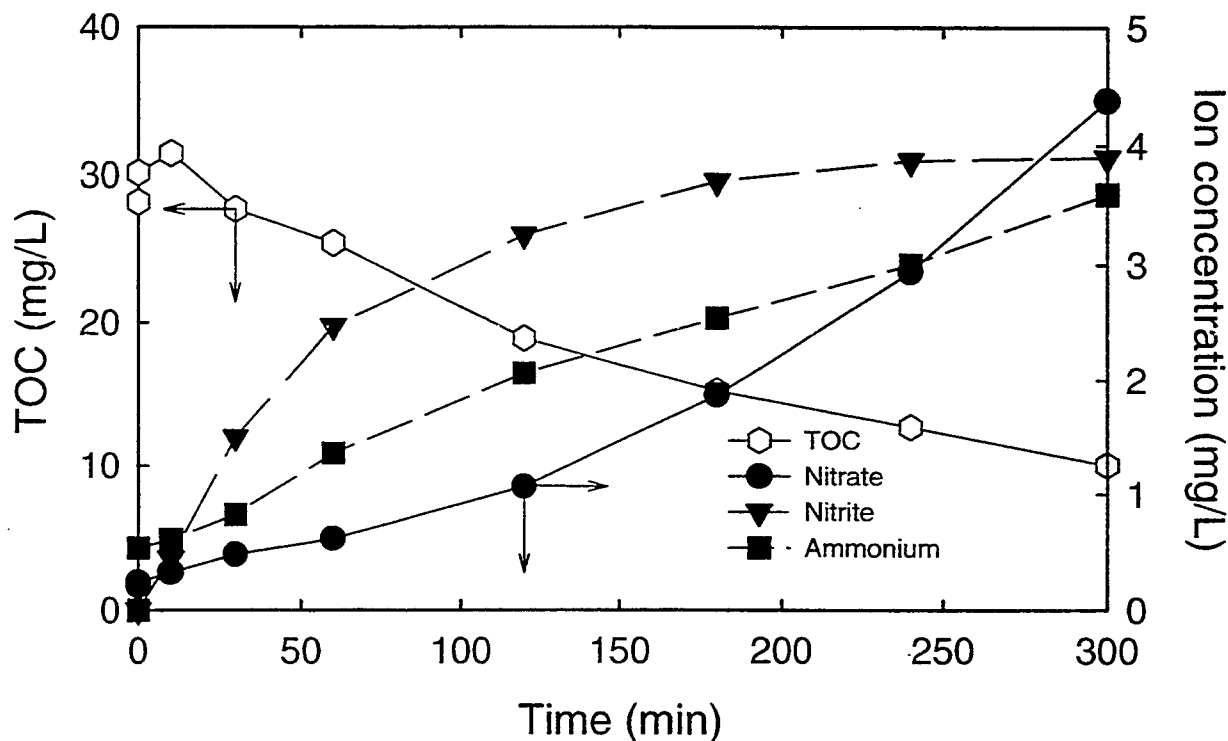


Figure 2.9b. TOC and inorganic nitrogen species formed during reductive destruction of TNT in the batch reactor.

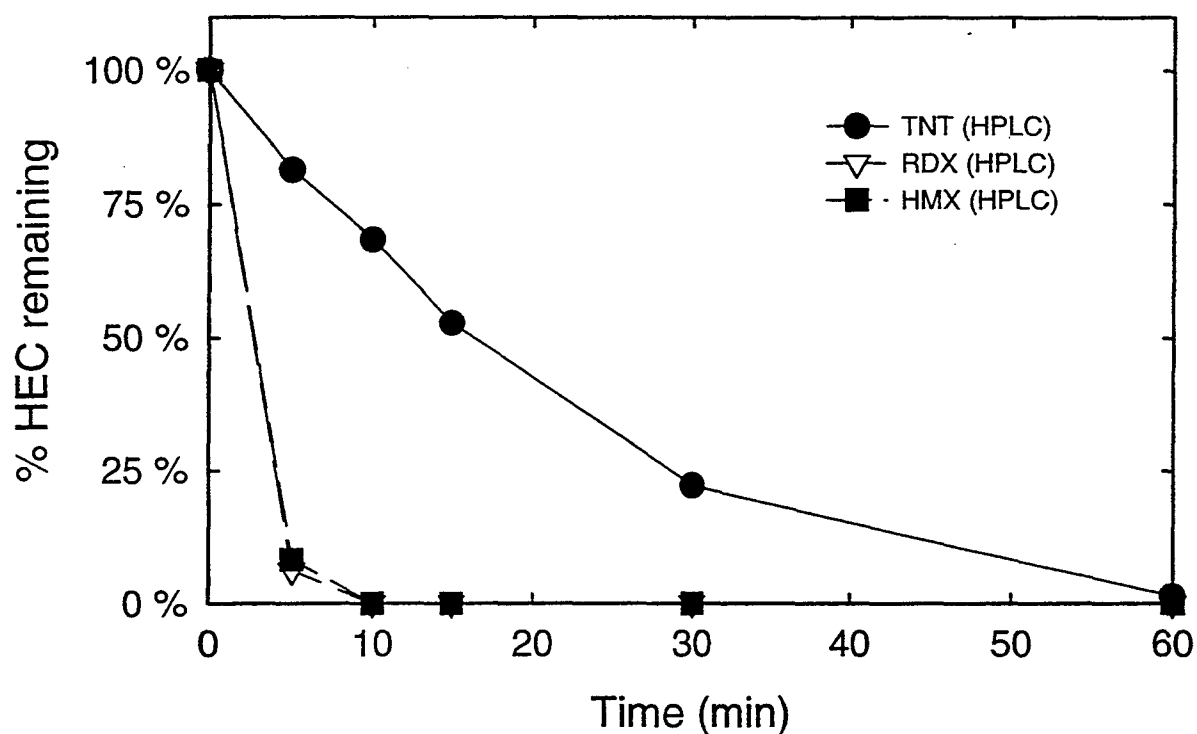


Figure 2.10 HPLC of HE destruction using 180-254 nm ozone generating spa lamp with TiO_2 in the recirculating reactor.

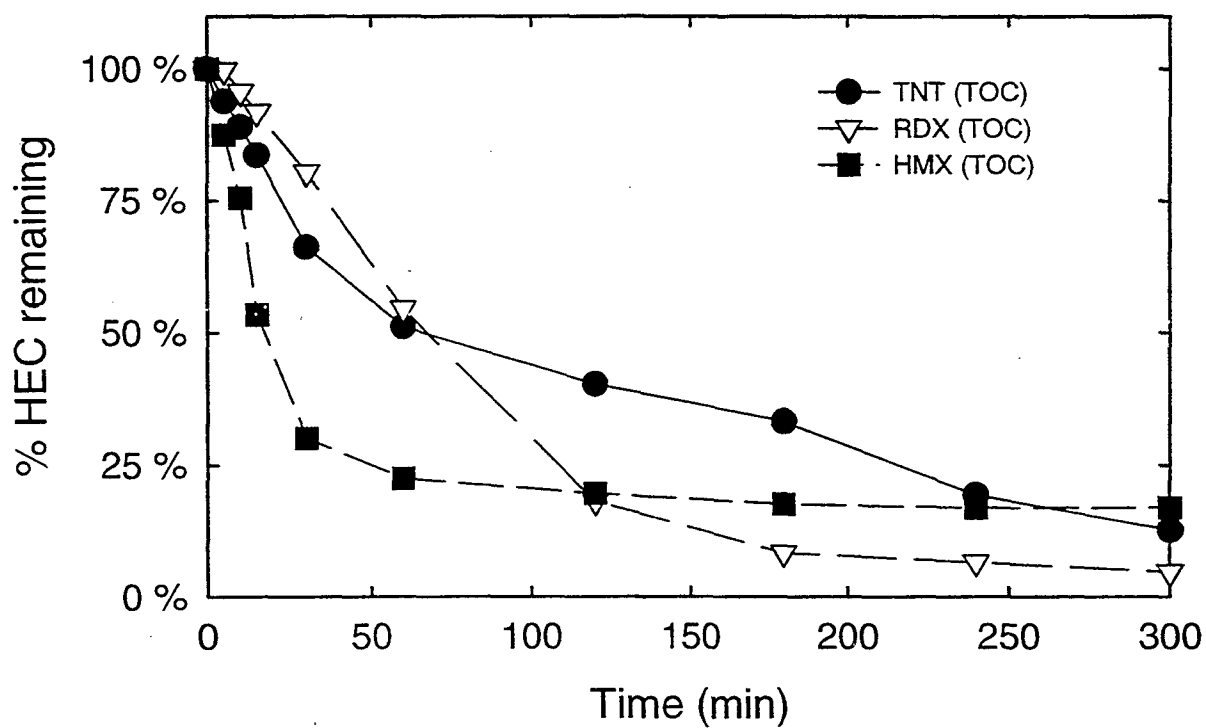


Figure 2.11 TOC of HE destruction using 180-254 nm ozone generating spa lamp with TiO_2 in the recirculating reactor.

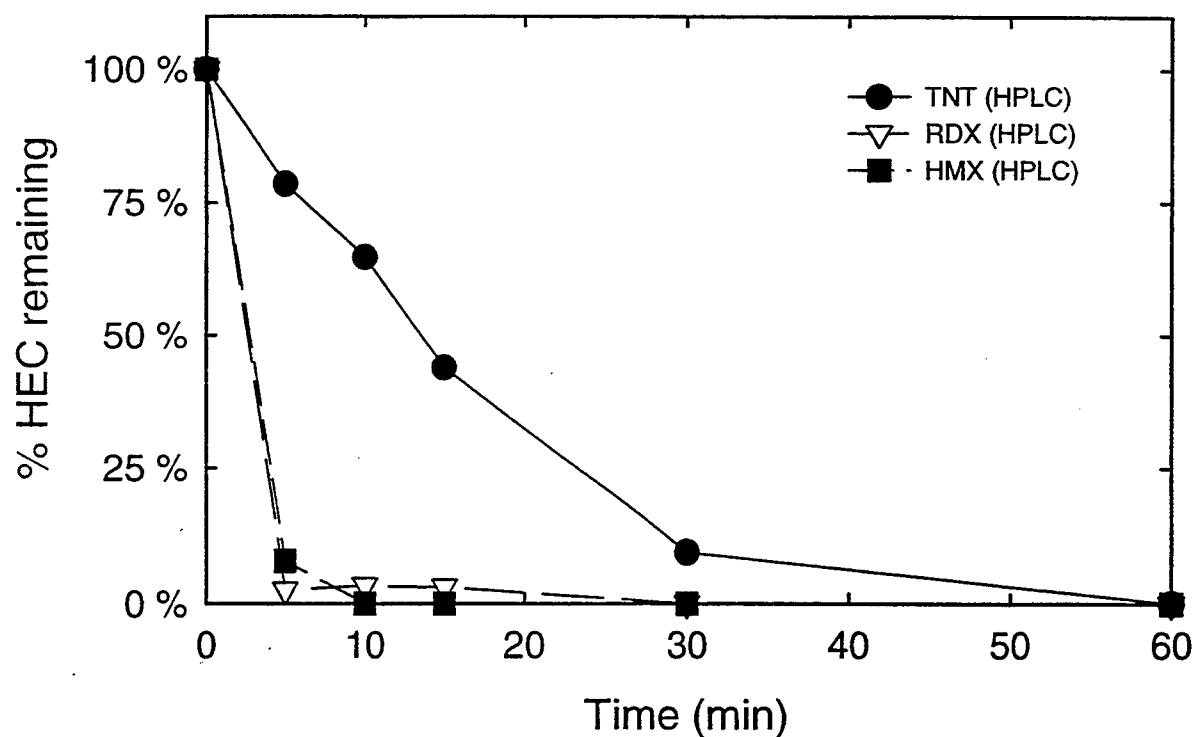


Figure 2.12 HPLC of HE destruction using 180-254 nm ozone generating spa lamp without TiO_2 in the recirculating reactor.

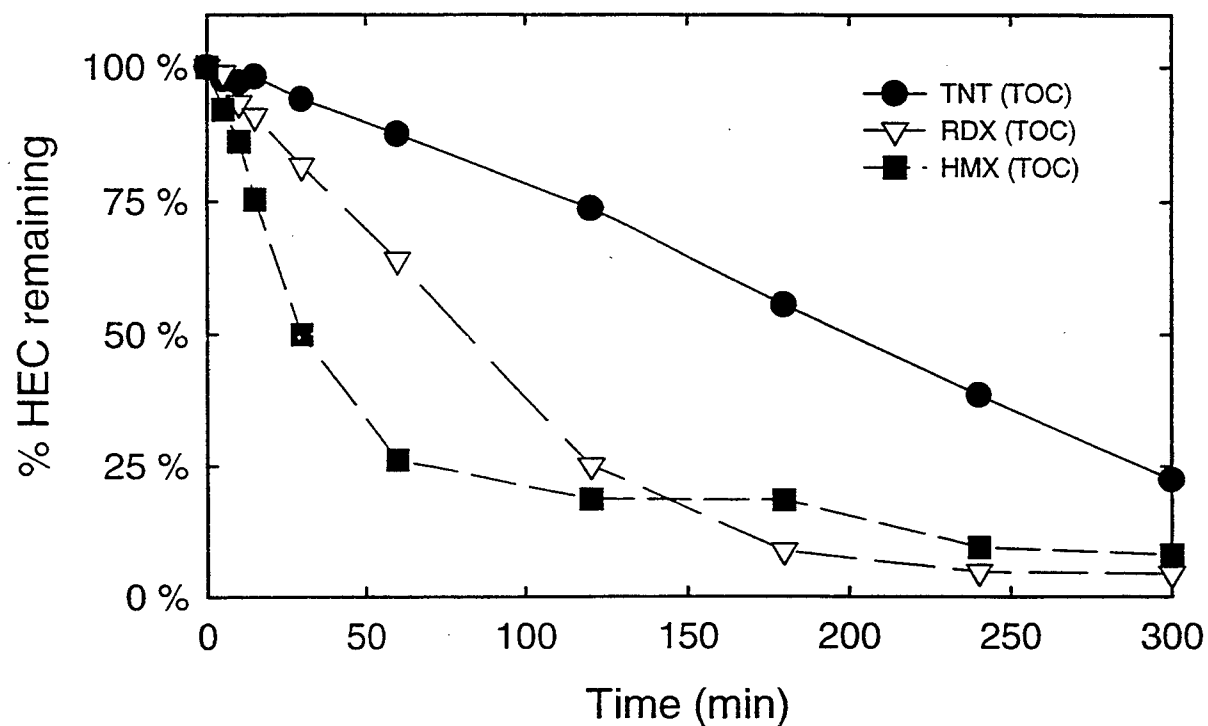


Figure 2.13 TOC of HE destruction using 180-254 nm ozone generating spa lamp without TiO_2 in the recirculating reactor.

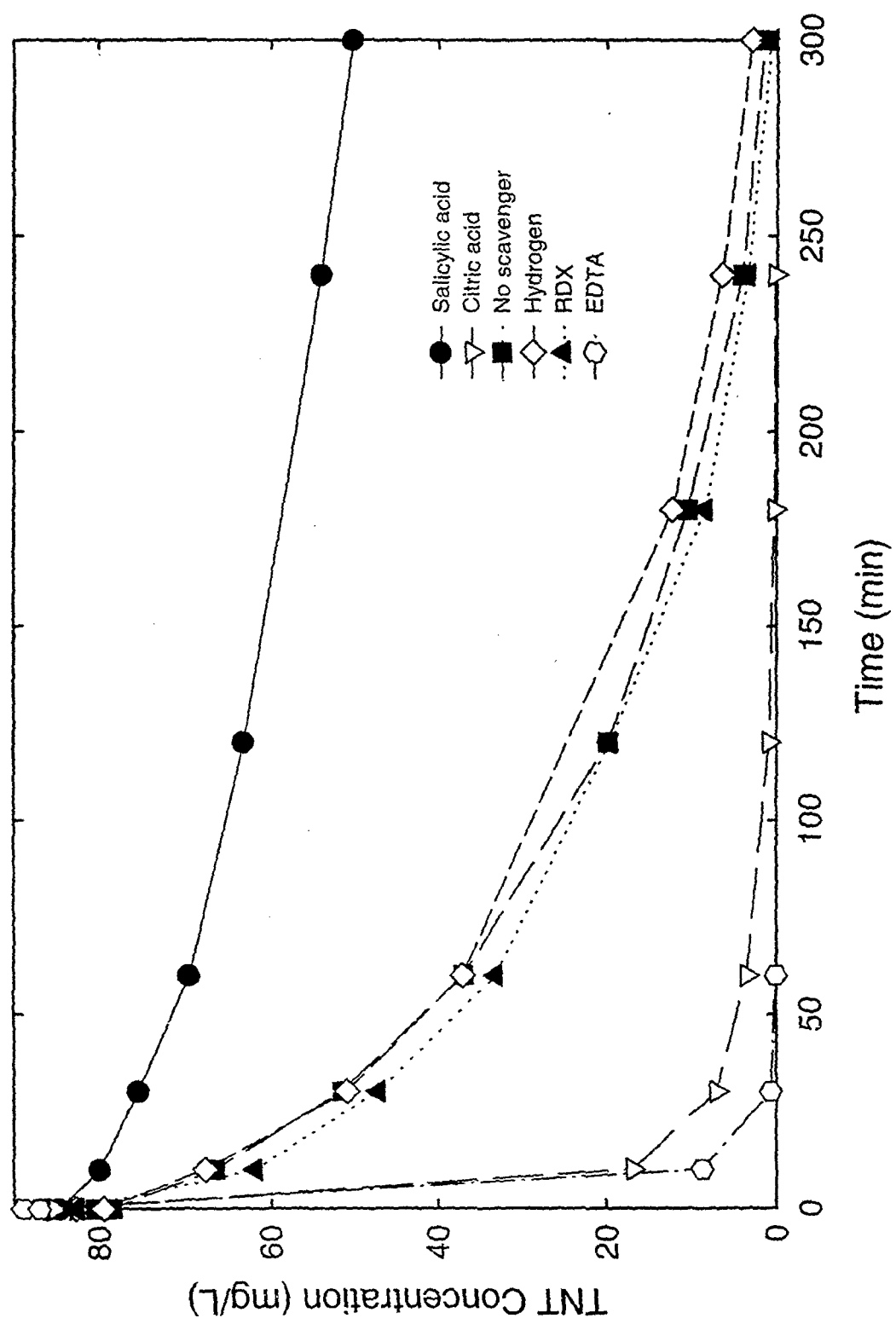


Figure 2.14. Reduction of TNT in the batch reactor with various added hole scavengers.

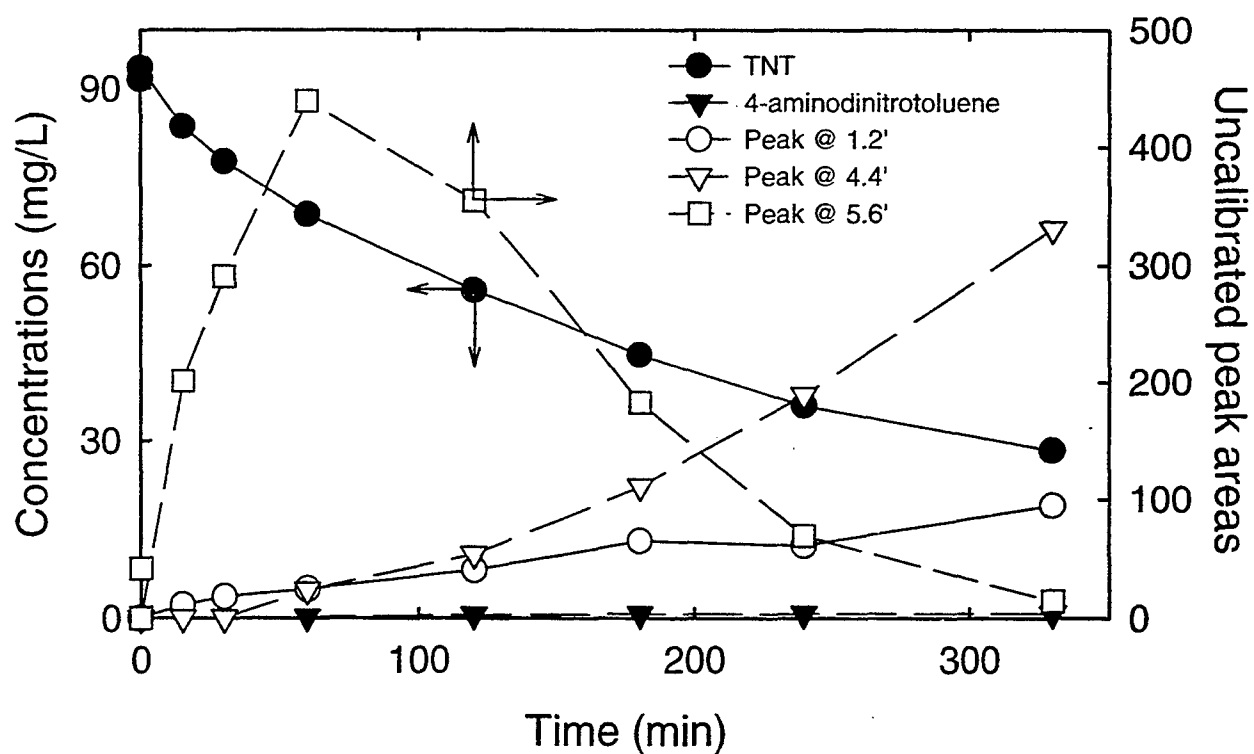


Figure 2.15. Intermediates and by-products detected in the photocatalytic reduction of TNT with EDTA at initial pH = 6.85.

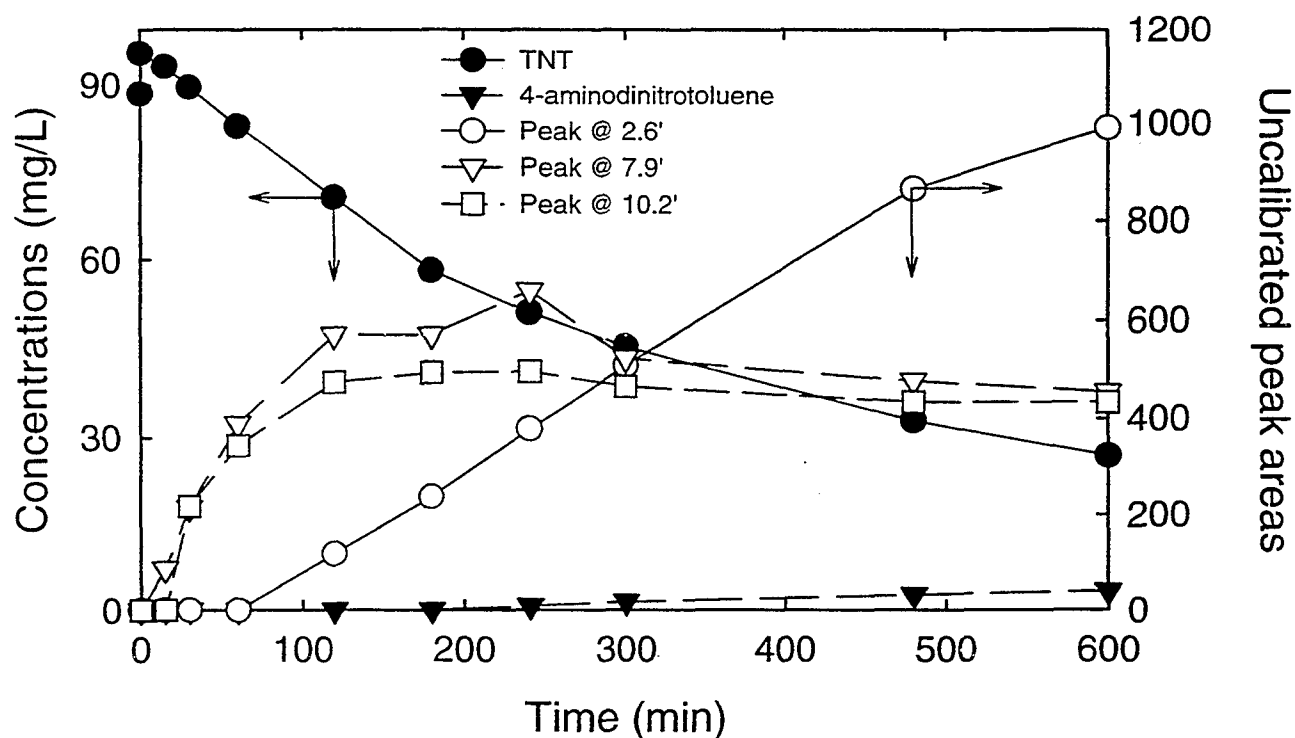


Figure 2.16. Intermediates and by-products detected in the photocatalytic reduction of TNT with EDTA held at pH = 7.0, using a bicarbonate buffer.

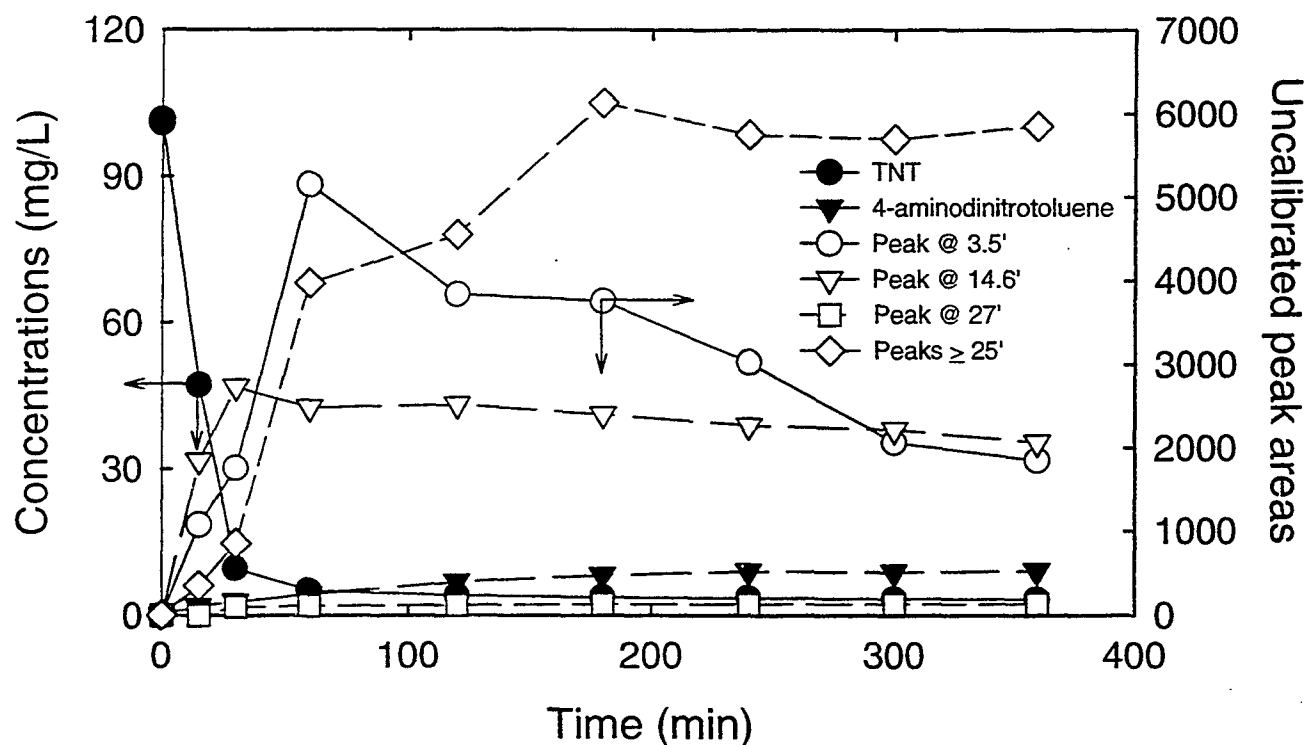


Figure 2.17. Intermediates and by-products detected in the photocatalytic reduction of TNT with EDTA at an initial pH = 5.15.

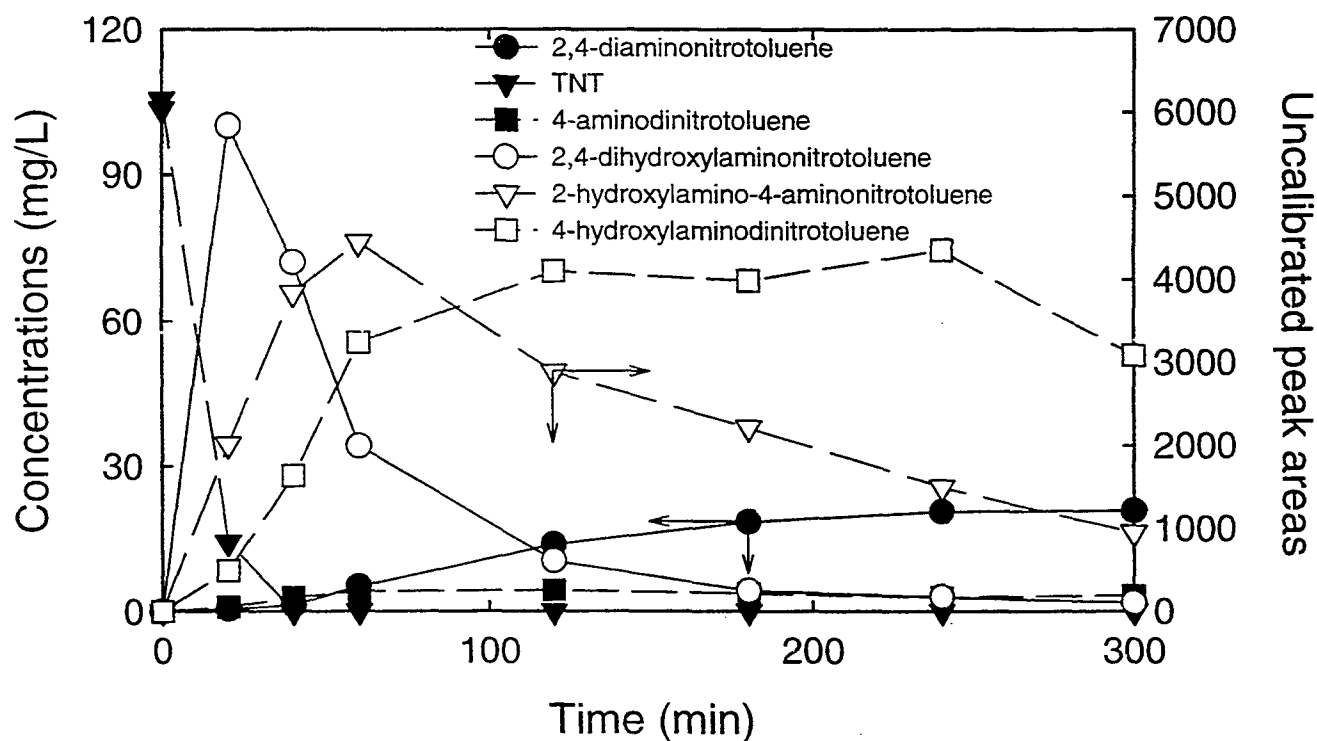


Figure 2.18. Intermediates and by-products detected in the photocatalytic reduction of TNT with EDTA at initial pH = 3.05.

