

REPORT DOCUMENTATION PAGE			Form Approved OMB No. 0704-0188	
Public reporting burden for this collection of information is estimated to average 1 hour per response, including the time for reviewing instructions, searching existing data sources, gathering and maintaining the data needed, and completing and reviewing the collection of information. Send comments regarding this burden estimate or any other aspect of this collection of information, including suggestions for reducing this burden, to Washington Headquarters Services, Directorate for Information Operations and Reports, 1215 Jefferson Davis Highway, Suite 1204, Arlington, VA 22202-4302, and to the Office of Management and Budget, Paperwork Reduction Project (0704-0188), Washington, DC 20503.				
1. AGENCY USE ONLY (Leave blank)		2. REPORT DATE 24-6-96	3. REPORT TYPE AND DATES COVERED Technical Report	
4. TITLE AND SUBTITLE Room Temperature Cyclopentadiene Elimination Reaction for the Synthesis of Diethylgallium-Amides, -Phosphides and -Thiolates. Crystal and Molecular Structures of $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ and $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$			5. FUNDING NUMBERS Grant: N00014-96-1-0483 R&T Code: 4135002	
6. AUTHOR(S) O. T. Beachley, Jr., Daniel B. Rosenblum, Melvyn Rowen Churchill, Charles H. Lake and Laurence M. Toomey				
7. PERFORMING ORGANIZATION NAME(S) AND ADDRESS(ES) Department of Chemistry State University of New York at Buffalo 516 NSM Complex Buffalo, NY 14260-3000			8. PERFORMING ORGANIZATION REPORT NUMBER Technical Report No. 44	
9. SPONSORING / MONITORING AGENCY NAME(S) AND ADDRESS(ES) Office of Naval Research 800 N. Quincy Street Arlington, VA 22217-5000			10. SPONSORING / MONITORING AGENCY REPORT NUMBER n/a	
11. SUPPLEMENTARY NOTES Accepted for publication - Organometallics				
12a. DISTRIBUTION / AVAILABILITY STATEMENT This document has been approved for public release and sale; its distribution is unlimited. Reproduction in whole or in part is permitted for any purpose of the United States government.			12b. DISTRIBUTION CODE n/a	
13. ABSTRACT (Maximum 200 words) The compounds $[\text{Et}_2\text{GaNET}_2]_2$ , $[\text{Et}_2\text{GaN}(\text{H})(\text{Me})]_2$ , $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$ , $[\text{Et}_2\text{GaP}(\text{i-Pr})_2]_2$ , $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ , and $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ have been prepared in high yields at room temperature by the elimination of cyclopentadiene from $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$ and the corresponding amine, phosphine or thiol. The three diethylgallium amides and $[\text{Et}_2\text{GaP}(\text{i-Pr})_2]_2$ are liquids whereas the other phosphide and the thiolate are crystalline solids at room temperature. All compounds were fully characterized by elemental analyses, $^1\text{H}$ and $^{31}\text{P}$ NMR spectroscopy and cryoscopic molecular weight studies in benzene, as appropriate and are dimeric in solution. The crystalline compounds $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ and $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ were characterized by X-ray structural studies. Even though $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$ and $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ were prepared at room temperature and isolated in high yield, $^1\text{H}$ NMR studies revealed that neither compound was formed in high yield in solution. Removal of the cyclopentadiene by either dimerization or distillation, and/or dimerization the gallium product are necessary to prevent the back reaction to reform the reactants.				
14. SUBJECT TERMS Organogallium compounds, gallium-nitrogen, phosphorus and sulfur compounds, X-ray structure studies			15. NUMBER OF PAGES 22	
			16. PRICE CODE n/a	
17. SECURITY CLASSIFICATION OF REPORT Unclassified	18. SECURITY CLASSIFICATION OF THIS PAGE Unclassified	19. SECURITY CLASSIFICATION OF ABSTRACT Unclassified	20. LIMITATION OF ABSTRACT UL	

19960703 047

OFFICE OF NAVAL RESEARCH

Contract N-00014-96-1-0483

R&T Code 4135002

TECHNICAL REPORT NO. 44

Room Temperature Cyclopentadiene Elimination Reaction for the Synthesis of  
Diethylgallium-Amides, -Phosphides and -Thiolates. Crystal and Molecular Structures of  
 $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$

by

O. T. Beachley, Jr.\*, Daniel B. Rosenblum,  
Melvyn Rowen Churchill\*, Charles H. Lake and Laurence M. Toomey

Prepared for Publication

in

Organometallics

State University of New York at Buffalo  
Department of Chemistry  
Buffalo, New York 14260-3000

24 June 1996

Reproduction in whole or in part is permitted for  
any purpose of the United States Government

\*This document has been approved for public release  
and sale; its distribution is unlimited

Contribution from the Department of Chemistry  
State University of New York at Buffalo, Buffalo, NY 14260

Room Temperature Cyclopentadiene Elimination Reaction  
for the Synthesis of Diethylgallium-Amides, -Phosphides and  
-Thiolates. Crystal and Molecular Structures of  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$   
and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$

O. T. Beachley, Jr.\*, Daniel B. Rosenblum, Melvyn Rowen Churchill\*,  
Charles H. Lake and Laurence M. Toomey

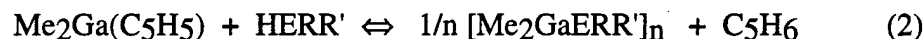
Abstract

The compounds  $[\text{Et}_2\text{GaNEt}_2]_2$ ,  $[\text{Et}_2\text{GaN}(\text{H})(\text{Me})]_2$ ,  $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$ ,  $[\text{Et}_2\text{GaP}(\text{i-Pr})_2]_2$ ,  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ , and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  have been prepared in high yields at room temperature by the elimination of cyclopentadiene from  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  and the corresponding amine, phosphine or thiol. The three diethylgallium amides and  $[\text{Et}_2\text{GaP}(\text{i-Pr})_2]_2$  are liquids whereas the other phosphide and the thiolate are crystalline solids at room temperature. All compounds were fully characterized by elemental analyses,  $^1\text{H}$  and  $^{31}\text{P}$  NMR spectroscopy and cryoscopic molecular weight studies in benzene, as appropriate and are dimeric in solution. The crystalline compounds  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ , and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  were characterized by X-ray structural studies. Even though  $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$  and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  were prepared at room temperature and isolated in high yield,  $^1\text{H}$  NMR studies revealed that neither compound was formed in high yield in solution. Removal of the cyclopentadiene by either dimerization or distillation, and/or dimerization the gallium product are necessary to prevent the back reaction to reform the reactants.

The preparation of electronic materials typically requires precursors which have an ultra-high level of purity. The obvious impurities are those imported with the reagents used to prepare the precursors and those which are formed during the synthetic reactions. The simplest and most direct route to single source precursors  $R_2GaER'_2$  ( $E = N, P$ ) from readily available, high purity but simple reagents which do not involve atoms from potentially electronically damaging elements, is the hydrocarbon elimination reaction (Equation 1). After the impurities associated with the reagents have been minimized, the second source of impurities originates with the high temperatures typically needed for the hydrocarbon elimination reaction. The preparation of gallium-nitrogen precursors uses temperatures of 100 - 150°C<sup>1,2</sup> whereas gallium-phosphorus derivatives typically use higher temperatures of 150 - 250°C.<sup>3</sup> A goal of our research has been to discover and study hydrocarbon elimination reactions which occur at or below room temperature.



One novel reaction which satisfied these criteria is the cyclopentadiene elimination reaction<sup>4</sup> which occurs between  $Me_2Ga(C_5H_5)$  and primary and secondary amines and phosphines (Equation 2). Even though a series of dimethylgallium amides and phosphides were prepared at room temperature, amines with very bulky substituents such as 2,4,6-tri(*t*-butyl)phenylamine,  $NH_2[(t-Bu)_3C_6H_2]$ , did not give the desired products. The reactants were recovered. The explanation for this observation<sup>4</sup> was the elimination reaction actually occurred but the resulting monomeric gallium nitrogen product reacted readily with the cyclopentadiene to reform the reactants instead of associating to form gallium-nitrogen dimers or trimers. Association of the monomer would tie up the lone pair on the group 15 element and, in turn, prevent the back reaction of cyclopentadiene as a weak acid.



In this paper we describe the results of our investigations of the reactions of  $Et_2Ga(C_5H_5)$ <sup>5</sup> with some primary and secondary amines, several secondary phosphines and a thiol. This study enabled us to learn more about the nature of the reaction including the relative ease of the forward and reverse reactions, the effects of the multiple gallium reactants which are present in solution due to the occurrence of the redistribution equilibrium<sup>5</sup> (Equation 3) as well as the effects of the ethyl verses the methyl groups on the nature of this reaction.



