

# Silicon solar cell efficiency improvement employing the photoluminescent, down-shifting effects of carbon and CdTe quantum dots

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**Abstract** We report the synthesis and characterization of Carbon and CdTe quantum dots (QDs), as well as the observed improvement in the power conversion efficiency (PCE) of photovoltaic devices upon the incorporation of the synthesized aforementioned nanostructures. Even though C quantum dots were observed to have a relatively smaller influence on solar cell performance, they are considered to be a more attractive option due to their affordability and minimal impact in the environment that could ultimately promote their widespread utilization on photovoltaic structures.

**Keywords** Photoluminescence · Quantum dots · Photovoltaics · Solar cells

## Introduction

There is a generalized trend to demonstrate higher solar cell efficiency with more affordable devices to promote their widespread utilization. Silicon remains the most widely employed material in solar cell manufacturing. However, one of the main limitations of silicon solar cells is their relatively low capacity to absorb photons near the band gap energy level [1]. In silicon solar cells those high-energy photons represent intrinsic losses. Improving the capture of those photons that are otherwise considered lost could lead to a dramatic increase in the efficiency of photovoltaic structures [2]. One of the options explored involves the utilization of photoluminescent, down-shifting [3] quantum structures to enable the capture of high-energy photons that tend to interact with lattice phonons rather than generate electron-hole pairs [4, 5].

Therefore, we report the utilization of Carbon [6, 7] and CdTe [8] quantum dots [9] for down shifting effects to capture the aforementioned high-energy photons. These nano crystals are obtained from semiconductor materials [10] and are small enough to exhibit quantum effects.

The electrical properties of these materials are intermediate between bulk semiconductor and discrete molecules and are known to exhibit several favorable characteristics comprising (1) significant interactions between the nano crystals and wavelengths near the ultra-violet range of the solar spectrum, (2) an extraordinary photo stability and (3) a strong relationship of the size and shape of the nanoparticle with its emission and absorption spectrum [11]. Due to the anticipated downshifting effects

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of these nano structures, their utilization in a thin film matrix has the potential to shift the wavelength of the incident radiation into a more favorable solar cell absorption range [12], and achieve an increased power conversion efficiency [13].

## Experimental details

### Synthesis of C QDs

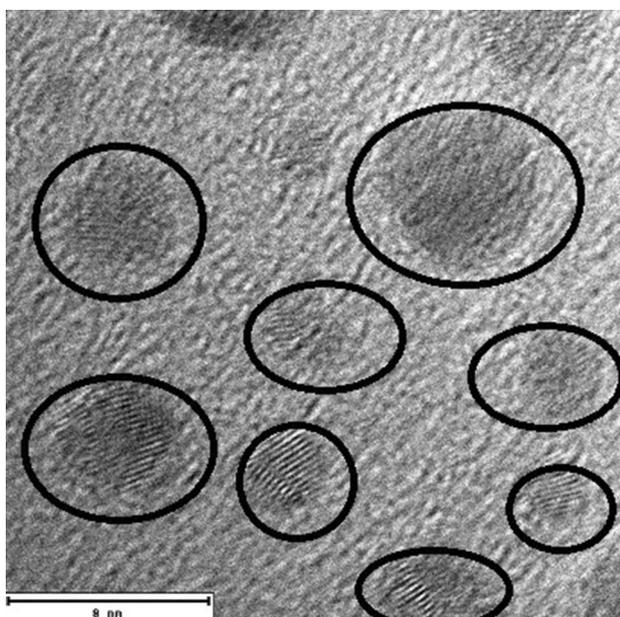
The C nanostructures were synthesized employing an alkali-assisted electrochemical fabrication method utilizing graphite rods for both the anode and the cathode [14, 15] while varying the applied current between 10 and 60 mA. The graphite rods employed had a diameter of 5 mm, a separation of anode to cathode of 25.4 mm and were submerged 30 mm in a 100-ml electrolyte solution composed of ethanol and water with a volume ratio of 99.5/0.5 to which 0.3 g of NaOH was added. Current was applied for 1 h immediately upon the submersion of the graphite rods within the specified current range. Subsequently, the samples were stored for 48 h at room temperature to stabilize them, and the produced solutions were evaporated until obtaining 5 ml for every 100 ml of quantum dot solution. Upon the completion of the evaporation step the samples were separated employing a silica-gel chromatography column with a 100-ml mixture of petroleum ether and diethyl ether with a volume ratio of 30/70. The final step

was to evaporate all the solvents in each vial to increase the C quantum dot concentration (see Fig. 1) [16, 17].