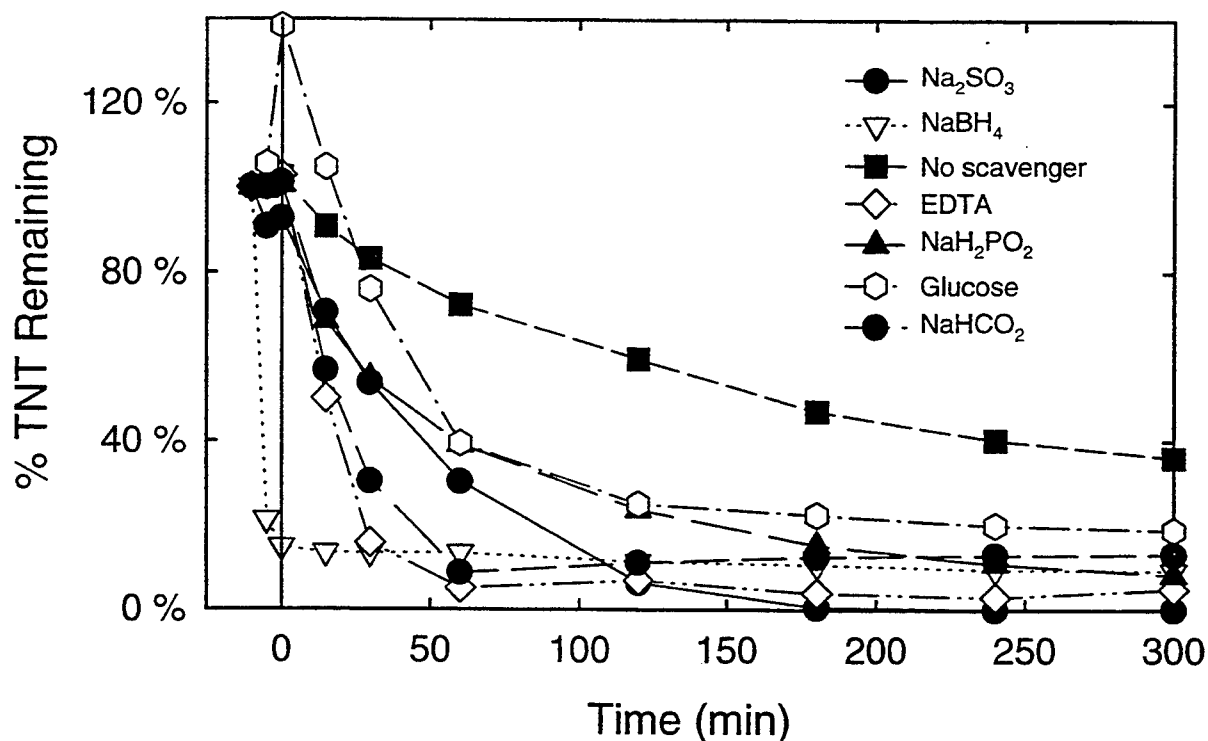


Figure 2.23a. Effect of hole scavengers upon the photocatalytic reduction of TNT in the recirculating reactor.

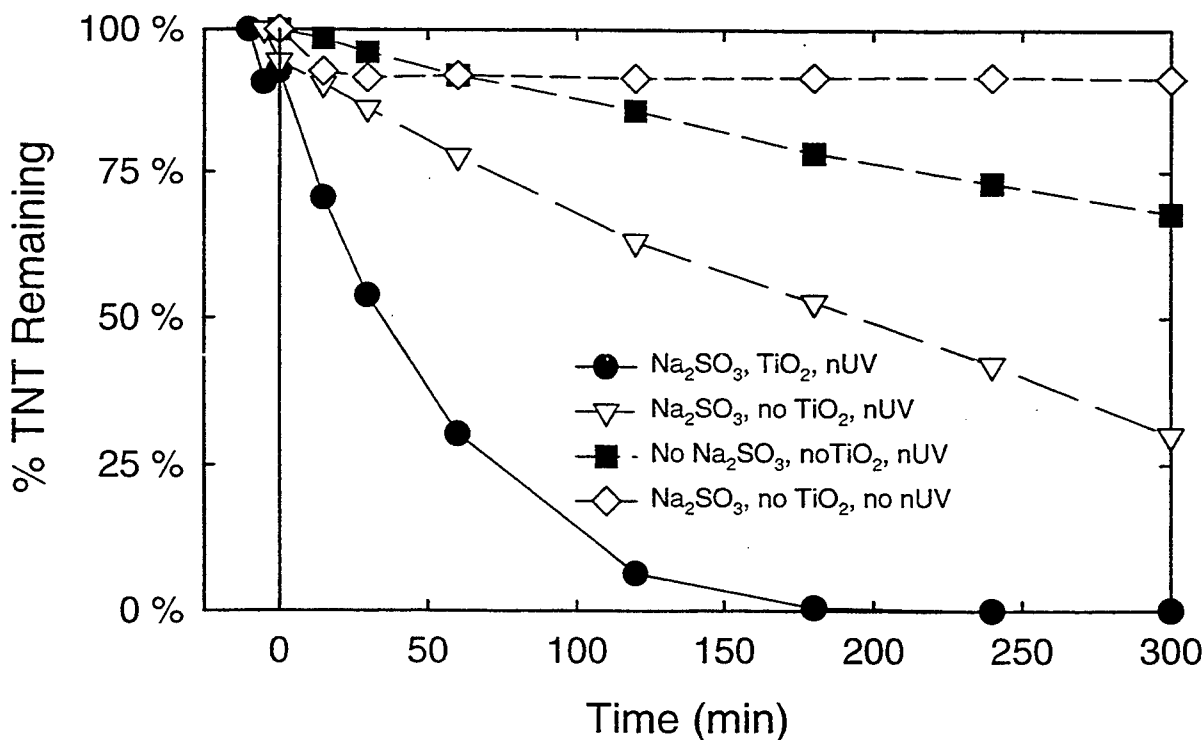


Figure 2.23b. Combined effects of sodium sulfite, TiO_2 , and UV illumination on the photocatalytic reduction of TNT in the recirculating reactor.

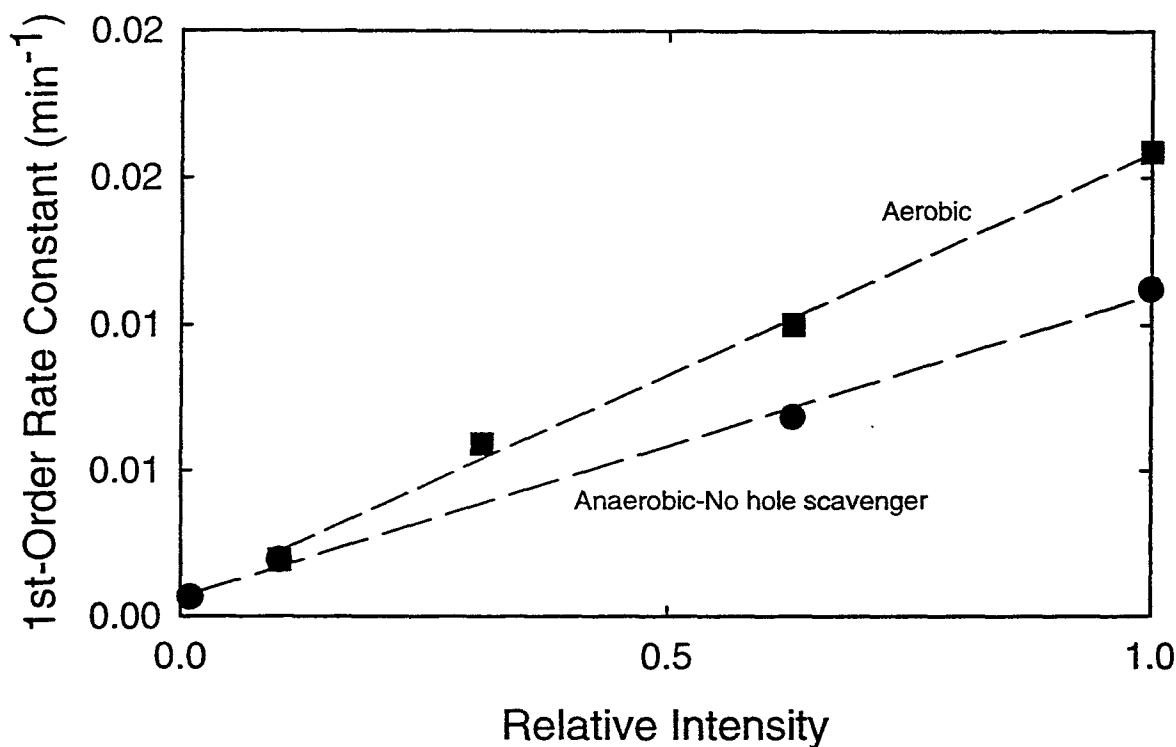


Figure 2.24a. First Order Rate Constant vs Relative Intensity for Anaerobic and Aerobic TNT Degradation (Rate constant from $\ln(c) = k t + b$)

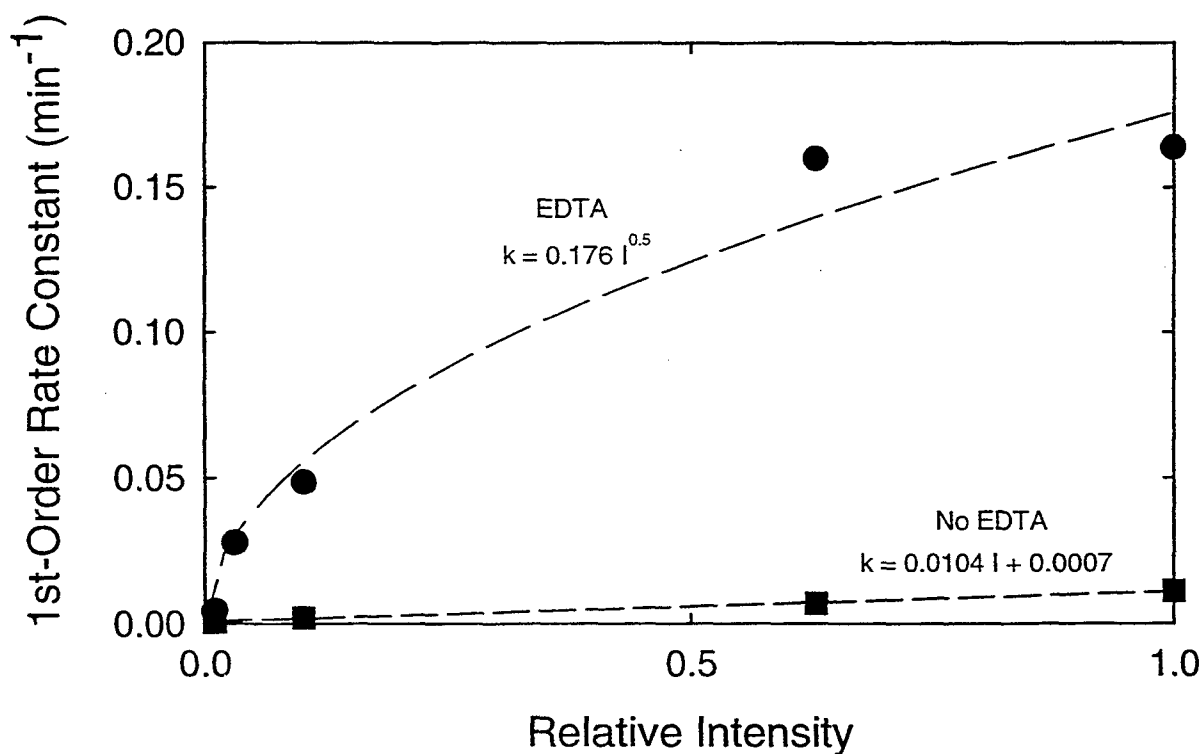


Figure 2.24b. First Order Rate Constant vs Relative Light Intensity for Anaerobic TNT Degradation with and without EDTA Present (Rate constant from $\ln(c) = k t + b$)

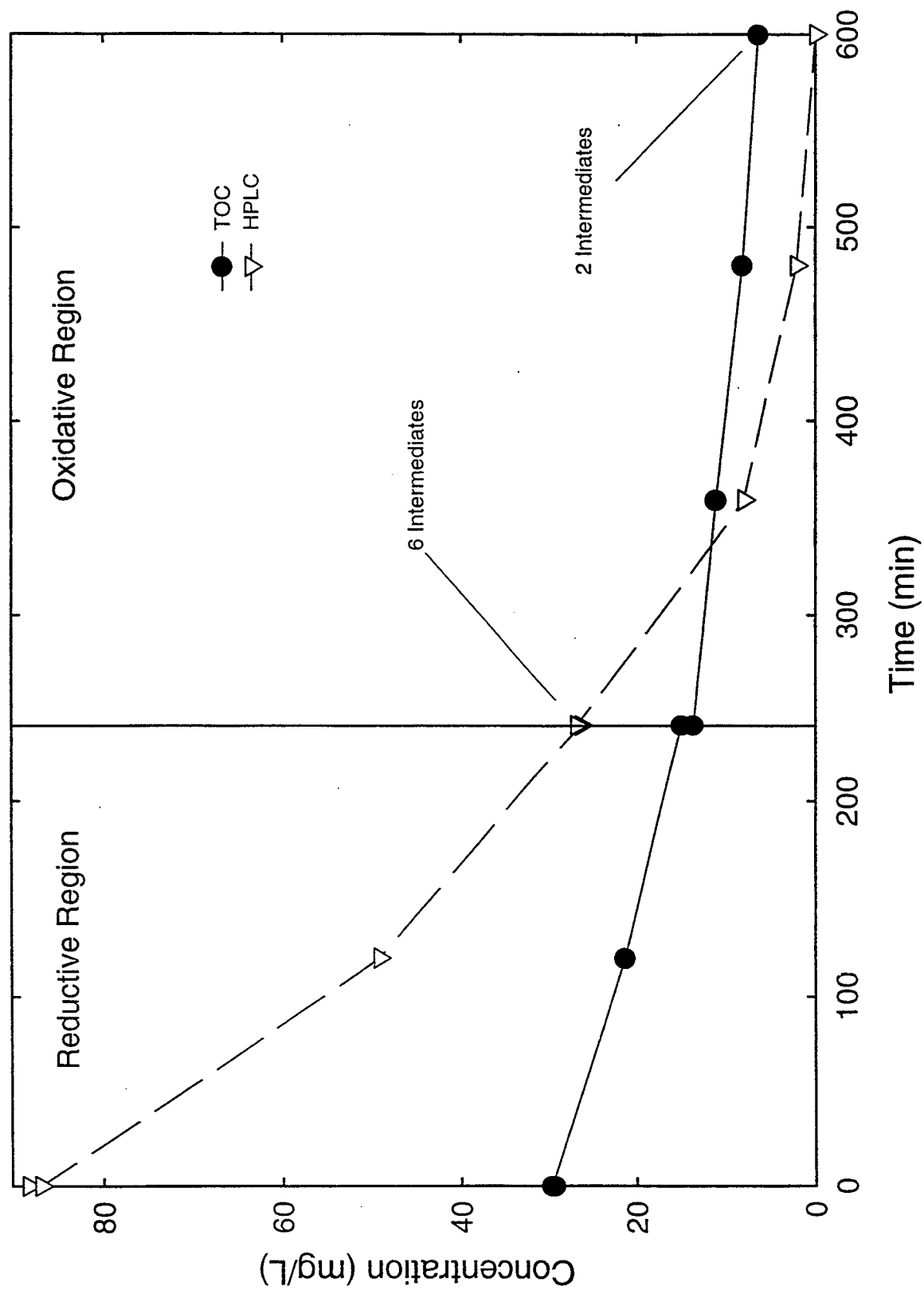


Figure 2.25. Plot of TOC and HPLC Results for Cycling of Oxidative and Reductive Conditions of TNT Degradation to Clean Catalyst Surface.

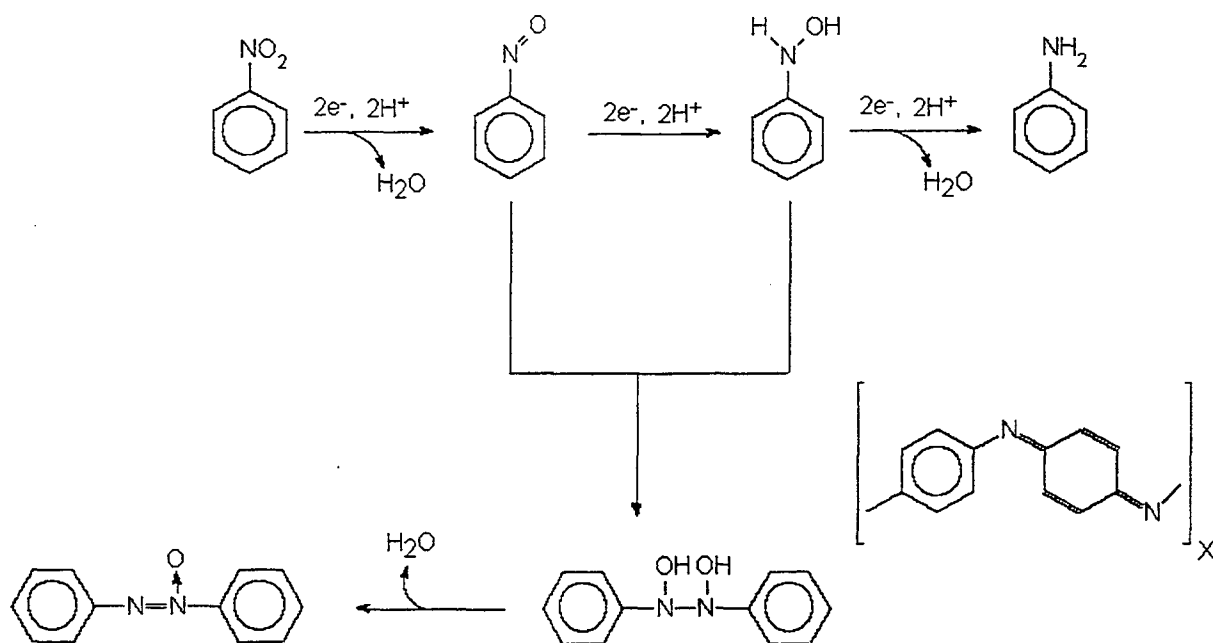


Figure 2.26. Simplified reaction scheme for reduction and coupling of the nitrogen functionality in nitroaromatic compounds.

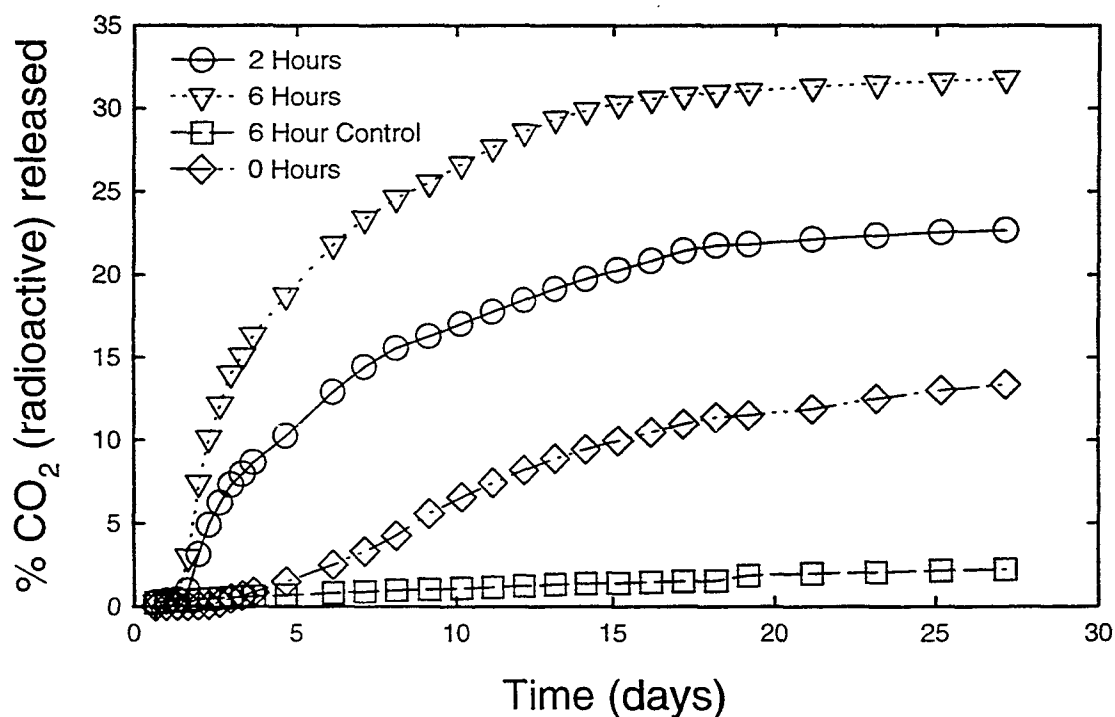


Figure 2.27. Evolution of ^{14}C - CO_2 over time from aerobic biological degradation of photocatalytic TNT degradation products by *P. cryosporium*. Graph shows four treatments: a non-inoculated control, no photocatalytic treatment (0 hours), and 2 and 6 hours of photocatalysis.

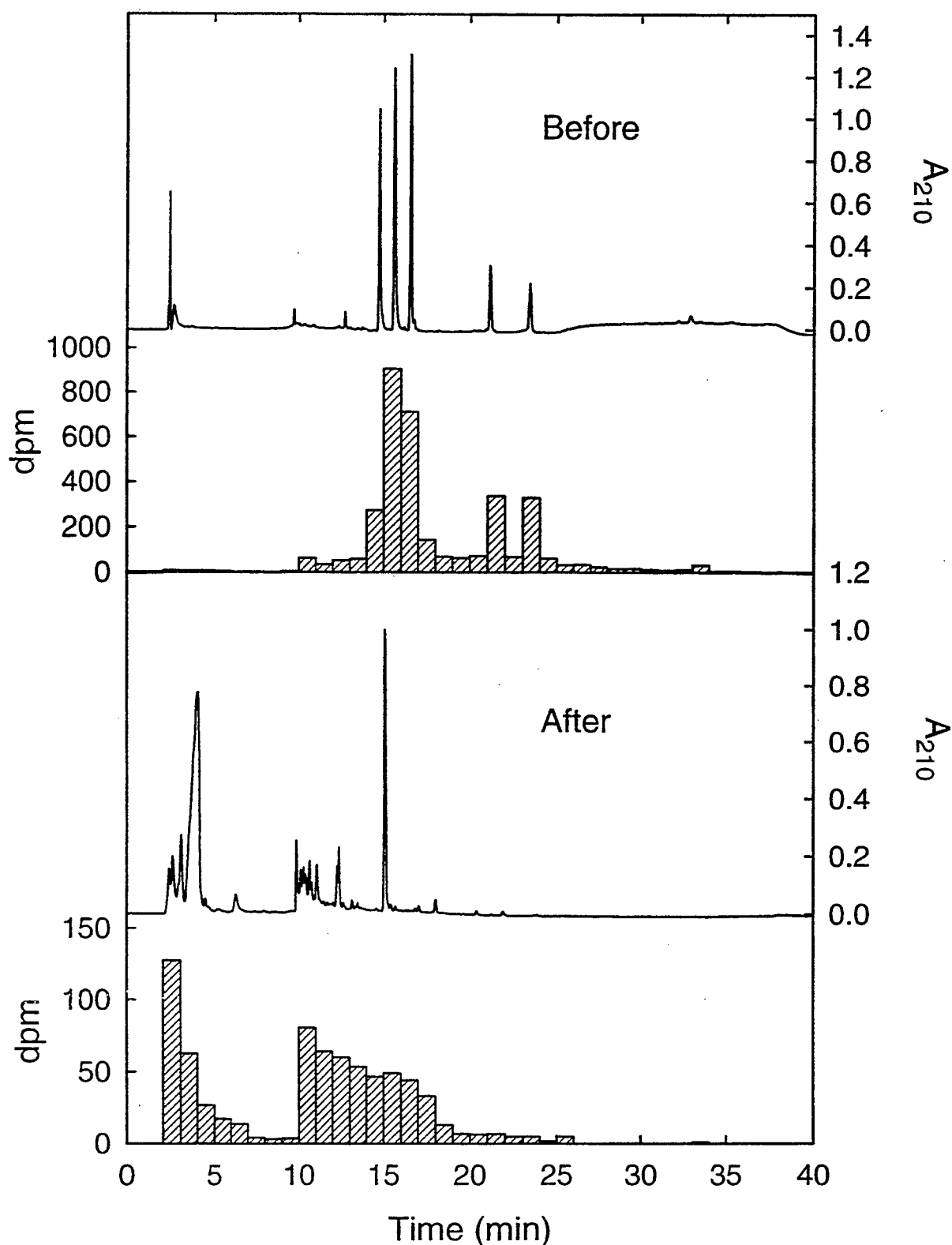


Figure 2.28. Chromatographic peaks (A_{210}) with associated ^{14}C fractionation (dpm) of photoreactor products from 2 hours of photocatalytic treatment of TNT before (top graph) and after (bottom graph) aerobic biological treatment by *P. cryosporium*.

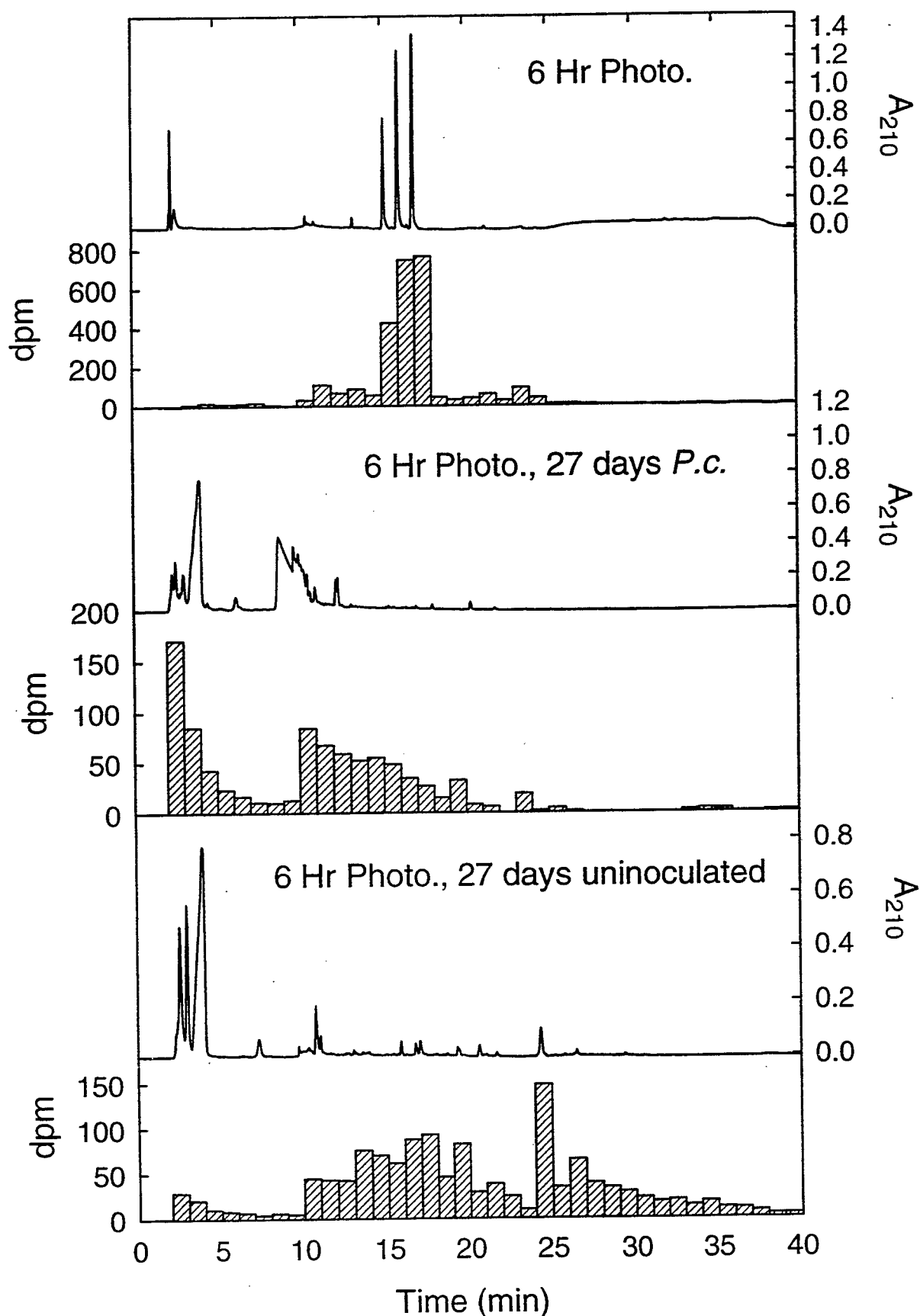


Figure 2.29. Chromatographic peaks (A_{210}) with associated ^{14}C fractionation (dpm) of photoreactor products from 6 hours of photocatalytic treatment of TNT before (top graph) and after (middle graph) aerobic biological treatment by *P. chrysosporium*. An abiotic control (bottom graph) is included for comparison.