Furthermore, this study provided us with a series of diethylgallium-group 15 and 16 compounds for which very few other examples are known. The other examples of which we are aware are  $\text{Et}_2\text{GaPEt}_2^6$ ,  $\text{Et}_2\text{GaAs}(\text{t-Bu})_2^7$ ,  $\text{Et}_2\text{GaN}_3^8$ ,  $\text{Et}_2\text{GaNH}_2^9$ ,  $\text{Et}_2\text{GaN}(\text{C}_5\text{H}_{10})^{10}$ ,  $\text{Et}_2\text{GaN}(\text{C}_2\text{H}_4)^{11}$ ,  $\text{Et}_2\text{GaN}(\text{i-Bu})_2^{12}$ ,  $\text{Et}_2\text{GaN}(\text{Cy})_2^{12}$ ,  $\text{Et}_2\text{GaN}=\text{C}(\text{H})(\text{Ph})^{13}$  and  $\text{Et}_2\text{GaN}=\text{C}(\text{H})(\text{t-Bu})^{13}$ .

The reactions of  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  with  $\text{HNEt}_2$ ,  $\text{HN}(\text{H})\text{Me}$ ,  $\text{HN}(\text{H})(\text{t-Bu})$ ,  $\text{HP}(\text{i-Pr})_2$ ,  $\text{HP}(\text{t-Bu})_2$ , and  $\text{HS}(\text{SiPh}_3)$  in benzene or pentane provide convenient room temperature routes to the preparation of  $[\text{Et}_2\text{GaN}(\text{H})(\text{Me})]_2$ ,  $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$ ,  $[\text{Et}_2\text{GaP}(\text{i-Pr})_2]_2$ ,  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ , and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  in high yields (70 - 90%). The insolubility of  $\text{HS}(\text{SiPh}_3)$  in pentane made benzene the preferred solvent for the synthesis of  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ . The only direct comparison between the temperature required for the synthesis of these compounds by the elimination of cyclopentadiene by using  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  with the temperature necessary for the elimination of ethane by using  $\text{GaEt}_3$  has been provided by this study. When  $\text{GaEt}_3$  was reacted with the thiol, a temperature of  $70^\circ\text{C}$  was necessary to eliminate ethane and form  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ . Only two other diethyl gallium compounds,  $\text{Et}_2\text{GaPEt}_2$  and  $\text{Et}_2\text{GaN}(\text{C}_2\text{H}_4)$ , have been prepared by hydrocarbon elimination reactions and both preparative reactions required temperatures in the range of  $100 - 150^\circ\text{C}^{6,11}$ . However, elimination reactions have been used for the preparation of several of the dimethylgallium derivatives of the amines and phosphines used in this study. For example,  $\text{GaMe}_3$  and  $\text{HN}(\text{H})\text{Me}$  required  $125^\circ\text{C}^1$ ;  $\text{HN}(\text{H})(\text{t-Bu})$  needed  $110^\circ\text{C}^2$  but  $\text{HP}(\text{t-Bu})_2$  was heated at  $220^\circ\text{C}$  for 10 days for the elimination of methane to form  $[\text{Me}_2\text{GaP}(\text{t-Bu})_2]_2^{14}$ .

All products prepared during this investigation were fully characterized by partial elemental analyses (C and H), cryoscopic molecular weight studies in benzene solution and  $^1\text{H}$  and  $^{31}\text{P}$  NMR spectroscopy, as appropriate. All group 13-15 products except  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  were liquids at room temperature and were isolated by vacuum distillation. The solids  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  were purified by recrystallization and were characterized additionally by their melting points and by X-ray structural studies.

The characterization data for benzene solutions of all compounds prepared in this study, cryoscopic molecular weight and NMR spectral data, confirm the presence of only dimers. The molecular weight data exhibited no concentration dependence and the  $^{31}\text{P}$  NMR spectra of the phosphorus compounds had only single lines. Thus, it is unlikely that multiple species with different degrees of association are present in solution. When the diethyl gallium derivatives are compared with the corresponding dimethyl species, differences in degrees of association are noted.

For example,  $\text{Me}_2\text{GaN(H)(Me)}$  is a trimer<sup>15</sup> in benzene solution and  $\text{Me}_2\text{GaP(i-Pr)}_2$  has been observed as an equilibrium mixture of dimers and trimers<sup>16</sup> by NMR spectroscopy but  $\text{Me}_2\text{GaN(H)(t-Bu)}_2$ ,<sup>17</sup>  $\text{Me}_2\text{GaNEt}_2$ <sup>4</sup> and  $\text{Me}_2\text{GaP(t-Bu)}_2$ <sup>14,18</sup> are dimers. It is also noteworthy that when direct comparisons are made between the physical properties of the methyl and ethyl gallium derivatives, the ethyl compounds have lower melting points and are typically liquids at room temperature whereas the methyl compounds are solids.

Proton NMR spectroscopy clearly shows that  $[\text{Et}_2\text{GaN(H)(t-Bu)}]_2$  and  $[\text{Et}_2\text{GaN(H)(Me)}]_2$  exist as mixtures of cis and trans isomers in benzene solution. The cis isomer has two magnetically different ethyl groups on gallium. As a result, two triplets and two quartets of equal intensity are observed. In contrast, the ethyl groups in the trans isomer are equivalent but magnetically different from the ethyl protons in the cis isomer. The methylene protons of the gallium ethyl groups in the trans isomer are also magnetically nonequivalent and therefore their resonance appears as a doublet of quartets. The methylene protons for both compounds exhibited the same geminal coupling constant, 13.6 Hz. The integrations of the two lines for the t-butyl groups on nitrogen due to the two isomers suggest approximately 77% trans and 23% cis isomer. In contrast, a sample of  $[\text{Me}_2\text{GaN(H)(t-Bu)}]_2$  prepared by either the analogous cyclopentadiene elimination reaction<sup>4</sup> or a metathesis reaction at  $-78^\circ\text{C}$ <sup>17</sup> had resonances for only the trans isomer (benzene solution) whereas when the compound was prepared by the elimination of methane from  $\text{GaMe}_3$  and  $\text{HN(H)(t-Bu)}$  at  $110^\circ\text{C}$ , the  $^1\text{H}$  NMR spectrum of a toluene solution demonstrated the presence of both isomers (67% trans and 33% cis). A benzene solution of  $[\text{Et}_2\text{GaN(H)(Me)}]_2$  had equal quantities of the cis and trans isomers according to the intensities of the N-Me resonances. Since the corresponding methyl gallium compound<sup>4</sup> is a trimer  $[\text{Me}_2\text{GaN(H)(Me)}]_3$ , comparisons are inappropriate.

The  $^1\text{H}$  and  $^{31}\text{P}$  NMR spectra of the two phosphides prepared in this investigation,  $[\text{Et}_2\text{GaP(i-Pr)}_2]_2$  and  $[\text{Et}_2\text{GaP(t-Bu)}_2]_2$ , are consistent with the presence of only dimers. The  $^{31}\text{P}$  NMR spectrum of each compound shows a single resonance. The chemical shifts of the lines are comparable with those observed for the corresponding dimeric dimethyl gallium derivatives. The compound  $[\text{Et}_2\text{GaP(i-Pr)}_2]_2$  has its line at -5.72 ppm whereas  $[\text{Me}_2\text{GaP(i-Pr)}_2]_2$  is at -11.0 ppm.<sup>16</sup> Similarly, the chemical shift for  $[\text{Et}_2\text{GaP(t-Bu)}_2]_2$  is 35.22 ppm whereas  $[\text{Me}_2\text{GaP(t-Bu)}_2]_2$  is 28.41 ppm.<sup>14,18</sup> The  $^1\text{H}$  NMR spectra of both diethyl gallium phosphides exhibit virtual coupling between the protons of the substituents on phosphorus and the two ring phosphorus atoms. Thus, the t-butyl resonance for  $[\text{Et}_2\text{GaP(t-Bu)}_2]_2$  is split into an apparent triplet ( $J_{\text{P-H}} = 7.0$  Hz) whereas the C-H proton for the methine hydrogen of the isopropyl group in  $[\text{Et}_2\text{GaP(i-Pr)}_2]_2$  appears as a septet of triplets ( $^3J_{\text{P-H}} = 7.2$  Hz). It is also noteworthy that the methylene protons for the ethyl groups bound to gallium in  $[\text{Et}_2\text{GaP(i-Pr)}_2]_2$  also exhibit virtual coupling as a quartet of triplets ( $J_{\text{P-H}} = 2.4\text{Hz}$ ) is observed.