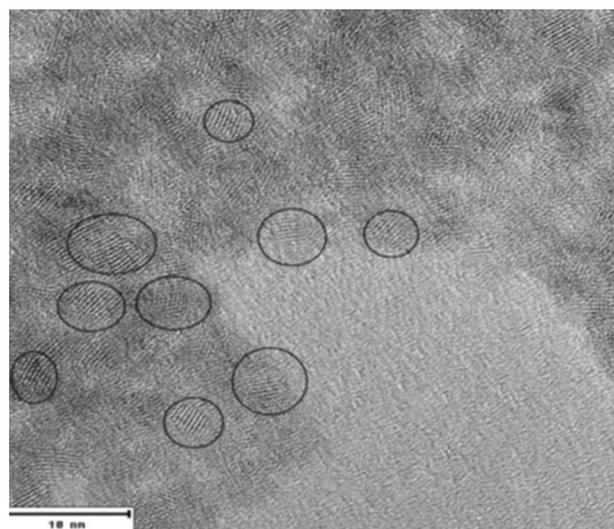
### Synthesis of CdTe QDs

CdTe QDs of measured size between 2 and 18 nm (see Fig. 2) were obtained employing a chemical synthesis scheme. Specifically, 0.0533 g of cadmium acetate dihydrate ( $\text{Cd}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$ , 99.5 %) was dissolved in 50 ml of deionized (DI) water in a 125-ml Erlenmeyer flask; subsequently 18  $\mu\text{l}$  of thioglycolic acid (TGA, 90 %) was added, and the pH was adjusted with 1 M sodium hydroxide (NaOH) set solution until reaching a value between 10.5 and 11 in the pH scale and stirred for 5 min. Separately, 0.0101 g of potassium tellurite ( $\text{K}_2\text{TeO}_3$ , 95 %) was dissolved in 50 ml of DI water in a 125-ml Erlenmeyer flask, stirred for 5 min and 0.0101 g  $\text{K}_2\text{TeO}_3$  was added to this second solution. Subsequently, the previously prepared solutions were mixed, 0.08 g of sodium borohydride ( $\text{NaBH}_4$ , 99.99 %) was added to the mixture and the reaction was allowed to proceed for 5 min. The mixed solution was then transferred to a single-neck, round-bottom flask that was attached to a Liebig condenser, which was stirred at 500 rpm while being refluxed. During the refluxing time, the flask remained submerged in laboratory oil whose temperature was maintained at 100 °C. QD size and photoluminescent emission wavelength were determined by the refluxing time (15, 30 min, 1, 3, 5, 6, 8 and 12 h) [18, 19].

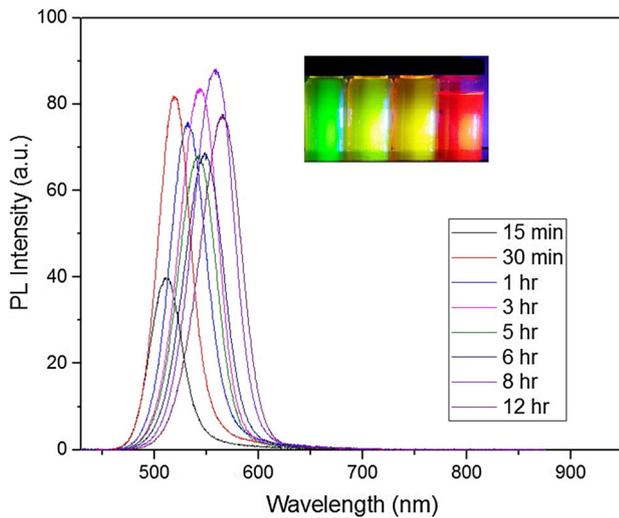
In order to incorporate the QDs on functional solar cells, they must be dispersed in a matrix. For this purpose, polymers are attractive materials for dispersing the nanoparticles for several reasons including affordability,



