

50 AD AO 5885 **PROJECT SQUID** TECHNICAL REPORT MIT-90-PU **EXPERIMENTAL AND THEORETICAL STUDIES** OF CHEMICAL DYNAMICS AND INSTABILITIES **IN IRREVERSIBLE PROCESSES** BY **JOC** FILE COPY JOHN ROSS CHEMISTRY DEPARTMENT MASSACHUSETTS INSTITUTE OF TECHNOLOGY 2 Colemn CAMBRIDGE, MASSACHUSETTS 02139 · SEP 18 19 **PROJECT SQUID HEADQUARTERS** CHAFFEE HALL PURDUE UNIVERSITY WEST LAFAYETTE, INDIANA 47907 **AUGUST 1978** Project SQUID is a cooperative program of basic research relating to Jet Propulsion. It is sponsored by the Office of Naval Research and is administered by Purdue University through Contract N00014-75-C1143, NR-098-038. This document has been approved for public release and sale; its distribution is unlimited 78 09 18 054

H SQUID-MIT-90-PU Technical Report MIT-90-PU PROJECT SQUID A COOPERATIVE PROGRAM OF FUNDAMENTAL RESEARCH AS RELATED TO JET PROPULSION OFFICE OF NAVAL RESEARCH, DEPARTMENT OF THE NAVY CONTRACT NO0014-75-C-1143 / NR-098-038 EXPERIMENTAL AND THEORETICAL STUDIES OF CHEMICAL DYNAMICS AND INSTABILITIES IN IRREVERSIBLE PROCESSES . by John/Ross SEP 18 1978 Chemistry Department Massachusetts Institute of Technology Cambridge, Massachusetts 02139 Final rept. 1 Oct 67-31 Dec 77, Project SQUID Headquarters Chaffee Hall -Purdue University West Lafayette, Indiana (12/2pp./ Augure 1978 This document has been approved for public release and sale; its distribution is unlimited 78 09 18 054 403 627 Lu

# ABSTRACT

The final report summarizes the work accomplished under the subcontract. The overall objectives of the investigation were as follows: The determination of molecular properties of chemical dynamics for reactions of importance to combustion and propulsion. Molecular beam techniques were used for the experimental part of this work and were accompanied by theoretical studies in chemical dynamics. The second purpose was the study of the interaction of chemical reactions with transport processes and flows in gases in which instabilities may occur.

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## PROJECT SQUID FINAL REPORT

# A. Identification

Principal Investigator: John Ross, F.G. Keyes Professor of Chemistry

Contractor:	Massachusetts Institute of Technology	
Contract No.:	Sub 4965-10 under Contract N00014-67-0226-0005	
Title:	Experimental and Theoretical Studies of Chemical Dynamics and Instabilities in	
	Irreversible Processes	

B. Duration:

October 1, 1967 - December 31, 1977

## C. Participation

Other Support:

Work has been supported in part by the National Science Foundation (30%) and M. I. T. (20%).

Names of Investigators who contributed to research:

Robert K. Brown	"Jennifer Makowski
<sup>*</sup> Randolph H. Burton	David L. McFadden
Rashmikant C. Desai	Charles Mims
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1

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## SQUID FINAL REPORT Page 2

#### D. Object

<u>Purpose of Research</u>: The determination of molecular properties of chemical dynamics for reactions of importance to combustion and propulsion. Molecular beam techniques were used for the experimental part of this work and were accompanied by theoretical studies in chemical dynamics. The second purpose was the study of the interaction of chemical reactions with transport processes and flows in gases in which instabilities may occur.

#### E. Achievements

1. <u>Transport Processes</u>: We developed a new method of solving the Boltzmann equation which describes transport processes in dilute gases (47). The method is similar to a WKB-type solution and at any stage of approximation is better than a sonine polynomial expansion. Experimental work was completed on measurements of the viscosity of gases as a function of temperature and pressure. The precision and accuracy of these measurements has not been superceded (54).

2. <u>Theory of Chemical Dynamics</u>: We have made progress in this field along a number of lines. We developed the theory of optical potentials for reactive systems (43, 46); we treated reactions by distorted wave approximations (51), semiclassical analysis (56), direct interaction and complex formation approaches (65, 72). The optical potential method has proven best where applicable for the determination of total collision cross sections. In an analysis of symmetry effects in chemical reactions (67) we derived the theoretical basis of the important Woodward-Hoffmann rules and showed their limitations.

3. <u>Molecular Beam Research</u>: In a number of publications (50, 66, 73, 76) we measured angular distributions of both reactants and products in a chemically reactive system and derived from that total reaction cross sections, probabilities of reactions, threshold conditions, such as activation energy and threshold distances necessary for reaction, and distribution of exothermicity in reaction products. This work contributed to showing that a molecular beam approach yields valuable data not available by other techniques.

4. <u>Chemical Instabilities</u>: Chemical Instabilities occur when nonlinear reaction mechanisms, normally achieved with auto- or cross catalysis, are driven sufficiently far from chemical equilibrium. In that case a number of interesting events, such as multiple stationary

## SQUID FINAL REPORT Page 3

states, oscillations and formation of macroscopic spatial structures may occur. We have made pioneering contributions in a number of areas. We showed the existence of resonance-like phenomena due to boundary or external perturbations (69, 70); we showed the importance of local autocatalysis (71) and predicted new cooperative phenomena in a field of local sites at which autocatalytic reactions may occur (89); we investigated the interaction of mechanical (sound) modes with chemical reactions and showed how chemical reactions may amplify sound waves (75, 79), which is of particular importance in combustion; we have predicted the occurrence of instabilities in illuminated systems (82, 90) and proceeded to confirm our predictions with experiments; we analyzed a variety of waves in oscillatory chemical reactions (80, 93, 101); and we discussed the connection between fluctuations and transitions in chemical instabilities as compared to those in phase transitions. Much of this work is reviewed in Ref. 114. Instabilities and oscillatory phenomena occur in flames and other combustion processes and we believe that we have contributed to an understanding of these phenomena.

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"The reaction of photo-excited NO<sub>2</sub> with Cyclopropane," to be submitted to J. Chem. Phys. (with N. Presser, H. Petek and G. Eadens).

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