The  $^1\text{H}$  NMR spectrum of  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  in benzene has only one set of lines for the protons of the ethyl groups. Thus, the molecule exists either only as the trans isomer or both isomers are present but a rapid isomerization process with breaking of a gallium-sulfur ring bond is occurring. Similar observations were recently reviewed by Oliver.<sup>19</sup>

The crystal and molecular structure of  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  is consistent with the observations for solutions. The molecular geometry for  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  is shown in Figure 1. Interatomic distances and angles are collected in Table 1. The molecule has precise  $C_i$  symmetry and a planar  $\text{Ga}_2\text{P}_2$  core in which  $\text{Ga}(1)\text{-P}(1) = \text{Ga}(1\text{A})\text{-P}(1\text{A}) = 2.482(2)\text{\AA}$  and  $\text{Ga}(1)\text{-P}(1\text{A}) = \text{Ga}(1\text{A})\text{-P}(1) = 2.483(2)\text{\AA}$ . The P-Ga-P angles are  $86.9(1)^\circ$ , while the Ga-P-Ga angles are  $93.1(1)^\circ$ . Each gallium(III) atom is linked to two ethyl groups with  $\text{Ga}(1)\text{-C}(11) = 1.999(6)\text{\AA}$  and  $\text{Ga}(1)\text{-C}(21) = 1.998(6)\text{\AA}$  and an interligand angle of  $\text{C}(11)\text{-Ga}(1)\text{-(C}21) = 107.1(3)^\circ$ . The angles at the  $\alpha$ -carbon atoms of the ethyl groups are larger than the normal tetrahedral value of  $109.5^\circ$ , with  $\text{Ga}(1)\text{-C}(11)\text{-C}(12) = 119.9(6)^\circ$  and  $\text{Ga}(1)\text{-C}(21)\text{-C}(22) = 118.0(6)^\circ$ . The carbon-carbon distances within the ethyl groups are anomalously short as a result of artificial librational contraction, with  $\text{C}(11)\text{-C}(12) = 1.380(12)$  and  $\text{C}(21)\text{-C}(22) = 1.418(13)\text{\AA}$ . Note that there are large anisotropic displacement coefficients for these atoms (*viz.*,  $U_{22}$  for C(11) is  $0.125(7)\text{\AA}^2$  while  $U_{22}$  for C(12) is  $0.228(12)\text{\AA}^2$ ;  $U_{22}$  for C(21) is  $0.113(6)\text{\AA}^2$  while  $U_{22}$  for C(22) is  $0.285(15)\text{\AA}^2$ ). The bridging di(*t*-butyl)phosphide ligands are associated with bond lengths of  $\text{P}(1)\text{-C}(31) = 1.904(6)$  and  $\text{P}(1)\text{-C}(41) = 1.900(6)\text{\AA}$  and an interligand angle of  $\text{C}(31)\text{-P}(1)\text{-C}(41) = 110.2(3)^\circ$ ; the distances of the C-CH<sub>3</sub> linkages are in the expected range with  $\text{C}(31)\text{-Me} = 1.490(11) \rightarrow 1.517(11)\text{\AA}$  and  $\text{C}(41)\text{-Me} = 1.493(11) \rightarrow 1.521(9)\text{\AA}$ .

A comparison of the major structural features of  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  with other closely related molecules of the type  $(\text{R}_2\text{GaPR}')_2$  suggests some differences but also many similarities. The other molecules  $[(\text{n-Bu})_2\text{GaP}(\text{t-Bu})_2]_2$ <sup>18</sup>,  $[\text{Me}_2\text{GaP}(\text{SiMe}_3)_2]_2$ <sup>20</sup>, and  $[(\text{t-Bu})_2\text{GaP}(\text{H})(\text{C}_5\text{H}_9)]_2$ <sup>21</sup> have planar  $\text{Ga}_2\text{P}_2$  cores; whereas  $[(\text{Me}_3\text{CCH}_2)_2\text{GaPPh}_2]_2$ <sup>22</sup> has a puckered core. Although the halide derivatives  $[\text{Cl}_2\text{GaP}(\text{SiMe}_3)_2]_2$ <sup>23</sup> and  $[\text{Br}_2\text{GaP}(\text{SiMe}_3)_2]_2$ <sup>24</sup> contain a planar  $\text{Ga}_2\text{P}_2$  core with equal Ga-P bond lengths, it is noteworthy that  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  is the only molecule containing two organic groups bonded to gallium that has both a planar  $\text{Ga}_2\text{P}_2$  core and all four gallium-phosphorus distances equal within experimental error. The angle between the organic groups bound to gallium is similar to that in the other structurally characterized di(*t*-butyl)phosphide derivative  $[(\text{n-Bu})_2\text{GaP}(\text{t-Bu})_2]_2$ <sup>18</sup> but it is smaller than those observed for  $[\text{Me}_2\text{GaP}(\text{SiMe}_3)_2]_2$ <sup>20</sup> ( $114.5(3)^\circ$ ),  $[(\text{Me}_3\text{CCH}_2)_2\text{GaPPh}_2]_2$ <sup>22</sup> ( $124.35(50)^\circ$  and  $121.98(49)^\circ$ ), and  $[(\text{t-Bu})_2\text{GaP}(\text{H})(\text{C}_5\text{H}_9)]_2$ <sup>21</sup> ( $121.0(2)^\circ$ ).

The compound  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  has also been subjected to an X-ray structural study and is shown in Figure 2. Interatomic distances and angles are provided in Table 2. This molecule has precise  $C_i$  symmetry (crystallography required) with a planar  $\text{Ga}_2\text{S}_2$  core. Most other structurally characterized molecules of this type including  $[(\text{t-Bu})_2\text{GaS}(\text{H})]_2$ <sup>25</sup>,  $[\text{Ph}_2\text{GaSEt}]_2$ <sup>26</sup>

and  $\{\text{Ph}_2\text{GaS}[\text{Sn}(\text{c-C}_6\text{H}_{11})_3]\}_2^{27}$  are also planar but  $[\text{Me}_2\text{GaS}(\text{C}_6\text{F}_5)]_2^{28}$  and  $[\text{I}_2\text{GaS}(\text{i-Pr})]_2^{29}$  are puckered. Bond lengths within the planar  $\text{Ga}_2\text{S}_2$  core of  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  alternate in length,  $\text{Ga}(1)\text{-S}(1) = \text{Ga}(1\text{a})\text{-S}(1\text{a}) = 2.450(2)\text{\AA}$  and  $\text{Ga}(1)\text{-S}(1\text{a}) = \text{Ga}(1\text{a})\text{-S}(1) = 2.396(2)\text{\AA}$ . Angles within this system are  $\text{S}(1)\text{-Ga}(1)\text{-S}(1\text{a}) = 88.1(1)^\circ$  and  $\text{Ga}(1)\text{-S}(1)\text{-Ga}(1\text{a}) = 91.9(1)^\circ$ . The ethyl ligands are associated with bond lengths of  $\text{Ga}(1)\text{-C}(1) = 1.934(14)$  and  $\text{Ga}(1)\text{-C}(3) = 1.971(11)\text{\AA}$  and a large interligand angle  $\text{C}(1)\text{-Ga}(1)\text{-C}(3) = 125.1(4)^\circ$ . Similar large interligand angles have been observed for the other thiolate derivatives. The angles at the  $\alpha$ -carbon atoms for  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  are  $\text{Ga}(1)\text{-C}(1)\text{-C}(2) = 119.6(8)^\circ$  and  $\text{Ga}(1)\text{-C}(3)\text{-C}(4) = 114.4(8)^\circ$ , while the  $\text{C}(\alpha)\text{-C}(\beta)$  bond lengths are  $\text{C}(1)\text{-C}(2) = 1.437(14)$  and  $\text{C}(3)\text{-C}(4) = 1.446(11)\text{\AA}$ . Libration is not as large a problem as in the previous structure. The largest relevant  $U_{ij}$  values are  $U_{33} = 0.143(10)\text{\AA}^2$  for  $\text{C}(1)$  versus  $U_{33} = 0.186(13)\text{\AA}^2$  for  $\text{C}(2)$ , and  $U_{22} = 0.140(10)\text{\AA}^2$  for  $\text{C}(3)$  versus  $U_{22} = 0.115(9)\text{\AA}^2$  for  $\text{C}(4)$ .