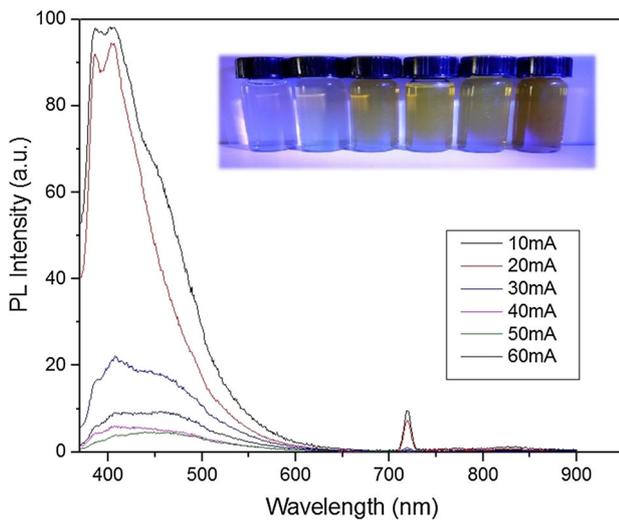
**Fig. 1** TEM image of Carbon QDs synthesized with an applied current of 20 mA



**Fig. 2** TEM image of CdTe QDs synthesized with 8 h reflux

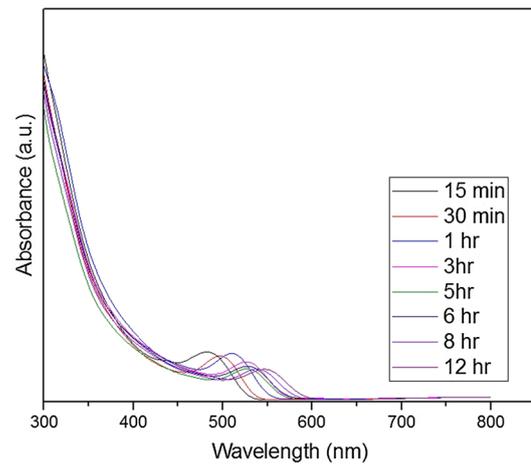


**Fig. 3** Photoluminescence spectra of the synthesized CdTe QDs and the observed color variation under a violet lamp

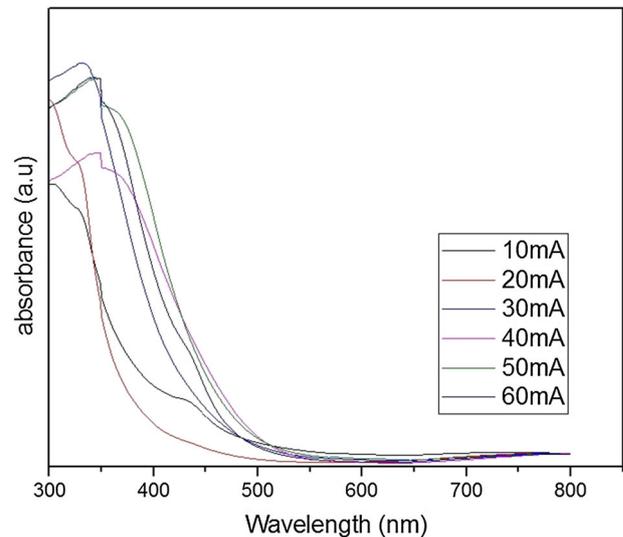


**Fig. 4** Photoluminescence spectra of the synthesized Carbon QDs and the observed color variation under a violet lamp

mechanical flexibility and ease of processing into thin films. Furthermore, most polymers are highly transparent at wavelengths longer than 400 nm; thus, nanocomposites based on nanoparticle dispersion in polymers are a promising option for integrated optical devices. In this effort we selected Polymethylmethacrylate (495 PMMA A2 from Microchem) as the matrix in which to disperse the Carbon and CdTe QDs on the basis of its high transparency (93 %), low fragility, high weather and UV resistance and excellent thermal insulation. Upon nanostructure dispersion, the PMMA + QD thin films were spin cast on the window side of 52 × 38 mm, commercially available