3.0 New Process Chemistry for Removal of TNT and Other HE Compounds from Water

A new process for the removal of TNT and other HE compounds from process water or contaminated groundwater has been identified in this project. The process is based on the use of novel reduction chemistry that does not require light or other unusual conditions. The chemistry has been characterized sufficiently to bring the process to the proof-of-concept level of development. The process is sufficiently innovative to warrant maintaining it as NREL Protected Information to allow for prosecution of patent. Therefore only a general description of the performance of this process is included here.

Figure 3.1 shows the removal of TNT from water by the process as a function of time using four variations of the reagents appropriate to the process. These reagents are identified in the figure as R1 through R4. The TNT content is reduced below the detection limit, 20 ppb, by the process. The performance of the process is relatively insensitive to the pH and presence of ambient air but is accelerated by increasing temperature. Not all intermediates and by-products of the reaction chemistry have been identified. For this reason it is not known with certainty if the treated water will be without hazard. Additional work will determine if hazardous products remain in the water. If so, there are a number of options for removing them with little alteration of the process.

An economic analysis based on the current level of understanding indicates that the process will be less expensive to operate than the GAC systems now used to treat pink water at AAPs. The reductant is readily available in bulk quantities at low cost and neither it nor its reaction products are hazardous. The capital cost is expected to be similar to that for those systems. This new process has the potential to be a major improvement over the GAC system and will not require significant changes in process operations.

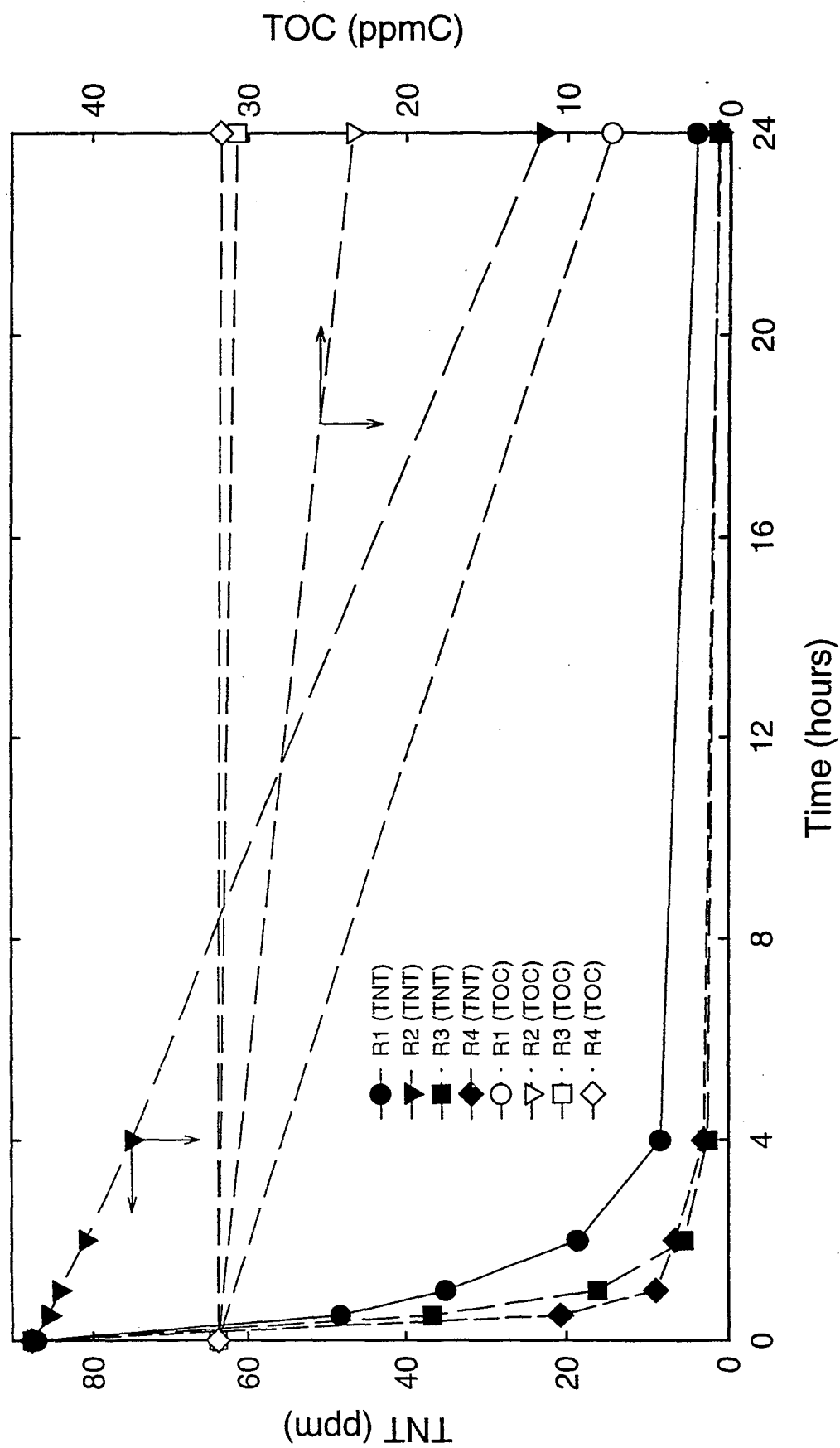


Figure 3.1. Effect of reagent type on the removal of TNT and TOC.

4.0 Conclusions

Photocatalytic treatment of pink water under oxidative conditions can mineralize TNT, HMX, and RDX but requires long residence times which translates to large power consumption by the process.

Photocatalytic treatment of pink water under reducing conditions will convert TNT to other substances. The most effective reducing agents are organic compounds such as EDTA, citric acid, glucose, and sodium formate. The power requirement and the reducing agent would contribute significantly to the process cost. However, the power required for the photocatalytic reduction process is significantly less than that for the oxidative process.

Biological treatment using *Phanerochaete cryosporium* following reductive treatment of pink water with EDTA resulted in partial mineralization of TNT. The degree of conversion of carbon content to CO₂ in the biological treatment step was dependent on the length of the photocatalytic pretreatment. The maximum extent was 32% which was achieved using a feed solution that had been exposed to 6 h of photocatalytic pretreatment.

This innovative new process for treating pink water is an effective method to reduce the level of TNT in pink water to below the detection limit (20 ppb). The rates of removal of HMX and RDX are lower than that for TNT under the treatment conditions that have been developed to date. It is possible that different treatment conditions will improve the rate of destruction of HMX and RDX. Parties interested in learning more about this process may contact the authors of this report at NREL.

The annual operating costs determined at this early stage of development for the new process process is \$179,000 for a 150 gpm flow of pink water. This is very favorable when compared to \$323,500 estimated for the carbon adsorption process. The capital costs are believed to be similar for the two processes.

5.0 Recommendations

The new process developed in this project has the potential to be a simple and low-cost method of treating pink water from LAP and demilitarization operations or for treatment of contaminated ground water. In order to develop this potential it is necessary to address gaps in the current understanding of the process. The following are recommended to provide a complete basis for assessing the process:

The nature and toxicity of the products from treatment of TNT in the new process must be determined.

The distribution of the organic and inorganic content of the treated water between solid and aqueous phases must be determined.

The amount of reagents assumed to be required, in order to estimate the cost of the process, was very conservative. Actual reagent consumption must be determined for the optimum process.

The reactivity of HMX and RDX must be improved for the process to be practical for pink water treatment. This may not be critical in the groundwater treatment application where HMX and RDX are not key factors.

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7.0 Appendices

- A. Report on Subcontract No. XCK-5-15163-01, "Combined Photocatalytic and Biological Treatment of Pink Water," University of Idaho, Moscow, ID, July 1996.**

FINAL REPORT

SUBCONTRACT NO. XCK-5-15163-01

Combined Photocatalytic and Biological Treatment of Pinkwater

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Abstract

A recirculating batch, annular photoreactor in combination with cultures of the ligninolytic fungus, *P. chrysosporium*, was used to investigate degradation and ultimate mineralization of aqueous solutions of 2,4,6-trinitrotoluene (TNT), a "pink water" constituent. The photocatalytic pretreatments of TNT were performed under anaerobic conditions with the addition of EDTA as a reducing agent to promote TNT reduction. Follow-on biological treatment of the photoreactor degradation products was performed under aerobic conditions using mycelial preparations of *P. chrysosporium*. Radiotracer experiments showed up to 32 percent mineralization of the parent C-TNT, depending on length of phototreatment. Longer photocatalytic treatment of the TNT resulted in a greater extent of mineralization. An early reduction intermediate in the TNT degradation pathway was shown to be inhibitory to the fungal mineralization of the photoproducts, as shown by the radiotracer studies. The combined processes of TiO_2 assisted photocatalysis and follow-on biological mineralization have the potential for being an effective technology for the remediation of pink water constituents.

INTRODUCTION

This report summarizes experimental work on Subcontract No. XCK-5-15163-01 "Combined Photocatalytic and Biological Treatment of Pinkwater" related to photocatalysis and biological treatment of 2,4,6-trinitrotoluene (TNT), a constituent of pinkwater.

METHODS AND MATERIALS

I. Photocatalytic Reactor System Description: All photocatalysis experiments for TNT degradation were performed in a batch, recirculating annular photoreactor, similar in design to reactors used by NREL scientists in their research (Blake et al., 1995). A schematic of the photocatalytic reactor system is shown in Figure 1. The reactor system consisted of a reservoir with temperature-control bath, a magnetic-drive centrifugal pump, and an annular photoreactor. The three parts were connected with 3/8" corrugated flexible Teflon tubing.

Reservoir: The reservoir was a 1L round-bottomed flask with three 24/40 ground glass-fittings at the top and one 1/2" (OD) glass tube protruding from the bottom. One of the ground-glass fittings served as the return line from the annular reactor, one held a nitrogen sparging assembly, and the third held a sample-port/thermocouple assembly. The ground-glass fitting used ACE-THRED adapters for the inlet tube, gas sparger, and thermocouple.

Centrifugal Pump: The centrifugal pump was variable speed with magnetic drive and plastic impeller. The pump had a maximum flow rate of 10 lpm providing a residence time of approximately 1.5 seconds per pass in the annular reactor.

Annular Reactor: The annular reactor design was similar in size to a reactor used by the group at NREL in their research (Blake et al., 1995). The reactor consisted of six parts. Two identical end pieces, machined out of Teflon, two glass pipes which constituted the inner and outer walls of the reactor, and two flow dispersion disks, also made out of Teflon. The two end pieces and two glass tubes were held together by compression of the O-rings. The outer glass pipe was constructed from 54 mm standard-wall Pyrex tubing. Each pipe was 10-1/2" in length and the ends lightly fire-polished. The inner glass pipe was constructed of 35 mm standard wall Pyrex or Quartz tubing 14" in length and lightly fire-polished. The choice of Pyrex or Quartz tubing was dictated by the light source. A quartz tube is necessary only when using a UV (germicidal) lamp, since Pyrex is not transparent to UV light. The Pyrex sleeve was used with the near UV (blacklight) lamp since Pyrex is transparent to these wavelengths.

The Teflon dispersion disks were used to prevent any "short-circuiting" of the reaction mixture and to guarantee a well-distributed flow through the reactor. The dispersion disks had an outer diameter of 48 mm and an inner diameter of 35 mm. The inlet disk had 24, 2.25 mm diameter holes evenly distributed around the disk (on 15° centers) providing an evenly distributed stream flow. The outlet disk had 8, 3 mm diameter holes evenly distributed around the disk (on 45° centers). The disks were held by compression between the inner and outer glass tubes of the reactor.

Light Source: The reactor assembly was illuminated by placing a 1" diameter, 18" long low-pressure mercury arc lamp in the center of the annulus. The UV lamp (General Electric 15W G15T8-BLB) emitted in a band between 300 nm and 400 nm, with a maximum wavelength of approximately 346 nm.

II. Sampling method for the reactor: Solutions were mixed and stirred in the anaerobic chamber overnight prior to placement in the photoreactor. TNT, EDTA powder, and $\text{Na}_2\text{EDTA} \cdot 2\text{H}_2\text{O}$ were all mixed and stirred thoroughly. The reactor system was assembled and checked for any leaks by running water through it. TiO_2 was added just before the reactor run. The light was allowed to warm up and the reaction samples were collected at periodic intervals of time. A circulating bath was used to monitor any change in the reaction temperature (maintained at 25° C) and a pH probe was used to check the drop in pH with reaction time.

In all the experiments, a standard procedure was followed for collecting the reactor samples through the length of the photocatalytic reaction. All samples for HPLC analysis were collected in duplicate. The samples for each time period were taken using a 5 ml-syringe out of which 2 ml of the sample was wasted, the next 2 ml collected in two different HPLC vials with the final 1 ml again wasted. The samples were immediately taken for HPLC analysis avoiding instantaneous reaction within the vials leading to errors in analysis and compound identification.

III. Analytical techniques: The primary analytical techniques used in the experiments were High Performance Liquid Chromatography (HPLC), Mass Spectrometry (MS), and the scintillation spectroscopy for the radiolabeled TNT experiments.

HPLC: The concentrations of TNT and its reduction products were determined by HPLC with a Phenomenex (Torrance, Calif.) Spherex 5-m C₁₈ reverse-phase column (250 x 2 mm) and a 5 cm x 2 mm guard column. A Hewlett-Packard model 1090A instrument equipped with a diode array UV/VIS detector and a computerized data system was used for the analyses. The column was run with a number of mobile phase programs based on the desired chromatographic separation of degradation intermediates. For example, the program employed early in the experimentation used 10% acetonitrile and 90% 41M LiOH phosphate buffer (pH 4.0) for 2 min, and then the acetonitrile was increased to 70% over the next 15 min. The acetonitrile was then increased to 100% over 1 min and maintained at that level for 2 min. The acetonitrile was then decreased back to 10% over 2 min. The solvent flow rate was 0.25 ml/min, and the column temperature was 42° C. Subsequent programs decreased the percent acetonitrile and increased the overall run length (see Conclusions and Discussion section). TNT and possible transformation products were detected by use of the diode array UV/VIS detector, recording the A₂₁₀ with continuous scanning of the absorption spectrum of each peak from 190 to 600 nm. Samples for HPLC were filtered through 0.2-μm-pore-size syringe filters (diameter, 25 mm; Supor, Gelman Sciences, Ann Arbor, Mich.) before analysis by HPLC.

Mass Spectrometry: Mass spectrometry was performed with a Hewlett Packard model 1050 HPLC with a 2.0 x 25 cm C₁₈ column (Spherex) and a 2.0 x 5 cm guard column (Phenomenex, Torrance, Calif.). Acetonitrile (A) and 15 mM ammonium acetate, pH 7.1, (B) were used as

solvents used as solvents with a flow rate of 0.3 ml/min. The gradient program was 99% B (0-5 min) to 90% B (5-7 min) to 55% B (7-27 min, held 3 min), followed by return to starting conditions (32 min). Absorbance at 220 nm as well as scanning ion monitoring with a model 5989A quadrupole mass spectrometer equipped with a model 59980B particle beam interface allowed detection of reduction compounds of interest. An injection volume of 50 μ l was used in order to direct enough material to the mass spectrometer for these analyses.

IV. Sample holding for the HPLC measurements: Early in the experimentation it was observed that spontaneous reaction and precipitation of contents was occurring in HPLC sample vials when held overnight or under refrigeration. A procedure was developed where the HPLC system was coordinated with sampling such that samples were collected and analyzed within 30 minutes.

V. Fungal growth media composition: Spores of *P. chrysosporium* were cultivated from 3% malt extract, 0.3% yeast extract, 1.5% agar medium in tissue culture flasks incubated for approximately 30 days. Suspensions of spores were made using sterile 0.2% Tween 80 and filtered through sterile glass wool. Spore suspensions were stored at 4° C. Composition of one liter of media for cultivating the mycelia of *P. chrysosporium* (modified from Bonnamy et al., 1991) was the following:

- Mineral salts - 100 ml
- Mn (0.1 M) solution - 1.8 ml
- Glucose - 50 ml
- DMS (pH 4.55) - 10 ml
- Diammonia Tartarate - 5 ml.

These ingredients were added to 830 ml of sterile water and inoculated using 3 ml of spore suspension. The inoculated solution was placed in a shake flask and maintained in an incubator at 150 rpm at 37° C.

VI. Fungal growth methods: A 500-ml portion of the above mentioned growth media was placed in a 2L shake flask and 1.5 ml of the fungal culture was aseptically added. The flask contents, with sterile cotton stopper, were maintained in an incubator at 37° C and 150 rpm. The culture flasks were ready to be used as mycelia inoculants 2-3 days after placement in the incubator.

VII. Fungal inoculation into photoreactor products: Fungal cultures retrieved from the incubator were allowed to cool for 30 min before inoculation into the photodegradation products solutions. Equal numbers of fungal mycelial masses were added to each flask of photoreactor degradation products using a sterile pipette in a laminar flow hood. Inoculation occurred immediately after opening the photoreactor, thus providing little time for any dissolved nitrogen to exchange to the atmosphere.

VIII. 14 C experiments- 40 CFR 796.310 Procedure: This experimental procedure was used to quantitate 14 C-CO₂ evolution and a 14 C balance during mineralization of 14 C-TNT by *P. chrysosporium*. All fungal cultures were made according to methods described above.

The CO₂ evolution shake flask apparatus consisted of 1 ml of 1 N NaOH in an open container (CO₂ trap) suspended over 100 ml of culture medium in a 250 ml Erlenmeyer flask. The NaOH container was a 10 ml glass tube suspended by means of a hole in the rubber stopper with the glass support rod attached to the container. The NaOH (with captured 14 CO₂) was removed from the tube at least daily and placed in scintillation vial for radioactivity counting. The tube was then rinsed twice with 2 ml of deionized water and refilled with fresh base. The wash solution was also transferred to scintillation vials and 15 ml of the scintillation cocktail was added to these samples. All samples were quenched in the dark overnight before being counted. The fungal cultures were

sparged with 10 psi oxygen (for 1-2 minutes) at each time period after the samples were collected. All flasks were maintained at 37° C in an incubator between sampling periods.

In order to perform the final carbon balance on the inoculated photoproduct solutions, the flask contents were first filtered through a glass fritted filter. One ml each of the samples before and after filtration were collected for radioactivity counts. The filtrate volume was noted in each case and the remaining filtrate was then collected in a scintillation vial which was used for the fraction collector for HPLC analysis. Duplicate wash samples were then collected for the counts after washes with 10 ml of M9 salts, 10 ml of cold trichloroacetic acid (TCA), and 10 ml of ethanol/ether mixture (1:1). Ten ml of hot TCA (brought to 70° C using a water bath) was then used for washing in a similar manner. The cells were then digested with 5 ml HCl and maintained at 50° C oven in centrifuge tubes. The digested cells were centrifuged at 7000 rpm at 4° C for 15 min and then decanted and centrifuged again. The final step involved the use of the tissue solubilizer and was repeated until all brownish color was removed from the fungal mycelium.

IX. Chemicals: Standards of 2,4-diamino-6-nitrotoluene (DANT) and 4-amino-2,6-dinitrotoluene (ADNT) were purchased from SRI International (Menlo Park, Calif.). 2,4,6-trinitrotoluene (TNT) was prepared from 2,4-dinitrotoluene (Eastman Organic Chemicals, Rochester, N.Y.) by nitration according to the procedure of Heertjes (1958) with slight modification. [Phenyl-U-¹⁴C] labeled TNT was prepared by two-step nitration from toluene and [ring-UL-¹⁴C]toluene (Sigma Chemical Company, St. Louis, Mo.). The low-temperature nitration was performed in acetonitrile solution using silver nitrate and fluorosulfonic acid, followed by a high temperature step with mixed acids (nitric acid-sulfuric acid-sulfur trioxide) under optimized time and temperature conditions. The specific activity was 58.9 Ci/mmol and the radiochemical purity was >99% as verified by HPLC. All ¹⁴C-TNT and unlabeled TNT solutions were made up and stirred in an anaerobic atmosphere overnight and added to the photoreactor the next day.

X. ¹⁴C - fractionation from HPLC: The system used was a Beckman model 338 with 110B pumps, 406 analog interface with pump control, 507 autosampler with a 100- μ l sample loop, and a 167 detector with data analysis and system control provided by an IBM PS/2 with System Gold software (Beckman Instruments, Inc., San Ramon, Calif.). This method was also used with a Hewlett Packard model 1090 equipped with a diode-array detector and ChemStation software (Hewlett Packard Co., Avondale, Penn.). Radioactivity in the eluent fractions was quantified by collecting fractions on a Frac 100 fraction collector (Pharmacia, Uppsala, Sweden) loaded with vials containing 10 ml of Bio Safe II scintillation cocktail (Research Products International, Mount Prospect, Ill.), mixing, and counting in a Beckman LS 7000 scintillation counter. Disintegrations per minute (dpm) were calculated from counts per minute using an efficiency correction curve made by counting a set of quenched standards and plotting efficiency versus H numbers (Beckman). In cases where a Packard scintillation counter was used, this curve was generated by counting a different set of standards and plotting efficiency versus SIS numbers.

RESULTS

I. ANALYTICAL METHODOLOGY DEVELOPMENT

Three experiments were conducted for analytical methodology development and achievement of baseline conditions for photocatalytic degradation of TNT similar to NREL experimentation. Experiments 1 and 2 tested two HPLC mobile phase systems while experiment 3 tested TNT photocatalysis under anaerobic conditions.

Experiment 1 - The intent of this experiment was to test the experimental apparatus and compare two HPLC analysis methodologies, a methanol-water solvent system and an acetonitrile solvent system. Analyses were carried out on a HP-1050 HPLC system with a 25 cm x 2 mm Spherex C₁₈ column (Phenomenex) with a 5 cm x 2 mm guard column.

A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor using anaerobic conditions (N₂ purge) and in the presence of 0.3 g TiO₂. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 6.4 and initial temperature was 29° C. Approximately 10 minutes after lamp ignition and startup of the reactor, a slight reddish-brown color to the reactor contents was noted, similar to previous NREL results. Sampling continued, as outlined in the NREL report, for five hours. Final solution pH was 4.95 and final temperature was 27° C.

Analyses of the samples were conducted using two different solvent systems, a methanol-water solvent method (NREL method) and an acetonitrile-buffer solvent method (ACN method). Representative chromatograms from the NREL and ACN methods are shown in Figures 2 and 3, respectively. Both methods used a Spherex C₁₈ column and 7 µl injection volumes. As shown by the enclosed chromatograms of both methods, the separation of reaction products was superior with the acetonitrile solvent system.

Overall TNT degradation in this first experiment was similar to that seen in NREL experiments, approximately 81 percent. (Based on peak areas from representative chromatograms, data not shown). Formation of reaction intermediates occurred, however mass spectrometry on the compounds was successful in identifying only one of the reduction compounds, an amino-nitrotoluene. The spectral analyses of two of the intermediates are shown in Figures 4 and 5. Both of these compounds were identified as amino-dinitrotoluene isomers with a mw of 197. An injection volume of 50 µl was used in order to direct enough material to the mass spectrometer for these analyses. Even with such a high injection volume, other intermediates, at retention times of 2.63 and 29.27 minutes, could not be identified due to low recovery amounts sent to the particle beam.

Experiment 2. This experiment was conducted to quantitate TNT degradation and intermediate production. The HPLC analyses were done with a HP 1090 system, a Spherex C₁₈ column and a solvent delivery methodology specific for TNT and amino-nitrotoluene quantitation (Table 1). The HPLC method in Table 1 was used for the remainder of experimentation on the projects with only minor changes depending on run length and type of column used.

Initial and operating conditions for the experiment were similar to those of experiment 1. Sampling was conducted for five hours during which the reactor contents remained at 23.5° C. Initial and final reaction pH were 6.85 and 4.42, respectively. Similar to the previous experiment, a reddish-brown product was formed that remained in solution at the end of the experiment after five hours of phototreatment.

Overall TNT degradation was approximately 65 percent as shown in Figure 6. Four intermediates were produced, only one of which was positively identified as 4-amino, 2,6-dinitrotoluene (ADNT). The other three intermediates were not identified, similar to the results in experiment 1.

Experiment 3 - The intent of this experiment was to create a totally anaerobic environment for the photocatalytic reaction using a nitrogen gas purge and quenching of hydroxyl formation from the valence band hole reaction using bicarbonate ion. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO₂, phosphate buffer at pH 7 and the addition of enough Na₂CO₃ to produce 1 mM HCO₃⁻ at pH 7. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al.,

1995). Initial solution pH was 7.04 and initial temperature was 27° C. After lamp ignition and startup of the reactor, a pink color of the reactor contents was noted, similar to previous Experiments 1 and 2 results, only more intense. Sampling continued as per the previous experiments except for a total of ten hours. Final solution pH was 6.98 and final temperature was 24° C.

Overall TNT destruction in this third experiment was approximately 70 percent after ten hours of treatment, as shown in Figure 7. Four reaction intermediates were formed and quantified one of which, eluting near TNT, was identified as ADNT (similar to Experiment 2). The remaining intermediates were not identified from this analysis. A chromatogram from HPLC analysis of the last sample of experiment 3 (data which went to form Figure 7) is shown in Figure 8. In comparison to a chromatogram from the last sample of experiment 2 (not buffered, no hydroxyl radical quench) the experiment 3 chromatogram shows additional intermediate compound production during photocatalysis. These differences were hypothesized to be due either to the anaerobic, buffered environment and/or the difference in reaction times between the two experiments, 5 and 10 hours for experiments 2 and 3, respectively.

II. PHOTOCATALYSIS OF TNT AND AMINO-DINITROTOLUENE (ADNT) AND QUALITATIVE PHOTOCATALYTIC PRODUCT IDENTIFICATION EXPERIMENTS

Eight experiments were conducted in this section of research devoted to either quantitating photocatalysis of TNT and ADNT or qualitative breakdown product identification. Experiments 4, 5, 10, 11 and 16 were conducted on photocatalytic TNT destruction in the presence of ethylenediaminetetraacetic acid (EDTA) as a catalyst hole scavenger. Experiments 6 and 7 were conducted on photocatalytic ADNT destruction to imitate potential precedent biological degradation products that could be sent to a photocatalytic reactor in a combined biological/photocatalytic degradation sequence. Experiment 17 was a repeat of experiments 10 and 11 examining photo-destruction products from a longer ten-hour photocatalytic treatment.

TNT Photocatalysis Degradation Product Identification Experiments

Experiment 4 - The intent of this experiment was to create a totally anaerobic environment for the photocatalytic reaction requiring nitrogen gas purge and quenching of hydroxyl formation from the valence band hole reaction. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 , phosphate buffer at pH 7, 107 mg Na_2CO_3 to produce 1 mM HCO_3^- at pH 7 and 0.72 mM disodium EDTA. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 7.26 and initial temperature was approximately 24° C. After lamp ignition and startup of the reactor, a slight orange color of the reactor contents was noted. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 7.28 and final temperature was approximately 23° C.

Overall TNT destruction in this third experiment was approximately 96 percent after five hours of treatment, as shown in Figure 9. Four reaction intermediates were formed and quantified. Two of the intermediates, eluting at approximately 20 and 27 minutes, were identified as 4-amino,2,6-dinitrotoluene (ADNT) and 2,2',6,6'-tetranitro-4,4'-azoxytoluene (TNAzT), respectively, based on UV-vis spectral analysis from the diode array. Spectral analyses and library match confirmation are shown in Figures 10 and 11. The remaining intermediates were not identified from this analysis, however spectral analysis of the intermediate eluting at 14.6 minutes was similar to ADNT and the compound eluting at 3.5 minutes appeared to be a mixture of compounds.

Experiment 5 - Experiment 5 was designed to observe the "fast" anaerobic photocatalytic reaction at uncontrolled pH as compared to experiment 4, at pH 7. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.72 mM disodium EDTA. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 5.15 and initial temperature was approximately 22° C. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 8.02 and final temperature was approximately 21° C. Reactor contents ultimately turned to a brown color, but at a slower rate than coloration during experiment 4.

Overall TNT destruction in this experiment was approximately 97 percent after five hours of treatment, as shown in Figure 12. At this pH regime, 90 percent TNT destruction was seen within 30 minutes of lamp ignition. Similar to experiment 4, four major reaction intermediates were formed and quantified, also shown in Figure 12. Two of the intermediates were identified, similar to experiment 4, as 4-amino,2,6-dinitrotoluene (ADNT) and 2,2',6,6'-tetranitro-4,4'-azoxytoluene (TNAzT), eluting at 20 and 27 minutes, respectively. Spectral matches were conclusive for ADNT and inconclusive for TNAzT, due to noise. Additional intermediates formed, eluting near the previously identified TNAzT peak at 27.16 min., were lumped together and listed as compounds with elution times >25 minutes (Figures 12 and 13). These compounds were thought to be azoxy-isomers, although their identification was not confirmed. The remaining intermediates were not identified from this analysis, however spectral analysis of the intermediate eluting at 14.6 minutes was similar to ADNT and the compound eluting at 3.5 minutes appeared to be a mixture of compounds. Additionally, the compounds eluting at 2.37 and 3.5 min. may have been EDTA and its reduction product, but were not confirmed.

Experiment 10. Because the acid-form of EDTA was used in experiment 9, pH values were lower than when using disodium EDTA. This presumably avoided the formation of azoxy compounds. Experiment 10 was a repeat of experiment 9 to further quantify the effects of acid-form EDTA use. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.4 mM acid-form EDTA and 0.43 mM disodium EDTA. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 3.2 and initial temperature was approximately 27° C. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 6.7 and final temperature was approximately 26° C.

Results of experiment 10 are shown in Figure 14. Upon lamp ignition, TNT disappeared at a fast rate with no detectable concentration after 60 minutes. Both ADNT and 2,4-diamono-6-nitrotoluene (DANT) were detected as TNT degradation products, in addition to TNAzT. The azoxytoluene did disappear by the end of the experiment, while the amino- compounds continued to increase in peak size.

Problems with this experiment indicated that there may have been compounds precipitating out of sample solution prior to analysis. Problems also arose with one of the HPLCs used for analysis.

Experiment 11. Experiment 11 was a repeat of experiment 10, to quantify the effects of acid-form EDTA use and correct sample holding time that gave problems in experiment 10. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.4 mM acid-form EDTA and 0.43 mM disodium EDTA. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 3.05 and initial temperature was approximately 22.5° C. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 6.56 and final temperature was not measured, but assumed equal to initial temperature $\pm 2^\circ$.