The two  $\text{SiPh}_3$  substituents in  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  take up the same mutually trans configuration as has been observed for other gallium-thiolate derivatives.<sup>19</sup> Each  $\text{S}(\text{SiPh}_3)$  ligand is associated with a  $\text{S}(1)\text{-Si}(1)$  bond length of  $2.148(3)\text{\AA}$  and  $\text{Si-Ph}$  distances of  $\text{Si}(1)\text{-C}(11) = 1.861(9)$ ,  $\text{Si}(1)\text{-C}(21) = 1.864(11)$  and  $\text{Si}(1)\text{-C}(31) = 1.864(11)\text{\AA}$ . Angles about silicon range from  $105.2(3)^\circ$  to  $115.3(2)^\circ$ . The  $\text{S-Si-Ph}$  angles are rather irregular ( $\text{S}(1)\text{-Si}(1)\text{-C}(31) = 105.2(3)^\circ$ ,  $\text{S}(1)\text{-Si}(1)\text{-C}(11) = 107.1(3)^\circ$  and  $\text{S}(1)\text{-Si}(1)\text{-C}(21) = 115.3(2)^\circ$ ) while the  $\text{Ph-Si-Ph}$  angles are closer in value ( $\text{C}(21)\text{-Si}(1)\text{-C}(31) = 108.5(4)^\circ$ ,  $\text{C}(11)\text{-Si}(1)\text{-C}(21) = 109.8(4)^\circ$  and  $\text{C}(11)\text{-Si}(1)\text{-C}(31) = 110.9(3)^\circ$ ). The phenyl groups are well behaved with all 18  $\text{C-C}$  bonds in the range  $1.353(16) - 1.392(13)\text{\AA}$ . One point worthy of note is a consistent reduction from the ideal trigonal value for the internal angles at the ipso carbon atoms (*viz.*,  $\text{C}(12)\text{-C}(11)\text{-C}(16) = 115.2(8)^\circ$ ,  $\text{C}(22)\text{-C}(21)\text{-C}(26) = 115.4(10)^\circ$ ,  $\text{C}(32)\text{-C}(31)\text{-C}(36) = 116.7(6)^\circ$ ).

The reagents  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  and  $\text{HN}(\text{H})(\text{t-Bu})$  or  $\text{HSSiPh}_3$  have been used to synthesize  $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$  and  $[\text{Et}_2\text{GaSSiPh}_3]_2$  in high yields of the isolated products at room temperature. In order to learn more about the nature of these reactions,  $^1\text{H}$  NMR spectroscopy was used to probe the relative rates of the cyclopentadiene elimination reaction and to measure the extent of the overall reaction in solution with time. Stoichiometric quantities of the reagents were combined in sufficient  $\text{C}_6\text{D}_6$  to form solutions with concentrations which were comparable to those used in the synthetic reactions. When the  $^1\text{H}$  NMR spectrum of the reagents upon mixing was observed, the only observed lines were those of the adducts of  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$ ,  $\text{EtGa}(\text{C}_5\text{H}_5)_2$ , and  $\text{GaEt}_3$  with either the amine or the thiol. The spectrum was qualitatively similar to that observed when pure  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  was dissolved in benzene<sup>5</sup> but the chemical shifts were different due to the formation of adducts. Lines for products were not observed. These observations confirm the occurrence of the ligand redistribution reaction<sup>5</sup> (Equation 3) and the presence of multiple gallium species when solutions of  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  are in the presence of Lewis



bases with acidic protons but also demonstrate that the elimination reaction with this base is either suprisingly slow or the equilibrium constant for the reaction is very small. In the case of the reaction with  $\text{HN}(\text{H})(\text{t-Bu})$ , evidence of product formation was noted only after 3 1/2 h at room temperature by the occurrence of lines for the t-Bu groups of the cis and trans isomers of  $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$ , 1.02ppm and 1.00ppm, respectively. The lines for the gallium-nitrogen products continued to grow such that the reaction was approximately 37% complete after 2 days at room temperature, 43% complete after 3 days and 49% complete after 4 days. Then, heating of the NMR tube at 65°C was initiated. After approximately 15 h of heating, the spectrum was consistent with the presence of only two gallium species, the cis and trans isomers of the dimeric product. The chemical shifts of all lines were identical to those observed for the product isolated from the synthetic experiment. All lines for the three initial gallium species had disappeared. In the case of the thiol with the bulky  $\text{SiPh}_3$  group, the Lewis base with the more acidic proton, the spectrum did not change from that initially observed, even after heating at 45°C for 2 weeks. Proton NMR lines for the product were not observed. Thus, since  $[\text{Et}_2\text{GaN}(\text{H})(\text{t-Bu})]_2$  and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$  can be isolated in high yield after cyclopentadiene elimination reactions at room temperature, removal of the cyclopentadiene monomer by either dimerization or distillation and/or dimerization of the gallium product to prevent the back reaction of the cyclopentadiene must be necessary to produce the observed high yields of products.

## Experimental

All compounds described in this investigation were extremely sensitive to oxygen and moisture and were manipulated in a standard vacuum line or in a purified argon atmosphere. The starting compound  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  was prepared and purified by the literature method.<sup>5</sup> Amines were dried over KOH and distilled prior to use. Phosphines were vacuum distilled prior to use. The compound  $\text{HS}(\text{SiPh}_3)$  was purchased from Aldrich Chemical Co. and used as received. Solvents were dried by conventional procedures. Elemental analyses were performed by E+R Microanalytical Laboratories, Inc., Corona, NY. The  $^1\text{H}$  NMR spectra were recorded at 400MHz by using a Varian VXR-400 spectrometer. Proton chemical shifts are reported in  $\delta$  units (ppm) and are referenced to  $\text{SiMe}_4$  at  $\delta$  0.00 and  $\text{C}_6\text{H}_6$  at  $\delta$  7.15. The  $^{31}\text{P}$  NMR spectra were recorded at 161.9MHz by using a Varian VXR-400 spectrometer and are referenced to 85%  $\text{H}_3\text{PO}_4$  at  $\delta$  0.00. The following abbreviations are used to report the multiplicities of lines: s (singlet), d (doublet), t (triplet), q (quartet), br (broad), st (septet of triplets), dq (doublet of quartets), qt (quartet of triplets). All samples for NMR spectra were contained in sealed NMR tubes. Infrared spectra were recorded for either neat liquids, Nujol solutions of liquids, or as Nujol mulls for solids by using CsI plates and a Perkin-Elmer Model 683 spectrometer. Absorption intensities are reported with abbreviations w (weak), m (medium), s (strong), vs (very strong), and sh (shoulder). The spectroscopic data are available with the Supplementary Material. Melting points were observed in sealed capillaries and are uncorrected. Molecular weights were measured cryoscopically in benzene by using an instrument similar to that described by Shriver and Drezdon.<sup>30</sup> Mass spectra were obtained by Electron Impact with a VG Model 70-SE high resolution mass spectrometer.

### **Reaction of $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$ with HERR' and HE'R (E=N, P; E'=S)**

In a typical experiment, a quantity (1-8 mmol) of the appropriate amine or phosphine was vacuum distilled into a tared tube equipped with a Teflon valve and standard taper joint. Next, a stoichiometric quantity of  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  was placed in a Schlenk flask equipped with a magnetic stirbar. Then, approximately 50mL pentane was distilled onto the  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  at  $-196^\circ\text{C}$ . The amine or phosphine was then vacuum distilled onto the  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$ -solvent mixture. Finally, the flask was warmed from  $-196^\circ\text{C}$  to ambient temperature and stirred overnight. The reactions of  $\text{Et}_2\text{Ga}(\text{C}_5\text{H}_5)$  with  $\text{HS}(\text{SiPh}_3)$ ,  $\text{HP}(\text{i-Pr})_2$  and  $\text{HP}(\text{t-Bu})_2$  were carried out in benzene rather than pentane for reaction times of two hours, two weeks, and two days, respectively. Liquid products were quantitatively transferred to small flasks with stir bars by washing with pentane, and then the flask was attached to a short-path still connected to a preweighed receiving flask. The pentane was removed by vacuum distillation at room temperature and then heat was applied by means of an oil bath to distill the product.

**[Et<sub>2</sub>GaN(H)(t-Bu)]<sub>2</sub>:** Product purified by vacuum distillation, oil bath at 90°C; bp: 70°C (<10<sup>-3</sup> mmHg); yield 95%; mp: 17-20.5°C; <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 1.36 (t, CH<sub>2</sub>CH<sub>3</sub>, cis), 1.30 (t, CH<sub>2</sub>CH<sub>3</sub>, trans), 1.27 (t, CH<sub>2</sub>CH<sub>3</sub>, cis), 1.02 (s, C(CH<sub>3</sub>)<sub>3</sub>, cis, 2.0H), 1.00 (s, C(CH<sub>3</sub>)<sub>3</sub>, trans, 6.4H), 0.68 (dq (J=12Hz), CH<sub>2</sub>CH<sub>3</sub>, trans), 0.43 (dq, (J=12Hz), CH<sub>2</sub>CH<sub>3</sub>, trans), 0.66 (q, CH<sub>2</sub>CH<sub>3</sub>, cis), 0.41 (q, CH<sub>2</sub>CH<sub>3</sub>, cis); Anal. calcd. for C<sub>8</sub>H<sub>20</sub>GaN: C, 48.05%; H, 10.08%. Found: C, 48.11%; H, 10.11%. Cryoscopic molecular weight, benzene solution, fw 158 (obsd molality, obsd mol wt, association): 0.090, 384, 1.93; 0.068, 377, 1.88, 0.041, 344, 1.73.