**Fig. 5** CdTe QD absorption spectra for different refluxing times



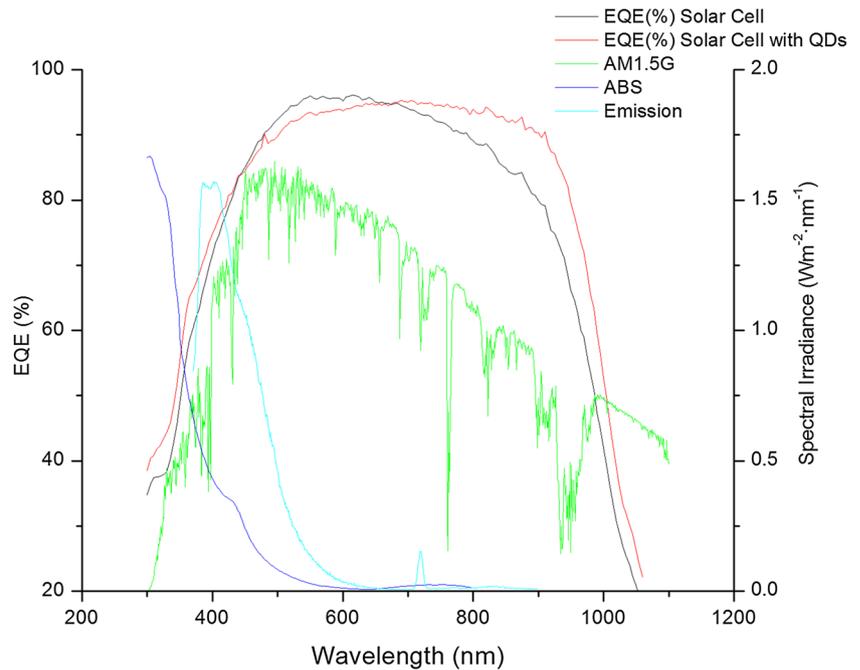
**Fig. 6** Absorption spectra of Carbon QDs obtained with different synthesis currents

polysilicon solar cells (Eco-worthy Company) with thickness of 200 μm.

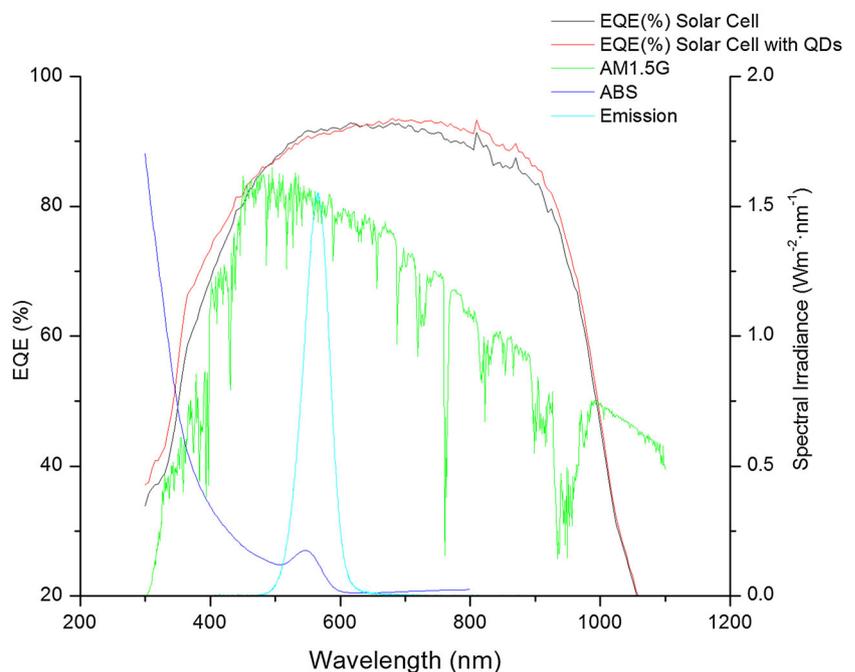
The dispersion of Carbon QDs in PMMA was obtained by adding to the dried nano crystals produced in the synthesis process previously described 1 ml of PMMA. In the case of the CdTe QDs, 3 ml of the synthesized nanostructures solution was mixed with acetone (volume ratio 1:1) and centrifuged at 10,000 rpm for 10 min to promote sedimentation. Upon removal of the liquid portion with a pipette, 1 ml of PMMA was added to the vial and sonicated for 5 min to disperse the quantum dots.

