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A CLASS OF TWO-PHASE STATE EQUATIONS

by

G. D. Kahl

March 1969

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BALLISTIC RESEARCH LABORATORIES

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G. D. Kahl

Exterior Ballistics Laboratory

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B A L L I S T I C R E S E A R C H L A B O R A T O R I E S

REPORT NO. 1429

GDKahl/so
Aberdeen Proving Ground, Md.
March 1969

A CLASS OF TWO-PHASE STATE EQUATIONS

ABSTRACT

The properties of a particular class of elementary, two-phase state equations are found to be simply related. The class includes some familiar equations, such as the Van der Waals, Berthelot, and Clausius types, as well as a form recently suggested by J. J. Martin. Conditions for stable equilibrium are given, and some examples are compared. A table of values is included which can be used with elementary transformations to find the equilibrium coexistence properties of any member of the set. A drawback for this type equation along the critical isotherm is noted.

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I. INTRODUCTION

There exists a substantial number of elementary pressure state equations which predict a liquid-vapor phase transition and a critical point. It is the purpose of this paper to show that many of these, though not all, belong to a single class, and that the various thermodynamic predictions between class members are simply related. Some familiar equations belong to this group, including the Van der Waals-Maxwell, [V-M], equation, which will serve as the basis for comparison. A direct method can be demonstrated relating the limiting values of thermodynamic quantities in the critical region.

The form of equation treated here is

$$P(V,T) = (RT/[V-b(T)]) - a(T)/[V+c(T)]^2 \quad (1)$$

where $V \geq b(T) > -c(T)$, and $T > 0$; P , V , and T are pressure, specific volume and temperature, respectively. R is the gas constant, and $a(T)$, $b(T)$ and $c(T)$ are analytic functions of positive T . Some restrictions are needed on the latter functions. For instance, $b(T)$ is assumed to be positive and larger than $-c(T)$; and $a(T)$ must be positive for all T less than some value exceeding the critical temperature, T_c , in order to produce a critical point. It is shown in the Appendix that $b(T)$ must be a constant, which is henceforth assumed.

Eq. (1) is expressed in reduced form by use of the scaled variables $\beta \equiv P/P_c$, $\gamma \equiv T/T_c$, $\eta \equiv 1/\alpha \equiv V/V_c$ (this notation corresponds to that of Reference [1], with subscript c denoting critical value) as

$$\beta(\eta,\gamma) = (1/Z_c) \{ (\gamma/[\eta-B]) - (9/8)A(\gamma)/[\eta+C(\gamma)]^2 \} \quad (2)$$

where $\eta \geq B > -C(\gamma)$ and $\gamma > 0$, and $Z_c \equiv P_c V_c / (RT_c)$ is the critical compressibility. The transformations from a , b , and c to A , B , and C are $(9/8) A(\gamma) = a(T)/[RT_c V_c]$, $B = b/V_c$ and $C(\gamma) = c(T)/V_c$. It follows that $A(\gamma)$ is positive for a range of γ larger than unity. The variables

β , γ , η and α all are unity at the critical point. It is seen that Eq. (1) approximates the ideal gas equation when V becomes very large.

In order that the state of the system be stable, certain relations must hold for the thermodynamic functions^{2†}. These are: $C_v > 0$, where C_v is the specific heat at constant volume; and $\left(\frac{\partial \beta}{\partial \eta}\right)_\gamma < 0$, except at the critical point. At the critical point, the conditions become $\left(\frac{\partial \beta}{\partial \eta}\right)_\gamma = 0 = \left(\frac{\partial^2 \beta}{\partial \eta^2}\right)_\gamma$.

Imposing these latter conditions at the critical point, one obtains the relations:

$$A_1 = 1+C_1; B_1 = (1/3)(1-2C_1); \text{ and } Z_c = (3/8)/(1+C_1) \quad (3)$$

where a subscript 1 on A, B, or C denotes the value at the critical temperature, $\gamma = 1$. Thus, the function $C(\gamma)$ fixes the critical compressibility, Z_c .

One observes that for $\gamma < 1$, there is a domain of (η, γ) where $\frac{\partial \beta}{\partial \eta}$ from Eq. (2) is positive, contrary to the stability requirement. Thus, Eq. (2) cannot represent a stable thermodynamic pressure function in this domain; instead, the vapor pressure function, $\beta_A(\gamma)$, independent of η , replaces Eq. (2) as the pressure function here. The usual method of defining $\beta_A(\gamma)$ so it is continuous with the $\beta(\eta, \gamma)$ of Eq. (2) is by means of the Maxwell rule: for any fixed $\gamma \leq 1$,

$$\int_{\eta_3(\gamma)}^{\eta_1(\gamma)} \beta(\eta, \gamma) d\eta = \beta_A(\gamma) \{ \eta_1(\gamma) - \eta_3(\gamma) \}, \quad (4)$$

with

$$\beta(\eta_1, \gamma) = \beta(\eta_3, \gamma) = \beta_A(\gamma) .$$

†References are listed on page 27.

The quantities $\eta_3(\gamma)$ and $\eta_1(\gamma)$ are points on the saturated liquid and vapor loci, respectively. For any fixed $\gamma \leq 1$, the domain $\eta_3 \leq \eta \leq \eta_1$ is the region where liquid and vapor coexist at the same pressure, temperature, and chemical potential. The boundaries of the coexistence region are also defined by Eq. (4). The use of the Maxwell rule implies that the Helmholtz free energy is a unique analytic function of V and T where Eq. (1) is the proper pressure function³. With this analyticity, Eq. (4) insures that the pressure and chemical potential are continuous functions of V and T across the coexistence boundaries. To avoid confusion, a subscript a will hereafter be used on a thermodynamic variable within the coexistence region, while a subscript b will refer to the value outside coexistence.