**[Et<sub>2</sub>GaN(H)(Me)]<sub>2</sub>:** Product purified by vacuum distillation, oil bath at 125°C; bp: 45°C (<10<sup>-3</sup> mmHg); 85% yield; <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 2.12 (d, NCH<sub>3</sub>, cis), 2.10 (d NCH<sub>3</sub>, trans), 1.28 (t, CH<sub>2</sub>CH<sub>3</sub>, cis), 1.23 (t, CH<sub>2</sub>CH<sub>3</sub>, trans), 1.19 (t, CH<sub>2</sub>CH<sub>3</sub>, cis), 0.38 (dq, (J=12Hz), CH<sub>2</sub>CH<sub>3</sub>, trans), 0.55 (q, CH<sub>2</sub>CH<sub>3</sub>, cis), 0.52 (dq, (J=12Hz), CH<sub>2</sub>CH<sub>3</sub>, trans), 0.35 (q, CH<sub>2</sub>CH<sub>3</sub>, cis). Approximate equal molar quantities of cis and trans isomers were observed. Anal. calcd. for C<sub>5</sub>H<sub>14</sub>GaN: C, 38.04%; H, 8.94%. Found: C, 38.56%; H, 8.62%. Cryoscopic molecular weight, benzene solution, fw 158 (obsd molality, obsd mol wt, association): 0.126, 327, 2.08; 0.098, 310, 1.97, 0.072, 316, 2.00; 0.063, 312, 1.98; 0.047, 314, 1.99.

**[Et<sub>2</sub>GaNEt<sub>2</sub>]<sub>2</sub>:** Product purified by vacuum distillation, oil bath at 100°C; bp: 70° (<10<sup>-3</sup> mmHg); 90% yield; <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 2.72 (q, NCH<sub>2</sub>CH<sub>3</sub>, 2H), 1.31 (t, GaCH<sub>2</sub>CH<sub>3</sub>, 3H), 0.78 (t, NCH<sub>2</sub>CH<sub>3</sub>, 3H), 0.55 (q, GaCH<sub>2</sub>CH<sub>3</sub>, 2H); Mass Spectrum (m/e, relative intensity): 371 (Dimer - Et<sup>+</sup>), 91; 198 (Monomer - H<sup>+</sup>), 43; 170 (Monomer - Et<sup>+</sup>), 91; 127 (Et<sub>2</sub>Ga<sup>+</sup>), 93; 100 (EtGa<sup>+</sup>), 66; 69 (Ga<sup>+</sup>), 100. Anal. calcd. for C<sub>4</sub>H<sub>20</sub>GaN: C, 48.05%; H, 10.08%; N, 7.00%. Found: C, 48.52%; H, 9.76%; N, 7.39%. Cryoscopic molecular weight, benzene solution, fw 200 (obsd molality, obsd mol wt, association): 0.086, 424, 2.12; 0.053, 414, 2.07, 0.034, 422, 2.11.

**[Et<sub>2</sub>GaP(i-Pr)<sub>2</sub>]<sub>2</sub>:** Viscous product purified by first washing with pentane at -78°C and then vacuum distillation, oil bath at 72°C; bp: 70° (<10<sup>-3</sup> mmHg); 72% yield; <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 2.18 (st, J<sub>P-H</sub>= 2Hz, J<sub>H-H</sub>= 7.2Hz, PC(CH<sub>3</sub>)<sub>2</sub>H, 1H), 1.40 (t, GaCH<sub>2</sub>CH<sub>3</sub>, 3H), 1.16 (q, J=7.2Hz, PC(CH<sub>3</sub>)<sub>2</sub>H, 6H), 0.84 (qt, <sup>3</sup>J<sub>P-H</sub>= 2.4Hz, J<sub>H-H</sub>= 8.4H, GaCH<sub>2</sub>CH<sub>3</sub>, 2H). <sup>31</sup>P{<sup>1</sup>H} -5.72 (s); Mass Spectrum (m/e, relative intensity): 461 (Dimer - Et<sup>+</sup>), 11; 373 (Dimer - P(i-Pr)<sub>2</sub><sup>+</sup>), 14; 303 (Dimer - GaP(i-Pr)<sub>2</sub><sup>+</sup>), 6; 215 (EtGaP(i-Pr)<sub>2</sub><sup>+</sup>), 35; 127 (Et<sub>2</sub>Ga<sup>+</sup>), 70; 118 (P(i-Pr)<sub>2</sub><sup>+</sup>), 21; 99 (GaP<sup>+</sup>), 10; 69 (Ga<sup>+</sup>), 93. Anal. calcd. for C<sub>10</sub>H<sub>24</sub>GaP: C, 49.03%; H, 9.87%. Found: C, 48.97%; H, 9.83%. Cryoscopic molecular weight, benzene solution, fw 245 (obsd molality, obsd mol wt, association): 0.076, 505, 2.06; 0.061, 497, 2.03, 0.040, 497, 2.03.

**[Et<sub>2</sub>GaP(t-Bu)<sub>2</sub>]<sub>2</sub>:** Product purified by recrystallization from pentane at -40°C; 64% yield; mp: 160.8-162°C (Dec); <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 1.43 (t, GaCH<sub>2</sub>CH<sub>3</sub>, 3H), 1.33 (t, J=7.0Hz, PC(CH<sub>3</sub>)<sub>3</sub>, 9H), 1.00 (q, GaCH<sub>2</sub>CH<sub>3</sub>, 2H); <sup>31</sup>P{<sup>1</sup>H} 35.22 (s); Anal. calcd. for

$C_{12}H_{28}GaP$ : C, 52.79%; H, 10.34%. Found: C, 52.84%; H, 10.41%. Insufficient solubility in benzene for cryoscopic molecular weight studies.

**[Et<sub>2</sub>GaS(SiPh<sub>3</sub>)]<sub>2</sub>:** Product purified by recrystallization from methylcyclohexane at -40°C; 81% yield; mp 157.8-158.1°C; <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ) 7.75-7.78 (m, *o*-Ph-H, 6H), 7.10-7.12 (m, *m,p*-Ph-H, 9H), 1.07 (t, CH<sub>2</sub>CH<sub>3</sub>, 6H), 0.68 (q, CH<sub>2</sub>CH<sub>3</sub>, 4H); Anal. calcd. for C<sub>22</sub>H<sub>25</sub>GaSSi: C, 63.02%; H, 6.01%. Found: C, 62.98%; H, 6.11%. Cryoscopic molecular weight, benzene solution, fw 419 (obsd molality, obsd mol wt, association): 0.049, 811, 1.93; 0.038, 812, 1.94, 0.026, 833, 1.99.

**Reaction of GaEt<sub>3</sub> and HS(SiPh<sub>3</sub>) in a 1:1 mole ratio.** A break-seal tube, which was charged with 0.23g (0.78 mmol) HS(SiPh<sub>3</sub>) and 0.12g (0.78 mmol) GaEt<sub>3</sub>, was sealed by fusion and then warmed from -196°C to room temperature. The tube which initially contained a light yellow solid was gradually heated to 70°C over the period of two hours. The resulting colorless solid and condensable gas, presumed to be ethane by its physical properties (0.025g, 0.83mmol, 100% yield) and <sup>1</sup>H NMR spectrum (<sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 0.80 (s)) were separated by vacuum distillation. The crude solid was identified as [Et<sub>2</sub>GaS(SiPh<sub>3</sub>)]<sub>2</sub> (0.28g, 0.67 mmol, 87% yield) by comparison of its properties with those for the compound prepared by the other route. Mp: 155.2-155.7°C. <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, δ): 7.77-7.79 (m, *o*-Ph-H, 6H), 7.12-7.14 (m, *m,p*-Ph-H, 9H), 1.09 (t, CH<sub>2</sub>CH<sub>3</sub>, 6H), 0.70 (q, CH<sub>2</sub>CH<sub>3</sub>, 4H).

#### Studies of Relative Rates of Reaction by NMR Spectroscopy:

**a) Et<sub>2</sub>Ga(C<sub>5</sub>H<sub>5</sub>) and HS(SiPh<sub>3</sub>)** An NMR tube was charged with 0.024g (0.08 mmol) HSSiPh<sub>3</sub>, 0.016g (0.08 mmol) Et<sub>2</sub>Ga(C<sub>5</sub>H<sub>5</sub>), and approximately 1mL C<sub>6</sub>D<sub>6</sub> at -196°C and then sealed by fusion. The tube was used for an <sup>1</sup>H NMR study as described in the Results and Discussion section.

**b) Et<sub>2</sub>Ga(C<sub>5</sub>H<sub>5</sub>) and HN(H)(t-Bu)** An NMR tube was charged with 0.049g (0.26 mmol) Et<sub>2</sub>Ga(C<sub>5</sub>H<sub>5</sub>), 0.037g (0.26 mmol) HN(H)(t-Bu), and approximately 1mL of C<sub>6</sub>D<sub>6</sub> in order to investigate the relative rate of formation of [Et<sub>2</sub>GaN(H)(t-Bu)]<sub>2</sub> (see Results and Discussion).