Spin casting was carried out by starting at 300 rpm for 10 s followed by a set speed of 4000 rpm and holding for 45 s. Upon completion of the spin coating process, the

**Fig. 7** Overlay of the AM1.5G spectrum and solar cell EQE before and after the addition of Carbon QDs. The absorption and emission spectra are also included as a reference



**Fig. 8** Overlay of the AM1.5G spectrum and solar cell EQE before and after the addition of CdTe QDs. The absorption and emission spectra are also included as a reference



solar cells with the incorporated thin film were annealed at 180 °C for 60 s on a hot plate to allow the evaporation of the solvents. The optical and electrical properties of the photovoltaic devices produced were collected and compared with the data obtained before the application of the PMMA + QDs thin films.

### Quantum dot characterization

The size of the synthesized CdTe and Carbon QDs was controlled by the duration of the refluxing step for the former and the applied current in the latter case. Size was monitored by collecting their absorption and

**Table 1** Average performance of solar cells with and without the addition of Carbon QDs

Poly-silicon solar cell	Voc (mV)	Jsc (mA/cm <sup>2</sup> )	FF (%)	PCE (%)
Sample 1 w/o QDs	616.62	36.92	57.92	13.21 ± 0.22
Sample 1 QDs 10 mA	619.91	37.27	58.73	13.56 ± 0.28
Sample 2 w/o QDs	613.91	36.13	56.14	12.45 ± 0.63
Sample 2 QDs 20 mA	616.99	36.53	57.67	13.03 ± 0.58
Sample 3 w/o QDs	618.76	36.93	58.49	13.37 ± 0.36
Sample 3 QDs 30 mA	621.93	37.32	59.38	13.71 ± 0.36
Sample 4 w/o QDs	619.81	36.77	60.25	13.72 ± 0.44
Sample 4 QDs 40 mA	622.48	37.12	60.59	14.00 ± 0.21
Sample 5 w/o QDs	620.85	36.90	59.48	13.62 ± 0.34
Sample 5 QDs 50 mA	625.23	37.12	60.57	14.06 ± 0.26
Sample 6 w/o QDs	617.15	36.90	57.72	13.13 ± 0.15
Sample 6 QDs 60 mA	622.96	36.49	59.44	13.51 ± 0.08

photoluminescence spectra. Ostensibly the color of CdTe QDs can be adjusted from green to orange-yellow by the refluxing time. Thus, a refluxing time of 15 min exhibits a characteristic emission peak at 510 nm, which shifts to 585 nm for a refluxing time of 12 h, both collected with an excitation wavelength of 395 nm (see Fig. 3).

Variations in the applied current determine the size of the synthesized carbon nanoparticles such that the colors observed [20, 21], with an excitation with a wavelength of 360 nm, vary from violet (400 nm) to green (550 nm) [22] with the maximum at approximately 420 nm (see Fig. 4) [23]. Fluorescence measurements were carried using an AMINCO-Bowman Series 2 at room temperature.