II. TRANSFORMATIONS

For compactness, let

$$f(\gamma) = 3[B+C(\gamma)]/A(\gamma) . \quad (5)$$

For $\gamma > 1$, $\frac{\partial \beta_a}{\partial \eta}$ must never be positive, or the stability requirement is violated. From Eq. (2),

$$Z_c \frac{\partial \beta_a}{\partial \eta} = - (\gamma/[\eta-B]^2) \{1-G(\eta,\gamma)\}, \quad (6)$$

where

$$G(\eta,\gamma) = (9/4) A(\gamma) [\eta-B]^2 / (\gamma[\eta+C]^3) . \quad (7)$$

Since both Z_c and γ are positive, Eq. (6) shows $\frac{\partial \beta_a}{\partial \eta}$ is negative, providing $G(\eta,\gamma)$ is less than 1 for any $\eta \geq B$. This latter condition is easily satisfied when $A(\gamma)$ is negative; but $A(\gamma)$ was previously required to be positive for all γ less than some value γ_0 exceeding one. In this range of γ , i.e., ($0 < \gamma < \gamma_0$), $G(\eta,\gamma)$ is positive and has a maximum value with respect to η , $G_m(\gamma)$, at $\eta = 3B+2C(\gamma)$, where

$$G_M(\gamma) = A(\gamma)/[3\gamma(B+C)] = 1/[\gamma f(\gamma)]. \quad (8)$$

Forcing $\gamma f(\gamma) > 1$ for $\gamma_0 > \gamma > 1$ will insure that $\frac{\partial \beta_B}{\partial \eta}$ is non-positive.

For $\gamma < 1$, one wishes to satisfy Eq. (4) in the (η, γ) domain of coexistence. The mathematical function $\frac{\partial \beta(\eta, \gamma)}{\partial \eta}$ of Eq. (2) must therefore have zeros in this domain, so $\gamma f(\gamma) < 1$ for $\gamma < 1$. At $\gamma = 1$, $\gamma f(\gamma)$ is unity, as may be seen from the relations already given for $\gamma = 1$.

Defining

$$X(\gamma) \equiv (3/8)/[Z_c f^2(\gamma)A(\gamma)], \quad (9)$$

and the transformations

$$u(\gamma) \equiv \gamma f(\gamma), \quad (10)$$

$$v(\eta, \gamma) \equiv [\eta + C(\gamma)]/[f(\gamma)A(\gamma)], \quad (11)$$

$$w(\eta, \gamma) \equiv \beta(\eta, \gamma)/X(\gamma), \quad (12)$$

$$w_A(\gamma) \equiv \beta_A(\gamma)/X(\gamma), \quad (13)$$

with the convention that $w(\eta, \gamma)$ has a subscript a when $\beta(\eta, \gamma)$ does, one finds that Eq. (2) transforms into

$$w(v, u) = \{8u/(3v-1)\} - 3/v^2. \quad (14)$$

Moreover, for any fixed $u \leq 1$, Eq. (4) transforms into

$$\int_{v_3(u)}^{v_1(u)} w(v, u) dv = w_A(u) \{v_1(u) - v_3(u)\} \quad (15)$$

with $w(v_1, u) = w_B(v_1, u) = w(v_3, u) = w_B(v_3, u) = w_A(u)$.

Eq. (14) has exactly the form of the reduced Van der Waals pressure function, providing one interprets w , v , and u as reduced pressure, volume, and temperature respectively. All the functions w_B , w_A , u , and v , are unity at the critical point where $\eta = 1$ and $\gamma = 1$.

The Jacobian, $J = \frac{\partial(u,v)}{\partial(\gamma,\eta)}$ of the transformations of the independent variables γ and η to the new independent variables u and v reduces to $\frac{\partial v}{\partial \eta} \frac{\partial u}{\partial \gamma}$ because u is independent of η . Using Eqs. (10) and (11) this Jacobian is

$$J = \{1+\gamma \frac{df}{d\gamma}/f(\gamma)\}/A(\gamma), \quad (16)$$

and where J or J^{-1} is neither zero nor infinite, the mapping is single valued. It will be assumed in the further development that for $\gamma \leq 1$ this mapping is single valued, so the necessary conditions on $A(\gamma)$, B and $C(\gamma)$ to accomplish this purpose will be assumed henceforth. The thermodynamic properties of any state equation of this class for $\gamma \leq 1$ can now be expressed in terms of those corresponding properties known for the [V-M] system, and the given transformation Eqs. (10-13).

The method for finding the properties in the coexistence region, including their limits at the critical point, is that demonstrated by Barieau¹. For fixed $u \leq 1$, Eq. (14) becomes a cubic polynomial in v ; one finds from it the largest and smallest values of v ($v_1(u)$ and $v_3(u)$, respectively) which simultaneously satisfy Eq. (15). Letting $\xi \equiv 1/v$, one then obtains, for $u \leq 1$,

$$w_A(\xi_1, \xi_3) = \xi_1 \xi_3 (3 - \xi_3 - \xi_1), \quad (17)$$

$$8u(\xi_1, \xi_3) = (\xi_1 + \xi_3)(3 - \xi_1)(3 - \xi_3). \quad (18)$$

Differentiating both sides of Eq. (15) by u results in the differential equation:

$$u \frac{dw_A}{du} = w_A + 3\xi_1 \xi_3 = \xi_1 \xi_3 (6 - \xi_3 - \xi_1); \quad (19)$$

similarly differentiating Eqs. (17) and (18), and using Eq. (17) gives one

$$u \frac{d\xi_1}{du} = \xi_1(3-\xi_1)/(3-2\xi_1-\xi_3), \quad (20)$$

$$u \frac{d\xi_3}{du} = \xi_3(3-\xi_3)/(3-2\xi_3-\xi_1). \quad (21)$$

The relation

$$u \frac{d^2 w_A}{du^2} = 3 \frac{d(\xi_1 \xi_3)}{du}, \quad (22)$$

follows directly from Eq. (19). Eqs. (17-22) have the same form as the corresponding ones of Reference [1] used to display properties of the [V-M] fluid, providing w_A , u , and ξ here are replaced by β , γ , and α to conform with the notation there.

Proceeding analogously, and defining $y \equiv \xi_3 - 1$ and $x \equiv 1 - \xi_1$, one can use Eqs. (14), (15), (17) and (18) to obtain y as a function of x ; this function can be expanded in the power series

$$y = \sum_{i=1}^{\infty} a_i x^i, \quad (23)$$

where Barieau¹ finds $a_1=1$, $a_2=1/5$, and $a_3=1/25$. The value of a_4 was subsequently evaluated by the present author to be 19/350.

All quantities ξ_1 , ξ_3 , w_A , etc., (as functions of u) have been tabulated previously for the [V-M] system. They may be used to find limiting thermodynamic values in the critical region for any equation of the class considered here; a brief table is in the Appendix.