#### Collection of X-Ray Diffraction Data

In each case crystals were sealed into thin-walled glass capillaries in an argon atmosphere inside a dry-box kept under strictly anaerobic and moisture-free conditions. The crystals were inspected under a binocular polarizing microscope to ensure that they were single; they were then centered accurately on a Siemens R3m/V automated four-circle diffractometer. The unit cell parameters and Laue symmetry were determined as described previously.<sup>31</sup> Intensity data

(Mo  $K\alpha$ ,  $\lambda = 0.710730$ ) were collected at ambient temperatures (23-25°C) with graphite-monochromatized radiation. Data were corrected for absorption and for Lorentz and polarization factors. Details are provided in Table 3.

**a) [Et<sub>2</sub>GaP(t-Bu)<sub>2</sub>]<sub>2</sub>.** The crystal was of dimensions 0.35 x 0.30 x 0.30 mm. The crystal belonged to the monoclinic system (2/m diffraction symmetry); the systematic absences  $h0l$  for  $h+l = 2n+1$  and  $0k0$  for  $k = 2n+1$  uniquely define the centrosymmetric monoclinic space group  $P2_1/n$  (No. 14 var.). One hemisphere of data, (i.e., two equivalent forms,  $\bar{h}\bar{k}l = \bar{h}\bar{k}l$  and  $hkl = hkl$ ) was collected. The 4075 reflections thus collected were merged to 1960 independent, space-group allowed data with  $R_{int} = 2.45\%$ .

**b) [Et<sub>2</sub>GaS(SiPh<sub>3</sub>)]<sub>2</sub>.** The crystal was rather small but equidimensional (0.2 x 0.2 x 0.2 mm). It was found to belong to the triclinic system ( $\bar{1}$  diffraction symmetry). Possible space groups are the noncentrosymmetric space group  $P1$  (No. 1) and the centrosymmetric space group  $P\bar{1}$  (No. 2). The centrosymmetric possibility is far more probable,<sup>32</sup> particularly for synthetic materials; the space group  $P\bar{1}$  was thus assumed and was confirmed by the successful solution of the structure.

### Determination of the Crystal Structures

All crystallographic calculations were carried out on a VAXstation 3100 computer with use of the Siemens SHELXTL PLUS (Release 4.11 (VMS)) program package.<sup>33</sup> The scattering factors used were the analytical values for neutral atoms; these were corrected for both components ( $\Delta f'$  and  $i\Delta f''$ ) of anomalous dispersion.<sup>34</sup> The structures were solved by direct methods and difference-Fourier syntheses. All non-hydrogen atoms were located and their positional and anisotropic thermal parameters were refined. Hydrogen atoms were included in calculated positions based upon  $d(C-H) = 0.96\text{\AA}$  and the appropriate idealized trigonal or tetrahedral geometry.<sup>35</sup> Refinement was continued until  $\Delta/\sigma < 0.003$  for each parameter. The correctness of each structure was confirmed by means of a final difference-Fourier synthesis.

**a) [Et<sub>2</sub>GaP(t-Bu)<sub>2</sub>]<sub>2</sub>.** The molecule is centered on the inversion center at 0, 0, 1/2. The Ga<sub>2</sub>P<sub>2</sub> ring is therefore required to be planar. Final discrepancy indices are  $R = 4.33\%$  for those 1170 data with  $F_0 > 6\sigma(F_0)$  and  $R = 7.16\%$  for all 1960 reflections. (Note that only 59.7% of the data are above  $6\sigma(F_0)$ .)

**b) [Et<sub>2</sub>GaS(SiPh<sub>3</sub>)]<sub>2</sub>.** The molecule lies about the inversion center at 0, 0, 1/2. The Ga<sub>2</sub>S<sub>2</sub> ring is required, by symmetry, to be strictly planar. The final discrepancy indices are  $R = 4.36\%$  for those 1261 reflections with  $F_0 > 6\sigma(F)$  and  $R = 11.56\%$  for all 2793 reflections. (The unattractively high  $R$ -value for all data simply reflects the problem of a very weak data set - only 45.1% of the data are above  $6\sigma(F_0)$ .)

**Acknowledgments.** This work was supported in part by the Office of Naval Research (O.T.B.). The upgrading of the X-ray diffractometer was made possible by Grant 89-13733 from the Chemical Instrumentation Program of the National Science Foundation.

**Supplementary Material Available.** Complete tables of positional parameters, interatomic distances and angles, anisotropic thermal parameters and calculated positions for hydrogen atoms (10 pages); listing of infrared absorptions for each compound (1 page). For ordering information see any current masthead page.

### References

- (1) Coates, G. E. *J. Chem. Soc.*, **1951**, 2003.
- (2) Park, J. T.; Kim, Y.; Kim, J.; Kim, K.; Kim, Y. *Organometallics*, **1992**, *11*, 3320.
- (3) (a) Coates, G. E.; Graham, J. *J. Chem. Soc.*, **1963**, 233. (b) Beachley, O. T., Jr.; Coates, G. E. *J. Chem. Soc.*, **1965**, 3241.
- (4) Beachley, O. T., Jr.; Royster, T. L., Jr.; Arhar, J. R.; Rheingold, A. L. *Organometallics*, **1993**, *12*, 1976.
- (5) Beachley, O. T., Jr.; Rosenblum, D. B.; Churchill, M. R.; Lake, C. H.; Krajkowski, L. M. *Organometallics*, **1995**, *14*, 4402.
- (6) (a) Maury, F.; Combes, M.; Constant, G.; Carles, R.; Renucci, J. B. *Journal De Physique*, **1982**, *43*, C1-347. (b) Maury, F.; Constant, G. *Polyhedron*, **1984**, *3*, 581.
- (7) Miller, J. E.; Mardones, M. A.; Nail, J. W.; Cowley, A. H.; Jones, R. A.; Ekerdt, J. G. *Chem. Mater.*, **1992**, *4*, 447.
- (8) Muller, J.; Dehnicke, K. *J. Organometal. Chem.*, **1968**, *12*, 37.
- (9) Andrews, J. E.; Littlejohn, M. A. *J. Electrochem. Soc.*, **1975**, 1274.
- (10) Sen, B.; White, L. *J. Inorg. Nucl. Chem.*, **1973**, *35*, 2207.
- (11) Storr, A.; Thomas, B. S. *J. Chem. Soc. (A)*, **1971**, 2850.
- (12) Nutt, W. R.; Murray, K. J.; Gulick, J. M.; Odom, J. D.; Dang, Y.; Lebioda, L. *Organometallics*, **1996**, *15*, 1728.
- (13) Jennings, J. R.; Wade, K. *J. Chem. Soc. (A)*, **1967**, 1222.
- (14) Arif, A. M.; Benac, B. L.; Cowley, A. H.; Kidd, K. B.; Nunn, C. M. *New J. Chem.*, **1988**, *12*, 553.
- (15) Storr, A. J. *J. Chem. Soc. A.*, **1968**, 2605.

- (16) Cowley, A. H.; Jones, R. A.; Mardones, M. A.; Nunn, C. M. *Organometallics*, **1991**, *10*, 1635.
- (17) Atwood, D. A.; Jones, R. A.; Cowley, A. H.; Bott, S. G.; Atwood, J. L. *J. Organometal. Chem.*, **1992**, *434*, 143.
- (18) Arif, A. M.; Benac, B. L.; Cowley, A. H.; Geerts, R. L.; Jones, R. A.; Kidd, K. B.; Power, J. M.; Schwab, S. T. *J. Chem. Soc. Chem. Commun.*, **1986**, 1543.
- (19) Oliver, J. P. *J. Organometal. Chem.*, **1995**, *500*, 269 (See references therein).
- (20) Dillingham, M. D. B.; Burns, J. A.; Byers-Hill, J.; Gripper, K. D.; Pennington, W. T.; Robinson, G. H. *Inorg. Chem. Acta.*, **1994**, *216*, 267.
- (21) Heaton, D. E.; Jones, R. A.; Kidd, K. B.; Cowley, A. H.; Nunn, C. M. *Polyhedron*, **1988**, *7*, 1901.
- (22) Banks, M. A.; Beachley, O. T., Jr.; Buttrey, L. A.; Churchill, M. R.; Fettingner, J. C. *Organometallics*, **1991**, *10*, 1901.
- (23) Wells, R. L.; Self, M. F.; McPhail, A. T.; Aubuchon, S. R. *Organometallics*, **1993**, *12*, 2832.
- (24) Aubuchon, S. R.; McPhail, A. T.; Wells, R. L.; Giambra, J. A.; Bowser, J. R. *Chem. Mater.*, **1994**, *6*, 82.
- (25) Power, M. B.; Barron, A. R.; *J. Chem. Soc. Chem. Comm.*, **1991**, 1315.
- (26) Hoffman, G. G.; Burschka, C. *J. Organomet. Chem.*, **1984**, *267*, 229.
- (27) Ghazi, S. U.; Heeg, M. J.; Oliver, J. P. *Inorg. Chem.*, **1994**, *33*, 4517.
- (28) Hendershot, D. G.; Kumar, R.; Barber, M.; Oliver, J. P. *Organometallics*, **1991**, *10*, 1917.
- (29) Hoffman, G. G.; Burschka, C. *Angew. Chem. Int. Ed. Engl.*, **1985**, *24*, 970.
- (30) Shriver, D. F. "The Manipulations of Air-Sensitive Compounds"; McGraw-Hill: New York, 1968, p. 38.
- (31) Churchill, M. R.; Lashewycz, R. A.; Rotella, F. J. *Inorg. Chem.*, **1977**, *16*, 265.