CdTe QDs solutions exhibit an absorption spectra extending from approximately 450 nm to as much as 600 nm for those samples with the longest refluxing time (see Fig. 5). Carbon QD solutions shows broad UV/Vis absorption with a characteristic curve at 340 nm (see Fig. 6). The UV/Vis absorption spectra were recorded with a UV-VIS-NIR Cary 5000 spectrometer.

### Solar cell characterization

Solar cell performance was quantified with an Oriel Sol2A solar simulator under standard testing conditions. Specifically, measurements were collected before and after the deployment of the PMMA + QD layer for the samples of Carbon QDs synthesized with different current values from 10 to 60 mA, and the CdTe QDs were obtained with refluxing times of 15, 30 min, 1, 3, 5, 6, 8 and 12 h, respectively. The external quantum efficiency (EQE) was measured with an Oriel Quantum Efficiency Measurement kit (QE-PV-SI) using a spot size of approximately 2 mm<sup>2</sup>.

Solar cells with Carbon or CdTe PMMA + QD layers exhibited higher EQE values across the range of interest of the solar spectrum [24] (see Figs. 7, 8).

Table 1 summarizes the measured Power Conversion Efficiency without and with carbon QDs. The collected measurements indicate that there is a modest but not negligible increase in the PCE when the CQDs are added, for all values of the current employed during QD synthesis.

**Table 2** Average performance of solar cells with and without CdTe QDs

Poly-silicon solar cell	Voc (mV)	Jsc (mA/cm <sup>2</sup> )	FF (%)	PCE (%)
Sample 1 w/o QDs	620.76	36.38	58.82	13.30 ± 0.37
Sample 1 QDs 15 min	623.40	36.81	58.88	13.48 ± 0.04
Sample 2 w/o QDs	620.51	37.09	56.59	13.03 ± 1.95
Sample 2 QDs 30 min	624.12	37.72	59.06	13.91 ± 0.22
Sample 3 w/o QDs	621.91	37.19	57.00	13.17 ± 0.67
Sample 3 QDs 1 h	625.18	37.16	59.27	13.77 ± 0.49
Sample 4 w/o QDs	618.73	37.09	56.72	13.01 ± 0.50
Sample 4 QDs 3 h	620.23	37.03	57.75	13.28 ± 0.04
Sample 5 w/o QDs	612.92	36.51	58.55	13.11 ± 0.65
Sample 5 QDs 5 h	615.51	36.66	59.46	13.42 ± 0.33
Sample 6 w/o QDs	607.31	36.32	59.02	13.02 ± 0.30
Sample 6 QDs 6 h	610.71	36.14	61.66	13.62 ± 0.07
Sample 7 w/o QDs	615.35	37.13	57.65	13.18 ± 0.45
Sample 7 QDs 8 h	623.45	37.26	59.06	13.73 ± 0.15
Sample 8 w/o QDs	618.41	37.53	58.53	13.58 ± 0.45
Sample 8 QDs 12 h	625.26	37.56	59.2	13.90 ± 0.15

Table 2 summarizes the measured Power Conversion Efficiency with and without the incorporation of CdTe QDs. The collected measurements indicate that there is a modest but not negligible increase in the PCE when the CdTe QDs are added, regardless of the refluxing time employed during QD synthesis.

## Conclusions

The synthesis and characterization of Carbon quantum dots produced with different applied currents (10–60 mA) and CdTe QDs using various refluxing times (15 min–12 h) corroborated the absorption of relatively high energy

photons and the emission of lower energy photons in the range of 350–550 nm for carbon QDs and between 480 and 620 nm for CdTe QDs. Ostensibly, the size and the corresponding emission spectra of the aforementioned QDs can be adjusted by varying the current for Carbon and the reflux time for the colloidal chemical synthesis of CdTe. These schemes are considered to be relatively simple, affordable and reproducible methods in both cases.

The experimental results of the incorporation of PMMA thin films with downshifting Carbon CdTe quantum dots on the window side of Poly-Silicon solar cells indicate in all cases a non-negligible increase in the power conversion efficiency (PCE). These results corroborate that the utilization of luminescent down shifting nanoparticles may

enable affordable, high-efficiency devices in the near future.

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