Some inverse transformations are

$$\eta_j = 1/\alpha_j = v_j f(\gamma) A(\gamma) - C(\gamma), \quad (j=1,3) \quad (24)$$

$$\beta_B(\eta, \gamma) = w_B(v, u)X(\gamma) \quad (25)$$

$$\beta_A(\gamma) = w_A(u)X(\gamma) \quad (26)$$

$$\left\{ \frac{\partial \beta_B(\eta, \gamma)}{\partial \gamma} \right\}_{\eta} = X(\gamma) \left\{ \left(\frac{\partial w_B}{\partial v} \right)_u \left(\frac{\partial v}{\partial \gamma} \right)_{\eta} + \left(\frac{\partial w_B}{\partial u} \right)_v \left(\frac{\partial u}{\partial \gamma} \right)_{\eta} \right\} + w_B(v, u) \frac{dX}{d\gamma} \quad (27)$$

$$\frac{d\beta_A(\gamma)}{d\gamma} = X(\gamma) \left\{ \frac{dw_A}{du} \left(\frac{\partial u}{\partial \gamma} \right)_{\eta} \right\} + w_A(u) \frac{dX}{d\gamma} \quad (28)$$

From Eqs. (27) and (28), it can be seen that $\left\{ \frac{\partial \beta_B(1,1)}{\partial \gamma} \right\}_{\eta}$ is equal to $\frac{d\beta_A(1)}{d\gamma}$, because it is known from the Van der Waals function that at the point $(v, u) = (1, 1)$: (a), w_B and w_A are each equal to one; (b), $\left(\frac{\partial w_B(1,1)}{\partial u} \right)_v$ and $\frac{dw_A(1)}{du}$ are each equal to four; and (c), $\left(\frac{\partial w_B}{\partial v} \right)_u$ is zero.

Therefore, for any equation of this class, the temperature derivative of the vapor pressure function is continuous at the critical point with that of the critical isometric, in reduced variables. This continuity does not necessarily hold for higher order derivatives.

III EXAMPLES AND COMPARISONS

When $A(\gamma)$, B and $C(\gamma)$ are specified, one can recognize some well-known state equations. Some examples are shown in Table I. Example 6, with $A(\gamma)$ a linear function of γ , comes from an equation recently suggested by Martin^{4*}. When $C(\gamma)$ is a constant, as is true for all equations in Table I, one can use the definitions to show that the product $f(\gamma)X(\gamma)$ is unity. In the following, certain thermodynamic predictions for some of these examples will be compared.

A. Coexistence Boundary.

A comparison is made in Figure 1 of the reduced density on the coexistence boundary for the state equations numbered 2, 4, and 6 in

Table I. Example state equations and transformations

a. Shown in the form $P(V,T) = [RT/(V-b)] - a(T)/(V+c)^2$, with constants a_j , b_j and c_j for the number (j) equation. For all shown {except No. (1), where both b and a(T) are zero}, $b_j = V_c [1-1/(4Z_c)]$ and $c_j = V_c [-1+3/(8Z_c)]$; $n_j \equiv (9/64)(RT_c)^2/P_c$.

No. (j)	Name	P(V,T)	a_j	Z_c
(1)	Ideal Gas	RT/V	- Undefined -	
(2)	Van der Waals	$[RT/(V-b_2)] - a_2/V^2$	$3 n_2$	$= 3/8$
(3)	(translated) Van der Waals	$[RT/(V-b_3)] - a_3/(V+c_3)^2$	$3 n_3$	$> 1/4$
(4)	Berthelot	$[RT/(V-b_4)] - a_4/(TV^2)$	$3 n_4 T_c$	$= 3/8$
(5)	Clausius	$[RT/(V-b_5)] - a_5/[T(V+c_5)^2]$	$3 n_5 T_c$	$> 1/4$
(6)	Martin ⁴ (A)	$[RT/(V-b_6)] - a_6(4 T_c - T)/(V+c_6)^2$	n_6/T_c	$> 1/4$
(7)	Martin (B)	$[RT/(V-b_7)] - a_7(5 T_c - 2T)/(V+c_7)^2$	n_7/T_c	$> 1/4$

b. Same state equations {omitting No. (1)} in the reduced form $\beta = (1/Z_c) \{ [\gamma/(\eta-B)] - (9/8)A(\gamma)/(\eta+c)^2 \}$, where $\beta = P/P_c$, $\gamma = T/T_c$ and $\eta = T/T_c$. For all shown here, $B = 1 - 1/(4Z_c)$, $C = -1 + 3/(8Z_c)$ {independent of γ }, and $X(\gamma) = 1/f(\gamma)$.

No. (j)	$Z_c A(\gamma)$	Z_c	$f(\gamma)$	$\left(\frac{d\beta_A}{d\gamma}\right)_c \left\{ = \left(\frac{\partial\beta_A}{\partial\gamma}\right)_c \right\}$
(2)	3/8	$= 3/8$	1	4
(3)	3/8	$> 1/4$	1	4
(4)	$3/(8\gamma)$	$= 3/8$	γ	7
(5)	$3/(8\gamma)$	$> 1/4$	γ	7
(6)	$(4-\gamma)/8$	$> 1/4$	$3/(4-\gamma)$	5
(7)	$(5-2\gamma)/8$	$> 1/4$	$3/(5-2\gamma)$	6