Results of experiment 11 are shown in Figure 15. Upon lamp ignition, TNT disappeared at a fast rate with no detectable concentration after 40 minutes. Both ADNT and DANT were detected as TNT degradation products, in addition to their precursor, 4-hydroxylamino-2,6-dinitrotoluene (HADNT). Two unidentified products were formed with retention times of 4.91 and 5.46 minutes. No azoxytoluenes were identified in this experiment. Both ADNT and DANT were confirmed by HPLC spectral analyses. The HADNT was confirmed by HPLC and mass spectral analysis (Figures 16 and 17).

Experiment 16 - This experiment was conducted similar to experiments 10 and 11, except with a ten-hour duration. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.4 mM acid-form EDTA and 0.43 mM disodium EDTA. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 3.30 and initial temperature was approximately 16.2° C rising to a constant 25° C after one hour of treatment. Final solution pH was 6.92. Results of the experiment are shown in Figure 18, with an ultimate TNT destruction of greater than 95 percent.

ADNT Photocatalytic Degradation Experiments

Experiment 6 - The intent of this experiment was to observe the photodegradation of a "model" biological product of TNT degradation, 4-amino-2,6-dinitrotoluene (ADNT). A one-liter quantity of 0.44mM ADNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.72 mM disodium EDTA. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 4.70 and initial temperature was approximately 25° C. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 8.07 and final temperature was approximately 25° C. Reactor contents turned to a light brown color after approximately 2 hours of reaction time.

Based on the results shown Figure 19, ADNT was degraded approximately 80 percent in five hours. Two major degradation products appeared, one of which was identified as DANT; the other was not identified. Although not positively confirmed, the suspected EDTA chromatographic peak declined throughout the experiment similar to that of ADNT, as in the previous experiments with TNT.

Experiment 7 - This experiment was a repeat of experiment 6 using similar sampling and analytical procedures. A 0.44mM ADNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.72 mM disodium EDTA. Initial pH was 4.51 and final pH was 7.76. Initial reactor temperature was approximately 29° C dropping to a constant 22° C after one hour of reaction time. The results of this experiment are shown in Figure 20. Based on results in the figure, approximately 88 percent of the original ADNT was degraded in five hours. The unidentified degradation product (retention time of 17.04 min) from experiment 6 did not appear in this experiment.

Partial Anaerobic Conditions Experiment

Experiment 17 - The final experiment under this section was conducted under partial anaerobic conditions. The preparation of the initial TNT solution was not done under anaerobic conditions as in all previous experiments but rather in natural atmospheric conditions. In addition, the experiment was not conducted with a continuous nitrogen gas purge, but rather, nitrogen was purged initially into the reactor, the reactor was sealed and then the experiment was run without further attempts at keeping the contents anaerobic. This kind of scenario was thought to better imitate true treatment process conditions and would result in overall cost savings by avoiding much of the gas purge. A 0.44 mM TNT solution was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.72 mM disodium EDTA. Initial and ending solution pH was 3.40 and 5.59,

respectively and the reactor contents were held at 25° C throughout the experiment. As can be seen in Figure 21, TNT was degraded to greater than 95 percent in 60 minutes with appearance of five photo-treatment products, two unidentified and the remaining identified as ADNT, DANT and TNazT. The azoxytoluene did disappear with time, similar to experiment 10.

III. PHOTOCATALYTIC DEGRADATION MECHANISM

Three experiments in this section of experimentation were intended to elucidate the photocatalytic degradation mechanism similar to procedures of Glaze et al. (1993). Experiments 8 and 13 used tetranitromethane (TNM) and periodate, respectively, as electron scavengers. Experiment 14 used methanol as a hydroxyl radical quencher.

Experiment 8 - The intent of this experiment was to identify the mechanism of TNT degradation. Because tetranitromethane (TNM) competes in the photocatalytic reaction for electrons, it was hypothesized that its addition would effectively stop reductive TNT degradation. Once the TNM was exhausted (completely reduced), the TNT would then reductively react as in previous experiments. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO₂, 0.72 mM disodium EDTA and 5.1 mM TNM. Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 4.97 and initial temperature was approximately 19° C. The pH dropped to 3.3 during the first 30 minutes of the experiment and then increased to experiment's end. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 4.76 and final temperature was approximately 20° C. Reactor contents initially were green, due to the TNM, but ultimately turned to a brown color, similar to previous anaerobic experiments.

TNT degradation was initially slowed in this experiment, presumably due to the presence of TNM (see Figure 22). Final TNT degradation amounted to only 75 percent, considerably less than the >95 percent as in previous experiments 10 and 11 without TNM. None of the major reduction products 4-hydroxylamino-2,6-dinitrotoluene (HADNT), 4-amino,2,6-dinitrotoluene (ADNT) nor 2,4-diamino-6-nitrotoluene (DANT) were found on the chromatograms. Neither was 2,2',6,6'-tetranitro-4,4'-azoxytoluene (TNazT) found. Two unidentified products were formed with retention times of 4.91 and 5.46 minutes. While not conclusively identified in the chromatograms, TNM was thought to be one of the compounds that disappeared at 15 minutes of treatment (HPLC retention time of 18.7 or 28.5 minutes). As seen in Figure 22, TNT did not appreciably degrade until after the 15 minute sample, possibly indicating that TNM did successfully compete for electrons, effectively stopping the TNT reduction reaction. The absence of typical TNT reduction products also leads to the hypothesis that TNT did not degrade reductively.

Experiment 13 - The intent of this experiment was to help identify the mechanism of TNT degradation. Because periodate competes in the reaction for electrons, it was thought that its addition would effectively stop reductive TNT degradation. A one-liter quantity of 100 ppm TNT solution was treated in the photoreactor in the presence of 0.3 g TiO₂, 0.72 mM disodium EDTA and 26.4 mM periodate (sodium form). Reactor startup and sampling was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995). Initial solution pH was 4.55 and initial temperature was approximately 20° C. Intermediate pH was not available due to problems with the reactor pH probe. Sampling continued as per the previous experiments for a total of five hours. Final solution pH was 8.90 and final temperature was approximately 22° C. Reactor contents initially were clear and remained so, unlike previous anaerobic experiments without an electron scavenger that turned a reddish-brown color during the course of the experiment.

Surprisingly, TNT disappearance was rapid, with TNT concentration dropping to zero by the 180 minute sample (Figure 23). No reduction products, however, either monoamino-nitrotoluene, diamino-nitrotoluene or azoxy-toluenes were detected throughout the experiment.

Experiment 14. Experiment 14 was designed to quench any hydroxyl radical formation using methanol in a methanol:TNT molar ratio of 70:1. Experimental conditions were similar to experiment 11, using 0.44 mmol TNT, 0.4 mmol Na₂EDTA·H₂O and 0.4 mmol acid form EDTA with the addition of 30.8 mmol methanol. Beginning pH was 3.08 and ending pH was 6.20. Temperature was held constant at 15.5 ± 2.5 °C. Results are shown in Figures 24 and 25. From the chromatogram in Figure 25, it can be seen that the results are similar to the experiment without methanol (experiment 11) with formation of ADNT and DANT and no azoxy- compounds.

IV. COMBINED PHOTOCATALYTIC AND BIOLOGICAL DEGRADATION EXPERIMENTS

The combined photocatalytic and biological degradation experiments comprised i) three preliminary experiments to identify a likely biological treatment organism and qualitative aspects of photoreactor treatment product degradation and ii) two large quantitative experiments using ¹⁴C-TNT in a precedent photocatalytic degradation step with follow-on biological mineralization of the photodegradation products. The ¹⁴C-radiotracer experiments allowed quantitation of ultimate TNT mineralization in addition to an overall carbon balance in the system.

Biological Experiments 1 and 2. Based on the assumption that reaction intermediates from the above photocatalytic reduction degradation experiment were possibly large dimers or polymeric compounds, it was decided to use *Phanerochaete chrysosporium* as a likely biological agent for follow-on photoreactor product degradation. Both experiments were conducted as simply presence or absence of biological activity as seen by decoloration of the compounds.

Fresh hyphae, harvested from *Phanerochaete* previously grown up on a glucose/dimethyl succinic acid media (see Materials and Methods for fungal growth media and techniques descriptions), were added to varying amounts of the leftover fluid from experiment 3. Biological experiment 1 used a ratio of hyphae:fluid as 5:1, on a volume to volume basis. The fluid was decolorized, or bleached white (the TiO₂ showed white) within 3 days. The second biological experiment used a hyphae:fluid ratio of 1:5. A color change from reddish-brown to light brown was observed at 3 days with no further decoloration up to 10 days. No analyses of the *Phanerochaete*-amended fluid was done at this time, although it is possible that the first biological experiment was successful in degrading the polymeric materials left from anaerobic chemical photocatalysis.

Biological Experiment 3 - *Phanerochaete chrysosporium* was inoculated into varying concentrations of the final, unfiltered products from experiment 5. Aliquots containing 0, 20, 40, 60, 70 and 83 percent photo-products were made, inoculated with fungal hyphae and allowed to react. Samples were withdrawn at three day intervals and analyzed for breakdown products from TNT degradation. Small concentrations of 4-amino-2,6-dinitrotoluene (ADNT) and 2,4-diamino-nitrotoluene (DANT) were detected in time 0 samples from higher photo-product concentration aliquots with their complete disappearance three days later. Identification of possible azoxy-compounds was not made during these experiments as well as inconclusive identification of other extraneous chromatographic peaks. Because the photo-products from experiment 5 were refrigerated, it was thought they may have precipitated out of solution prior to the start of this biological experiment. The fungal inoculants did, however, appear to grow, increasing in size, and were not affected by the EDTA nor the TiO₂ in the reactor effluent. Peroxidase may have been expressed by the organisms, indicative of their brown color at approximately three days growth time.

Experiment 9 - The intent of this experiment was to quantitate the carbon fractions incorporated into TNT breakdown products for use in follow-on biological experiments. A one-liter quantity of 98 ppm TNT solution (^{14}C =28,600,00 dpm) was treated in the photoreactor in the presence of 0.3 g TiO_2 and 0.4 mM acid-form EDTA and 0.43 mM disodium EDTA. Reactor startup was identical to the methodology outlined in the NREL Six Month Progress Report (Blake et al., 1995) but sampling was limited to initial and final samples only. The experiment was run for a total of 5.5 hours. Temperature and pH profiles were not taken to avoid undue radioactive contamination, but assumed similar to previous anaerobic experiments.

Based on differences between initial and final samples, the overall ^{14}C balance accounted for 94 percent of total radiolabeled TNT in solution (Table 3, difference between original 28,592,000 dpm applied and 26,882,712 dpm recovered). Table 3 shows relative minor amounts of radioactivity associated with products 1) adsorbed on reactor glass (2073 dpm/ml removed in a reactor rinse); 2) presumed ^{14}C - CO_2 in carbosorb (945 dpm/ml); and 3) a nonquantitative amount (volume of sample unknown due to a simple "wipe" test) in the solids from the reactor (26677 dpm from inner tube and 21714 dpm from outer tube). Figure 26 shows the relative amounts of radioactivity associated with each of the TNT breakdown products overlain with identified products from the chromatogram. Radioactivity was associated with TNT, ADNT and DANT in addition to two other peaks thought to be diamino-nitrotoluene isomers. It is interesting to note that no detectable amounts of azoxy- compounds were formed in this experiment, presumably due to the low pH arising from use of acid-form EDTA. Additional loss of radioactivity was presumably due to compounds retained by filtration as shown by information in Table 3. Fifteen to 16 percent of final reactor contents were lost on filtration depending on filter type, Teflon vs. cellulose acetate. Additionally, between 7 and 11 percent of radioactivity was lost upon filtration due to either adsorption of soluble material to the filter or adsorption of material to TiO_2 and loss upon filtration, respectively.

Experiment 12. Experiment 12 used reactor contents from experiment 9, photocatalysis of ^{14}C -TNT. The experiment followed protocol similar to that of 40 CFR 796.310 Aerobic Aquatic Biodegradation Shake Flask Protocol. Replicate 50 ml quantities of mixed reactor contents from photocatalysis of ^{14}C -TNT were inoculated with *Phanerochaete chrysosporium* on 10/9/95 and incubated on a shake table. The experiment used one control and three concentrations of an exogenous nitrogen source, 0 mM NH_4 , 0.25 mM NH_4 , and 2.5 mM NH_4 . Removal and counting of the base solution contained within the flasks was done every two to three days for the 21-day experiment, to quantitate the ^{14}C - CO_2 evolved. Previous preliminary results from the experiment showed that radioactive counts occurred within the first two days, presumably due to aerobic, biodegradative attack of the photocatalysis byproducts by the *Phanerochaete*. Further results showed varying amounts of ultimate ^{14}C - CO_2 production (Figure 27), depending on supplemental NH_4 added. Previous research has shown that addition of the exogenous nitrogen was necessary for expression of the peroxidase enzyme in *Phanerochaete*, although that was not the case in this study, presumably enough nitrogen was supplied by the degradation products. A mass balance on ^{14}C within the system, based on 1) evolved ^{14}C - CO_2 , 2) ^{14}C bound to fungal hyphae and 3) MeCN extracted ^{14}C from the TiO_2 solids showed only 60-70 percent recovery of original added radioactivity (Table 4).

The decline in soluble radioactive compounds with time is shown in Figure 28. Presumably, the photocatalysis products contained within the reactor fluid either sorbed to the TiO_2 solids within the matrix or the glass of the storage flask or a spontaneous reaction between molecular oxygen and amino- compounds occurred.

Further ^{14}C analyses of the previous reactor contents before and after degradation by *Phanaerochaete* is shown in Figure 29. The top graph shows photocatalytic byproducts prior to biological treatment for qualitative comparison. The first three sharp peaks represent reduction

products, only the last of which was identified as diamino-nitrotoluene (DANT). The bottom two graphs show analysis of extracts of the fungal hyphae from experimental treatments of the fungus with 0.25 mM NH_4 or zero NH_4 . The important feature of these graphs is the presence of the DANT peak with the 0.25 mM NH_4 and the absence of the peak with no NH_4 added. Exogenous NH_4 was previously thought necessary for fungal metabolism of xenobiotic compounds.

Experiment 15. Experiment 15 was a repeat of the previous experiment 9, photocatalysis of ^{14}C -TNT and experiment 12, treatment of photocatalytic TNT breakdown products using *Phanerochaete chrysosporium*. The photocatalytic portion of the experiment was identical to previous experiment 9. The biological portion of the experiment followed protocol similar to that of 40 CFR 796.310 Aerobic Aquatic Biodegradation Shake Flask Protocol (1992). Replicate 50 ml quantities of mixed reactor contents from photocatalysis of ^{14}C -TNT at time 2 h and 6 h and in addition to a flask with 100 mg/l ^{14}C -TNT (no photocatalytic treatment) were inoculated with *Phanerochaete chrysosporium* and incubated on a shake table. An abiotic control was also run on the reactor contents. Removal and counting of the base solution contained within the flasks was done every two days for the 27 day experiment, to quantitate the ^{14}C - CO_2 evolved. Final results from the experiment showing accumulated percent ^{14}C - CO_2 released are presented in Figure 30. Ultimately more ^{14}C - CO_2 was released from the photocatalytic treatment for six hours than from that treated for two hours. This probably indicates that either longer reductive photocatalytic treatment is necessary for follow-on oxidative biological treatment or the destruction of a biologically toxic intermediate occurs after two or more hours of photocatalytic treatment. Additionally, the comparison of liquid and evolved ^{14}C - CO_2 counts for the 2 hr, 6 hr and TNT flasks is shown in Figure 31.

The overall carbon mass balance is shown in Table 5. All recoveries of radio-tracer were good and in all cases greater than 92 percent. As can be seen in the lower portion of Table 5, almost twice as much radio-labeled material was associated with the biomass of the fungi (numbers in the HCl digest and tissue solubilizer rows) in the 2 hr treatment as with the biomass in the six hour treatment. This data agreed with the hypothesis that longer photocatalytic treatment produced more polar products with less tendency to adsorb.

Chromatograms showing TNT reduction products, before and after biological treatment, with overlain radioactive counts for the 2 h and 6 h sample are shown in Figures 32 and 33. Figure 33 also shows a chromatogram and overlain radioactive counts for the uninoculated control. Distribution of photocatalytic products in both the 2 and 6 hour treatments were similar with the exception of two peaks between 20 and 25 minute retention being at higher concentration in the 2 hr treatment. This was probably the toxic fraction of photoproducts that inhibited mineralization in the 2 hr samples. Distribution of radiolabel was similar in both the 2 and 6 hr treatments after fungal treatment.

CONCLUSIONS AND DISCUSSION

I. ANALYTICAL METHODOLOGY DEVELOPMENT

The following method for analysis of TNT and amino-nitrotoluenes on the HPLC was superior to the NREL method using a water/methanol gradient. Chromatographic peaks were better separated and quantified using the resultant method. (The method listed below is somewhat different than the method listed in Experiment 2 due to longer run time, 23 vs 18 minutes.)

Hewlett-Packard LC 1090				
Pump (DR5)				
Stop Time				23.00 min
Flow				0.25 ml/min
Min. Pressure				0 bar
Max. Pressure				400 bar
Oven Temp.				Off
Solvent A				5.0 % (acetonitrile)
Solvent B				95.0 % (PO ₄ buffer)
Injector				
Injection Volume				10.0 µl
Draw speed				83.3 µl/min
Time Table				
Time (min)	%A	%B	Flow (ml/min)	
0.0	5	95	0.25	
0.5	5	95	0.25	
8.0	50	50	0.25	
9.0	50	50	0.25	
11.0	100	0	0.25	
15.0	100	0	0.25	
18.0	5	95	0.25	

Sample holding time was also found to be critical for accurate analyses. When samples were held for longer than 30 minutes or stored under reduced temperature, loss of compounds in solution occurred. This disappearance of the compounds under study was thought to be due to either adsorption onto the TiO₂ particles or glass of the holding vials or spontaneous reaction and conversion to other chemical forms.

II. PHOTOCATALYSIS OF TNT AND AMINO-DINITROTOLUENE (ADNT)

Three major reduction products of anaerobic photocatalysis of TNT were identified in the experimentation: 1) 4-hydroxylamino-2,6-dinitrotoluene (HADNT); 2) 4-amino,2,6-dinitrotoluene (ADNT); and 3) 2,4-diamono-6-nitrotoluene (DANT). Additionally a condensation product 2,2',6,6'-tetranitro-4,4'-azoxytoluene (TNAzT) was identified in experiments where pH became alkaline during the course of the experiment. Other isomers of the three reduction products were suspected to have occurred along with other azoxytoluenes, based on mass spectrometry results, however positive identification was not made.

The proven reductive photocatalytic destruction of TNT in this research followed the theoretical pathway similar to that given in the Appendix of NREL's Six Month Progress Report (Blake et al., 1995): TNT → 4-hydroxylamino-2,6-dinitrotoluene (HADNT) → 4-amino,2,6-dinitrotoluene (ADNT) → 2,4-diamono-6-nitrotoluene (DANT). (See Figure 34) Based on this pathway, a probable biological agent, *Phanerochaete chrysosporium*, was deemed adequate to mineralize the resultant photodegradation products. Because HADNT is toxic to *P. chrysosporium* through its inhibition of the veratryl alcohol oxidase activity of lignin peroxidase (Paszczynski and Crawford,

1995, Bumpus and Tartako, 1994, Michels and Gottschalk, 1994), it was also thought that photocatalysis of TNT had to proceed past HADNT and at least to ADNT to achieve adequate mineralization of the parent compound.

IV. COMBINED PHOTOCATALYTIC AND BIOLOGICAL DEGRADATION OF TNT

Based on results from this section of research the following conclusions were drawn:

1. Mineralization of TNT was possible using a combined photocatalytic and biological degradation scheme, with *P. chrysosporium* as the biological agent. The maximum extent of observed mineralization of the original C-TNT was approximately 32 percent.
2. The length of photocatalysis affected the final extent of C-TNT mineralization in the combined photocatalytic and biological degradation scheme. There was evidence from the varying lengths of photocatalysis used, two and six hours, that longer treatment in the photoreactor resulted in a greater extent of biological mineralization of the initial C-TNT, 23 percent vs. 32 percent, respectively. Longer photo-treatment (six hours) also produced effluent compounds that were more polar than the original influent compounds (TNT) or than effluent compounds from shorter photo-treatment (two hours).
3. It was previously thought that preparations of *P. chrysosporium* had requirements for an exogenous nitrogen source. This phenomenon was shown to be false in our research. Fungal preparations without additional nitrogen or with low nitrogen added (0.25 mM NH_4) showed greater overall mineralization than did those with higher added nitrogen (2.5 mM NH_4).

Others have investigated the sole use of *P. chrysosporium* to mineralize TNT (Fernando et al., 1990; Spiker et al., 1992; Michels and Gottschalk, 1994). While they all obtained mineralization, the ultimate extent of TNT converted to CO_2 was less than in the present research. TNT was also acknowledged to be toxic to *P. chrysosporium* at concentrations above 20 mg/l in aqueous solution, further limiting ultimate mineralization using only the fungal preparations.

The greater extent of TNT mineralization in samples receiving longer photocatalysis was probably due to either absence of any toxic intermediates in the degradation pathway or greater bioavailability of the carbon in the degradation products. Based on our results, the occurrence of HADNT, toxic to *P. chrysosporium*, was avoided by longer photocatalytic treatment. Also, longer treatment produced more polar and soluble intermediate compounds that were more bioavailable for the follow-on biological treatment, however an optimum length of photocatalytic treatment was not determined. In summary, the combined photocatalytic and biological treatment of aqueous TNT solutions appears to be a viable remediation technology if the process is properly optimized.

RECOMMENDATIONS FOR FUTURE WORK

Recommendations for future research into photocatalytic and biological destruction of pink water are the following:

1. Optimization of the combined photocatalytic and biological processes. While the combined photocatalytic and biological scheme for TNT destruction was not optimized, experimental evidence showed that better treatment could be had based on photocatalytic process length. The use of central composite experimental designs together with response surface methodologies has been successfully used to optimize response in various hazardous and non-hazardous waste treatment processes (Hess, et al., 1996; Watts et al.,

1994a). The overall combined photocatalytic and biological process should also be investigated for optimized configuration and resultant cost savings based on sequential vs. simultaneous treatment schemes.

2. Investigation of catalyst-coated solid support media. Many studies of photocatalytic treatment of waste streams are using TiO_2 -coated solid support media for ease of TiO_2 removal in a full-scale application (Glaze et al., 1993; Hofstadler et al., 1994). Previous research has shown that aqueous suspensions of TiO_2 are hard to remove, increasing the overall cost of a full-scale remediation process (Watts et al., 1994b).

3. Investigation of mixtures of explosives compounds for input into a combined photocatalytic and biological treatment scheme. It is likely that mixtures of nitroaromatic compounds will behave differently during both photocatalysis and biological degradation than does a single compound. Process optimization and resultant process cost minimization should be studied.

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Hewlett-Packard LC 1090				
Pump (DR5)				
Stop Time		35.00 min		
Flow		0.25 ml/min		
Min. Pressure		0 bar		
Max. Pressure		400 bar		
Oven Temp.		42° C		
Solvent A		1.0 % (acetonitrile)		
Solvent B		99.0 % (water)		
Time Table				
Time (min)	%A	%B	Flow (ml/min)	
0.0	1	99	0.25	
3.0	1	99	0.25	
16.0	50	50	0.25	
20.0	50	50	0.25	
26.0	0	0	0.25	
31.0	0	0	0.25	
33.0	1	99	0.25	

Table 1. HPLC method used in experiment 1.

Hewlett-Packard LC 1090				
Pump (DR5)				
Stop Time	18.00 min			
Flow	0.25 ml/min			
Min. Pressure	0 bar			
Max. Pressure	340 bar			
Oven Temp.	42° C			
Solvent A	10.0 % (acetonitrile)			
Solvent B	90.0 % (LiOH/PO ₄ , pH 4.0)			
Injector				
Injection Volume	10.0 µl			
Draw speed	83.3 µl/min			
Time Table				
Time (min)	%A	%B	Flow (ml/min)	
0.0	10	90	0.25	
1.0	10	90	0.25	
2.0	40	60	0.25	
8.0	60	40	0.25	
9.5	100	0	0.25	
10.5	100	0	0.25	
15.0	100	0	0.25	
16.5	10	90	0.25	

Table 2. HPLC method used in experiment 2.

¹⁴ C Source	DPM
Original ¹⁴ C added	28,592,000
Reactor contents after 5.5 h of photocatalysis	24,809,893
Reactor rinse	2,072,818
Total ¹⁴ C recovered	26,882,712
Total recovery	94.02 %

Table 3. ¹⁴C balance from photocatalysis of ¹⁴C-TNT in experiment 9.

¹⁴ C Source	Percent ¹⁴ C-CO ₂ evolved	Percent ¹⁴ C in reactor liquid	Percent ¹⁴ C bound to hyphae	Percent ¹⁴ C in MeCN extraction	Total percent ¹⁴ C recovered
Control	0.02±0	not done	NA	NA	0.02±0*
0 mM NH ₄	19.86±2.40	24.19±3.41	16.86±0.75	6.63±3.38	67.54±5.16
0.25 mM NH ₄	6.65±0.97	20.60±1.69	19.34±1.94	16.16±1.62	62.75±2.34
2.5 mM NH ₄	3.66±0.69	21.29±0.26	19.30±3.02	18.03±3.87	62.27±6.46

*recovered as ¹⁴C-CO₂ only

Table 4. ¹⁴C balance, experiment 12, from biological mineralization, by *P. chrysosporium*, of products from photocatalysis of ¹⁴C-TNT. Data are averages ± standard error for 0, 0.25 and 2.5 mM additions of NH₄ to the fungal/photoproduct suspensions in addition to a noninoculated control.

¹⁴ C source	0 h treatment	2 h treatment	6 Hr treatment	Control
CO ₂ evolved	13.7±2.5	22.7±1.3	31.8±0.8	2.2±0.1
Particulates in solute	17.7±4.1	2.0±0.6	2.1±0.4	21.4±2.1
Soluble	17.1±6.0	21.3±0.7	32.7±0.8	72.0±2.5
M9 Salts Wash	1.7±0.5	0.7±0.2	0.9±0.4	N/A
Cold TCA Wash (carbohydrates)	0.9±0.5	0.2±0.0	0.4±0.2	N/A
Ether: Ethanol Wash (lipids)	5.3±2.4	0.9±0.4	1.0±0.3	N/A
Hot TCA Wash (nucleic acids)	0.8±0.3	0.2±0.0	0.4±0.1	N/A
Residue	1.1±0.1	0.2±0.1	0.2±0.2	N/A
HCl digest (protein)	4.3±0.5	5.1±0.2	3.6±0.3	N/A
Tissue Solubilizer Digest (strongly sorbed, bound)	38.2±13.7	39.4±2.1	21.5±0.9	N/A
Total Recovery	100.7±11.6	92.7±2.1	94.6±1.6	95.6±3.5

Table 5. ¹⁴C balance showing average ± standard error of samples from biological mineralization, by *P. chrysosporium*, of products from photocatalysis of ¹⁴C-TNT, experiment 15. Data are from 0, 2, and 6 hours of precedent photocatalysis in addition to a noninoculated control.