- (32) Nowacki, W.; Matsumato, T.; Edenharter, A. *Acta. Cryst.*, **1967**, 22, 935.
- (33) Sheldrick, G. M.; SHELXTL PLUS, Release 4.11 (VMS), Siemens Analytical Instrument Corp.; Madison, WI, 1989.
- (34) International Tables for X-Ray Crystallography, Kynoch Press, Birmingham, England, 1974, Vol. 4, pp. 99-101 and 149-150.
- (35) Churchill, M. R. *Inorg. Chem.*, **1973**, 12, 1213.

Table 1. Important Interatomic Distances (Å) and Angles (°) for [Et<sub>2</sub>GaP(t-Bu)<sub>2</sub>]<sub>2</sub>.

## (A) Bond lengths (Å)

Ga(1)-P(1)	2.482 (2)	Ga(1)-C(11)	1.999 (6)
Ga(1)-C(21)	1.998 (6)	Ga(1)-P(1A)	2.483 (2)
P(1)-C(31)	1.904 (6)	P(1)-C(41)	1.900 (6)
P(1)-Ga(1A)	2.483 (2)	C(11)-C(12)	1.380 (12)
C(21)-C(22)	1.418 (13)		

## (B) Bond angles (°)

P(1)-Ga(1)-C(11)	114.8 (2)	P(1)-Ga(1)-C(21)	117.3 (2)
C(11)-Ga(1)-C(21)	107.1 (3)	P(1)-Ga(1)-P(1A)	86.9 (1)
C(11)-Ga(1)-P(1A)	117.1 (2)	C(21)-Ga(1)-P(1A)	113.1 (2)
Ga(1)-P(1)-C(31)	114.2 (2)	Ga(1)-P(1)-C(41)	113.2 (2)
C(31)-P(1)-C(41)	110.2 (3)	Ga(1)-P(1)-Ga(1A)	93.1 (1)
C(31)-P(1)-Ga(1A)	113.2 (2)	C(41)-P(1)-Ga(1A)	112.1 (2)
Ga(1)-C(11)-C(12)	119.9 (6)	Ga(1)-C(21)-C(22)	118.0 (6)

Table 2. Important Interatomic Distances (Å) and Angles (°) for [Et<sub>2</sub>GaS(SiPh<sub>3</sub>)]<sub>2</sub>.

## (A) Bond lengths (Å)

Ga(1)-S(1)	2.450 (2)	Ga(1)-C(1)	1.934 (14)
Ga(1)-C(3)	1.971 (11)	Ga(1)-S(1A)	2.396 (2)
S(1)-Si(1)	2.148 (3)	S(1)-Ga(1A)	2.396 (2)
Si(1)-C(11)	1.861 (9)	Si(1)-C(21)	1.864 (11)
Si(1)-C(31)	1.864 (6)	C(1)-C(2)	1.437 (14)
C(3)-C(4)	1.446 (11)		

## (B) Bond angles (°)

S(1)-Ga(1)-C(1)	112.3 (3)	S(1)-Ga(1)-C(3)	102.1 (3)
C(1)-Ga(1)-C(3)	125.1 (4)	S(1)-Ga(1)-S(1A)	88.1 (1)
C(1)-Ga(1)-S(1A)	108.8 (2)	C(3)-Ga(1)-S(1A)	113.8 (3)
Ga(1)-S(1)-Si(1)	124.0 (1)	Ga(1)-S(1)-Ga(1A)	91.9 (1)
Si(1)-S(1)-Ga(1A)	114.8 (1)	S(1)-Si(1)-C(11)	107.1 (3)
S(1)-Si(1)-C(21)	115.3 (2)	C(11)-Si(1)-C(21)	109.8 (4)
S(1)-Si(1)-C(31)	105.2 (3)	C(11)-Si(1)-C(31)	110.9 (3)
C(21)-Si(1)-C(31)	108.5 (4)	Ga(1)-C(1)-C(2)	119.6 (8)
Ga(1)-C(3)-C(4)	114.4 (8)	Si(1)-C(11)-C(12)	122.7 (8)
Si(1)-C(11)-C(16)	122.0 (7)	C(12)-C(11)-C(16)	115.2 (8)
Si(1)-C(21)-C(22)	121.6 (8)	Si(1)-C(21)-C(26)	122.9 (6)
C(22)-C(21)-C(26)	115.4 (10)	Si(1)-C(31)-C(32)	122.9 (6)
Si(1)-C(31)-C(36)	120.4 (5)	C(32)-C(31)-C(36)	116.7 (6)

Table 3. Data for X-Ray Structural Studies of [Et<sub>2</sub>GaP(t-Bu)<sub>2</sub>]<sub>2</sub> and [Et<sub>2</sub>GaS(SiPh<sub>3</sub>)]<sub>2</sub>.

	[Et <sub>2</sub> GaP(t-Bu) <sub>2</sub> ] <sub>2</sub>	[Et <sub>2</sub> GaS(SiPh <sub>3</sub> )] <sub>2</sub>
molec. formula	C <sub>24</sub> H <sub>56</sub> Ga <sub>2</sub> P <sub>2</sub>	C <sub>44</sub> H <sub>50</sub> Ga <sub>2</sub> S <sub>2</sub> Si <sub>2</sub>
crystal syst.	monoclinic	triclinic
space group	P2 <sub>1</sub> /n (No. 14)	P $\bar{1}$ (No. 2)
a, Å	8.8379 (11)	9.2715 (15)
b, Å	11.5032 (19)	9.5413 (13)
c, Å	14.6460 (28)	14.1872 (22)
$\alpha$ , deg	90.000	71.135 (11)
$\beta$ , deg	91.010 (10)	89.409 (13)
$\gamma$ , deg	90.000	65.247 (10)
V, Å <sup>3</sup>	1488.2 (4)	1067.0 (3)
Z	2	1
molec. wt	546.1	838.6
D(calc), g cm <sup>-3</sup>	1.219	1.305
$\mu$ (Mo K $\alpha$ ), mm <sup>-1</sup>	1.918	1.435
Transmission min/max	0.3808/0.5013	0.8527/0.9445
2 $\theta$ range, deg	5.0 $\leq$ 45.0	5.0 $\leq$ 45.0
index ranges	h -9 $\leq$ +9	h 0 $\leq$ +9
	k -12 $\leq$ +12	k -10 $\leq$ +10
	l -15 $\leq$ +0	l -15 $\leq$ +15

Table 3. Data for X-Ray Structural Studies of  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$  and  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ . (Cont.)