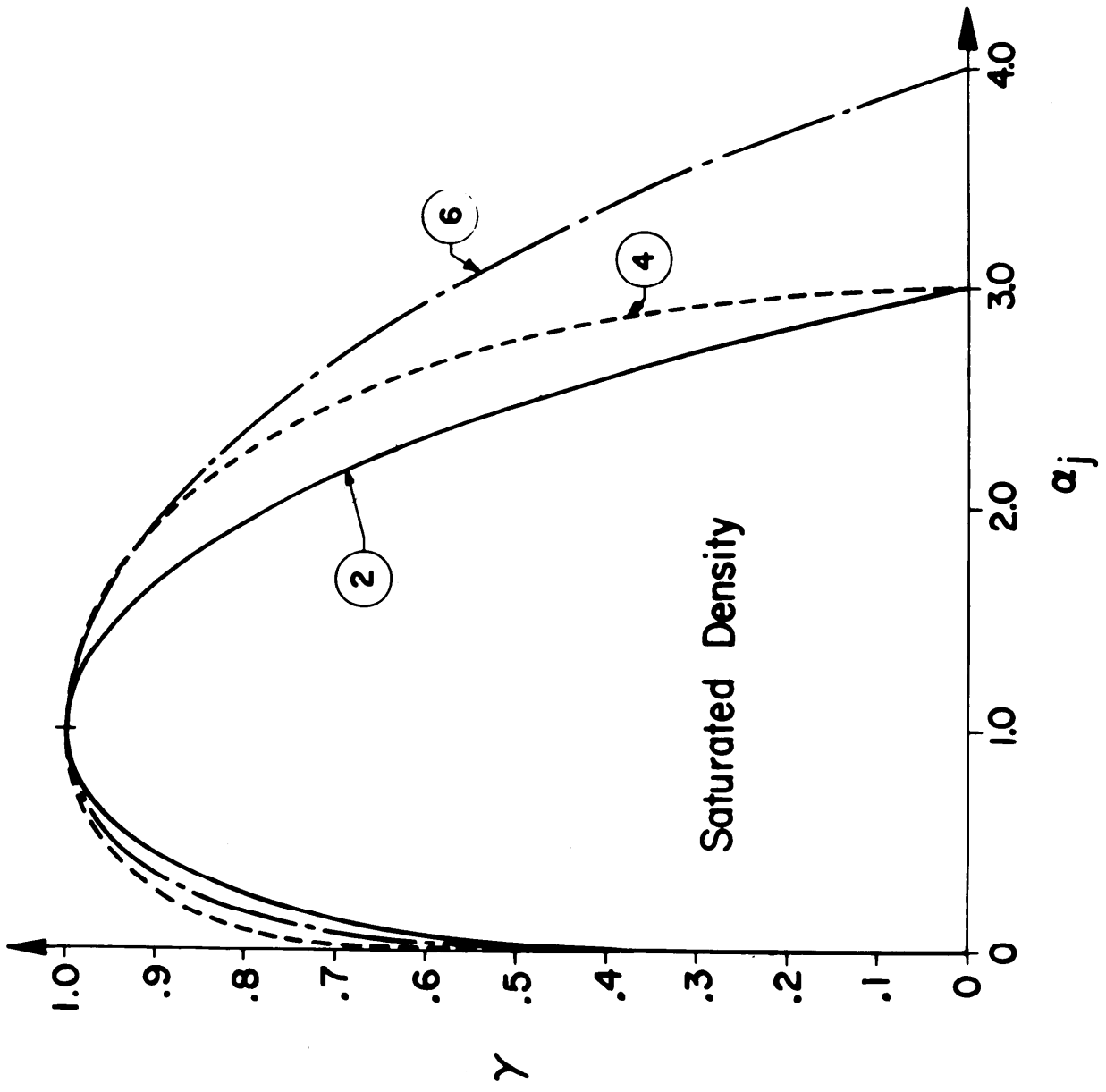


Figure 1. Reduced temperature vs. saturation densities. Numbers refer to state equations in Table I. For (6), the value of Z_c is assumed to be $1/3$.

Table I, using the given functions with Eq. (24); similar plots for examples 3 and 5 of Table I can easily be obtained from those of 2 and 4 respectively, when the value of C_1 is assigned.

B. Vapor Pressure, ($\gamma \leq 1$).

Eq. (26) gives the vapor pressure, $\beta_A(\gamma)$, as a function of $w_A(u)$, the [V-M] vapor pressure at the reduced temperature u , and the specific transforming function $X(\gamma)$. These are shown in Figure 2. For state equations having the same $f(\gamma)$ functions and the same $X(\gamma)$ functions, the vapor pressure is the same function of γ . Thus, the state equations numbered 2 and 3 in Table I have the same vapor pressure, as do those numbered 4 and 5.

C. Heat of Vaporization.

For $T \leq T_c$, the latent heat of vaporization $\Delta Q_A(T)$ is found by using the Clapeyron relation, $\Delta Q_A(T) = T \frac{dP_A}{dT} [V_1(T) - V_3(T)]$; in reduced variables this becomes

$$\Delta Q_A(T)/[RT_c] = Z_c \gamma \frac{d\beta_A}{d\gamma} [\eta_1(\gamma) - \eta_3(\gamma)] . \quad (29)$$

Using the inverse transformations,

$$\begin{aligned} \Delta Q_A(T)/[RT_c] = \{ \Delta Q_A(T)/[RT_c] \}_0 \left\{ \left(\frac{\partial u}{\partial \gamma} \right) \eta / f^2(\gamma) \right\} \\ + \{ u w_A(u) [v_1(u) - v_3(u)] \} \left\{ Z_c A \frac{dX}{d\gamma} \right\} \end{aligned} \quad (30)$$

where the first factor on the right is the latent heat of vaporization for the [V-M] equation at the reduced temperature u . The functions $w_A(u)$, $v_1(u)$ and $v_3(u)$ are also known; as with the vapor pressure function, the heat of vaporization is the same function of γ for state equations having the same $f(\gamma)$ and $X(\gamma)$. Comparisons are shown in Figure 3; the values for the Berthelot and Clausius equations become