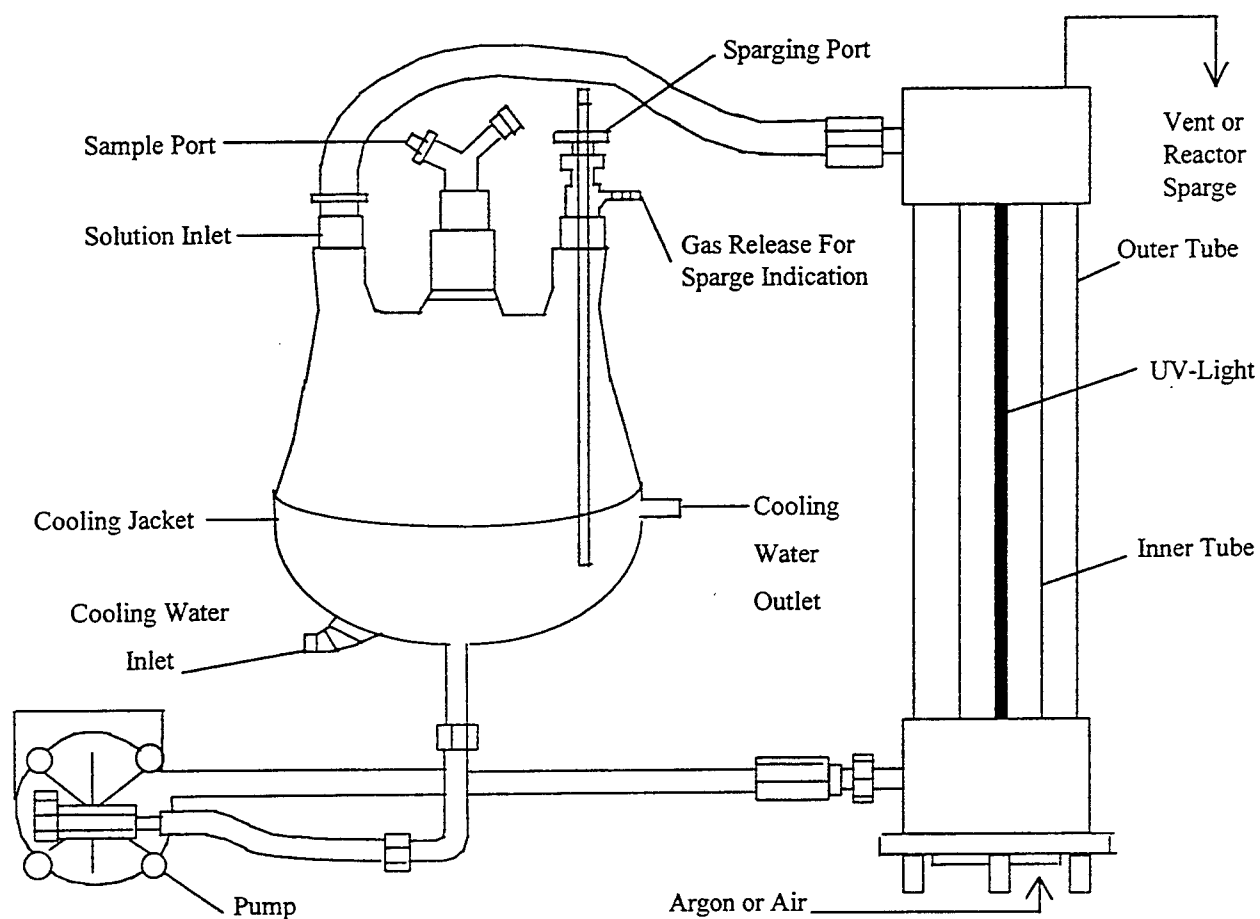
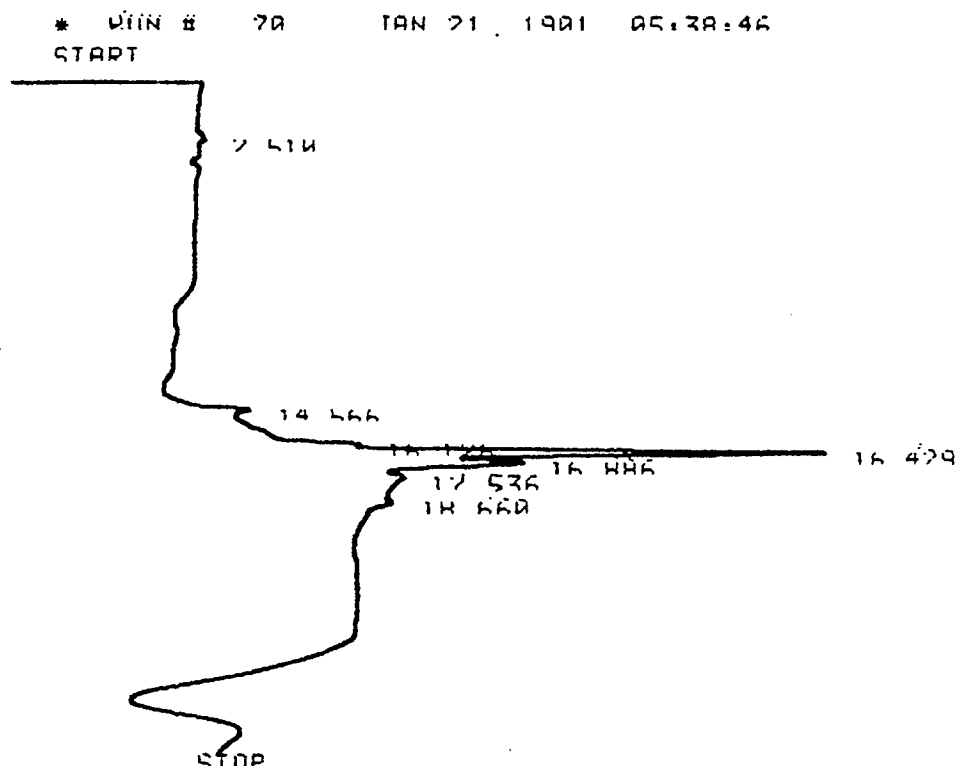


Figure 1. Schematic of annular ring, photocatalytic reactor used in all experimentation.



Closing signal file M:STGNAI AND

RUN# 70 JAN 21, 1991 05:38:46

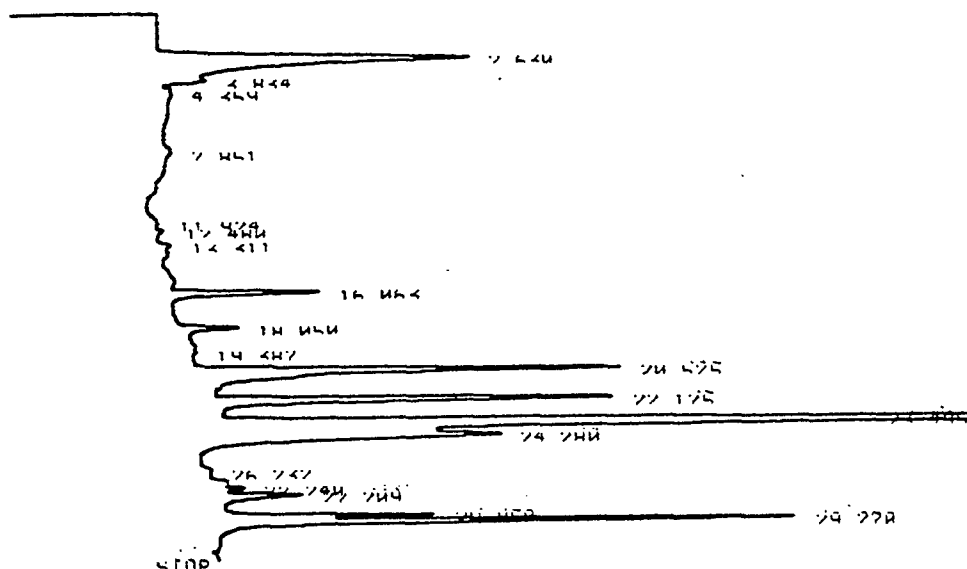
STGNAI FILE: M:STGNAI AND
AREA%

RT	AREA	TYPE	WIDTH	AREA%
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14.566	1342199	PII	373	3.79306
15.176	4910285	III	654	13.87647
15.479	10046496	III	309	28.39147
16.886	7244490	III	480	20.47294
17.536	6430275	III	825	18.17198
18.660	5237770	III	950	14.80194

TOTAL AREA=3.5306E+07
MIN FACTOR=1.0000E+00

Figure 2. Chromatogram of products from photocatalytic degradation of TNT using NREL method in Experiment 1.

* WIN # 97 JAN 23, 1991 04:03:10
START



RIIN# 97 JAN 23 1981 04:03:10

STANAI FTIF, MISTANAI AND
AREAX

PHZ	DT	REFR	TYPE	LITENTH	REFRAX
2	530	140000000	PH	543	12 14031
3	034	1137754	HP	370	90067
4	350	376601	PR	620	20170
7	051	1475400	AP	1 000	1 22914
11	074	200534	PH	564	24075
12	400	255015	HH	351	22050
13	311	206533	HH	312	25570
15	063	2732541	PR	260	2 35630
16	050	1210714	PH	204	1 04401
19	307	464571	HP	013	40060
20	525	11023440	PR	370	0 50563
21	175	03000000	PR	207	7 20051
23	405	46377037	PH	335	20 00710

Figure 3. Chromatogram of products from photocatalytic degradation of TNT using ACN method, Experiment 1.

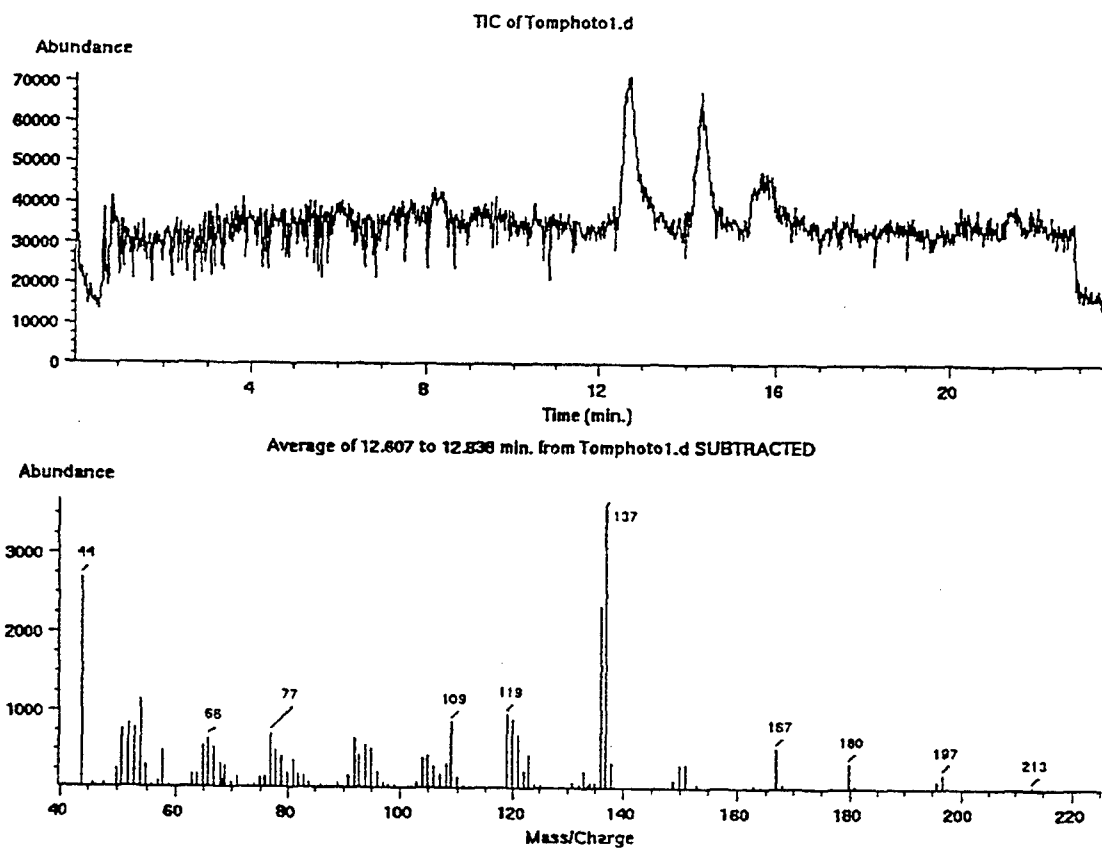


Figure 4. Chromatogram (top) and mass spectra (bottom) of an amino-nitrotoluene isomer formed during photocatalytic degradation of TNT in Experiment 1.

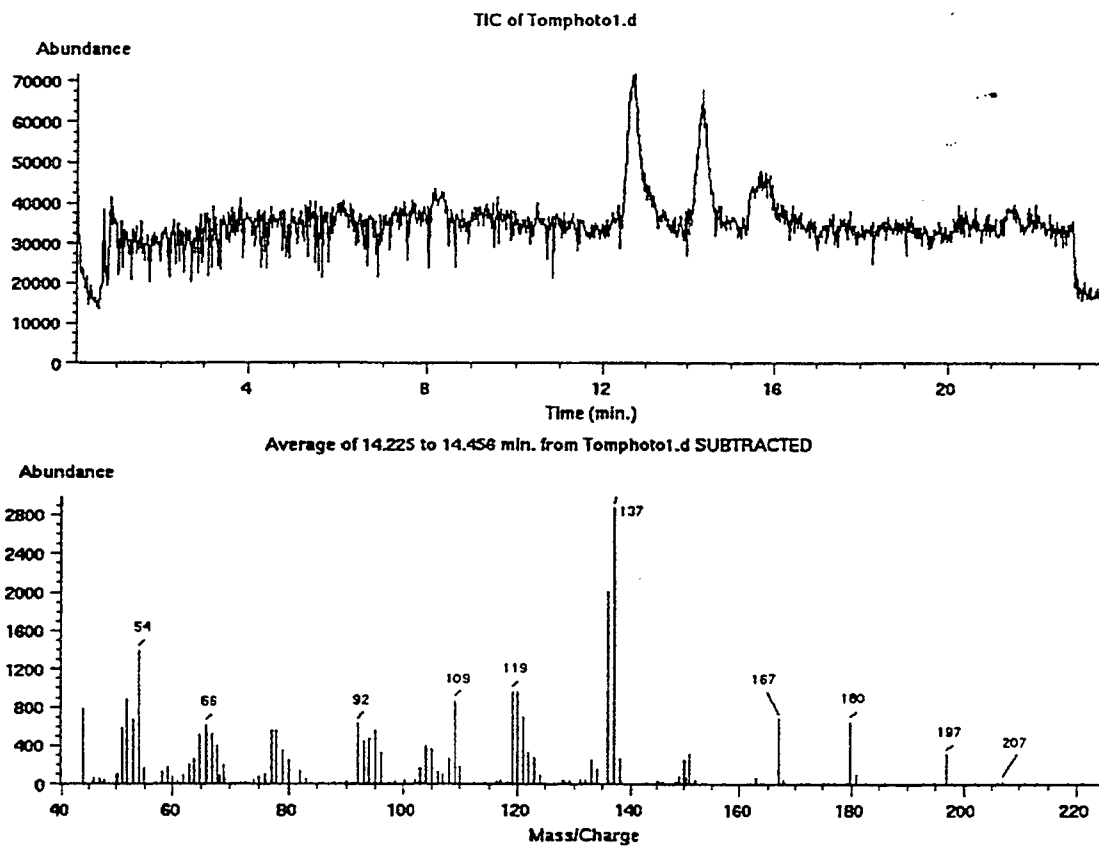


Figure 5. Chromatogram (top) and mass spectra (bottom) of an amino-nitrotoluene isomer formed during photocatalytic degradation of TNT in Experiment 1.

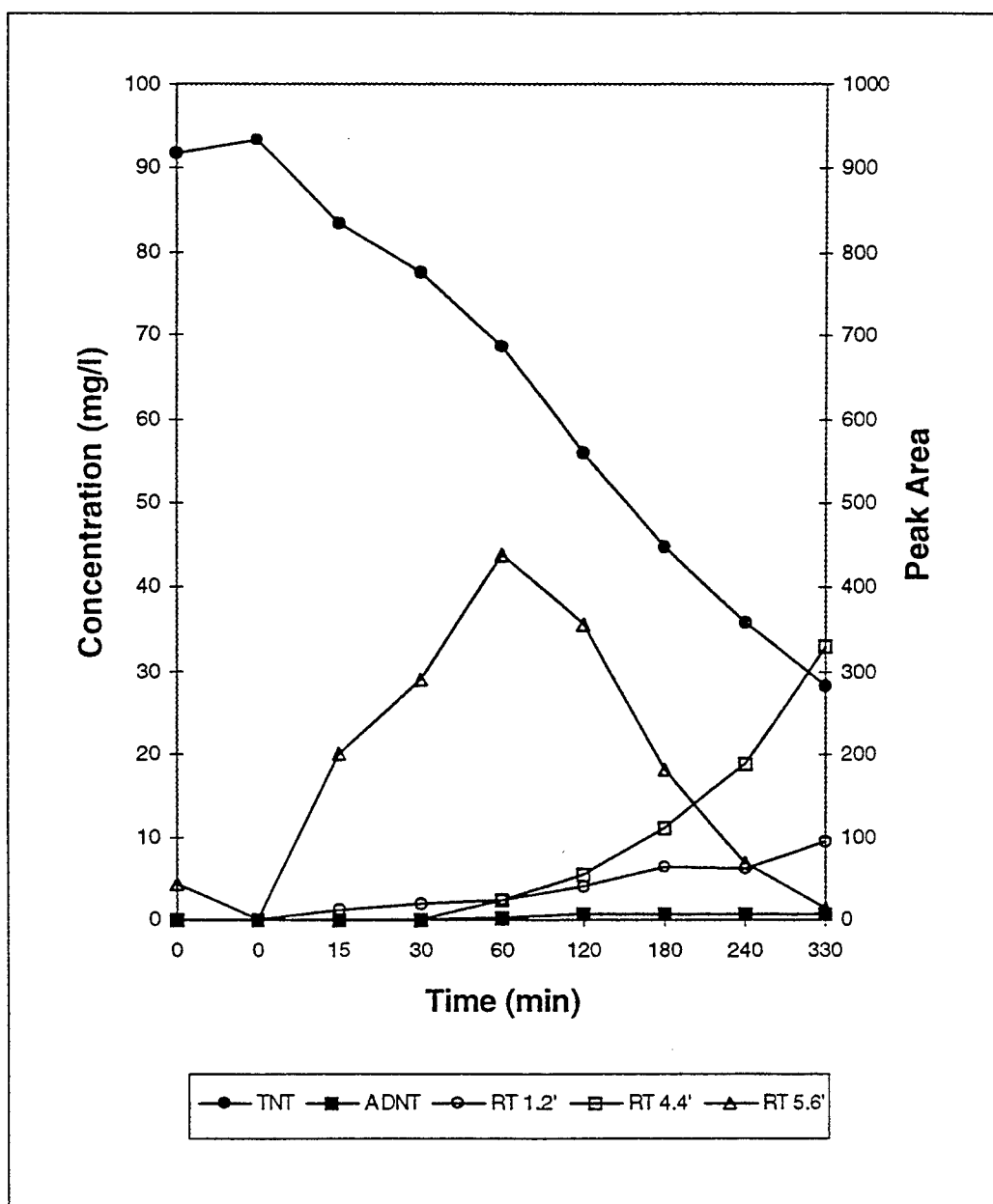


Figure 6. Anaerobic photocatalytic degradation of TNT in experiment 2. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

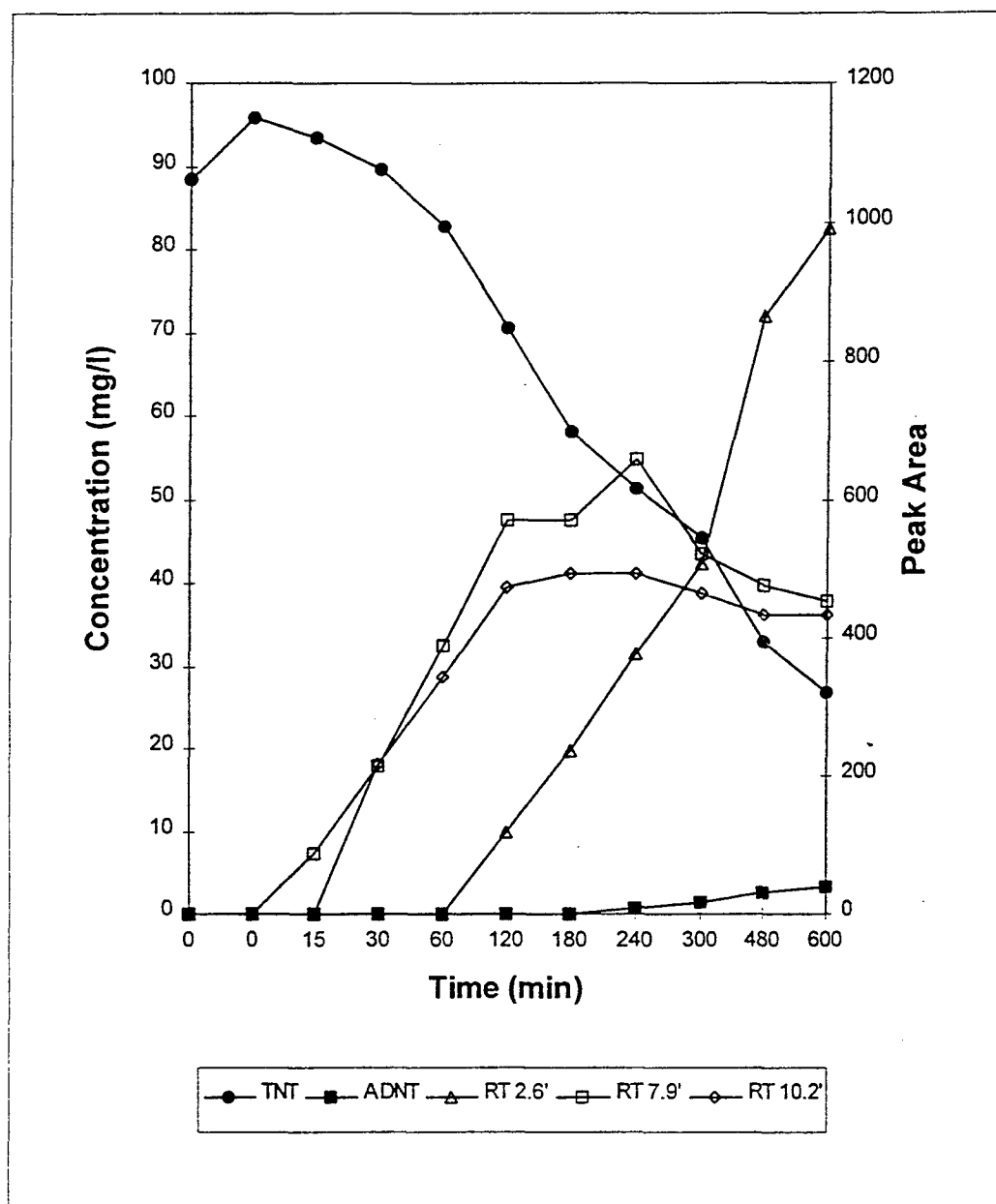


Figure 7. Anaerobic photocatalytic degradation of TNT in the presence of HCO_3^- in experiment 3. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

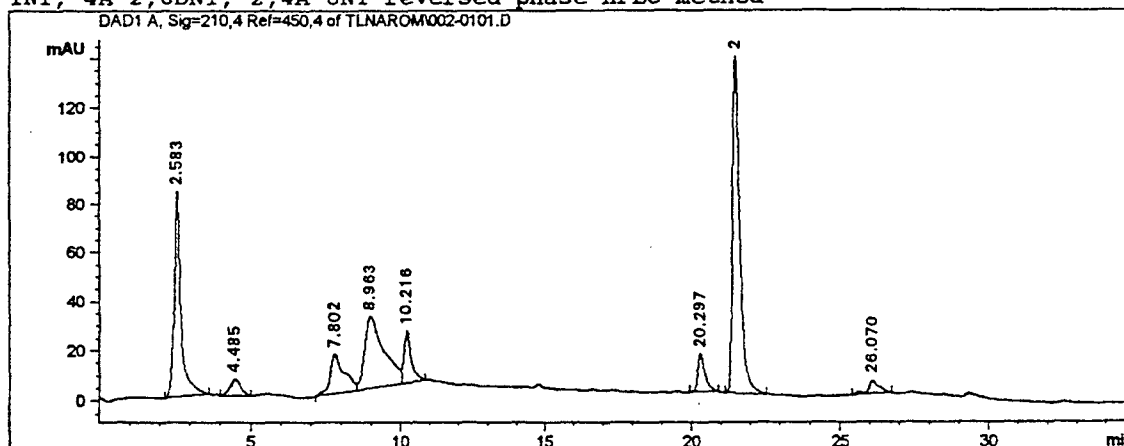
Sample Name : photo3-smp11

Inj Volume : 10 µl

Sequence File : C:\HPCHEM\1\SEQUENCE\TL42095.S

Analysis Method : C:\HPCHEM\1\METHODS\TAL.M

TNT, 4A-2,6DNT, 2,4A-6NT reversed phase HPLC method



External Standard Report

Sorted by Retention Time

Calib. Data Modified : Wednesday, April 19, 1995 1:51:30 AM

Multiplier : 1.000000

RF Uncal. Peaks : 0.000000

Signal 1: DAD1 A, Sig=210,4 Ref=450,4

RT [min]	Sig	Type	Area [mAU*sec]	Amt/Area	Amount [mg/L]	Grp	Name
2.583	1	BB	1271.49133	0.00000	0.00000	?	
4.485	1	BV	163.04066	0.00000	0.00000	?	
7.802	1	BV	515.74963	0.00000	0.00000	?	
8.963	1	VV	1319.96631	0.00000	0.00000	?	
10.216	1	VB	375.41434	0.00000	0.00000	?	
15.161	1	*	not found	*			2,4-DANT
20.297	1	BB	264.28940	1.24508e-2	3.29062		4A, 2,6DNT
21.489	1	BB	2383.32373	1.13362e-2	27.01776		TNT
26.070	1	BB	134.58821	0.00000	0.00000	?	

Totals : 30.30838

1 Warnings or Errors :

Warning : Calibrated compound(s) not found

Figure 8. Chromatogram from anaerobic photocatalytic degradation of TNT in the presence of HCO_3^- during Experiment 3.

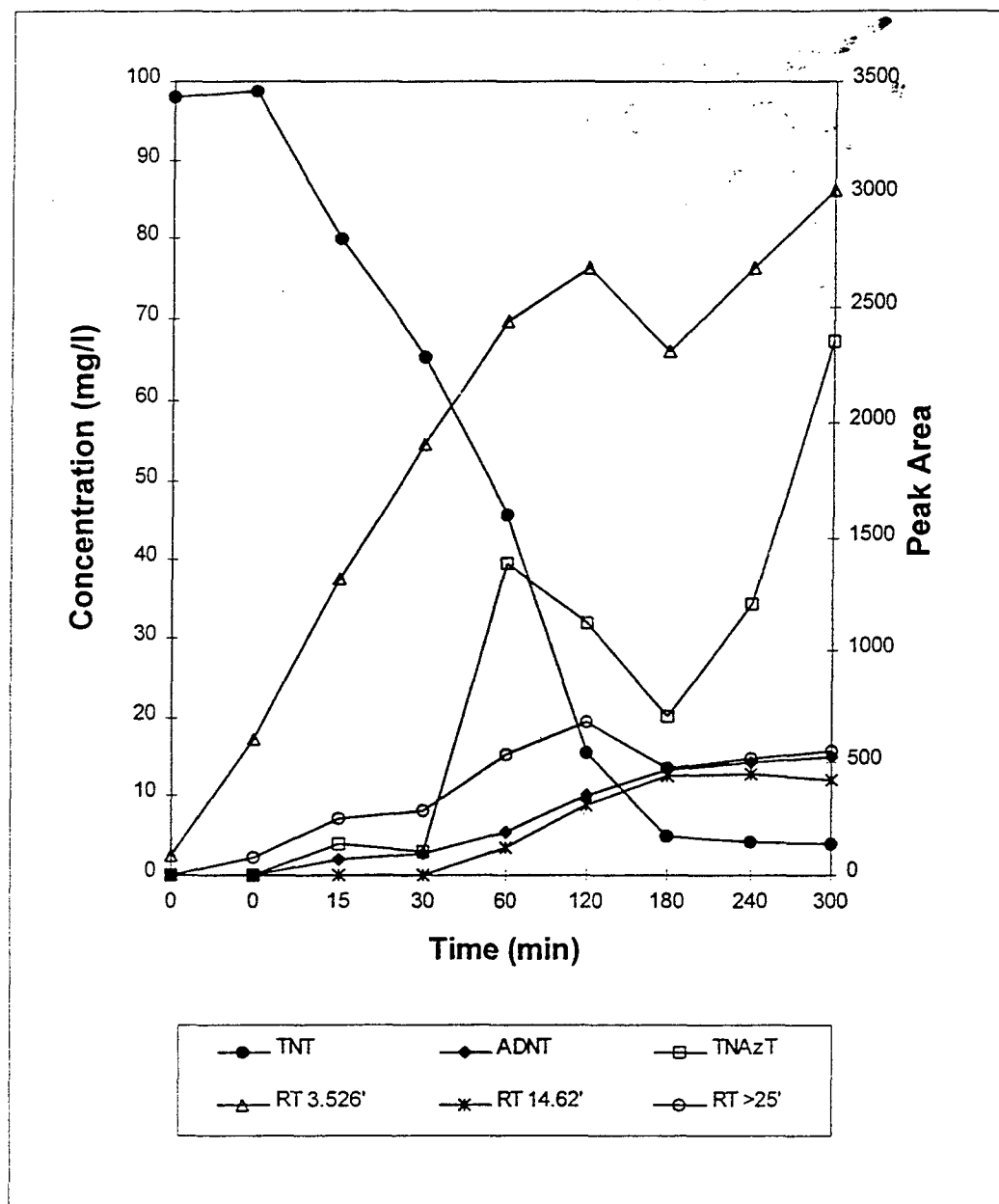


Figure 9. Anaerobic photocatalytic degradation of TNT in the presence of HCO_3^- and EDTA in experiment 4. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

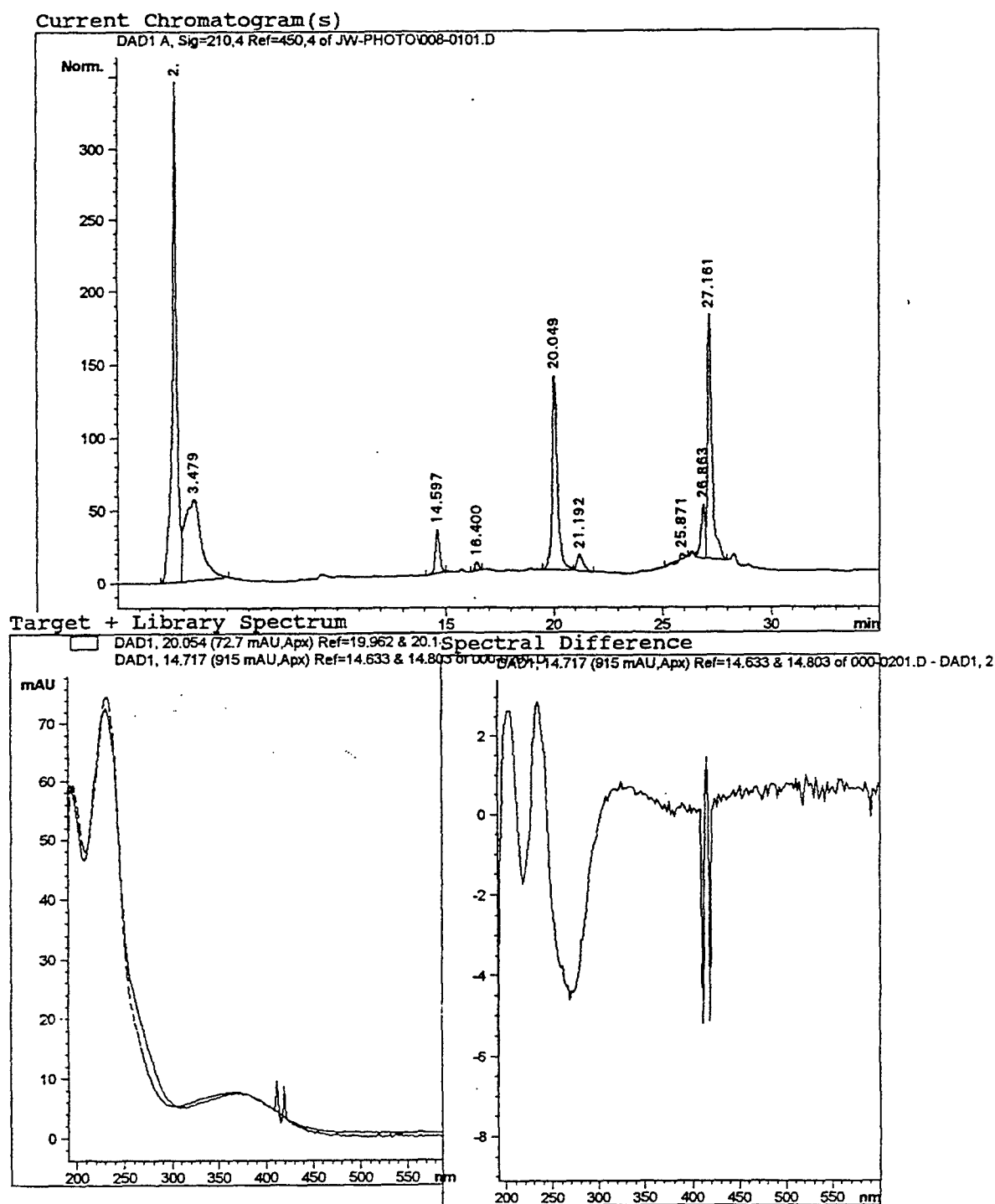


Figure 10. Chromatogram (upper) from Experiment 4 showing the spectral match (lower left) and spectral differences (lower right) for intermediate, RT = 20.049 min, identified as 4-amino-2,6-dinitrotoluene (ADNT).