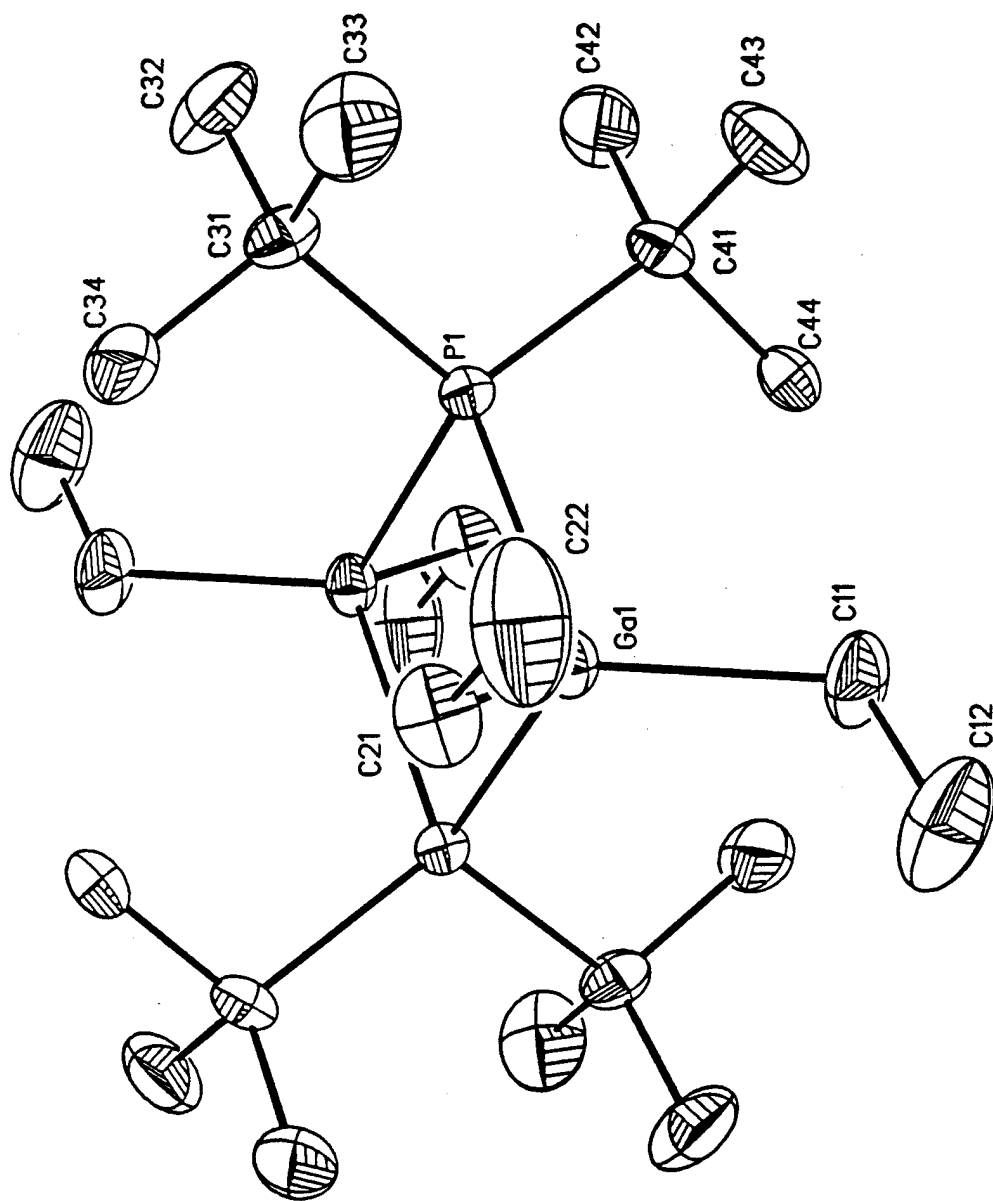
no. of reflns collected	4075	3002
no. of unique reflns	1960 ( $R_{\text{int}} = 2.45\%$ )	2793 ( $R_{\text{int}} = 1.70\%$ )
no. of reflns $> 6\sigma$	1170	1261
R indices ( $6\sigma$ data)	$R = 4.33\%$	$R = 4.36\%$
	$wR = 4.96\%$	$wR = 3.29\%$
R indices (all data)	$R = 7.16\%$	$R = 11.56\%$
	$wR = 7.66\%$	$wR = 4.27\%$
largest diff peak, $\text{e}/\text{\AA}^3$	0.80	0.35
deepest diff peak, $\text{e}/\text{\AA}^3$	-0.48	-0.30

CAPTIONS TO FIGURES

Figure 1. Molecular geometry of  $[\text{Et}_2\text{GaP}(\text{t-Bu})_2]_2$ . ORTEP2 diagram, 30% probability ellipsoids for non-hydrogen atoms and with hydrogen atoms omitted for clarity.

Figure 2. Molecular geometry of  $[\text{Et}_2\text{GaS}(\text{SiPh}_3)]_2$ . ORTEP2 diagram, 30% probability ellipsoids for non-hydrogen atoms with hydrogen atoms artificially reduced.

Fig 1



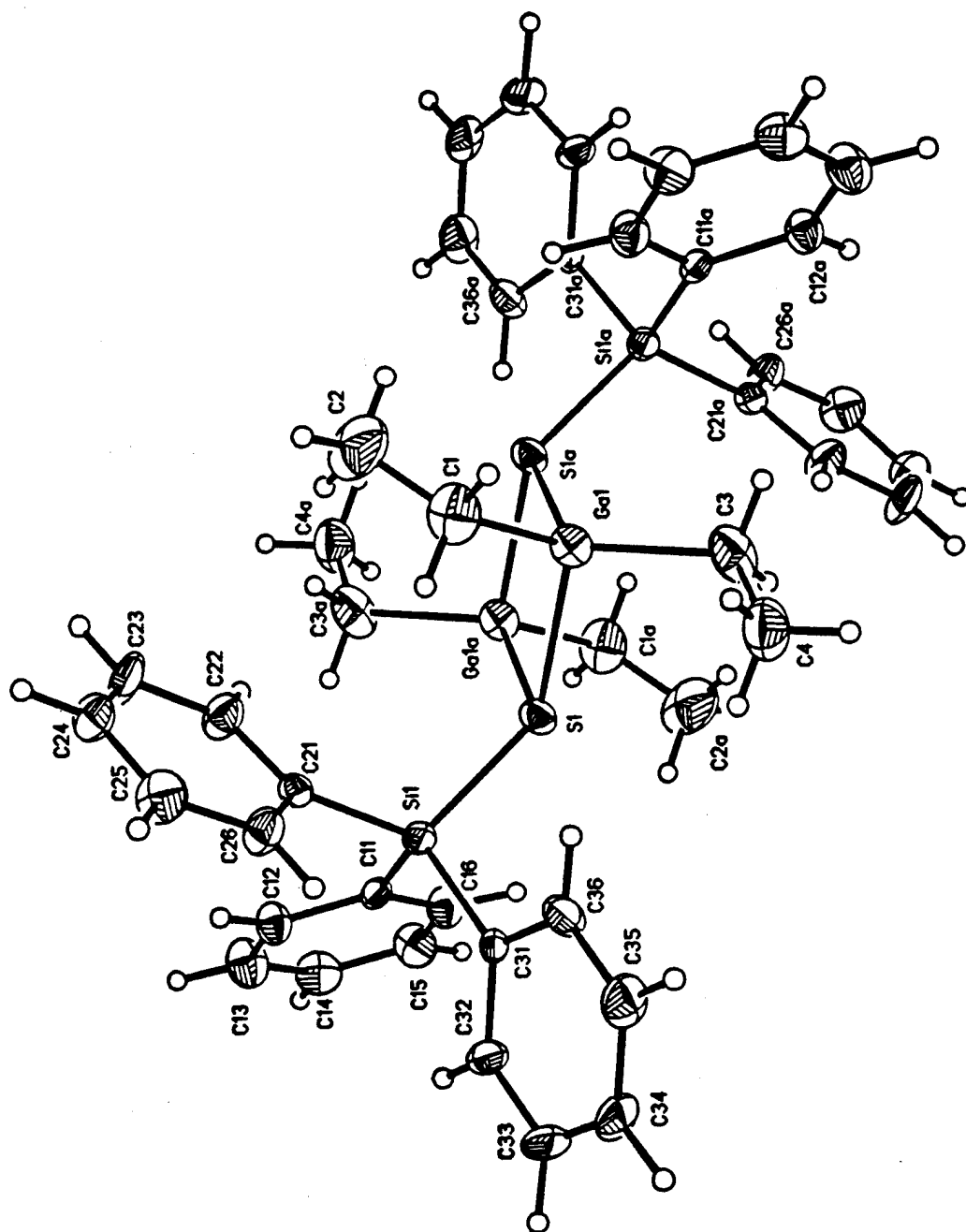


Fig 2



## TECHNICAL REPORT DISTRIBUTION LIST - GENERAL

Office of Naval Research (1)\*  
Chemistry Division, ONR 331  
800 North Quincy Street  
Arlington, Virginia 22217-5660

Dr. Richard W. Drisko (1)  
Naval Facilities & Engineering  
Service Center  
Code L52  
Port Hueneme, CA 93043

Defense Technical Information  
Center (2)  
Building 5, Cameron Station  
Alexandria, VA 22314

Dr. Eugene C. Fischer (1)  
Code 2840  
Naval Surface Warfare Center  
Carderock Division Detachment  
Annapolis, MD 21402-1198

Dr. James S. Murday (1)  
Chemistry Division, Code 6100  
Naval Research Laboratory  
Washington, D.C. 20375-5320

Dr. Bernard E. Douda (1)  
Crane Division  
Naval Surface Warfare Center  
Crane, Indiana 47522-5000

Dr. John Fischer, Director (1)  
Chemistry Division, C0235  
Naval Air Weapons Center  
Weapons Division  
China Lake, CA 93555-6001

Dr. Peter Seligman (1)  
Naval Command, Control and  
Ocean Surveillance Center  
RDT&E Division  
San Diego, CA 92152-5000

\* Number of copies to forward