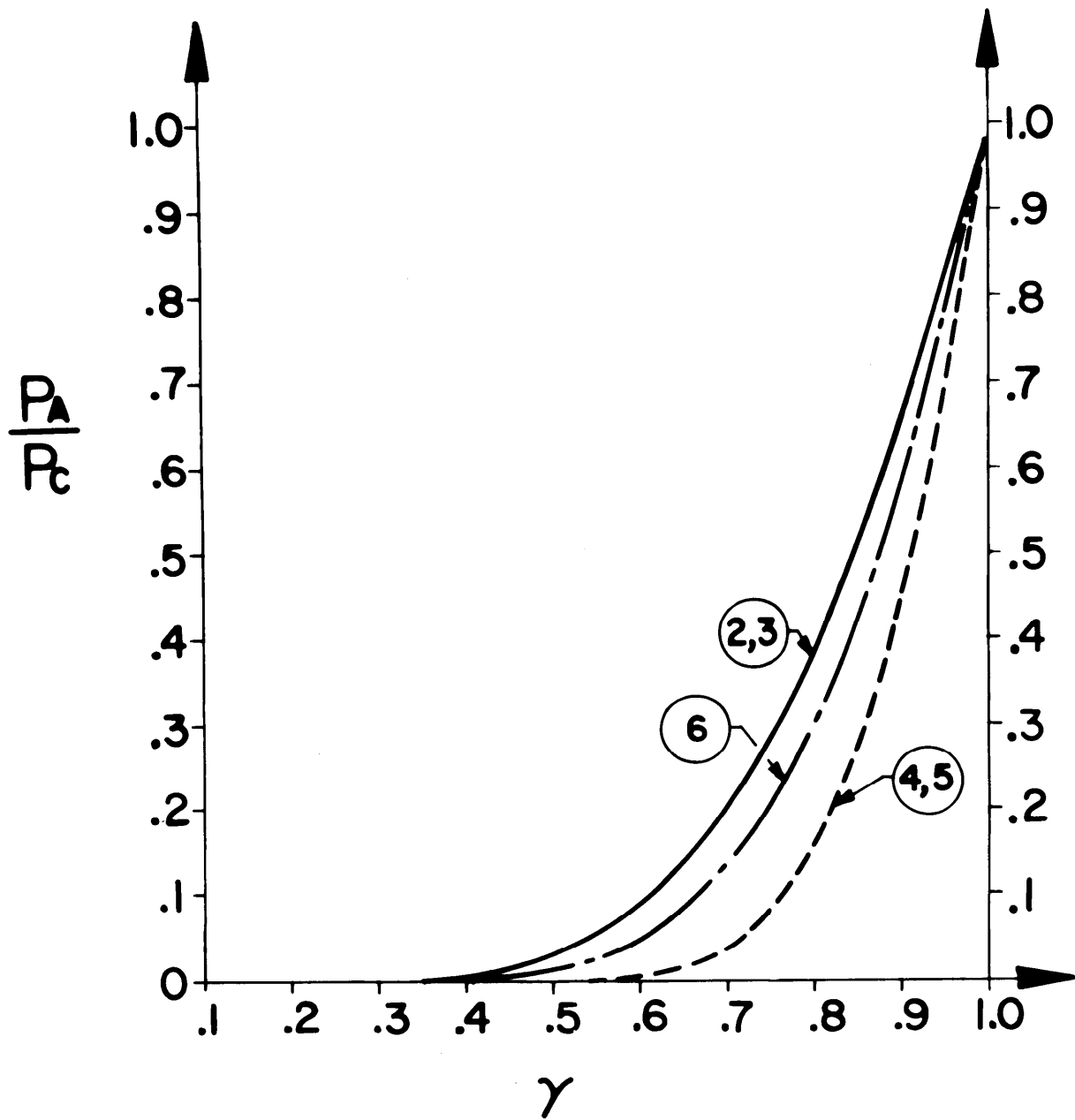


Figure 2. Vapor pressure vs. temperature for equations numbered in Table I. Vapor pressure is independent of Z_c for the examples shown.

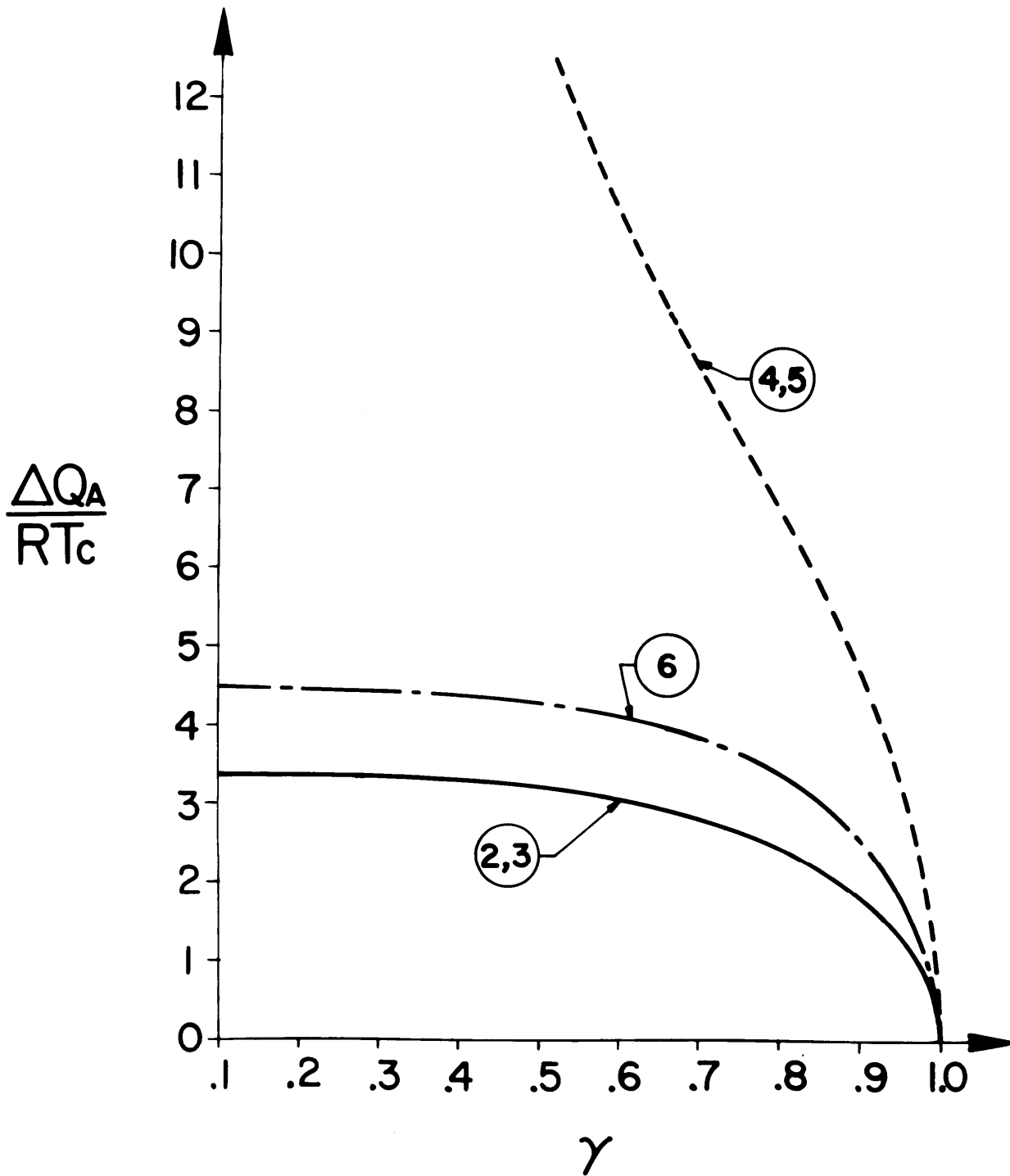


Figure 3. Latent heat of vaporization vs. reduced temperature for equations numbered in Table I. Latent heat is independent of Z_c for examples shown.

unrealistically large at small γ . Obviously, $\Delta Q_A(T)$ vanishes at $T = T_c$ for all equations. The entropy of vaporization, $\Delta S_A(T)$, is simply computed from $\Delta Q_A(T)$ by using $T \Delta S_A(T) = \Delta Q_A(T)$.