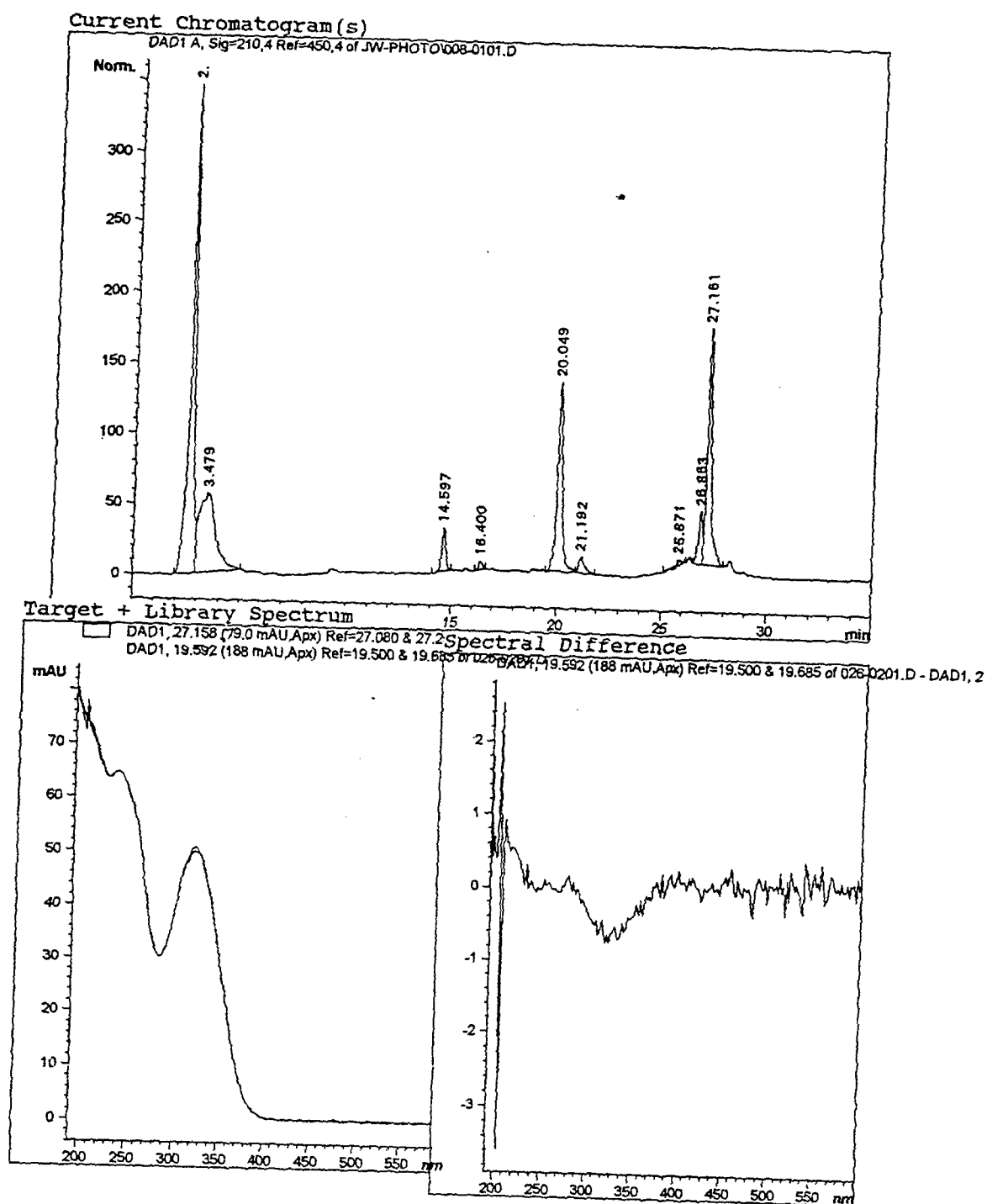


Figure 11. Chromatogram (upper) from Experiment 4 showing spectral match (lower left) and spectral difference (lower right) for intermediate, RT = 27.16 min, identified as 2,2',6,6'-tetranitro-4,4'-azoxy toluene (TNAzT).

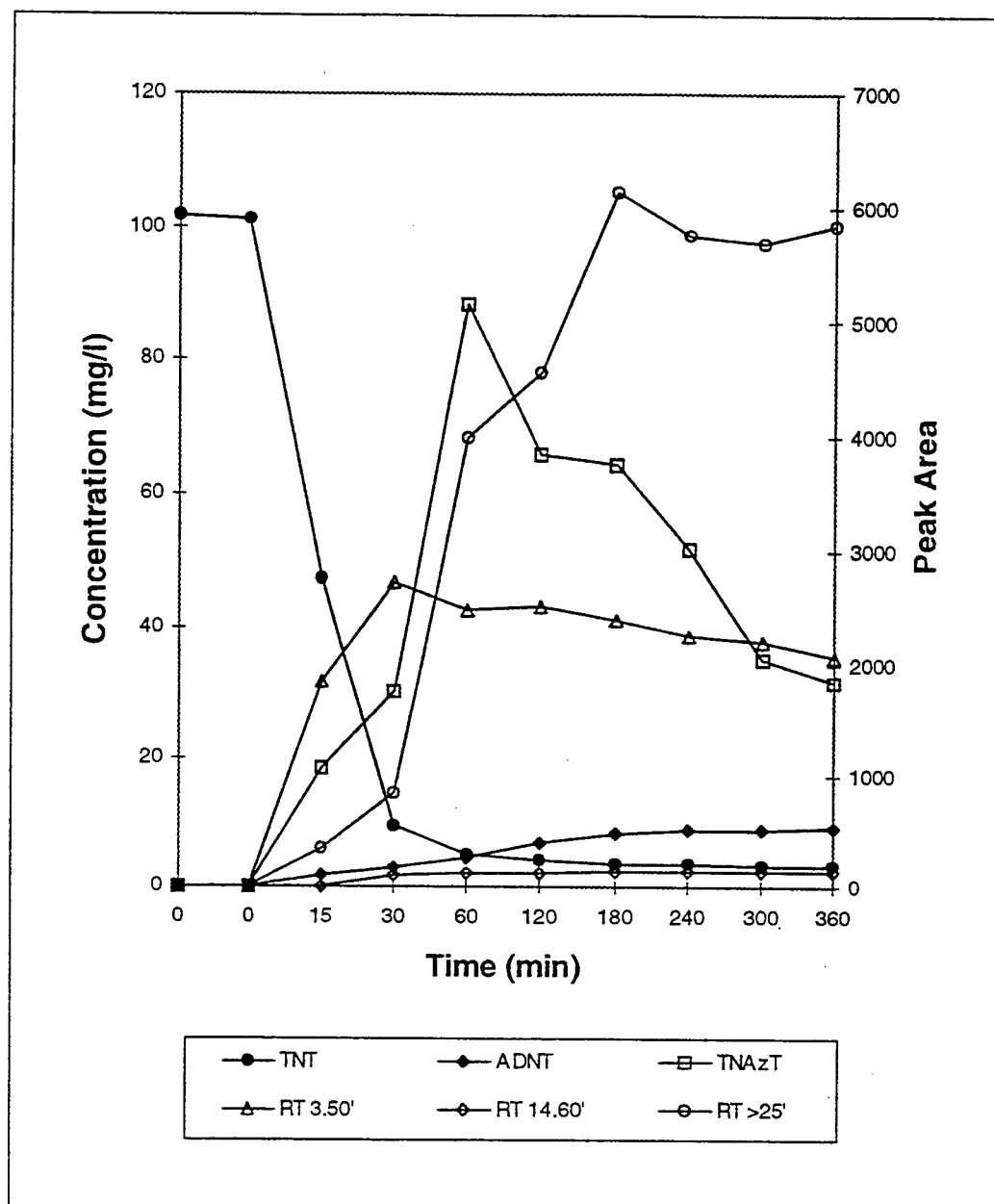
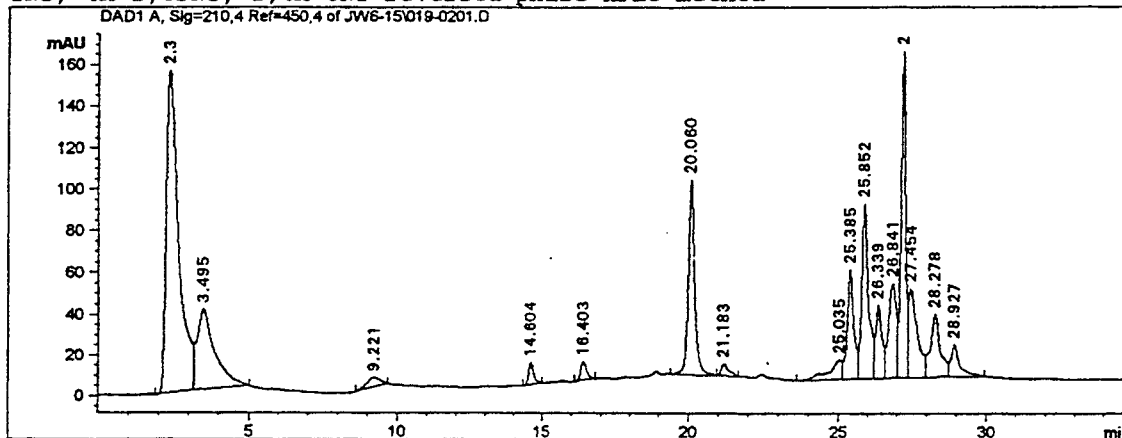


Figure 12. Anaerobic photocatalytic degradation of TNT in the presence of EDTA in experiment 5. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

Acq. Method : TAL612.M Seq. Line : 2
 Acq. Operator : jWells Vial : 19
 Injection Date : 6/17/95 1:02:59 AM Inj : 1
 Sample Name : 6-15smp1-10d Inj Volume : 10 µl

Sequence File : C:\HPCHEM\1\SEQUENCE\JW6-15.S
 Analysis Method : C:\HPCHEM\1\METHODS\TAL612.M
 TNT, 4A-2,6DNT, 2,4A-6NT reversed phase HPLC method



External Standard Report

Sorted by Retention Time

Calib. Data Modified : Wednesday, June 14, 1995 7:57:29 AM
 Multiplier : 1.000000
 RF Uncal. Peaks : 0.000000

Signal 1: DAD1 A, Sig=210,4 Ref=450,4

RT [min]	Sig	Type	Area [mAU*sec]	Amt/Area	Amount [mg/L]	Grp	Name
2.371	1	BV	4472.12207	0.00000	0.00000	?	
3.495	1	BB	1651.51758	0.00000	0.00000	?	
9.221	1	VB	150.13554	0.00000	0.00000	?	
14.604	1	BB	140.55016	0.00000	0.00000	?	
15.424	1	*	not found	*			2,4-DANT
16.403	1	BB	122.31805	0.00000	0.00000	?	
20.060	1	BB	1389.71936	6.60070e-3	9.17312		4A, 2,6DNT
21.183	1	BB	100.42084	3.25642e-2	3.27012		TNT
25.035	1	BV	290.10065	0.00000	0.00000	?	
25.385	1	VV	831.85468	0.00000	0.00000	?	
25.852	1	VV	1370.25330	0.00000	0.00000	?	
26.339	1	VV	547.01031	0.00000	0.00000	?	
26.841	1	VV	845.85693	0.00000	0.00000	?	
27.163	1	VV	1802.58618	0.00000	0.00000	?	
27.454	1	VV	899.49219	0.00000	0.00000	?	
28.278	1	VV	717.41400	0.00000	0.00000	?	
28.927	1	VB	328.31653	0.00000	0.00000	?	

Figure 13. Chromatogram from anaerobic photocatalytic degradation of TNT in the presence of EDTA in Experiment 5.

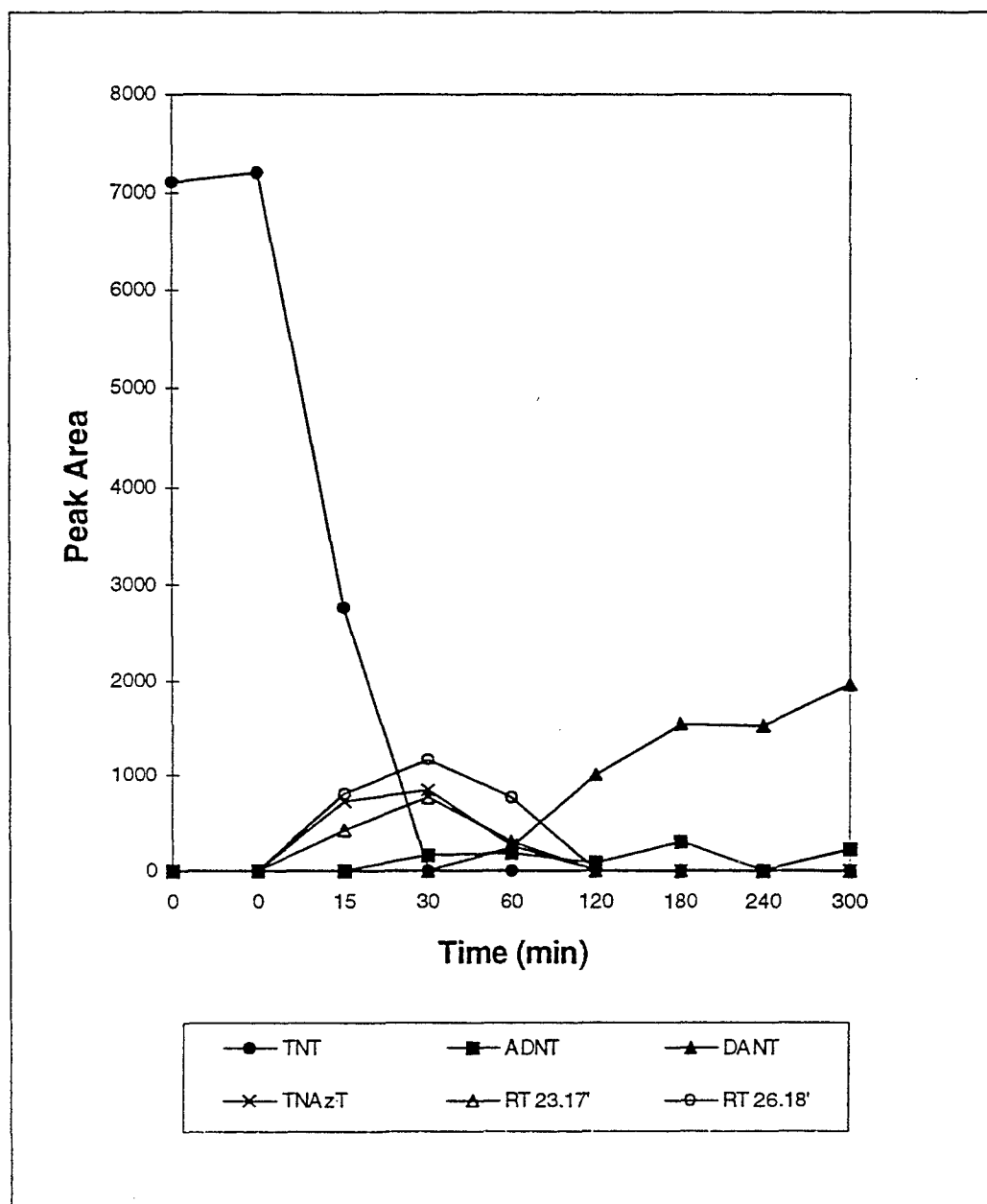


Figure 14. Anaerobic photocatalytic degradation of TNT in the presence of EDTA in experiment 10. All data are given as peak area from chromatograms.

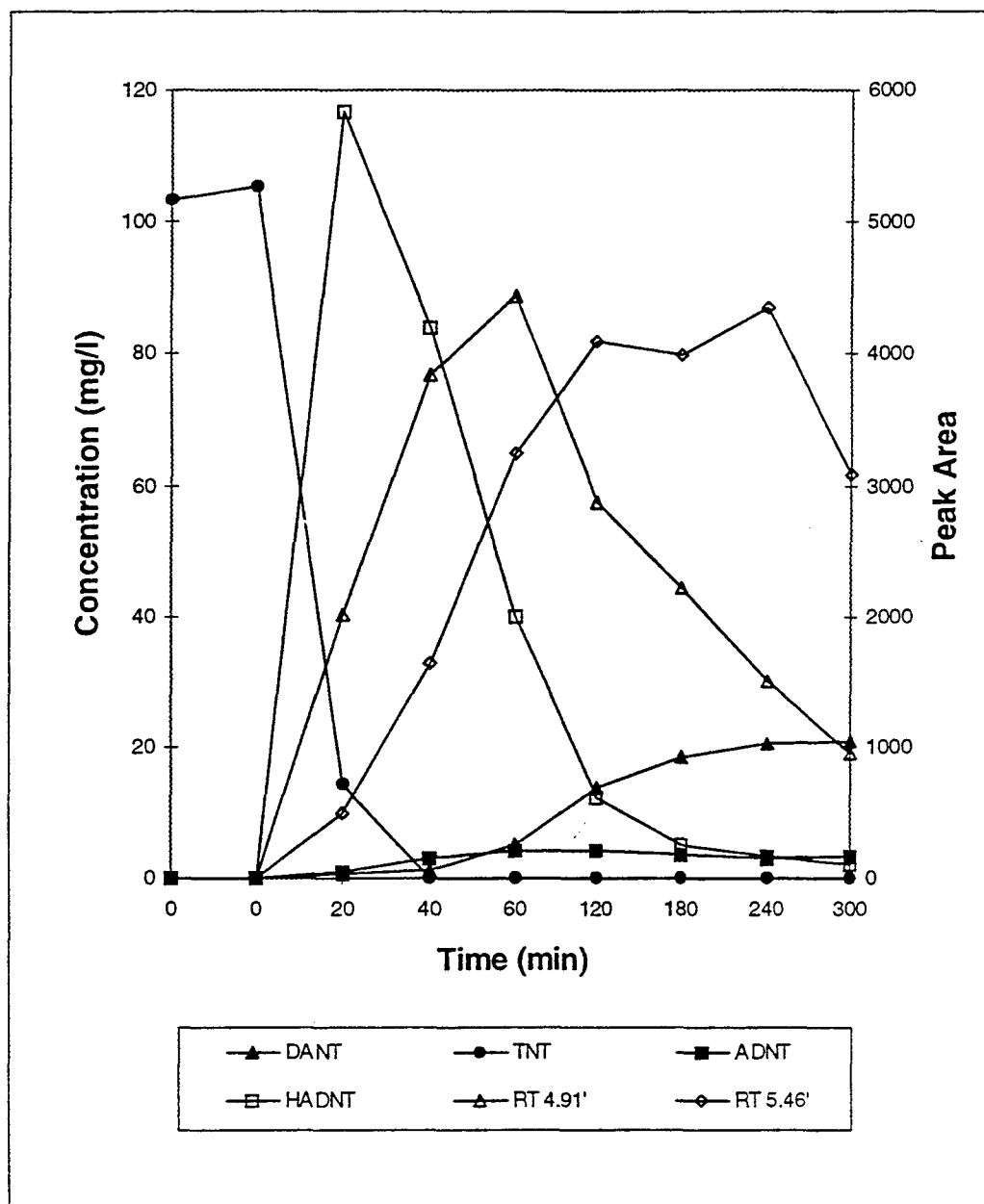


Figure 15. Anaerobic photocatalytic degradation of TNT in the presence of EDTA in experiment 11. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

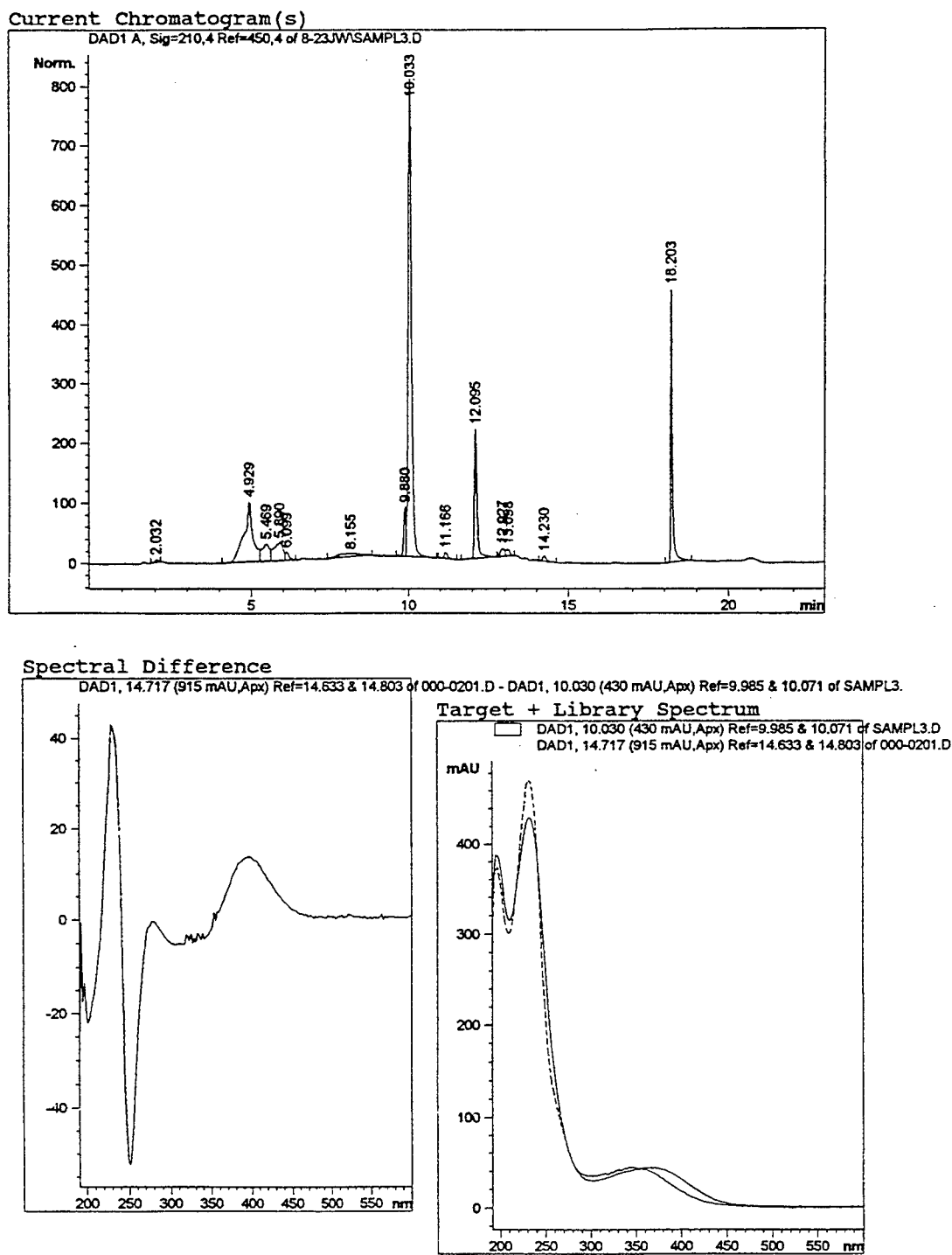


Figure 16. Chromatogram (upper) from Experiment 11 showing spectral match (lower right) and spectral difference (lower left) for intermediate, RT = 10.033 min, identified as 4-hydroxylamino-2,6-dinitrotoluene (HADNT).

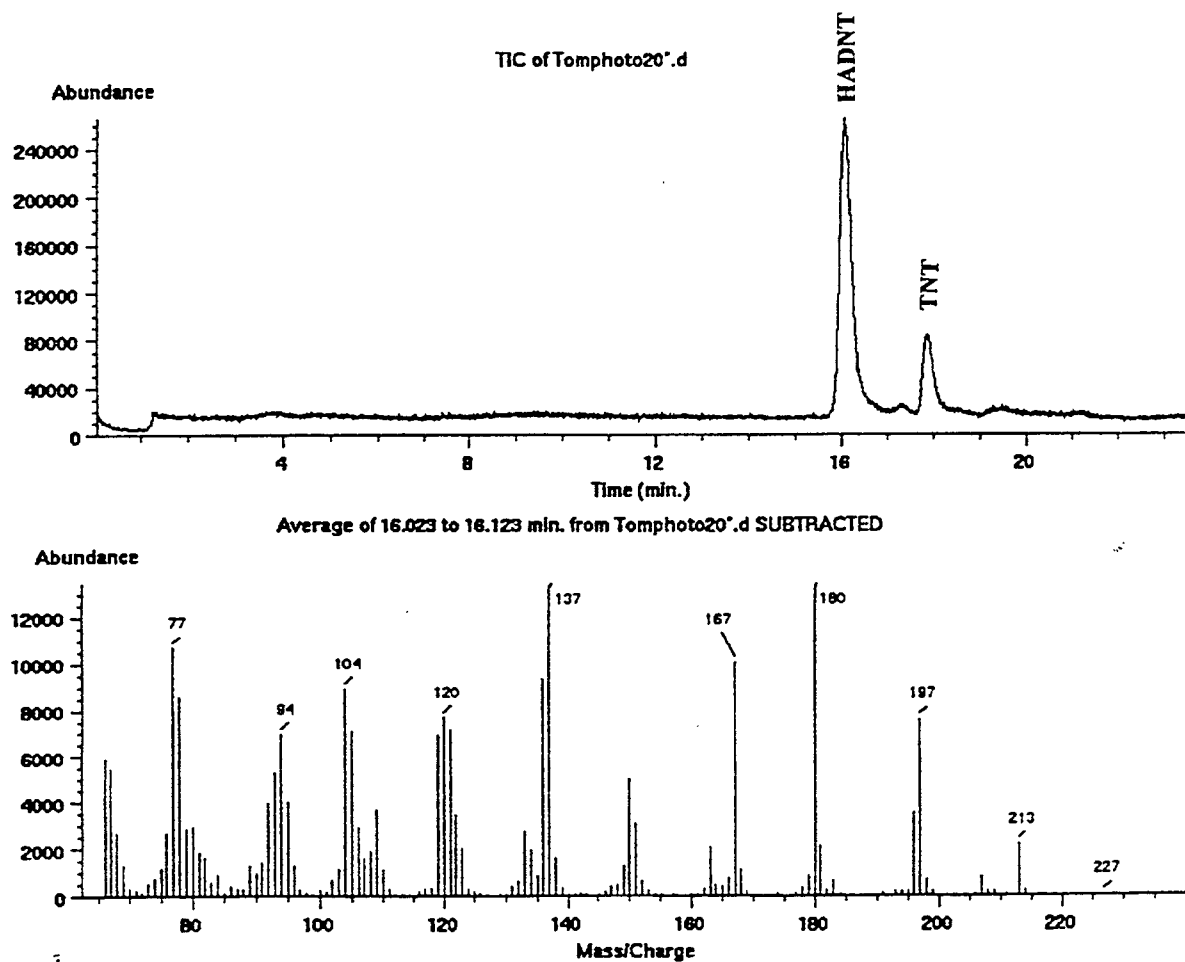


Figure 17. Chromatogram (top) and mass spectra (bottom) of 4-hydroxylamino-2,6-dinitrotoluene (HADNT), an intermediate formed during anaerobic photocatalytic degradation of TNT in the presence of EDTA, in Experiment 11.

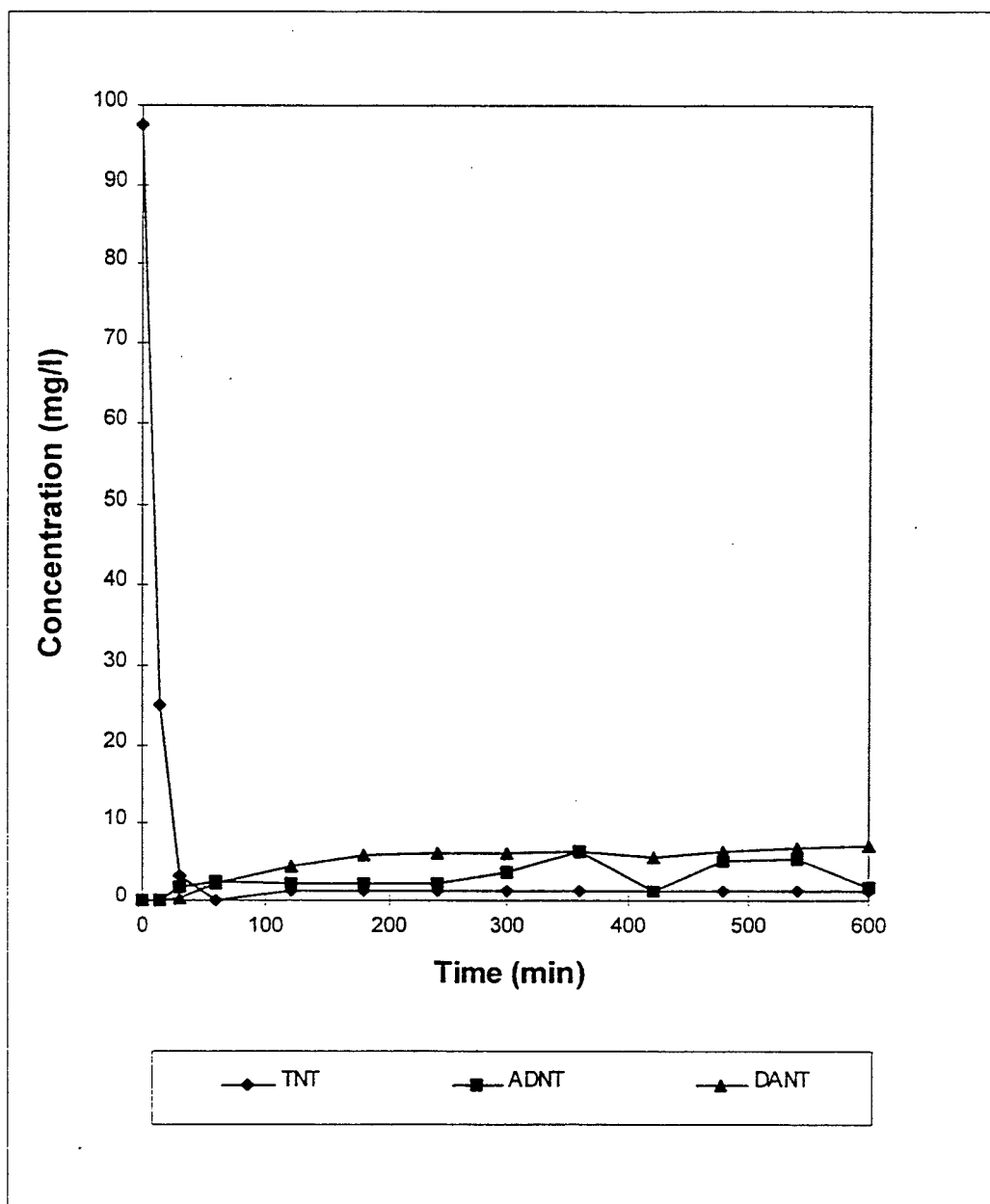


Figure 18. Anaerobic photocatalytic degradation of TNT the presence of EDTA in experiment 16.