D. Specific Heat at Constant Volume, C_V .

With $E(V,T)$ the internal energy for unit mass (here taken to be the molar mass) the constant volume specific heat is $C_V(V,T) = \left(\frac{\partial E}{\partial T}\right)_V$, so that $\left(\frac{\partial C_V}{\partial V}\right)_T = \frac{\partial^2 E}{\partial V \partial T}$. The form of C_V within coexistence is $C_{VA}(V,T)$, and it differs from that of the single phase, $C_{VB}(V,T)$. By differentiating the thermodynamic relation

$$\left(\frac{\partial E}{\partial V}\right)_T = T \left(\frac{\partial P}{\partial T}\right)_V - P, \quad (31)$$

one obtains $\frac{\partial^2 E}{\partial T \partial V} = T \left(\frac{\partial^2 P}{\partial T^2}\right)_V$; using reduced variables and equating the above mixed derivatives, one gets:

$$\frac{\partial (C_{VB}/R)}{\partial \eta} = \gamma Z_c \left(\frac{\partial^2 \beta_B}{\partial \gamma^2}\right)_\eta. \quad (32)$$

Performing a partial integration with respect to η , one obtains

$$C_{VB}(\eta, \gamma)/R = C_V^0(\gamma)/R + \gamma F(\eta, \gamma) \quad (33)$$

where, with a prime hereafter signifying $\left(\frac{d}{d\gamma}\right)$,

$$F(\eta, \gamma) = \left\{ (9/8) / [\eta + C(\gamma)]^3 \right\} \left\{ A'' [\eta + C(\gamma)]^2 - [2A'C' + AC''] [\eta + C(\gamma)] + 2A(C')^2 \right\}. \quad (34)$$

Here $C_V^0(T)$ is the specific heat of the single phase at very low density $\{\eta \rightarrow \infty, F(\eta, \gamma) \rightarrow 0\}$; for a monatomic vapor at moderate temperature $C_V^0(T)$ is $3R/2$. Since C_{VB} must always be positive in the single phase,

further restrictions are imposed on the permissible functions $A(\gamma)$ and $C(\gamma)$. It is for this reason that B is forced to be a constant rather than a function of γ , as shown in the Appendix. Because $Z_c \frac{\partial^2 \beta_A}{\partial \gamma^2} = \left(\frac{\partial F}{\partial \eta} \right)_\gamma$, from Eq. (34) one obtains,

$$Z_c \frac{\partial^2 \beta_A}{\partial \gamma^2} = \left\{ (9/8) / [\eta + C(\gamma)]^4 \right\} \left\{ -A'' [\eta + C(\gamma)]^2 + 2[2A'C' + AC''] [\eta + C(\gamma)] - 6A(C')^2 \right\}. \quad (35)$$

If $C'(\gamma)$ and $A''(\gamma)$ are identically zero, then $\frac{\partial^2 \beta_A}{\partial \gamma^2}$ and $F(\eta, \gamma)$ are identically zero in the single phase.

Within coexistence, $T \leq T_c$, $V_3(T) \leq V_1(T)$, the specific heat is $C_{V_A}(V, T)$; it can be found by using $E_A(V, T)$ and $P_A(V, T)$ in Eq. (31), integrating it at fixed $T \leq T_c$ from $V_3(T)$ to some V within coexistence, and differentiating the result by T . Both $E_A(V, T)$ and $C_{V_A}(V, T)$ are linear functions of V . Using the fact that the internal energy and pressure functions are continuous across the boundary of coexistence, one can obtain in reduced variables

$$\{C_{V_A}(V, T) - C_V^0(T)\} / R = \Gamma(\eta, \gamma), \quad (36)$$

where

$$\Gamma(\eta, \gamma) = \gamma \eta_3'(\gamma) \left\{ \left(\frac{\partial \beta_A}{\partial \gamma} \right)_{\eta = \eta_3} - \beta_A'(\gamma) \right\} + \gamma F(\eta_3, \gamma) + \gamma \beta_A''(\gamma) [\eta - \eta_3(\gamma)]. \quad (37)$$

In general $\left(\frac{\partial \beta_A}{\partial \gamma} \right)_\eta$ is not equal to $\beta_A'(\gamma)$ at (η_3, γ) , so the constant volume specific heat exhibits a discontinuity across the coexistence boundary.

Recalling that the ξ_j are functions of u , one can use the inverse transformations and arrange the terms to obtain:

$$\Gamma(\eta, \gamma) = \{\Gamma(v, u)\}_0 k_1(\gamma) + (3-\xi_3)(\xi_3-\xi_1)\{k_2(\gamma) - (3-2\xi_3-\xi_1)k_3(\gamma)\} \quad (38)$$

$$+ \{(\xi_3-\xi)/(\xi_3\xi)\}\{u \frac{dw_A}{du} k_4(\gamma) + uw_A k_5(\gamma)\} + \gamma F_2(\xi_3, \gamma).$$

Here the k_i are functions of γ , and:

$$k_1 \equiv \{u'/f\}^2; \quad k_2 \equiv 6Z_c u' C' X/f$$

$$k_3 \equiv 3Z_c X [C']^2 / [f^2 A]; \quad k_4 \equiv AZ_c [Xu'' + 2X'u']$$

$$k_5 \equiv AZ_c X''; \text{ and with } m = \xi_3/[fA],$$

$$F_2(\xi_3, \gamma) = (9/8) m \{A'' - [2A'C' + AC''] m + 2A[C']^2 m^2\}.$$

$\{\Gamma(v, u)\}_0$ is the corresponding quantity for the [V-M] system at the point (v, u) inside coexistence. The value of $\Gamma(\eta, \gamma)$ at the critical point is easily found because both ξ and ξ_1 become equal to ξ_3 there, and the second and third terms vanish. The critical value of $\{\Gamma(1, 1)\}_0$ is known to be 4.5, and neither $k_1(\gamma)$ nor $F_2(1, 1)$ is infinite there. Therefore, $\Gamma(\eta, \gamma)$ is always finite at the critical point, and state equations of this class cannot show an infinite C_{vA} here, as suggested by some experimental data. Along the saturated liquid locus ($\xi = \xi_3$), the third term of Eq. (38) vanishes. When $C' \equiv 0$ (as happens for all example equations in Table I), Eq. (38) simplifies along the saturated liquid locus, (η_3, γ) , to:

$$\Gamma(\eta_3, \gamma) = \{\Gamma(v_3, u)\}_0 \{u'/f\}^2 + (9/8) \xi_3(u) \{\gamma A''/[fA]\}. \quad (39)$$

E. Vaporizing Signal Speed

The speed of a small amplitude adiabatic expansion wave (with partial vaporization) propagating into the saturated liquid⁵ is denoted by $a_A(V_3, T)$, and is equal to $V_3 \frac{dP_A}{dT} \{T/C_{vA}(V_3, T)\}^{\frac{1}{2}}$; in reduced notation,

$$a_A(V_3, T)/[RT_c]^{3/2} = z_c \eta_3 \frac{d\beta_A}{d\gamma} \{ \gamma / [\Gamma(\eta_3, \gamma) + C_V^0/R] \}^{1/2}. \quad (40)$$

A comparison of these speeds is shown in Figure 4 for the sample equations numbered 2, 4, and 6 in Table I, with the assumption that C_V^0/R is 3/2. The limiting values at the critical point for certain coexistence properties of sample state equations are shown in Table II.

IV. DISCUSSION

The predicted thermodynamic properties of coexistence for this class of elementary state equations are seen to be conveniently done in terms of the known [V-M] system. The use of different functions $A(\gamma)$, B and $C(\gamma)$, might improve the agreement between prediction and observation. However, this class of state equation contains one serious drawback in the critical region which cannot be overcome by modifying these functions. This defect is that this entire class predicts $|P-P_c|$ to vary as the third power of $|V-V_c|$ on the critical isotherm; recent data suggests that this exponent is larger than three⁶, and perhaps as large as five⁷. With this restriction, however, this class equation is useful in giving a qualitative survey of some subcritical transition properties of liquids and gases.

Table II. Critical point values for designated quantities
 State equation numbers correspond to those
 of Table I. Value of C_V^0 was assumed
 $3R/2$ to compute last column.

Equation No.	$\frac{d^2 \beta_A}{d\gamma^2}$	$\frac{\partial^2 \beta_B}{\partial \gamma^2}$	$(C_{V_A} - C_V^0)/R$	$(C_{V_B} - C_V^0)/R$	$a_A/(RT_c)^{\frac{1}{2}}$
2	$9 \frac{3}{5}$	0	$4 \frac{1}{2}$	0	.612
3	$9 \frac{3}{5}$	0	$4 \frac{1}{2}$	0	1.633 Z_c
4	$32 \frac{2}{5}$	-6	$20 \frac{1}{4}$	$2 \frac{1}{4}$.563
5	$32 \frac{2}{5}$	-6	$20 \frac{1}{4}$	$2 \frac{1}{4}$	1.501 Z_c
6	$17 \frac{2}{5}$	0	8	0	1.622 Z_c
7	$26 \frac{2}{3}$	0	$12 \frac{1}{2}$	0	1.871 Z_c