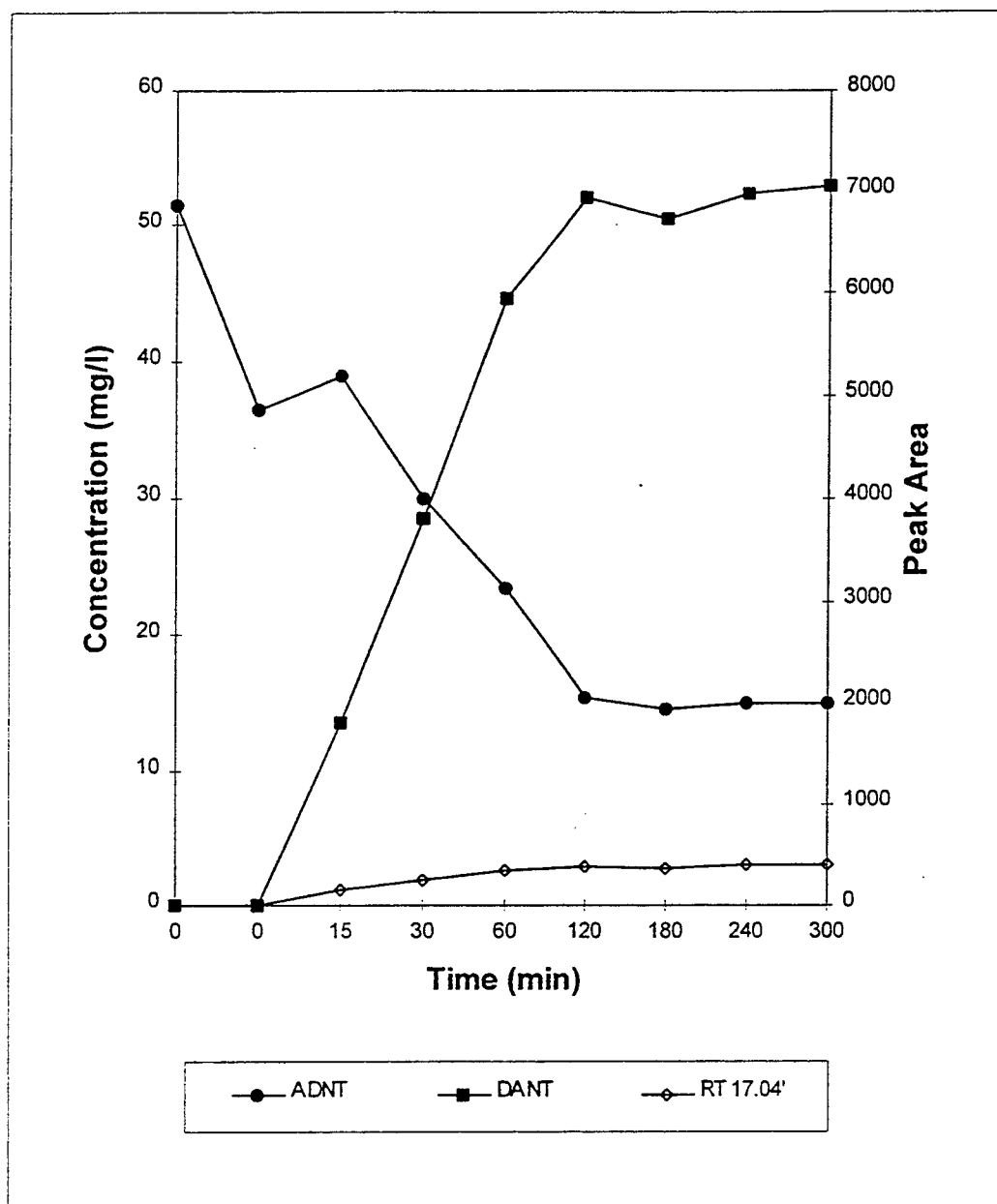


Figure 19. Anaerobic photocatalytic degradation of ADNT in the presence of EDTA in experiment 6. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

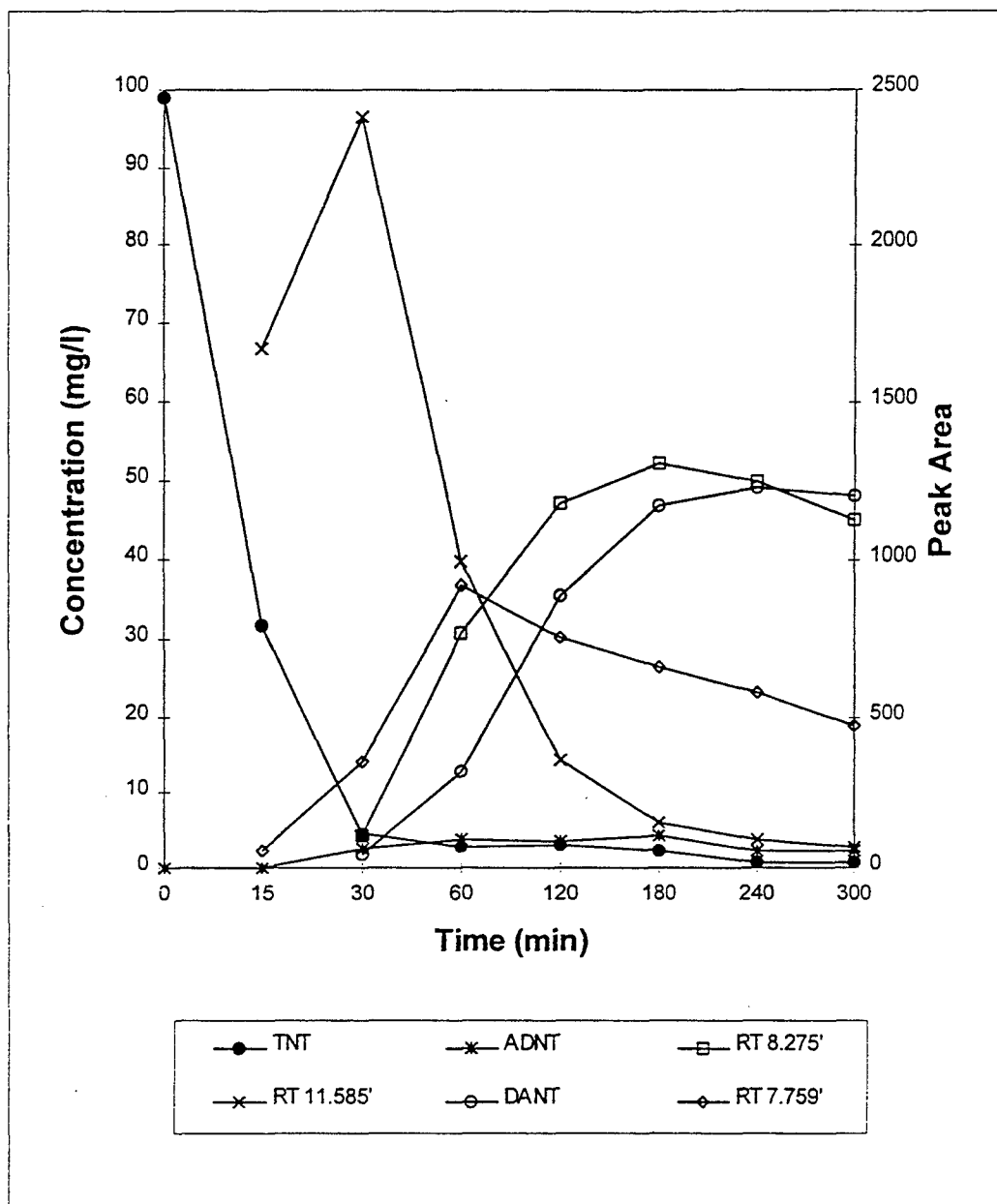


Figure 21. Partially anaerobic photocatalysis of TNT in the presence of EDTA in experiment 17. Only an initial nitrogen purge of the reactor contents was done.

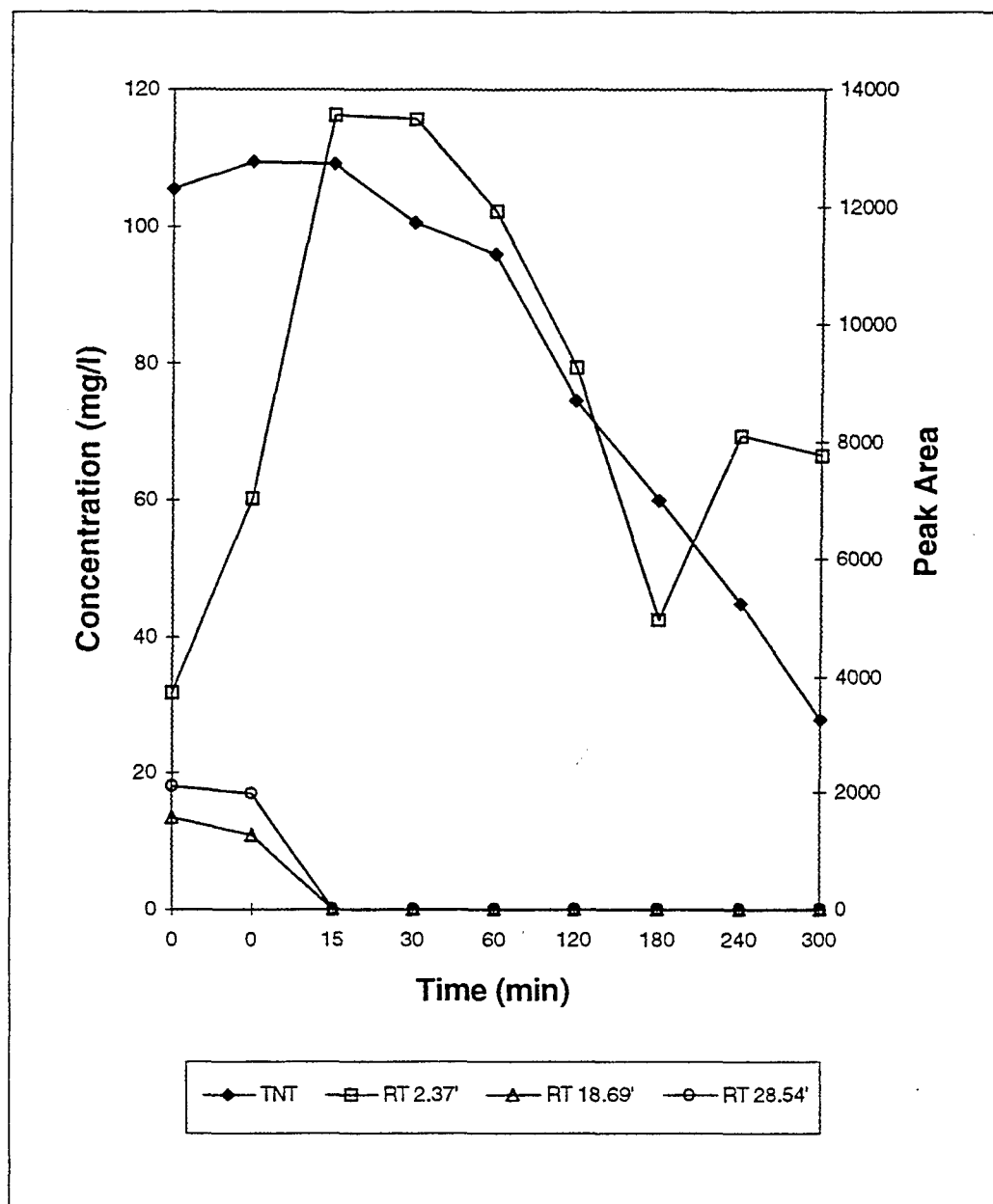
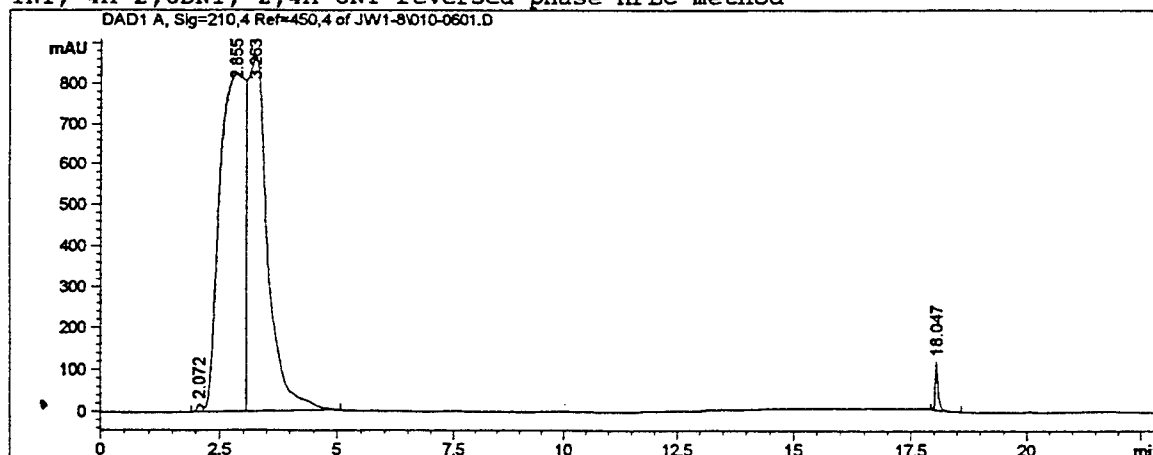


Figure 22. Anaerobic photocatalytic degradation of TNT in the presence of tetranitromethane and EDTA in experiment 8. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

Acq. Method : PROD2.M Seq. Line : 6
 Acq. Operator : JWells Vial : 10
 Injection Date : 1/8/96 9:30:43 PM Inj : 1
 Sample Name : samp6 Inj Volume : 10 µl

Sequence File : C:\HPCHEM\1\SEQUENCE\JW1-8.S
 Analysis Method : C:\HPCHEM\1\METHODS\PROD2.M
 TNT, 4A-2,6DNT, 2,4A-6NT reversed phase HPLC method



External Standard Report

Sorted by Retention Time
 Calib. Data Modified : Monday, January 08, 1996 7:01:55 PM
 Multiplier : 1.000000
 Uncalibrated Peaks : not reported

Signal 1: DAD1 A, Sig=210,4 Ref=450,4

RT [min]	Sig	Type	Area [mAU*sec]	Amt/Area	Amount [mg/L]	Grp	Name
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11.880	1	*	not found	*			4A,2,6-DNT
12.331	1	*	not found	*			TNT

Totals : 0.00000

1 Warnings or Errors :

Warning : Calibrated compound(s) not found

Figure 23. Chromatogram from anaerobic photocatalytic degradation of TNT in the presence of periodate in Experiment 13.

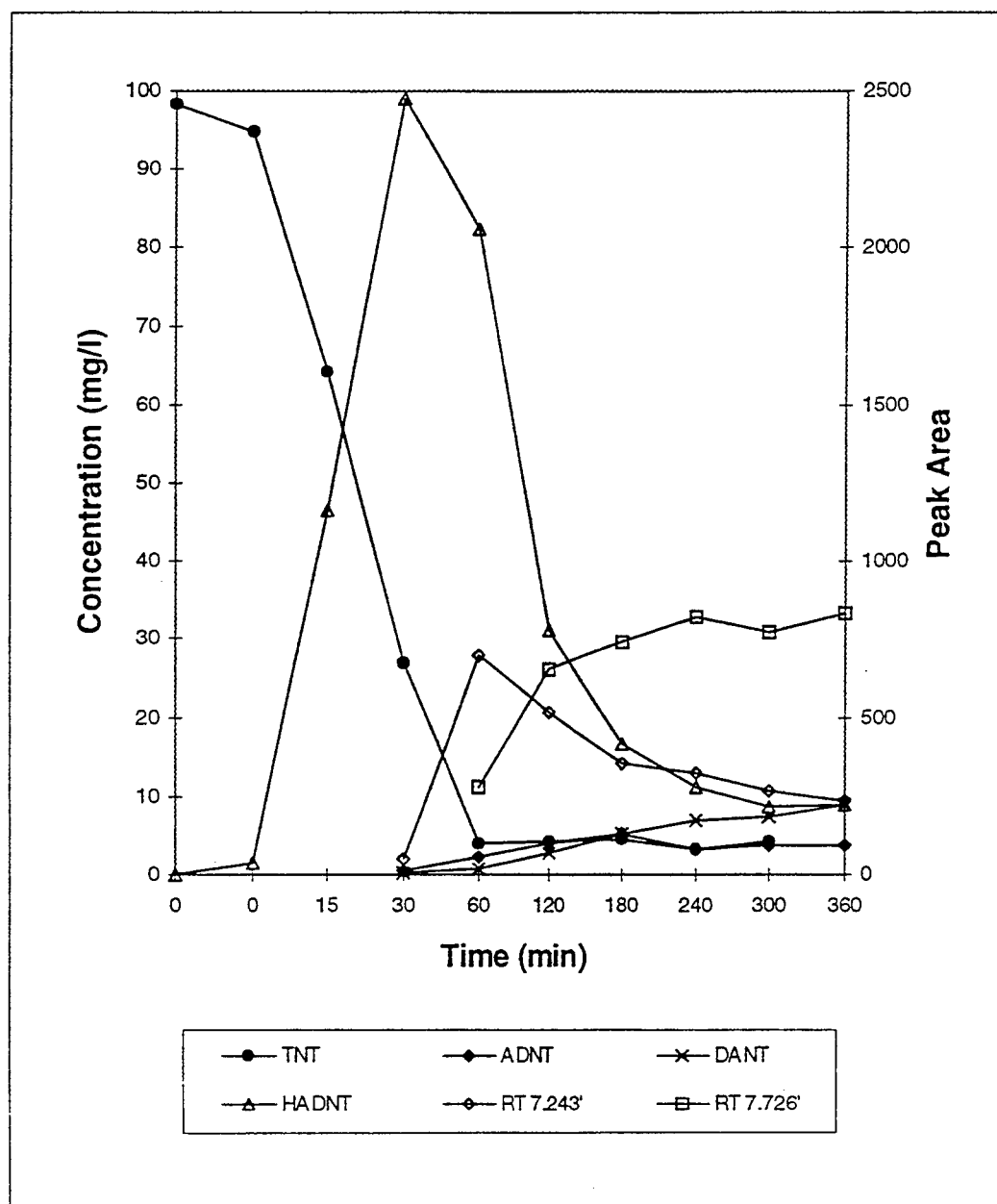
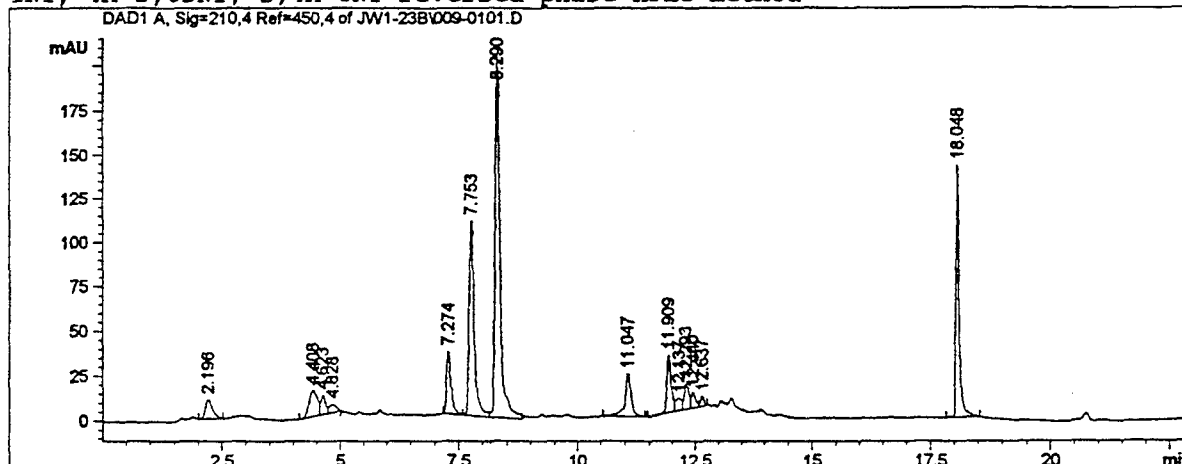


Figure 24. Anaerobic photocatalytic degradation of TNT in the presence of methanol and EDTA in experiment 14. Solid symbols represent data in mg/l, open symbols show data as peak area from chromatograms.

Acq. Method : PROD2.M Seq. Line : 1
 Acq. Operator : wells Vial : 9
 Injection Date : 1/24/96 2:51:13 AM Inj : 1
 Sample Name : sample 10 Inj Volume : 10 µl

Sequence File : C:\HPCHEM\1\SEQUENCE\JW1-23.S
 Analysis Method : C:\HPCHEM\1\METHODS\PROD2.M
 TNT, 4A-2,6DNT, 2,4A-6NT reversed phase HPLC method



External Standard Report

Sorted by Retention Time
 Calib. Data Modified : Tuesday, January 23, 1996 9:22:46 PM
 Multiplier : 1.000000
 RF Uncal. Peaks : 0.000000

Signal 1: DAD1 A, Sig=210,4 Ref=450,4

RT [min]	Sig	Type	Area [mAU*sec]	Amt/Area	Amount [mg/L]	Grp	Name
2.196	1	PV	120.77103	0.00000	0.00000	?	
4.408	1	BV	182.77753	0.00000	0.00000	?	
4.623	1	VV	80.90977	0.00000	0.00000	?	
4.828	1	VV	51.15357	0.00000	0.00000	?	
7.274	1	BV	237.92986	0.00000	0.00000	?	
7.753	1	VV	831.02484	0.00000	0.00000	?	
8.290	1	VV	1454.01855	6.08068e-3	8.84142		2,6DA4-NT
11.047	1	BB	224.28879	0.00000	0.00000	?	
11.909	1	BV	211.89180	1.73433e-2	3.67490		4A,2,6-DNT
12.137	1	VV	62.31731	0.00000	0.00000	?	
12.293	1	VV	105.04922	4.47929e-2	4.70546		TNT
12.440	1	VV	59.31694	0.00000	0.00000	?	
12.637	1	VV	31.83701	0.00000	0.00000	?	
18.048	1	PB	646.27380	0.00000	0.00000	?	

Totals : 17.22178

Figure 25. Chromatogram from anaerobic photocatalytic degradation of TNT in the presence of methanol in Experiment 14.

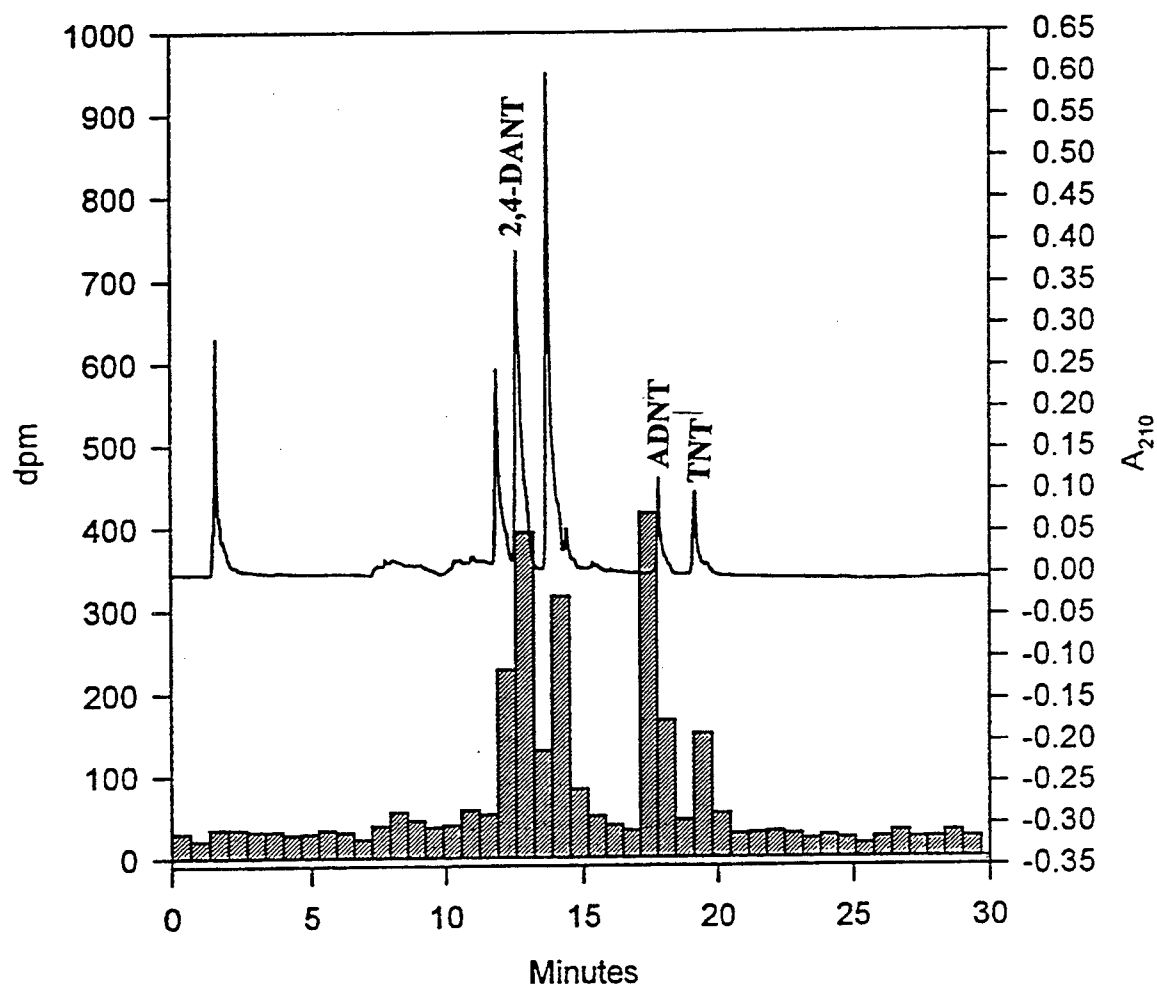


Figure 26. Chromatographic peaks (A_{210}) and associated ^{14}C fractionation (dpm) from photocatalytic degradation of TNT in the radiotracer study of Experiment 9.

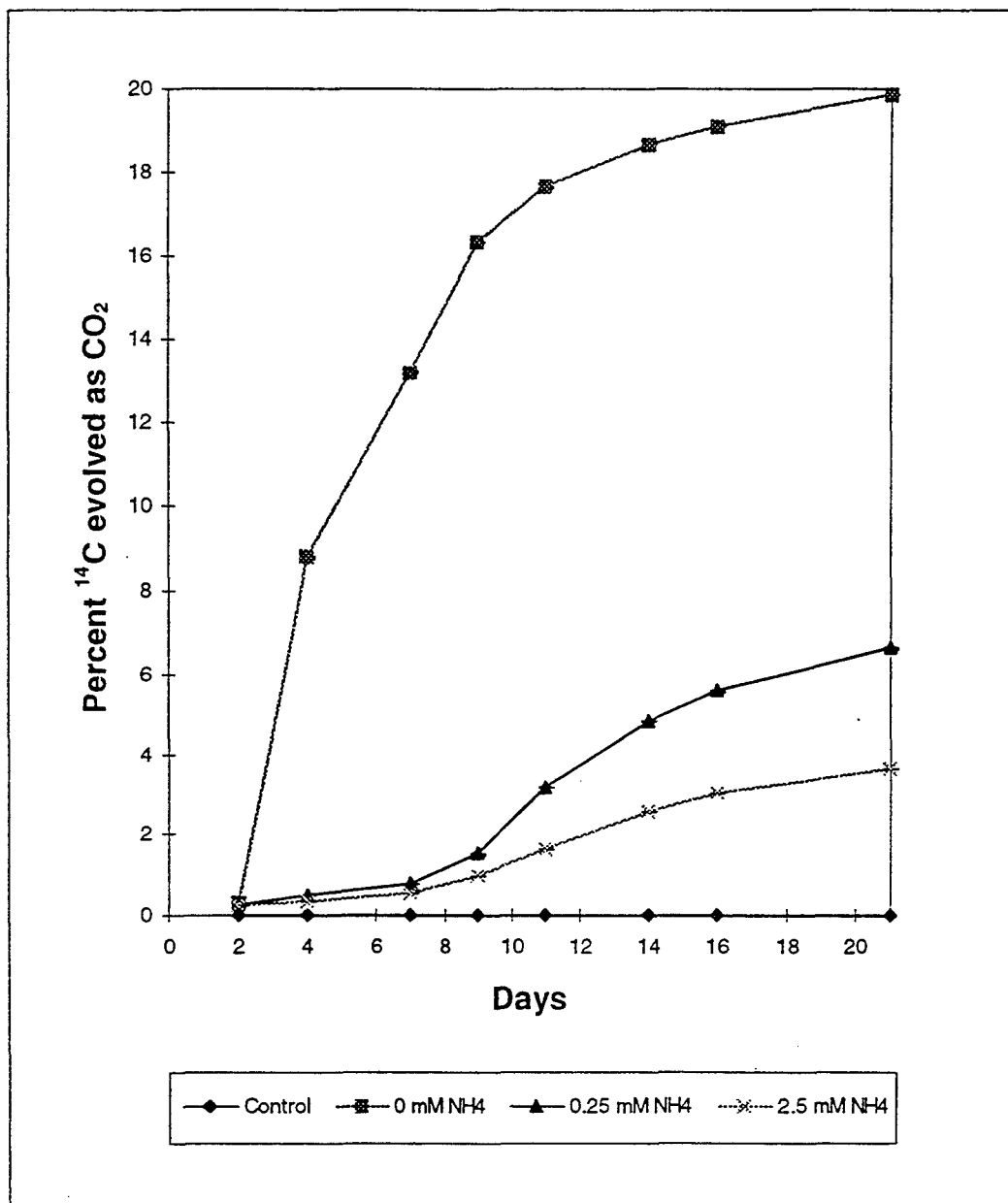


Figure 27. Evolution of ^{14}C - CO_2 over time from aerobic biological degradation of photocatalytic TNT degradation products by *P. chrysosporium* in experiment 12. Graph shows four treatments: a non-inoculated control, and exogenous additions of 0, 0.25 and 2.5 mM NH_4 .