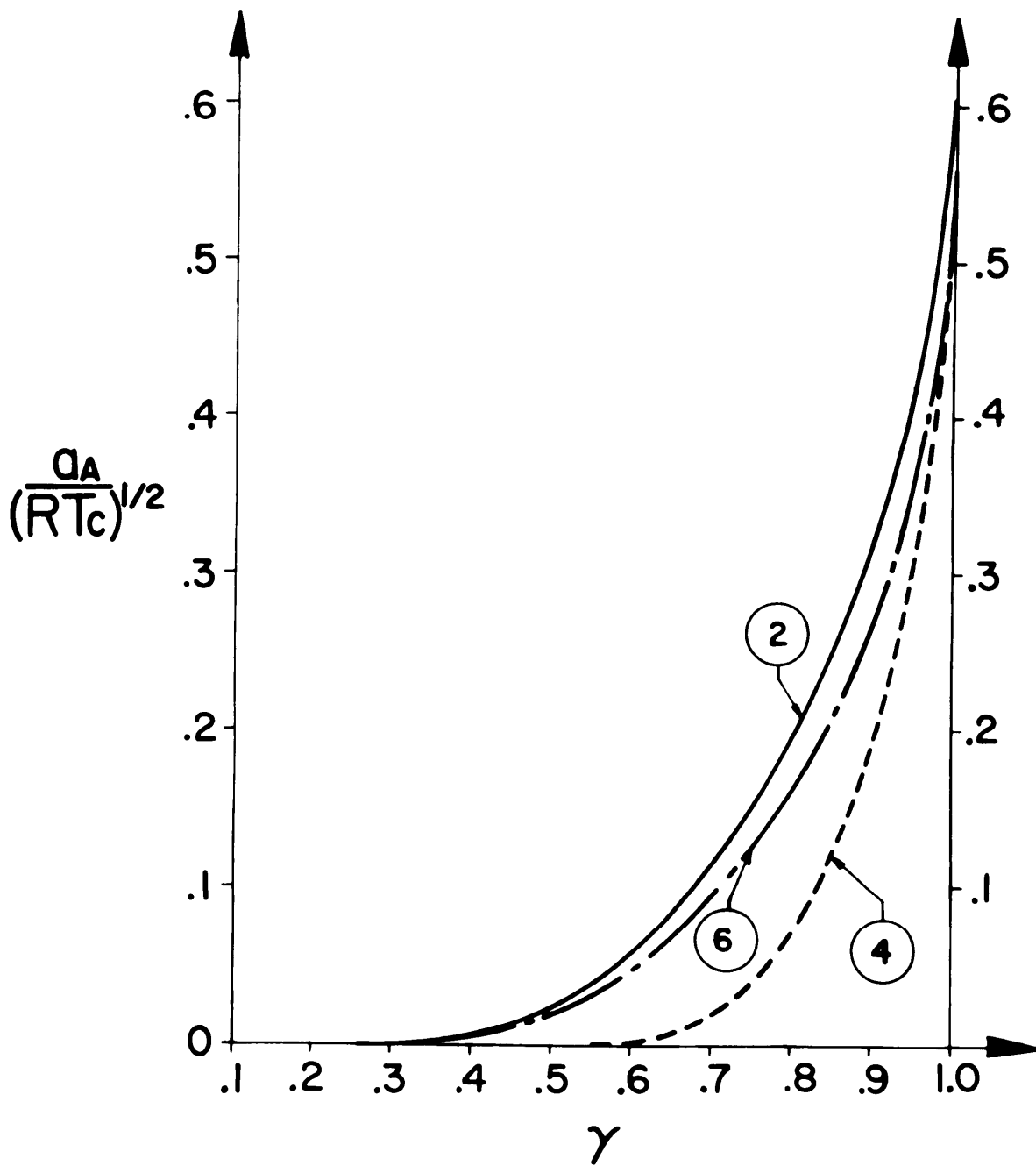


Figure 4. Vaporizing signal speed of the saturated liquid for equations numbered in Table I. The curve shown for No. 6 was computed for $Z_c = 1/3$. The value $3R/2$ was used for C_v^0 .

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- * Martin uses the forms {his Eq. (28) in present notation}
 $A(\gamma) = (4-\gamma)/8Z_c$, $B = .085/Z_c$ and $C(\gamma) = .04/Z_c$, and cautions that the critical conditions, $\gamma = \beta = 1$, $\frac{\partial \beta}{\partial \eta} = \frac{\partial^2 \beta}{\partial \eta^2} = 0$, will hold only for $Z_c = .335$. He sometimes gives up these critical conditions to obtain a better overall fit to experimental P, V, T data outside coexistence. A. V. Grosse uses this same form with $Z_c = .27$ to describe approximately the properties of cesium, {Inorg. Nucl. Chem. Letters 4, 261 (1968)}.
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APPENDIX

I. Showing B must be Independent of γ

If B were allowed to be a function of γ , there would be additional terms in $C_{vB}(V,T)$ of Eq. (33). These extra terms would be

$$- \left\{ (9/8) R\gamma/(\eta-B) \right\} \left\{ 2B' + \gamma(B')^2/(\eta-B) \right\}.$$

Since η can approach B in the single phase of this fluid, the last term would contribute a dominant negative value to $C_{vB}(V,T)$, unless $B'(\gamma)$ were identically zero. Because of the stability condition $C_v(V,T) > 0$, it is then necessary that $B(\gamma)$ be a constant. Restrictions are also required on $A(\gamma)$ and $C(\gamma)$, but they are not as simple. With $x = [\eta + C(\gamma)]^{-1}$, and the definitions $L(\gamma) \equiv 2A''(\gamma)$, $M(\gamma) \equiv 2A'C' + AC''$ and $N(\gamma) \equiv 2AC''$, then Eq. (33) becomes

$$C_{vB}/R = C_v^0/R + (9/8) \gamma x \{ L(\gamma) - xM(\gamma) + x^2N(\gamma) \}, \quad (A-1)$$

where, for $\gamma \leq 1$, the permissible x range over a positive interval from 0 to $x_1 = + [B + C(\gamma)]^{-1} < + \infty$; for $\gamma \geq 1$, permissible values of x range over a similar interval, excluding, however, a sub-interval corresponding to η values within the coexistence region. {For this sub-interval, the proper expression for C_v is C_{vA} , given by Eq. (36)}. Eq. (A-1) is a cubic polynomial in x , and can have a variety of shapes, depending on the coefficients $L(\gamma)$, $M(\gamma)$ and $N(\gamma)$. The functions $A(\gamma)$ and $B(\gamma)$ must be restricted so C_v is always positive. This requirement is satisfied for all the example equations shown in Table I.

II. Reduced Variables for the [V-M] Fluid

The set of values given in Table III for the [V-M] system can be used with the transformation equations to compare easily the sub-critical

Table III. Reduced values of coexistence temperature, saturation densities, and vapor pressure for the Van der Waals-Maxwell state equation

u	ξ_3	ξ_1	w_A
1.000	1.0000	1.0000	1.0000
.998	1.0902	.91140	.99202
.994	1.1571	.84773	.97617
.990	1.2035	.80454	.96048
.98	1.2894	.72669	.92191
.97	1.3558	.66844	.88429
.96	1.4121	.62042	.84762
.95	1.4617	.57901	.81188
.93	1.5482	.50931	.74318
.90	1.6573	.42574	.64700
.85	1.8071	.31973	.50449
.80	1.9327	.23967	.38336
.75	2.0424	.17721	.28246
.70	2.1404	.12802	.20046
.65	2.2296	.89475x10 ⁻¹	.13584
.60	2.3116	.59778x10 ⁻¹	.86869x10 ⁻¹
.55	2.3875	.37580x10 ⁻¹	.51580x10 ⁻¹
.50	2.4585	.21747x10 ⁻¹	.27789x10 ⁻¹
.45	2.5251	.11217x10 ⁻¹	.12134x10 ⁻¹
.40	2.5879	.49109x10 ⁻²	.51745x10 ⁻²
.35	2.6475	.16875x10 ⁻²	.15673x10 ⁻²
.30	2.7042	.39907x10 ⁻³	.31882x10 ⁻³
.25	2.7583	.51259x10 ⁻⁴	.34165x10 ⁻⁴
.20	2.8012	.22296x10 ⁻⁵	.11891x10 ⁻⁵
.15	2.8602	.41565x10 ⁻⁷	.43602x10 ⁻⁸
.10	2.9083	.21612x10 ⁻¹²	.57631x10 ⁻¹³
.04	2.9640	.56824x10 ⁻³⁴	.60612x10 ⁻³⁵
0.00	3.0000	0.00	0.00

thermodynamic predictions for any state equation belonging to this class. Five significant figures are given for every function except u , which is assumed to be exact. These values were computed at this laboratory with the assistance of D. C. Mylin and F. H. MacIntosh. R. Barieau¹ gives a much more extensive tabulation of various thermodynamic properties for the [V-M] fluid for $.25 \leq u \leq 1.0$. For these u , the above values agree with his tabulations.

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