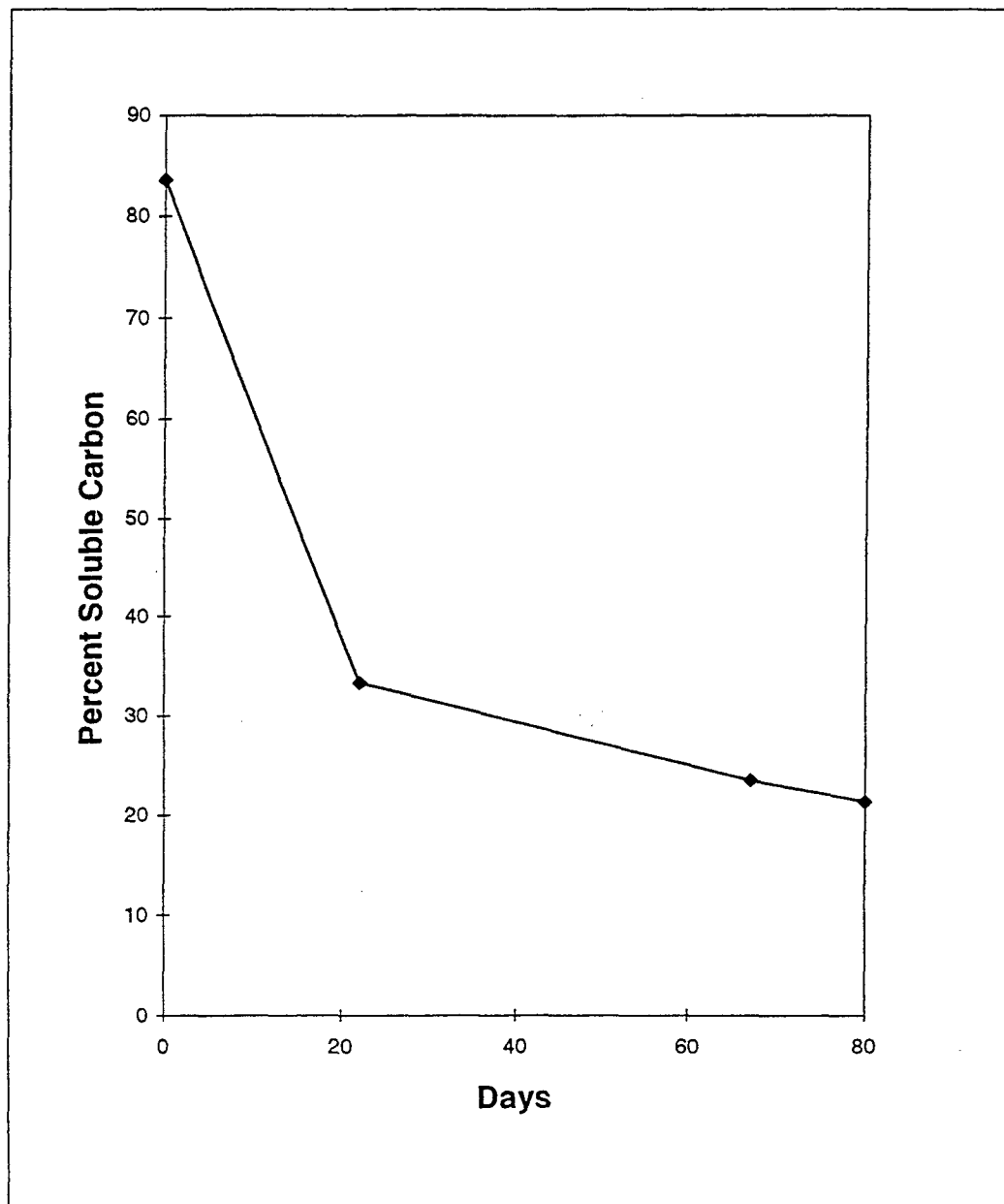


Figure 28. Loss of soluble photoproducts from solution over time.

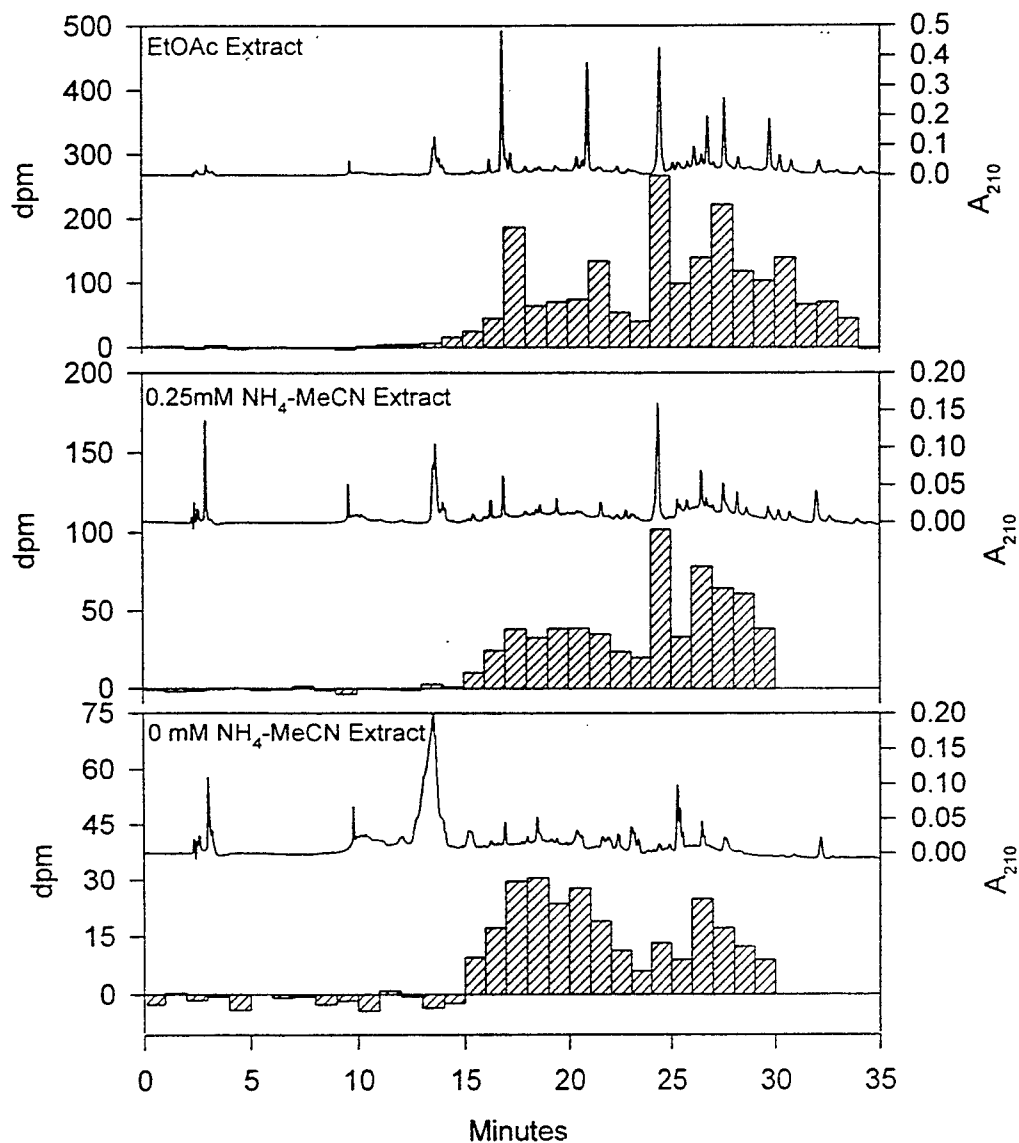


Figure 29. Chromatographic peaks (A_{210}) with associated ^{14}C fractionation (dpm) of photoreactor products after photocatalytic treatment of TNT prior to biological treatment (top graph) and after aerobic biological treatment by *P. chrysosporium* in the presence of 0.25 mM NH_4 (middle graph) and no NH_4 (bottom graph), all in experiment 12.

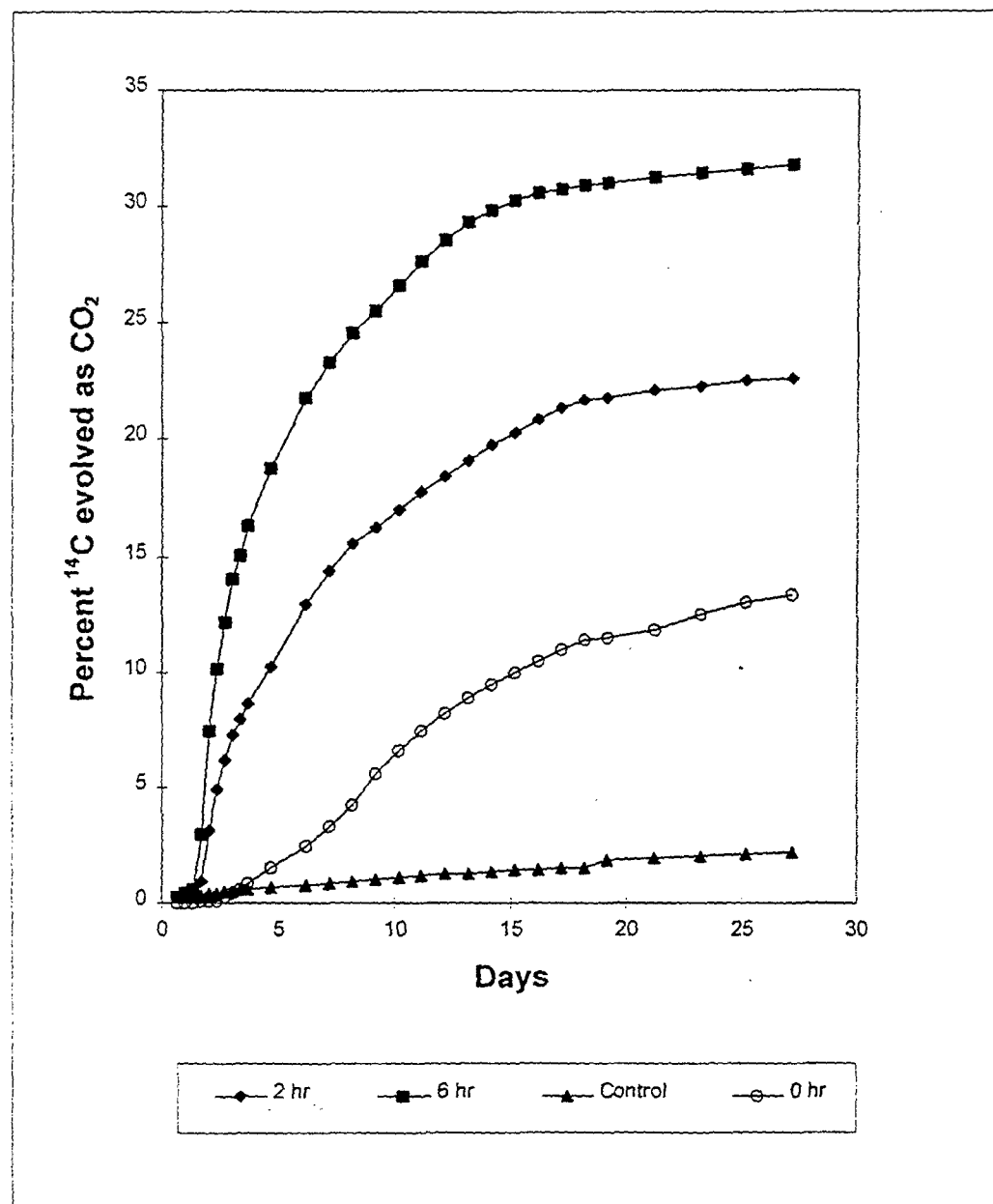


Figure 30. Evolution of ^{14}C - CO_2 over time from aerobic biological degradation of photocatalytic TNT degradation products by *P. chrysosporium* in experiment 15. Graph shows four treatments: a non-inoculated control, no photocatalytic treatment (0 hr) and 2 and 6 hours of photocatalysis.

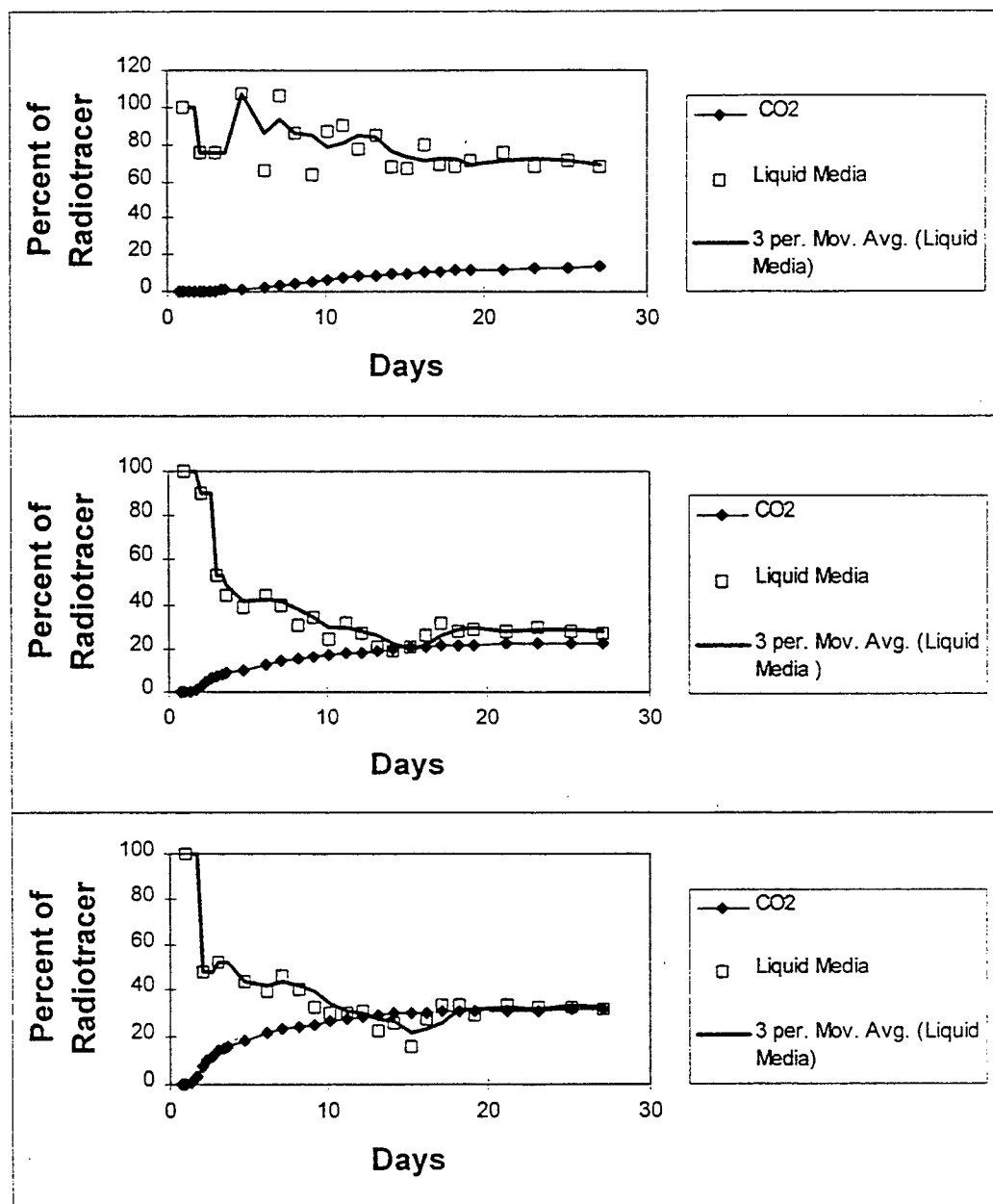


Figure 31. Distribution of ^{14}C radiotracer between evolved CO_2 and that soluble in liquid media during biological treatment of photoreactor contents by *P. chrysosporium* in experiment 15. Graphs show (top to bottom): 0, 2 and 6 hours of photocatalytic treatment, respectively, prior to biological treatment.

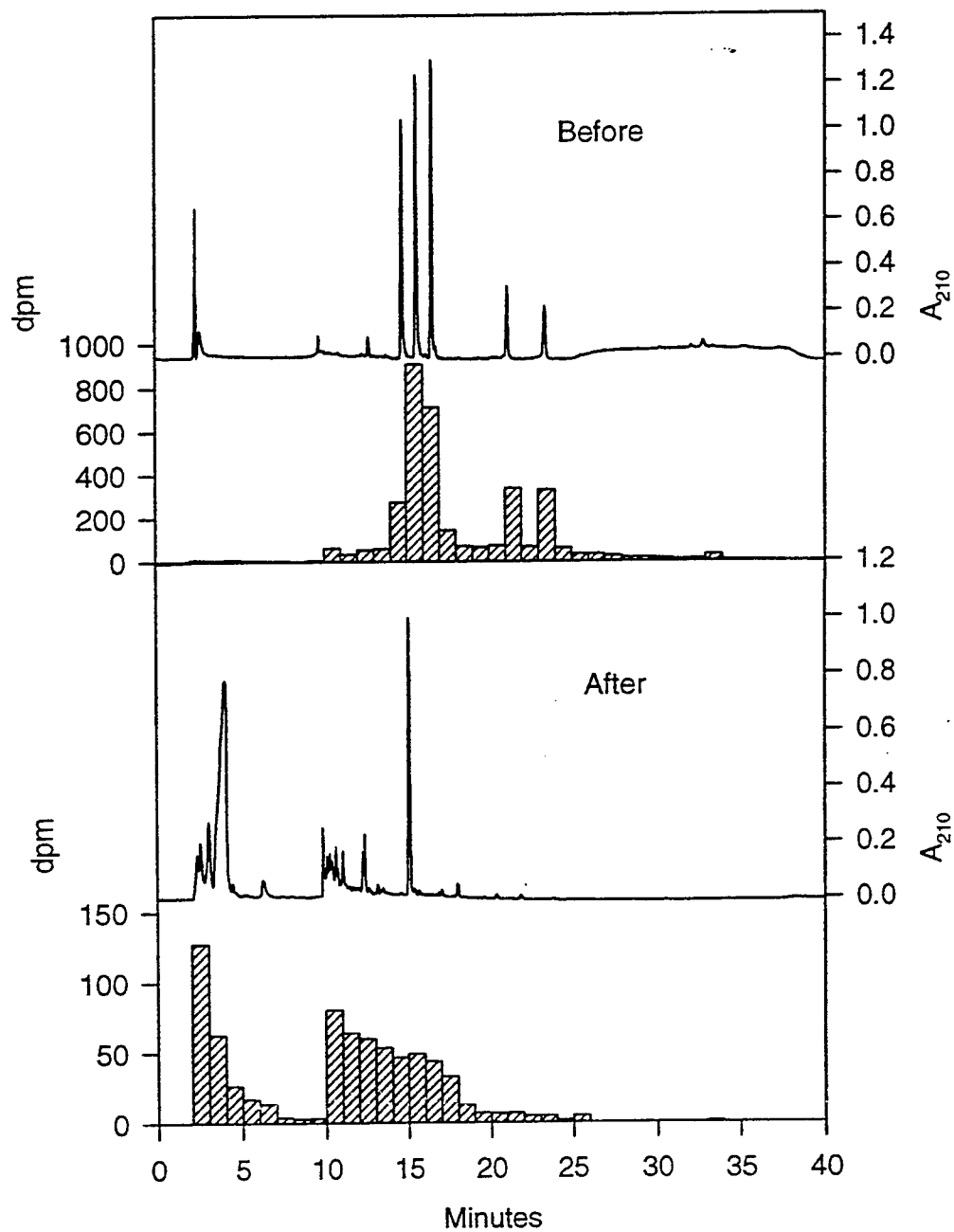


Figure 32. Chromatographic peaks (A_{210}) with associated ^{14}C fractionation (dpm) of photoreactor products from 2 hours of photocatalytic treatment of TNT before (top graph) and after (bottom graph) aerobic biological treatment by *P. chrysosporium* in experiment 15.

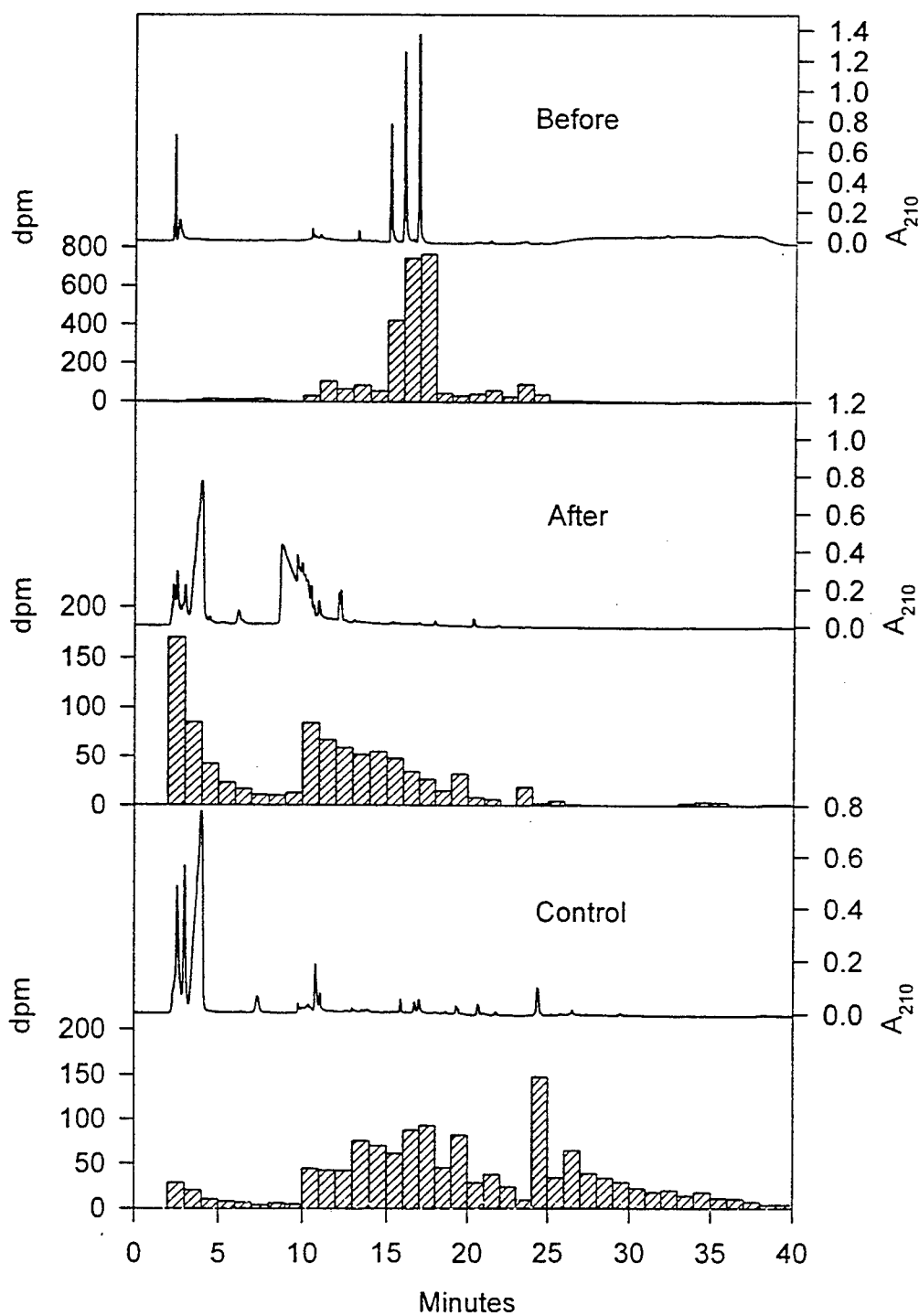


Figure 33. Chromatographic peaks (A_{210}) with associated ^{14}C fractionation (dpm) of photoreactor products from 6 hours of photocatalytic treatment of TNT before (top graph) and after (middle graph) aerobic biological treatment by *P. chrysosporium* in experiment 15. An abiotic control (bottom graph) is included for comparison.

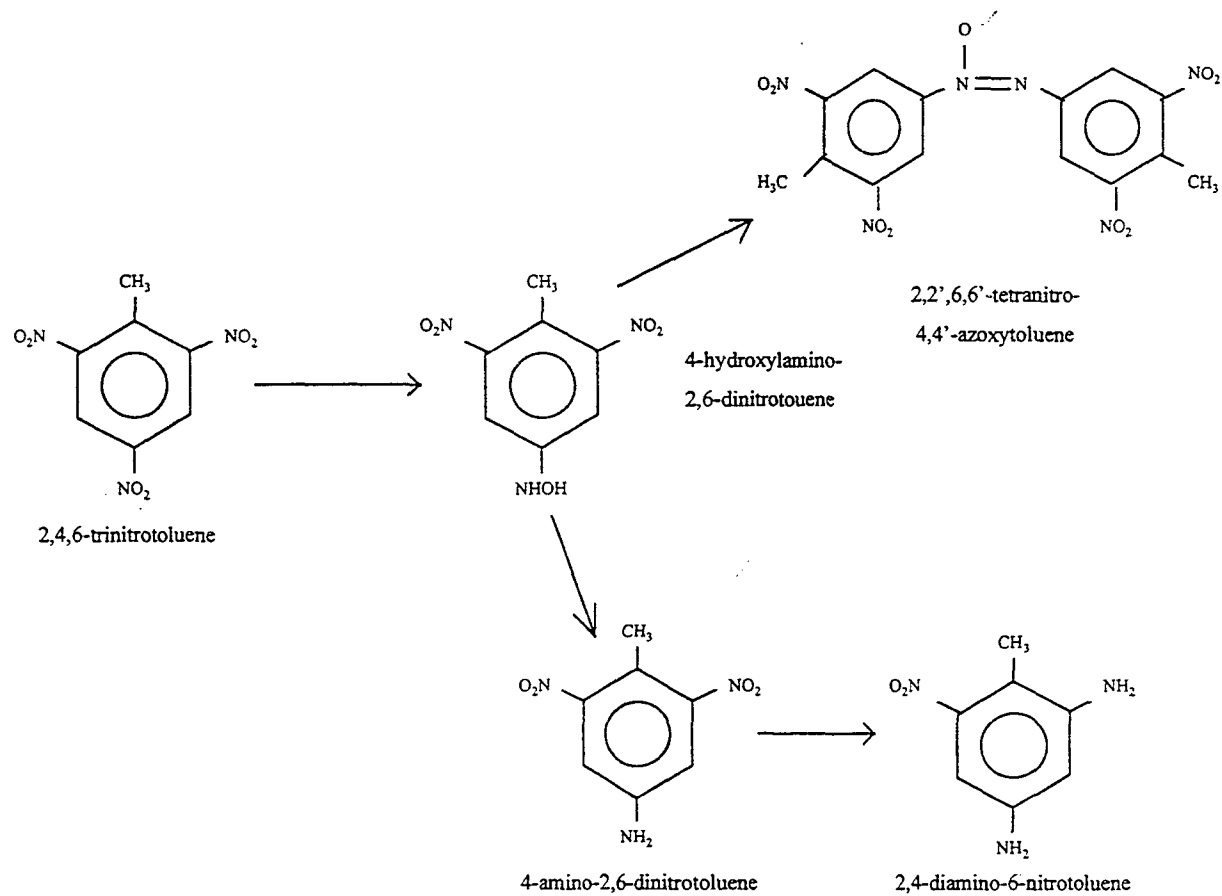


Figure 34. Proven pathway for anaerobic photocatalytic degradation of TNT.

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13. ABSTRACT (Maximum 200 words) Photocatalytic treatment of pink water under oxidative conditions can mineralize TNT, HMX, and RDX but requires long residence times and a large power consumption. Photocatalytic treatment under reducing conditions, using EDTA, citric acid, glucose, and sodium formate, converts TNT to other substances. Power costs are significantly less than for the oxidative process. Biological treatment using <i>Phanerochaete crysosporium</i> following EDTA reduction of pink water partially mineralized TNT. The degree of conversion of carbon to CO ₂ depended on the duration of the photocatalytic pretreatment. Six hours of pretreatment produced 32% mineralization. This innovative new process for treating pink water is an effective method for reducing the level of TNT below the detection limit of 20 ppb. HMX and RDX are removed at a somewhat lower rate. Parties interested in learning more about this process may contact the authors of this report at NREL. Annual operating costs are estimated at \$179,000 for a 150 gpm flow of pink water. The currently used carbon adsorption process costs \$323,500. Capital costs are believed to be similar for both processes.			
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